Supplementary Data

Highly Selective Fluorogenic Anion Chemosensors:

Naked-eye Detection of F⁻ and AcO⁻ ions in Natural Water using a Test Strip

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Table of Contents:

Fig. S9. Benesi-Hildebrand plot: UV-visible titration result of S2 with (a) fluoride & (b) acetate anion.

Fig. S10. Fluorescence emission spectra of S1 (5 x 10^{-5} M) in DMSO solution upon addition of 3 equiv. of different anions as their tetrabutylammounium salt respectively; (λ_{ex} at 419 nm).
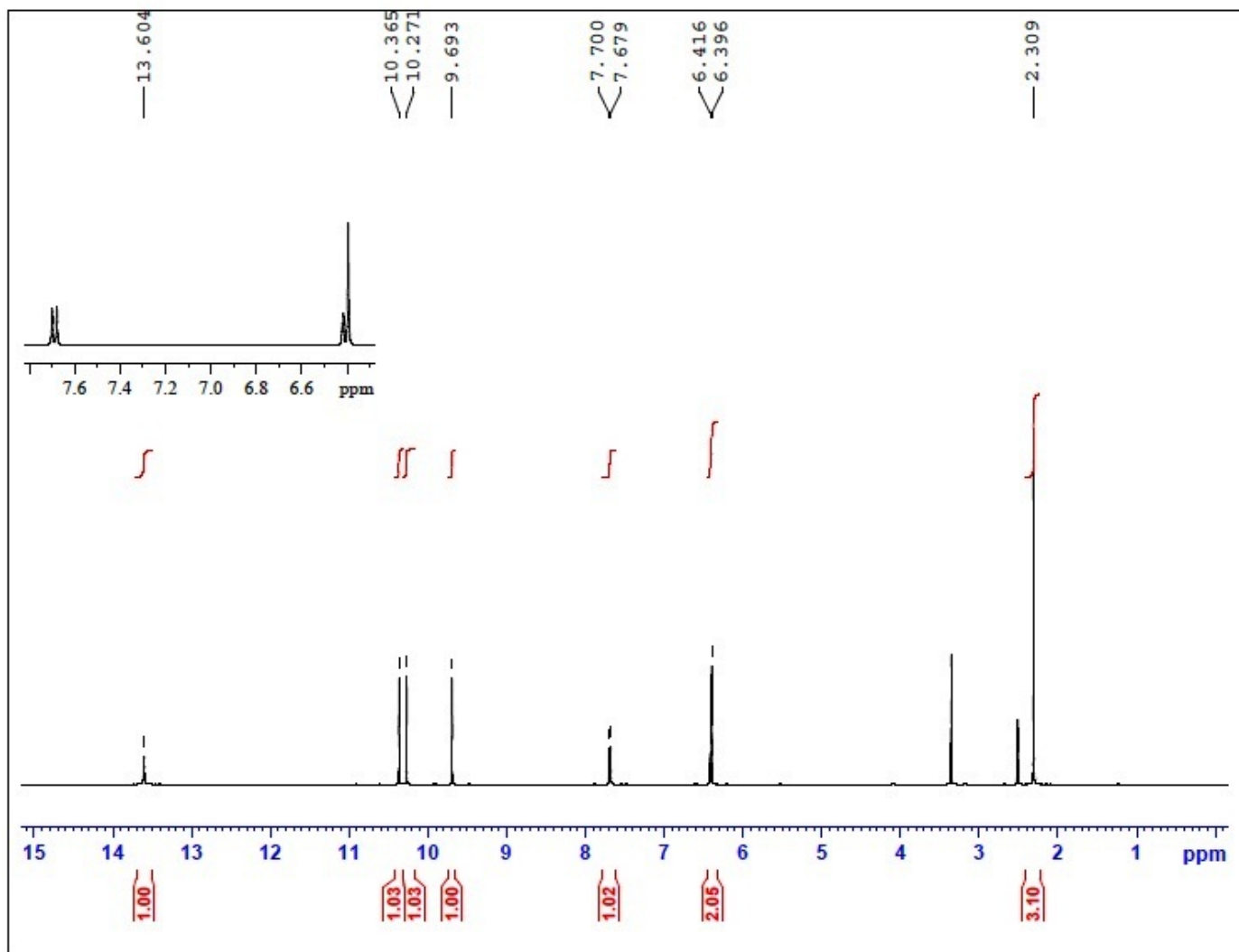
Fig. S11. Fluorescence emission spectra of S1 (5 x 10^{-5} M) in DMSO solution upon addition of 3 equiv. of different anions as their tetrabutylammounium salt, respectively; (λ_{ex} at 450 nm).

Fig. S12. Comparison bargraph of fluorescence emission spectra of S1(λ_{ex} at 419 nm) and S2 (λ_{ex} at 419 nm) in DMSO solution (5 x 10^{-5} M) towards F¯and AcO¯ (3 equiv) anions, respectively.

Fig. S13. Changes in emission spectra of S1 in DMSO (5 x 10^{-5} M) upon (a) addition of F¯ & (b) AcO¯ anions; λ_{ex} at 450 nm.
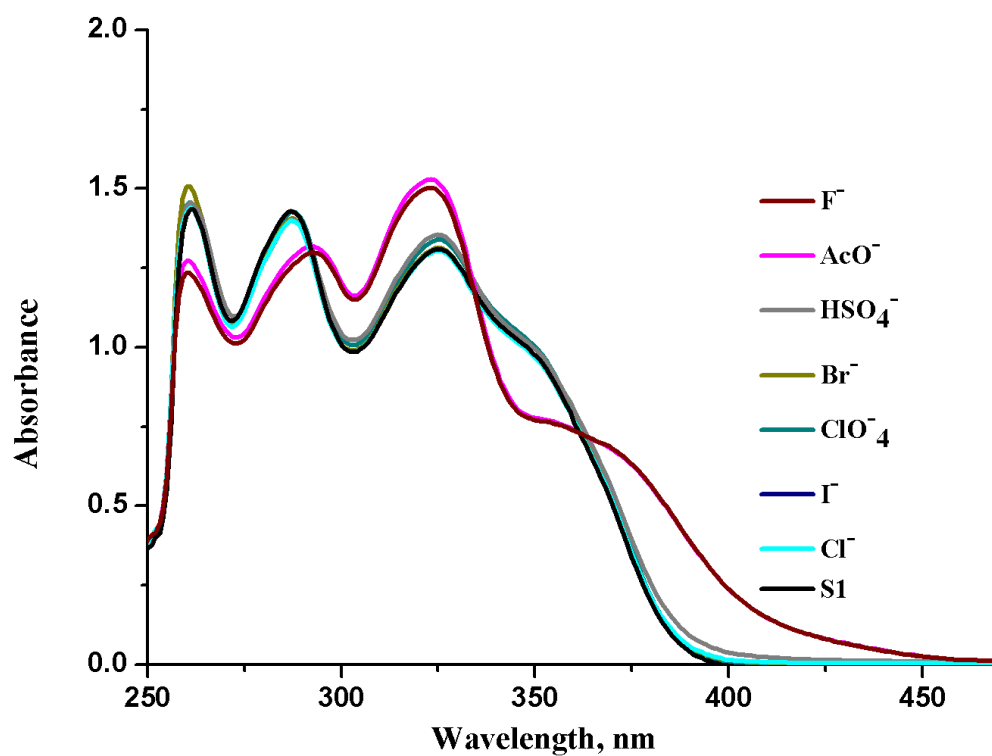
**Fig. S14.** Changes in emission spectra of S2 in DMSO (5 x 10^{-5} M) upon (a) addition of F¯ & (b) AcO¯ anions; λ_{ex} at 450 nm.

Fig. S1. ^{1}H NMR spectrum of the S1 (400 MHz) in DMSO-d_6 solvent. (Inset: Expansion of the aromatic proton region 6.3-7.8 ppm).
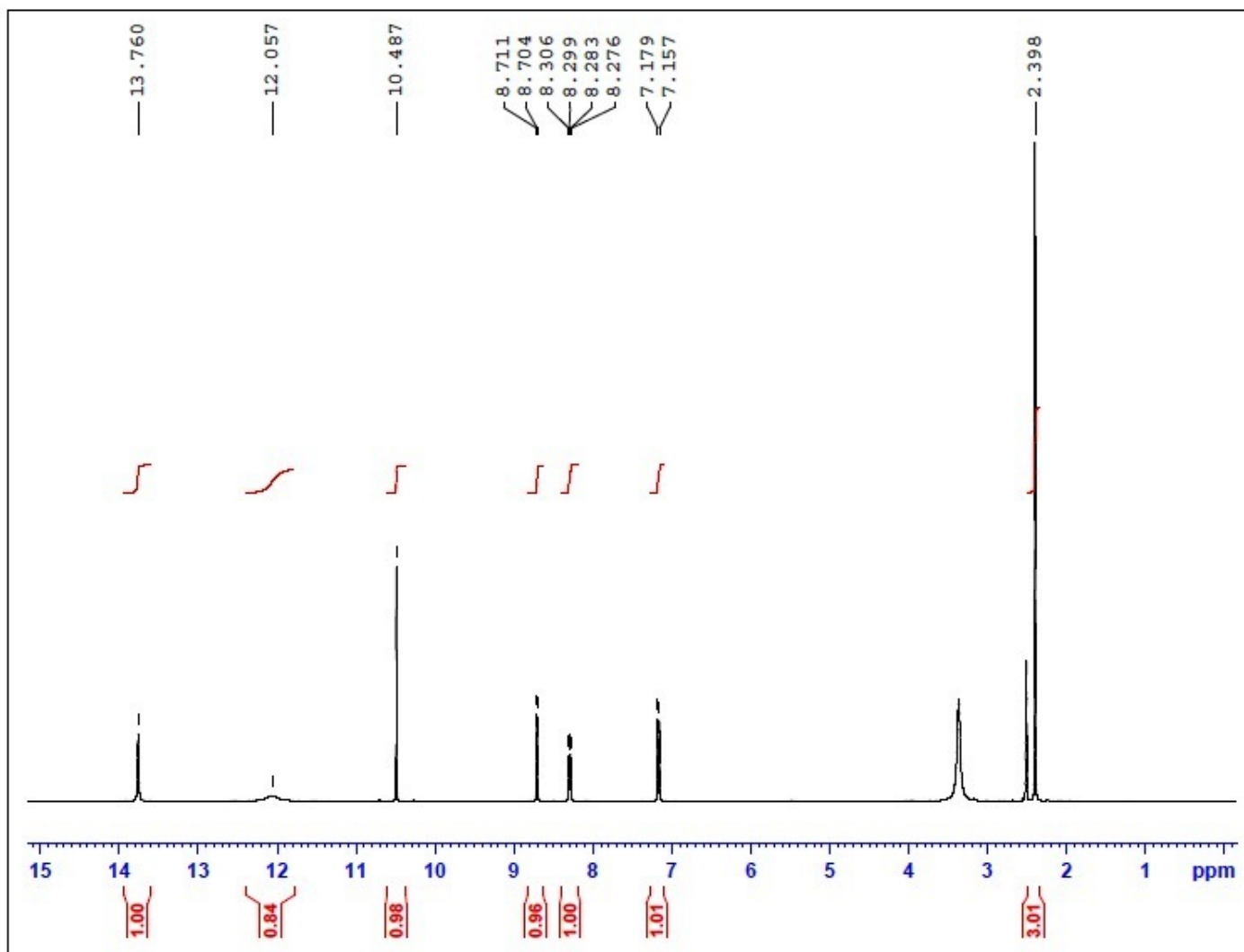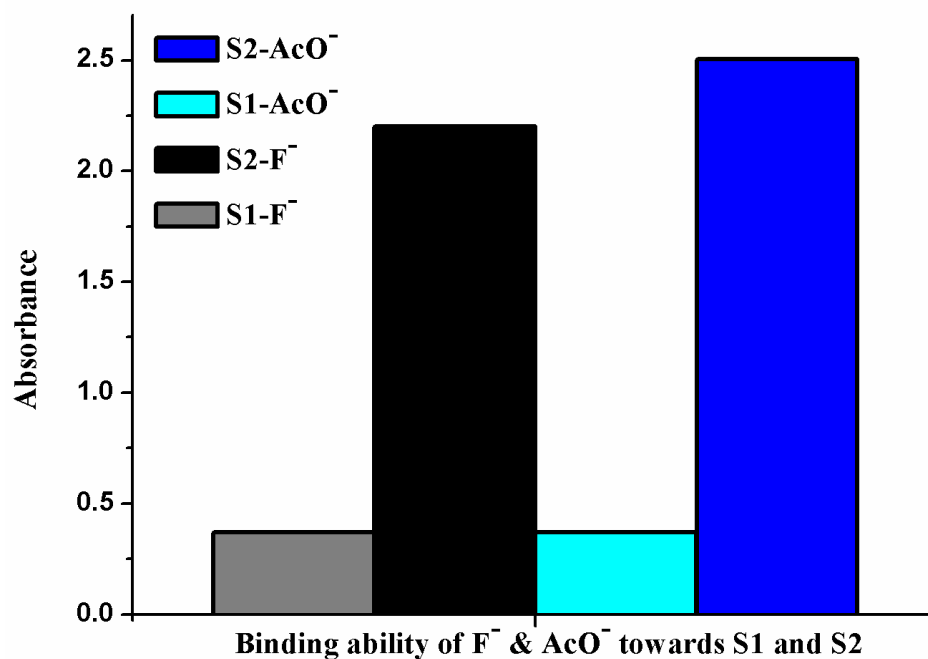
Fig. S2. ^{1}H NMR spectrum of the S2 (400 MHz) in DMSO-d_6 solvent.

Fig. S3. UV-visible spectra of S1 (5×10^{-5} M) in DMSO solution upon addition of 3 equiv. of different anions as their tetrabutylammounium salt respectively.
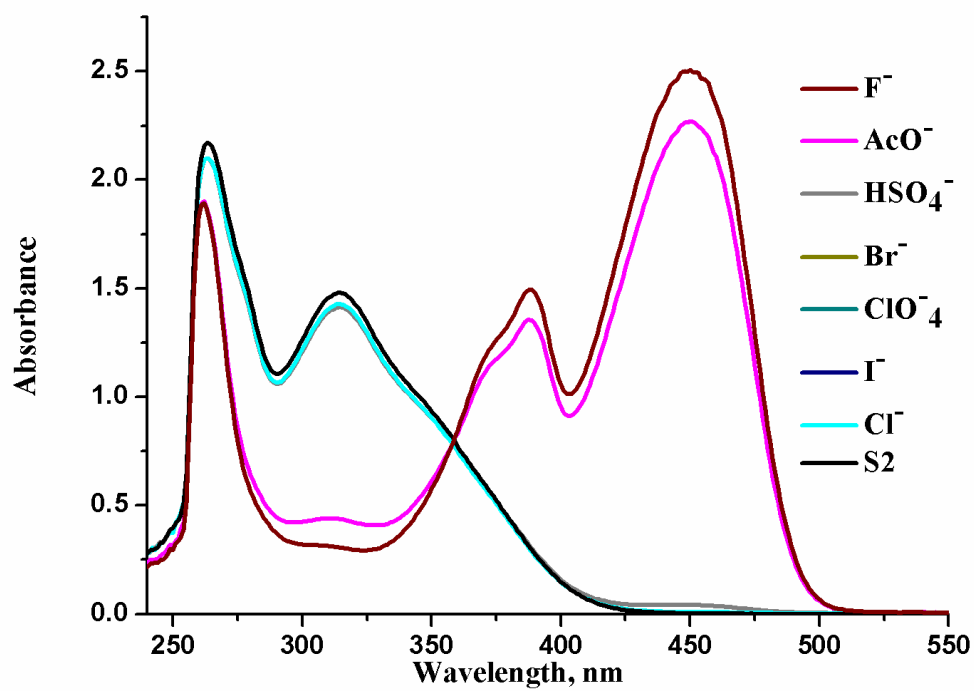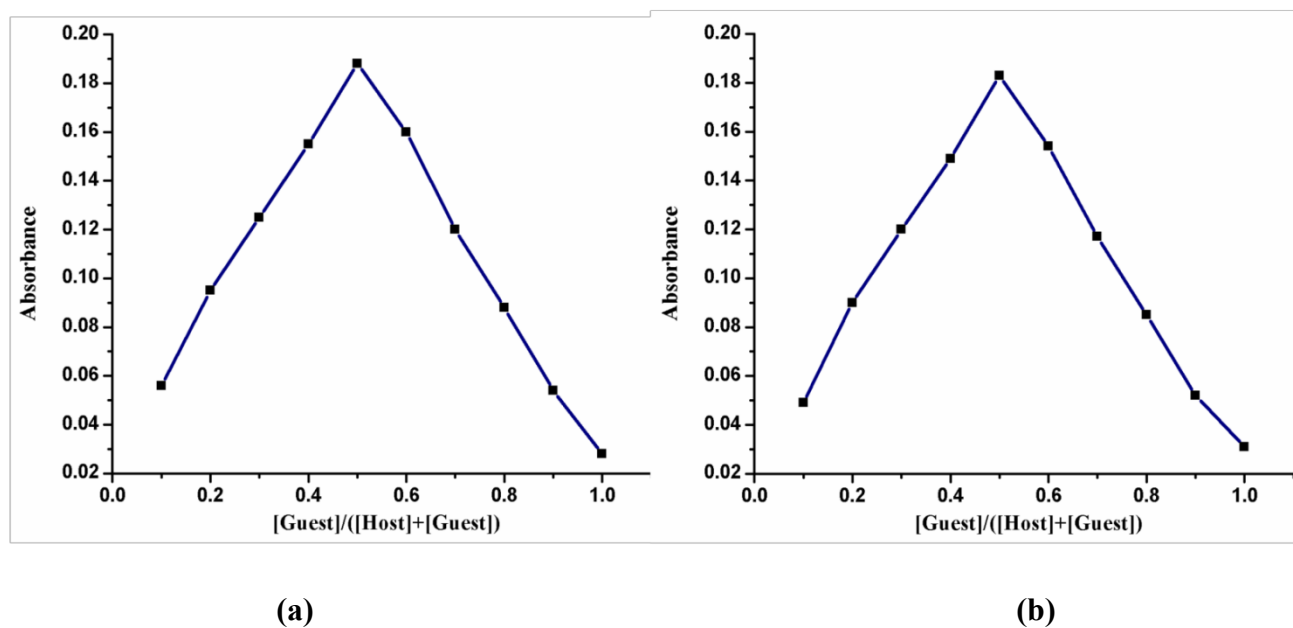
Fig. S4. UV-visible spectra of S2 (5×10^{-5} M) in DMSO solution upon addition of 3 equiv. of different anions as their tetrabutylammounium salt respectively.
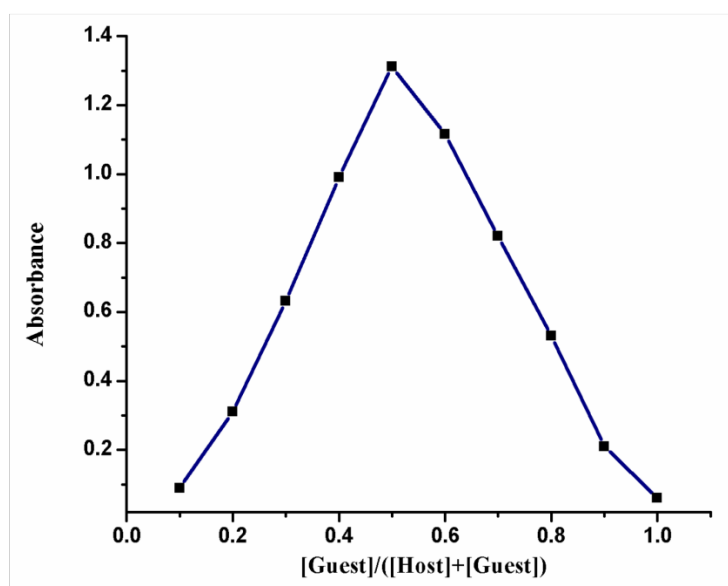
Fig. S5. Comparison bargraph of UV-visible spectra of S1 and S2 (5×10^{-5} M) in DMSO solution towards F^- and AcO^- (3 equiv) anions, respectively.
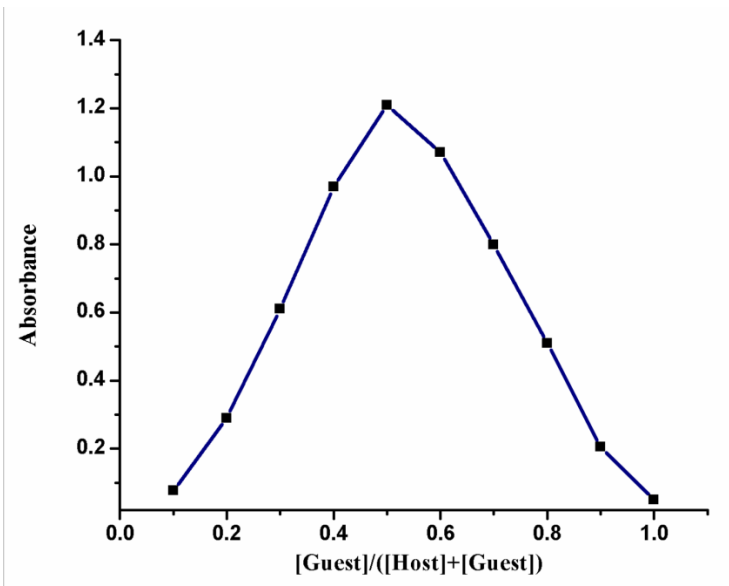
(a)   **(b)**

Fig. S6. Job plots for complexation of S1 with (a) fluoride & (b) acetate anions.
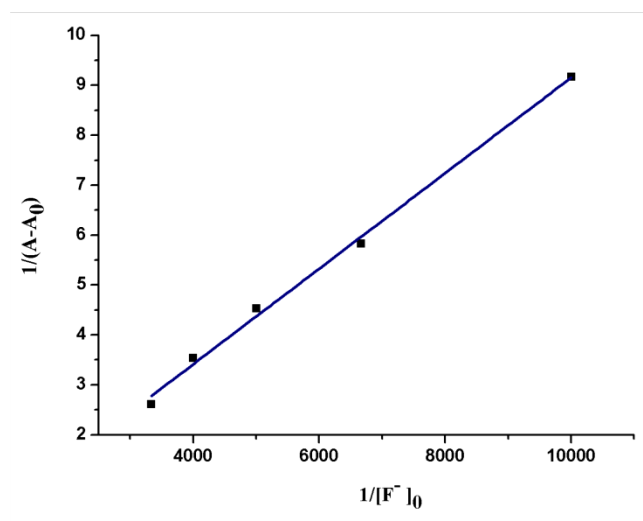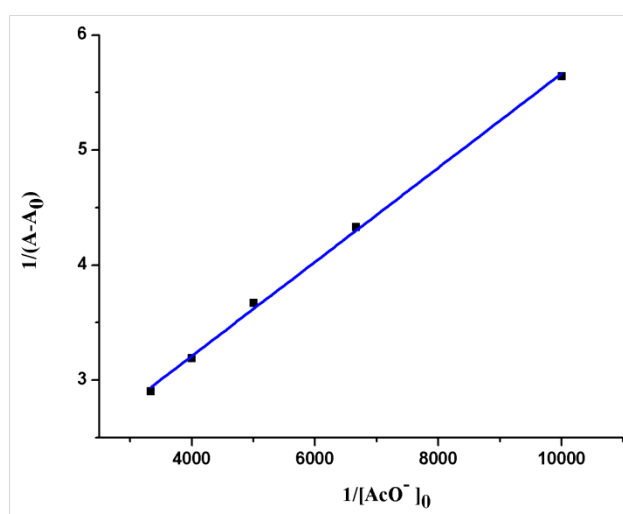
(a)

(b)

Fig. S7. Job plots for complexation of S2 with (a) fluoride & (b) acetate anions.

(a)   **(b)**

Fig. S8. Benesi-Hildebrand plot: UV-visible titration result of S1 with (a) fluoride & (b) acetate anion.

(a)   **(b)**
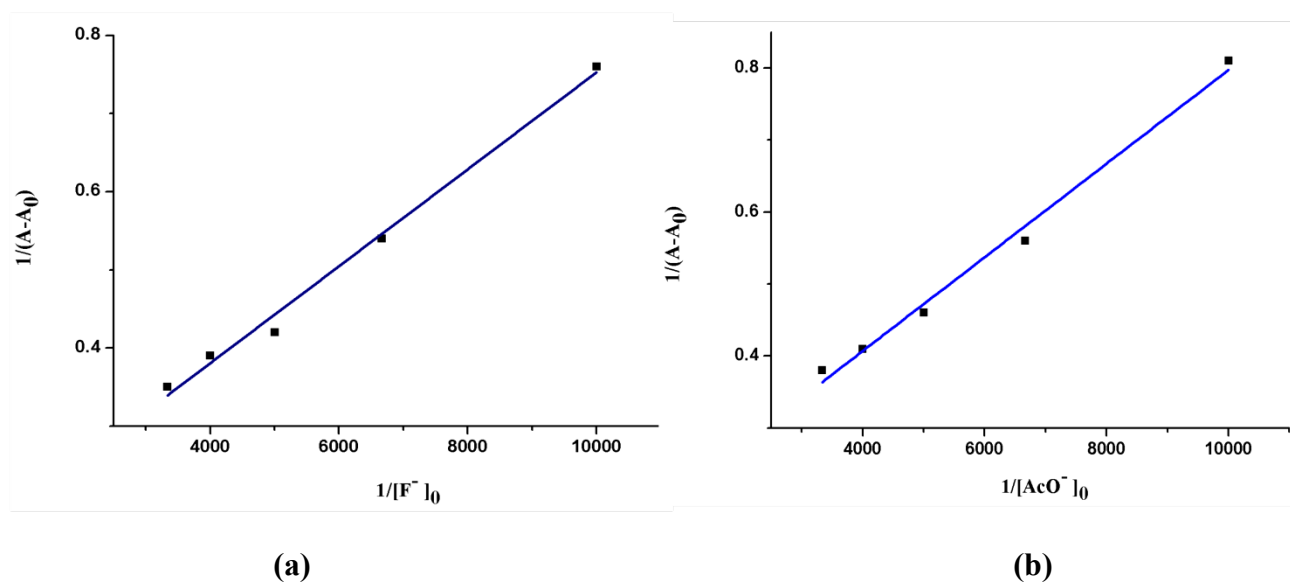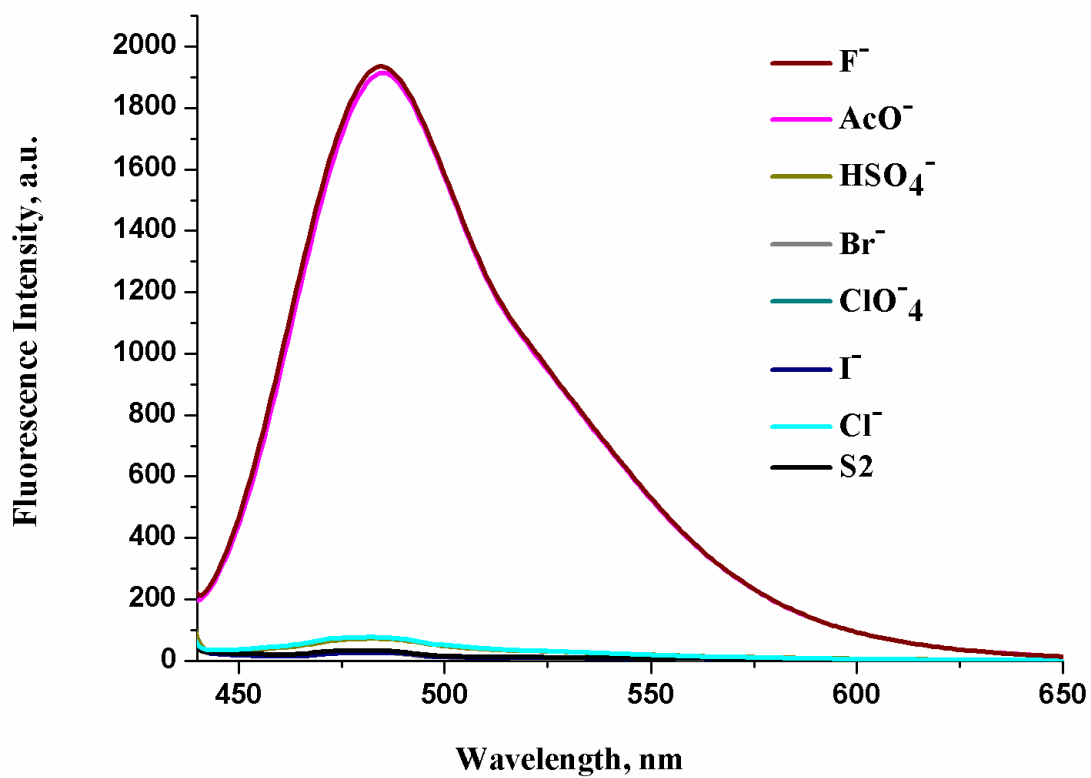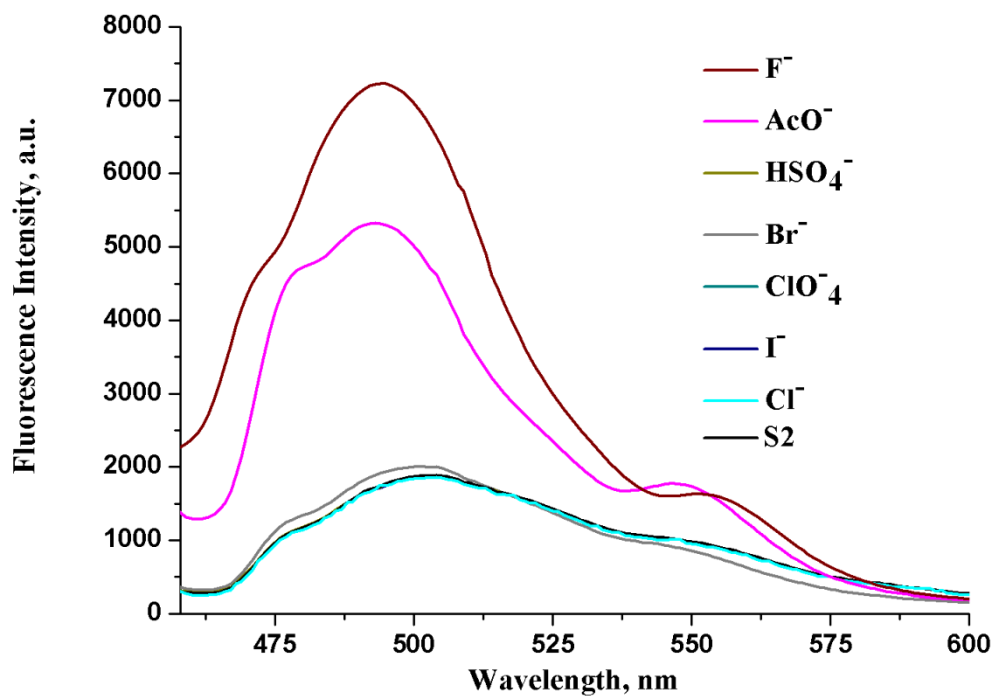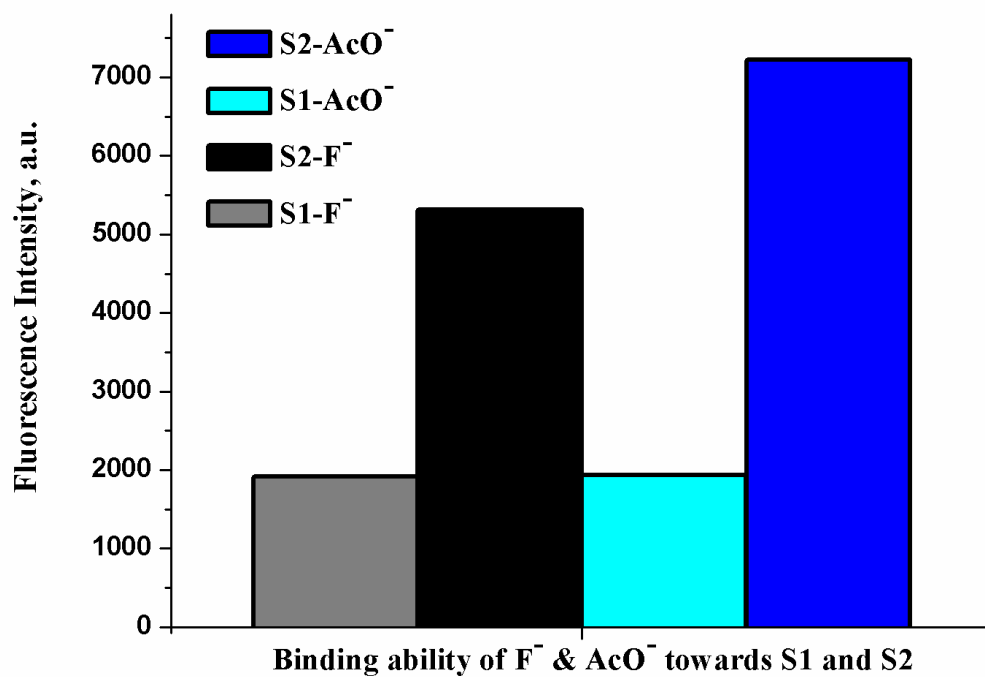
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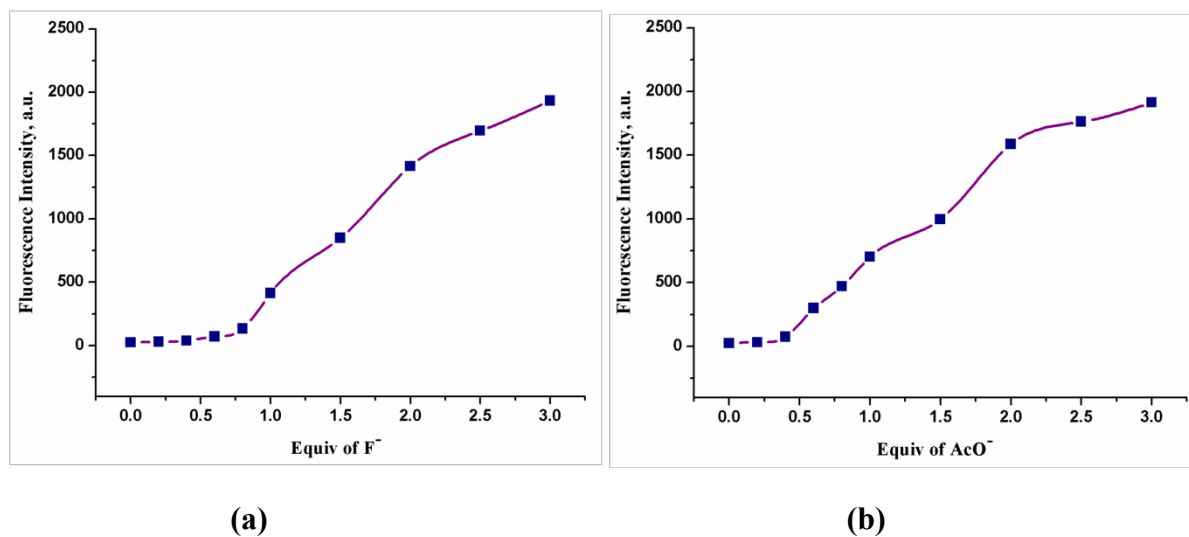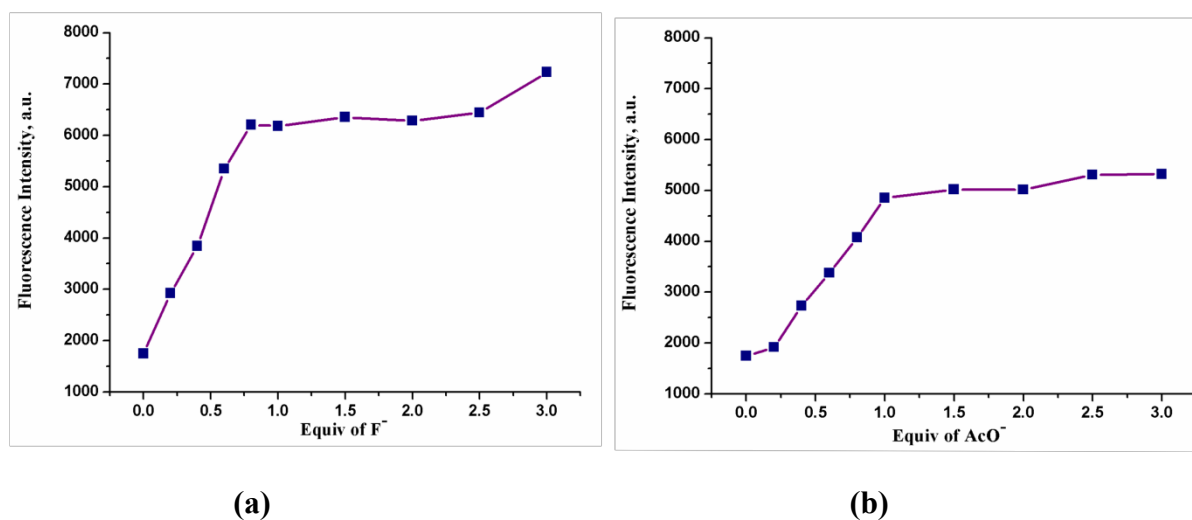
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(a)   (b)

Fig. S13. Changes in emission spectra of S1 in DMSO (5 x 10^{-5} M) upon (a) addition of F$^-$ & (b) AcO$^-$ anions; λ_{ex} at 450 nm.

(a)   **(b)**

Fig. S14. Changes in emission spectra of S2 in DMSO (5×10^{-5} M) upon (a) addition of F^- & (b) AcO^- anions; λ_{ex} at 450 nm.