

An ionic charge-transfer dyad prepared cost-effectively from tetrathiafulvalene carboxylate anion and TMPyP cation

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Chart S1 Structure of TTF(CONH₂)₂.

Fig. S1 Frontier molecular orbitals of **2** calculated using DFT B3LYP/6-311G** method.

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Fig. S4 Comparison of the fluorescence spectra of **2** (λ_{ex} = 488 nm) with that of TMPyP (λ_{ex} = 470 nm), recorded at 1.0 × 10⁻³ and 5.0 × 10⁻⁴ mol·L⁻¹ concentrations in DMF at room temperature.

Fig. S5 The relationship between the quenching degree of the fluorescence spectra at the peak of 652 nm and the concentration of **1** [Y-axis (quenching degree) = (fluorescence intensity of TMPyP – fluorescence intensity of **1**)/ fluorescence intensity of TMPyP].

Fig. S6 Fluorescence spectra of TMPyP, **1**, and **2** in solid state at room temperature.

Fig. S7 Changes of fluorescence spectra of TMPyP in DMF at room temperature, upon quantitative titration of TTF(CONH₂)₂.

Fig. S8 Changes of fluorescence spectra of TMPyP (2.0 × 10⁻⁵ mol·L⁻¹) in DMF at room temperature, upon addition of increasing amounts of L¹ and L².

Table S1. Crystal Data and Structural Refinement Parameters for **1** and **2**

	1	2
formula	C ₈₄ H ₇₂ N ₈ O ₁₉ S ₂₄	C ₆₄ H ₅₂ N ₈ O ₂₁ S ₈
fw	2267.07	1525.61
cryst size (mm ³)	0.25 × 0.26 × 0.75	0.05 × 0.12 × 0.20
cryst syst	monoclinic	monoclinic
space group	<i>P</i> 2 ₁ / <i>m</i>	<i>C</i> 2/ <i>c</i>
<i>a</i> (Å)	10.8302(4)	23.6681(8)
<i>b</i> (Å)	33.4895(16)	22.0587(6)
<i>c</i> (Å)	13.8319(6)	13.2272(4)
<i>α</i> (deg)	90.00	90.00
<i>β</i> (deg)	99.191(4)	111.589(3)
<i>γ</i> (deg)	90.00	90.00
<i>V</i> (Å ³)	4952.4(4)	6421.3(4)
<i>Z</i>	2	4
ρ_{calcd} (g cm ⁻³)	1.492	1.578
<i>F</i> (000)	2300	3152
μ (mm ⁻¹)	5.376	0.365
<i>T</i> (K)	293(2)	293(2)
reflns collected	20829	19610
unique reflns	9892	5651
observed reflns	6181	4629
no. params	638	486
GOF on <i>F</i> ²	1.182	1.034
<i>R</i> ₁ [<i>I</i> > 2 σ (<i>I</i>)]	0.0837	0.0401
<i>wR</i> ₂ [<i>I</i> > 2 σ (<i>I</i>)]	0.1767	0.1063

Table S2. The energy of the frontier orbitals and the energy gaps of **2**.

orbital	energy (eV)	energy gap
LUMO+1	-3.132	$E_{\text{LUMO+1}} - E_{\text{HOMO}} = 1.164$ eV (1065.2 nm)
LUMO	-3.259	$E_{\text{LUMO+1}} - E_{\text{HOMO-1}} = 1.171$ eV (1058.8 nm)
HOMO	-4.296	$E_{\text{LUMO}} - E_{\text{HOMO}} = 1.001$ eV (1238.6 nm)
HOMO-1	-4.303	$E_{\text{LUMO}} - E_{\text{HOMO-1}} = 1.044$ eV (1187.6 nm)

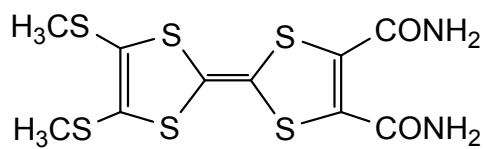


Chart S1 Structure of TTF(CONH₂)₂.

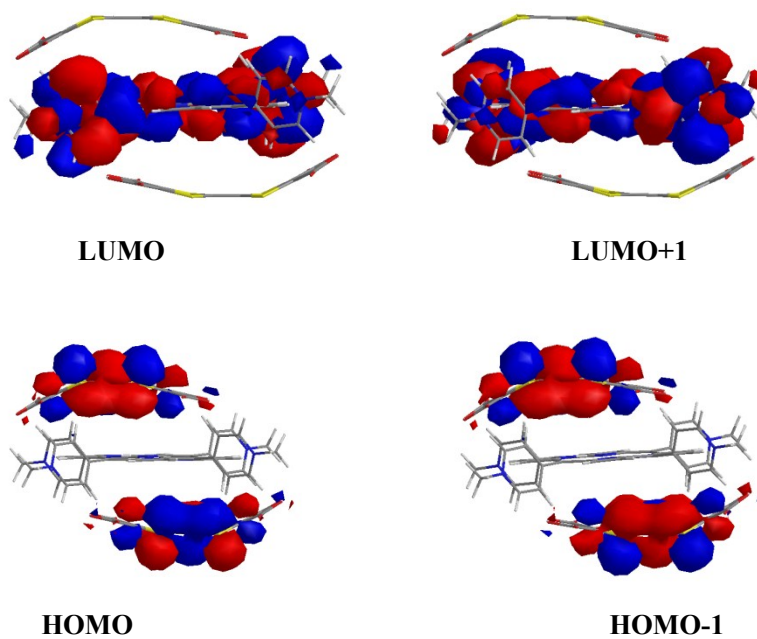


Fig. S1 Frontier molecular orbitals of **2** calculated using DFT B3LYP/6-311G** method.

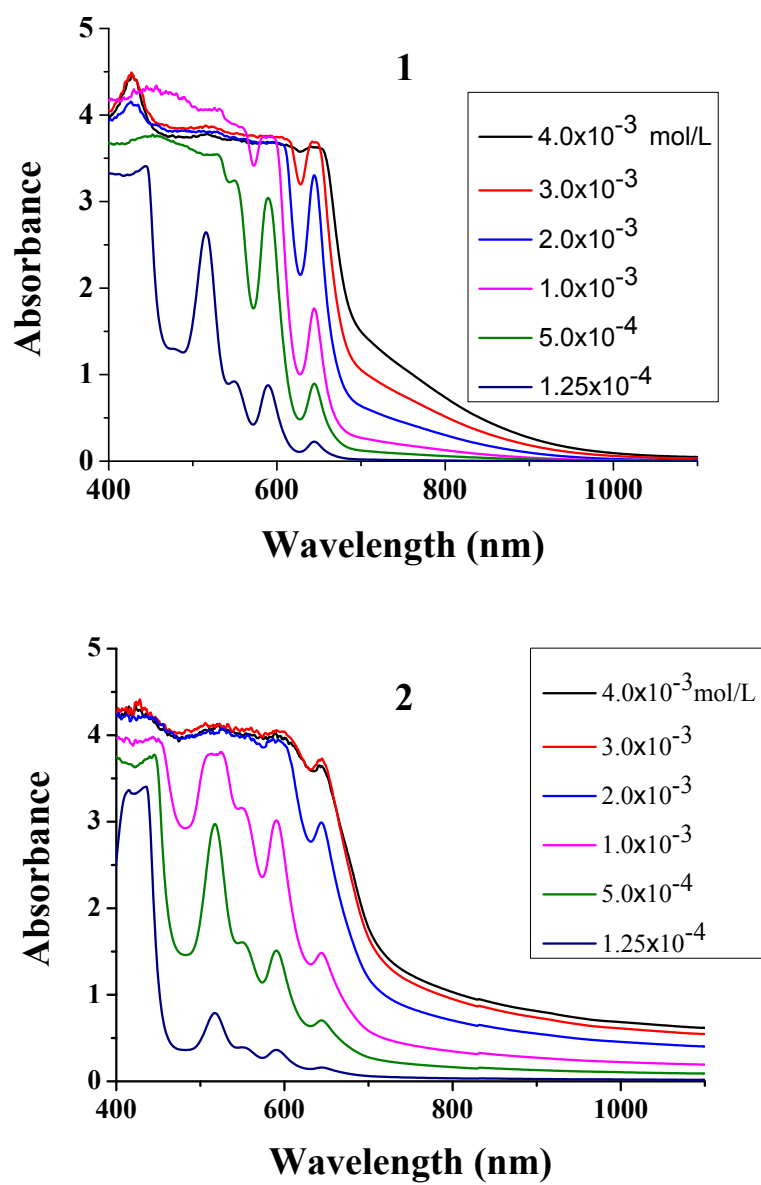


Fig. S2 UV-vis absorption of crystals **1** and **2** in DMF in different concentrations.

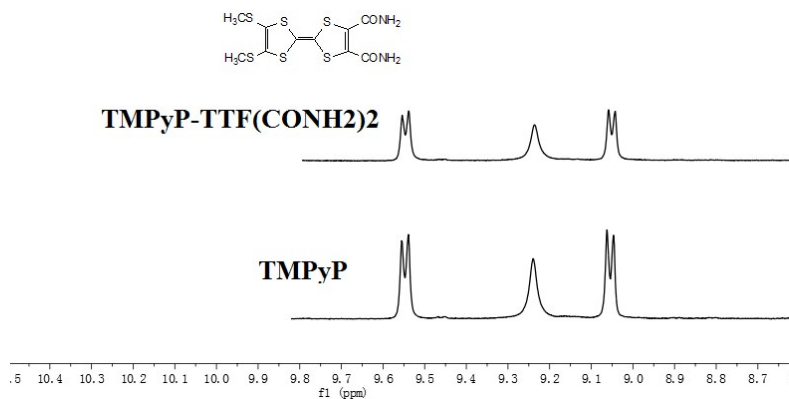


Fig. S3 Partial ^1H NMR spectra of TMPyP and the mixture of TMPyP and 4eq. TTF(CONH₂)₂ in *d*₆-DMSO (the concentration of TMPyP is 4×10^{-3} mol·L⁻¹).

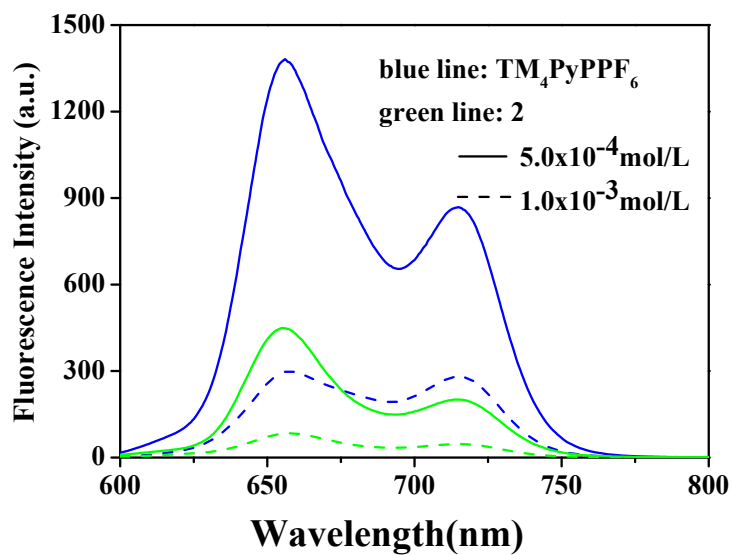


Fig. S4 Comparison of the fluorescence spectra of **2** ($\lambda_{\text{exc}} = 488$ nm) with that of **TMPyP** ($\lambda_{\text{exc}} = 470$ nm), recorded at 1.0×10^{-3} and 5.0×10^{-4} mol·L⁻¹ concentrations in DMF at room temperature.

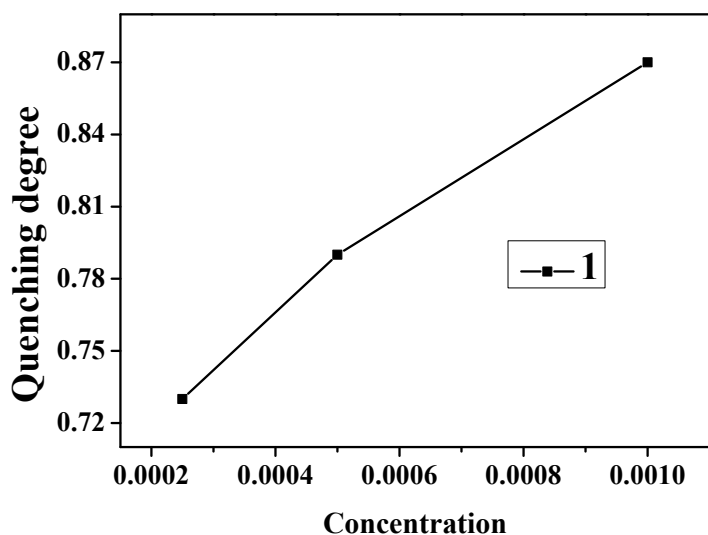


Fig. S5 The relationship between the quenching degree of the fluorescence spectra at the peak of 652 nm and the concentration of **1** [Y-axis (quenching degree) = (fluorescence intensity of TMPyP – fluorescence intensity of **1**)/ fluorescence intensity of TMPyP].

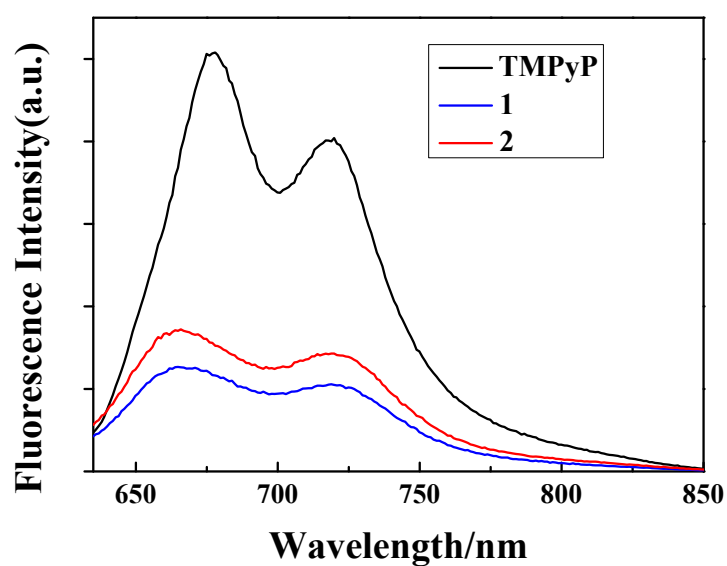


Fig. S6 Fluorescence spectra of TMPyP, **1**, and **2** in solid state at room temperature.

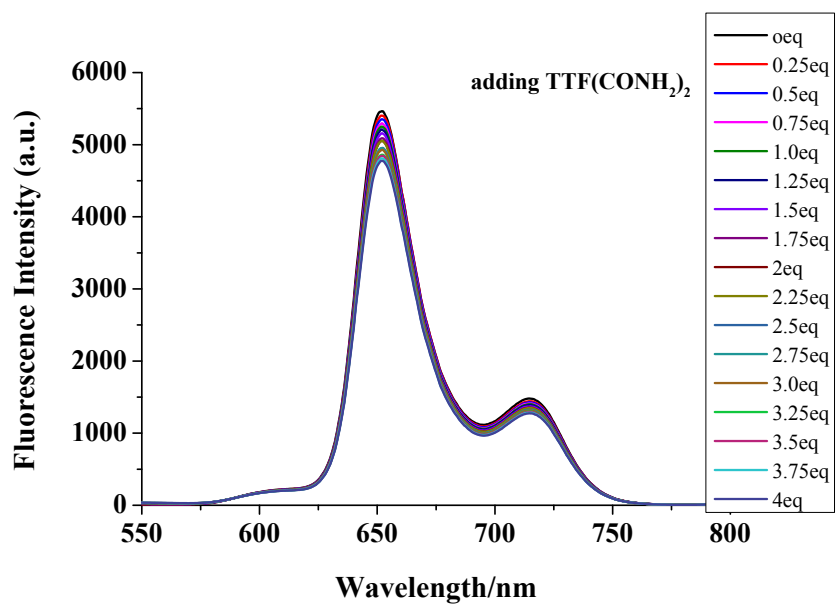
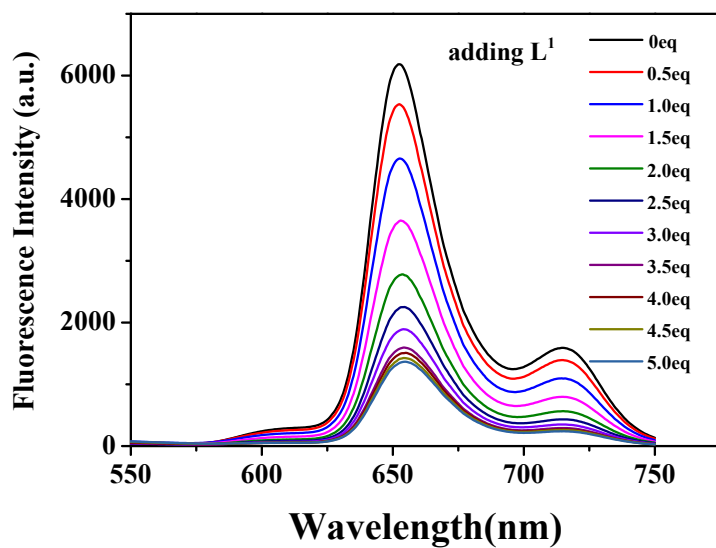


Fig. S7 Changes of fluorescence spectra of TMPyP in DMF at room temperature, upon quantitative titration of TTF(CONH₂)₂.



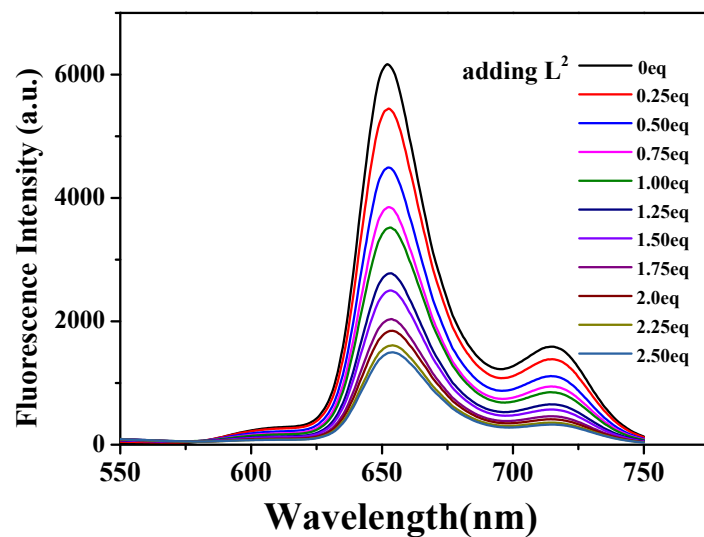


Fig. S8 Changes of fluorescence spectra of TMPyP ($2.0 \times 10^{-5} \text{ mol}\cdot\text{L}^{-1}$) in DMF at room temperature, upon quantitative titration of L¹ and L².