

SUPPLEMENTARY INFORMATION

Biosourced graphitic nanoparticles loaded hyperbranched polyurethane composites – Application as multifunctional high-performance coatings

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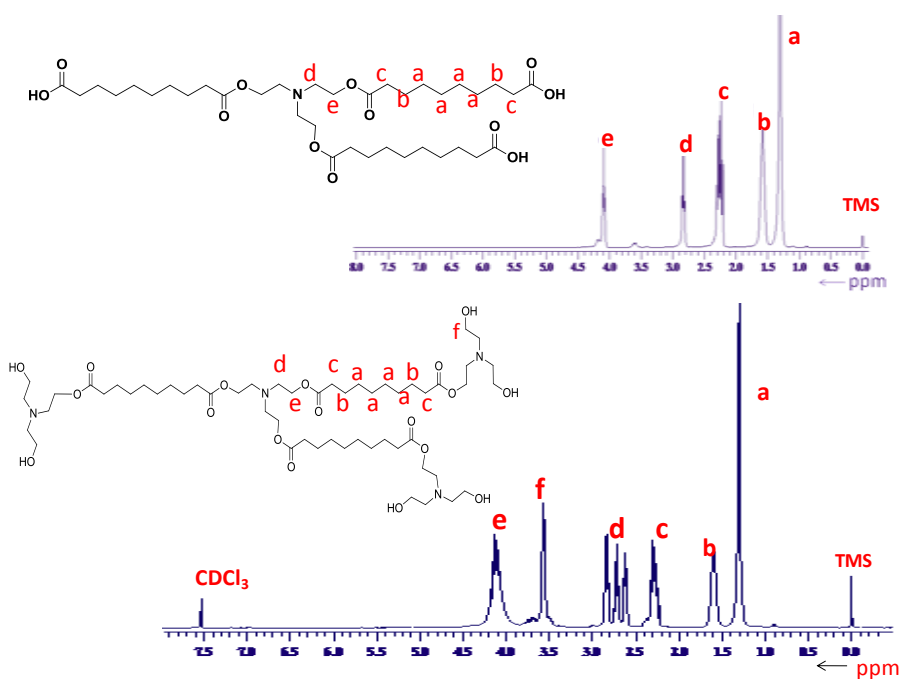


Fig. S1 ¹H NMR of SA-COOH and SA-OH

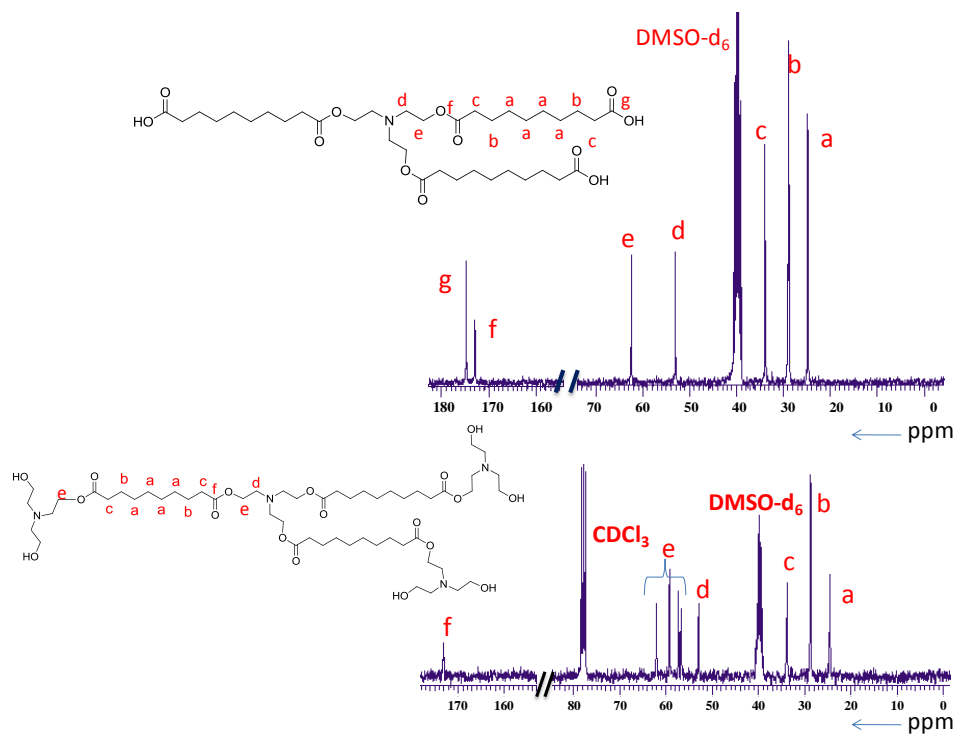


Fig.S2 ^{13}C NMR spectrum of SA-COOH and SA-OH

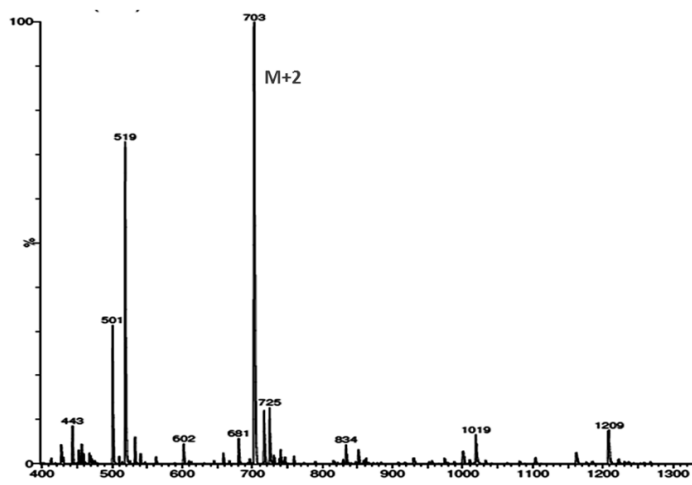


Fig. S3 ESI-MASS of SA-COOH

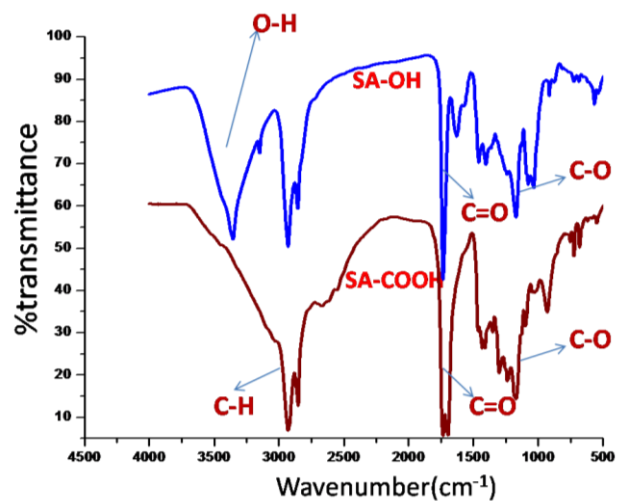


Fig. S4 FTIR overlay of SA-COOH and SA-OH

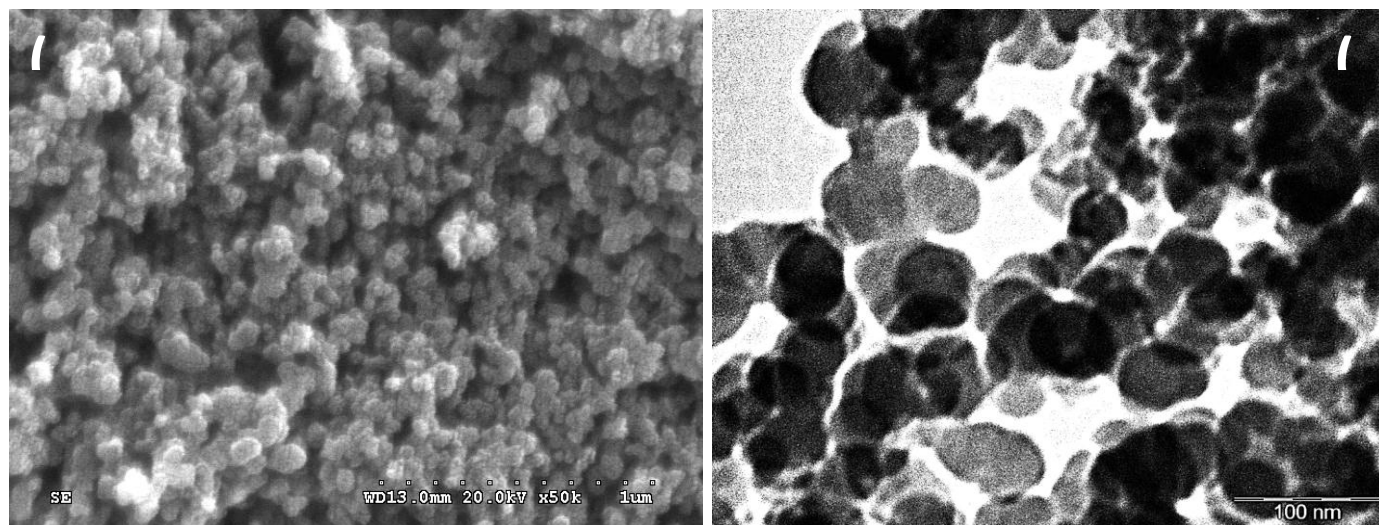


Fig. S5 (a) FESEM and (b) TEM images of GP-COOH

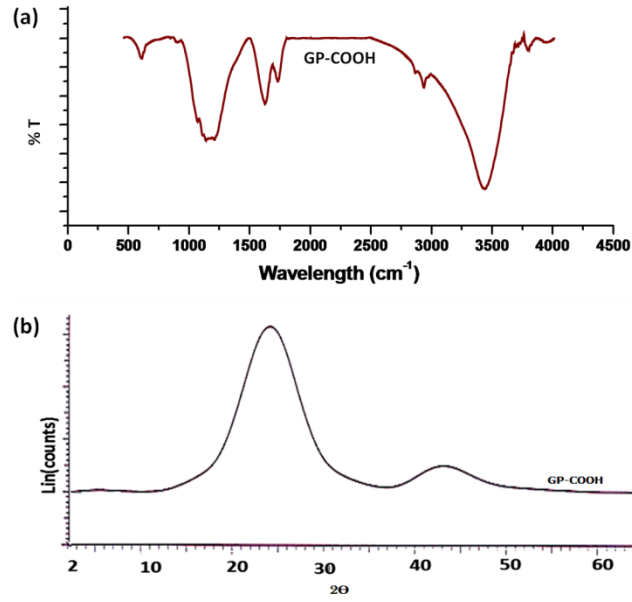


Fig. S6 (a) FTIR and (b) XRD profile of GP-COOH

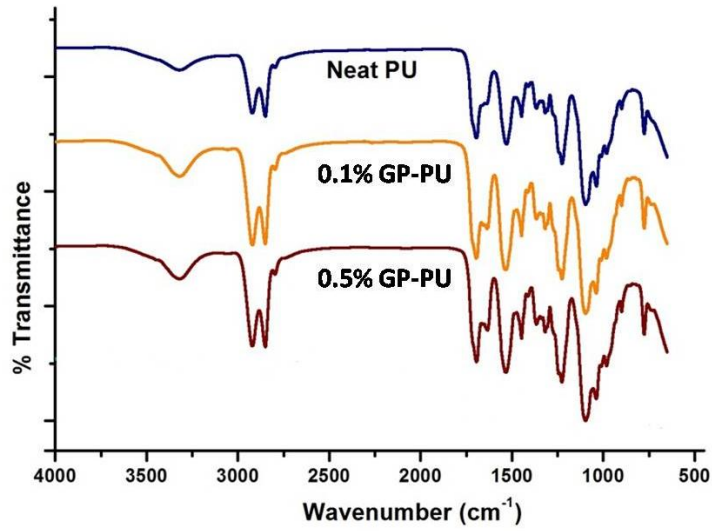


Fig. S7 FTIR overlay of different GP-PU hybrid composite films

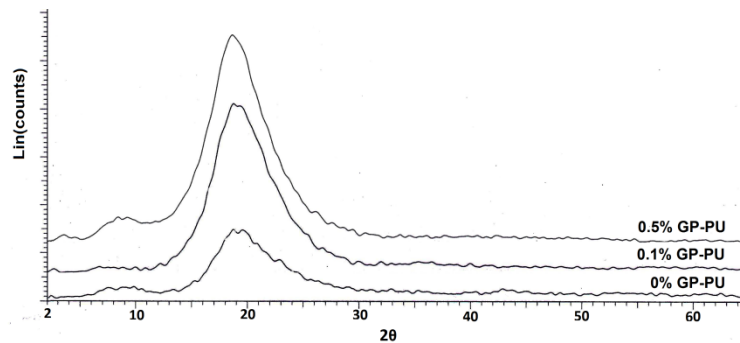


Fig. S8 XRD overlay of GP-PU hybrid films

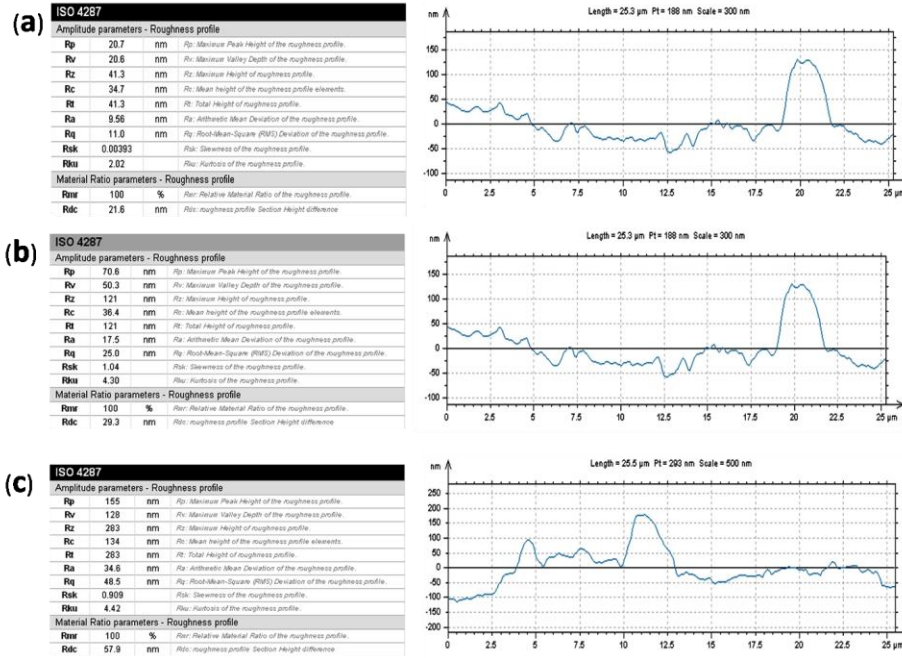


Fig. S9 Roughness profile of (a) 0% GP-PU, (b) 0.1% GP-PU and (c) 0.5% GP-PU

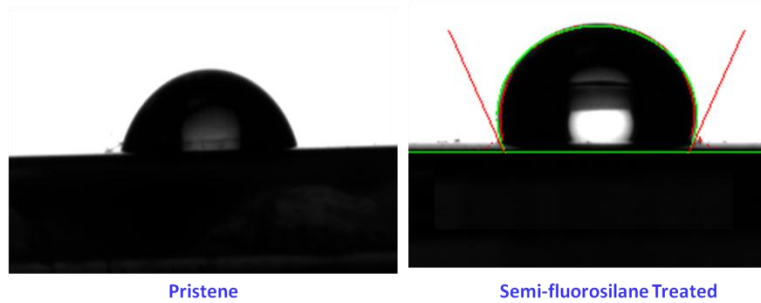


Fig. S10 Contact angle image of 0.5% GP-PU hybrid film

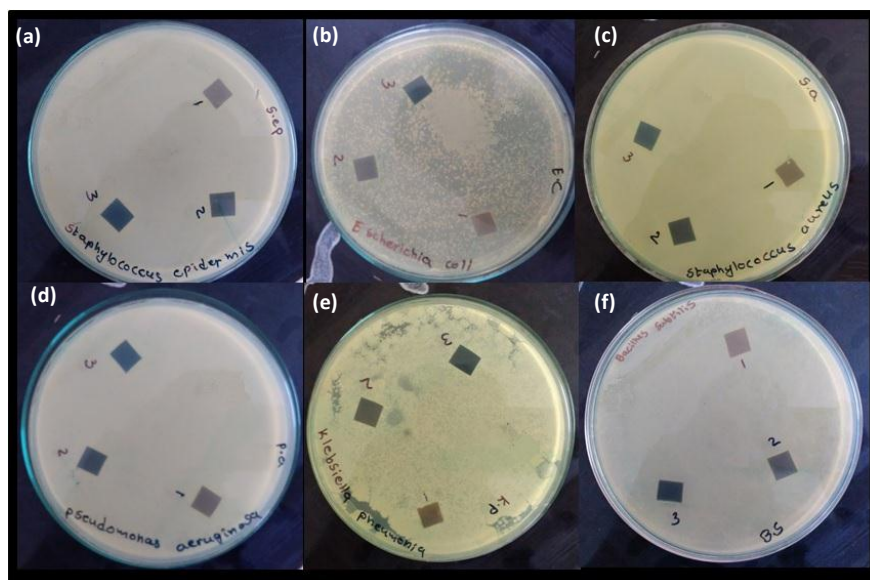


Fig. S11 Antibacterial activity of various GP-PU tested with the following bacterial strains; (a)*Staphylococcus epidermidis*, (b)*Escherichia coli*, (c)*Staphylococcus aureus*, (d)*Pseudomonas aeruginosa*, (e)*Klebsiella pneumoniae* and (f)*Bacillus subtilis*. (Within the figure 1, 2 and 3 represent 0%, 0.1% and 0.5% GP-PU respectively)

Calculation of the degree of branching:

The degree of branching is calculated by using the equation below

$$DB_{Frechet} = \frac{D + T}{D + L + T}$$

$$DB_{Frey} = \frac{2D}{2D + L}$$

Reference

A. R. Fornof, T. E. Glass and T. E. Long, *Macromolecular Chemistry and Physics.*, 2006, **207**, 1197-1206.

Calculation of the corrosion rate:

The experimental setup consisted of a conventional three-electrode cell containing the working electrode, a saturated calomel electrode (SCE) and a counter electrode. Using the potentiostat of an electro-chemical workstation, the tests were performed by monitoring the current density as a function of the free open-circuit potential. After a stabilizing the solution for few minutes at room temperature, the potentiodynamic polarization was carried out at a scan rate of 2 mV/s.

The corrosion current (i_{corr}) was estimated by linear fit and Tafel extrapolation to the cathodic and anodic part of the polarization curves. The corrosion rate depends on the kinetics of both anodic (oxidation) and cathodic (reduction) reactions. According to Faraday's law, there is a linear relationship between the metal dissolution rate or corrosion rate, R_c , and the corrosion current i_{corr}

$$R_c = \frac{tM}{nF\rho} i_{\text{corr}}$$

where t is the exposure time, M is the atomic weight of the metal, ρ is the density, n is the charge number which indicates the number of electrons exchanged in the dissolution reaction and F is the Faraday constant, (96.485 C/mol). The ratio M/n is also sometime referred to as equivalent weight.

In the present case we have used galvanised steel panels coated with GP-PU.

Reference

Witte, F., Fischer, J., Nellesen, J., Crostack, H. A., Kaese, V., Pisch, A., and Windhagen, H. (2006). In vitro and in vivo corrosion measurements of magnesium alloys. *Biomaterials*, 27(7), 1013-1018.