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Electronic Supporting Information for

Selective synthesis of functionalized pyrroles from

3-aza-1,5-enynes

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1. Copies of NMR Spectra for New Compounds





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S5



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$$- 162.82$$

$$- 145.64$$

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$$- 21.77$$

$$- 51.57$$

$$- 19.06$$



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S9



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S26



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S93




























































2. Crystal Structures of 3a and 5m

Molecular structure of **3a** with thermal ellipsoids at the 50% probability level.



CCDC 962702 (**3a**) contains the supplementary crystallographic data for this paper. These datacan be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif.

Molecular structure of **5m** with thermal ellipsoids at the 50% probability level.



CCDC 978572 (**5m**) contains the supplementary crystallographic data for this paper. These datacan be obtained free of charge from The Cambridge Crystallographic Data Centre via <u>www.ccdc.cam.ac.uk/data_request/cif</u>.

3. Procedures for Calculating the HPLC Yields

Take the procedure for calculating the HPLC yields of product **3a** (Table 3 in the main text) for example:

HPLC conditions: instrument: Agilent 1200 series; column: Extend-C18, 5 μ m, 4.6 × 150 mm; eluent: H₂O/MeCN = 45/55, 1 mL/min; UV detector wavelength: 254 nm; inject volume: 5 μ L; constant temperature: 25 °C.

Step 1: calculating the standard curve equation (the relationship between the ratio of HPLC peak area with the weight of the desired product). We prepared the pure product **3a** beforehand. To four reaction tubes, marked as No. 1, No. 2, No. 3, and No. 4, we added 10.0, 25.0, 40.0, and 55.0 mg **3a**, respectively. Then to each of these four tubes, we added 30.0 mg naphthalene as internal standard, and each sample was diluted with 10.5 mL acetonitrile. After homogeneous mixing, the samples were then analyzed with HPLC. The ratio of A_{3a}/A_{nap} could be calculated based on the peak areas of **3a** and naphthalene (A_{3a} : peak area of **3a**; A_{nap} : peak area of naphthalene). The results are described as the below table (m_{3a}: the weight of **3a**; m_{nap}: the weight of naphthalene):

Experiment No.	$m_{3a} (mg)$	$m_{nap} (mg)$	$m_{\mathbf{3a}}/m_{nap}$	A_{3a}/A_{nap}
No. 1	10.0	30.0	0.3333	0.3547
No. 2	25.0	30.0	0.8333	0.8632
No. 3	40.0	30.0	1.3333	1.3553

No. 4 55.0 30.0 1.8333 1.80

According the equation: $m_{3a}/m_{nap} = f \times (A_{3a}/A_{nap})$, and base on the results in the above table, we could calculate that f = 0.9780. Therefore, the standard curve equation was: $m_{3a}/m_{nap} = 0.9780 \times (A_{3a}/A_{nap})$. Thus, $m_{3a} = 0.9780 \times m_{nap} \times (A_{3a}/A_{nap})$.

Step 2: carrying out the experiments in Table 3 of the main text. After termination of the reactions, the samples were cooled down to room temperature. To each reaction tube, 30.0 mg naphthalene (as internal standard) and 10.0 mL acetonitrile were added. After homogeneous mixing, the samples were then analyzed with HPLC. The ratio value of A_{3a}/A_{nap} of each sample could be calculated based on the peak areas of **3a** and naphthalene.

Step 3: calculating the HPLC yields. After putting the numerical value of A_{3a}/A_{nap} in to the equation $m_{3a} = 0.9780 \times m_{nap} \times (A_{3a}/A_{nap})$, we could get the calculated weight of **3a**. Thus, the yield was calculated as following: **3a**% = (m_{3a}/m_{theo}) × 100% = 0.9780 × (m_{nap}/m_{theo}) × (A_{3a}/A_{nap}) × 100% (*note*: m_{theo} was the theoretical weight of **3a** in 100% yield).

One example: take the experiment of entry 2, Table 3 in the main text for example: the ration of A_{3a}/A_{nap} was 0.8463, $m_{nap} = 30.0$, $m_{theo} = 51.4$ (in 0.1 mmol scale). Thus, the HPLC yield of **3a** was following: $3a\% = 0.9780 \times (m_{nap}/m_{theo}) \times (A_{3a}/A_{nap}) \times 100\% = 0.9780 \times (30.0/51.4) \times 0.8463 \times 100\% = 48.3083\% \approx 48\%$.