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Supplementary Information

Metal-free yne-addition/aryl migration/decarboxylation cascade reaction of

alkynoates with C_{sp3}-H centers

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General Information:

All chemicals were obtained from commercial sources and were used as received unless otherwise noted. Alkynoate were prepared according to original or modified literature procedures. All reactions were carried out in flame-dried reaction vessels with Teflon screw caps under a nitrogen atmosphere by using standard Schlenk techniques. Anhydrous solvents were purified and dried following standard procedures. TLC analysis was performed on pre-coated, glass-backed silica gel plates and visualized with UV light. Column chromatography was performed on silica gel(200-300 mesh) using ethyl acetate (EA)/petroleum ether (PE).¹H-NMR and ¹³C-NMR Spectra were obtained on a Bruker 300 MHz, 400MHz or 500 MHz NMR spectrometer in the deuterated solvents indicated Chemical shifts are reported in ppm from tetramethylsilane with the solvent resonance as the internal standard. The following abbreviations were used to designate chemical shift multiplicities: s= singlet, d= doublet,t= triplet, q= quartet, h= heptet, m= multiplet,. All first-order splitting patterns were assigned on the basis of the appearance of the multiplet. Splitting patterns that could not be easily interpreted are designated as multiplet (m) or broad (br). Melting points were measured on Beijing Tech X-4 apparatus without correction. IR spectra were recorded on a Nicolet 6700 FT-IR spectrometer. HRMS Masss pectra were obtained using electrospray ionization (ESI) or atmospheric pressure chemical ionization (APCI) mass spectrometer.

2. Mechanistic studies

2.1 Radical scavenger effect studies

To gain insight into the pathway of the present reaction, we conducted some control experiments under the standard reaction conditions. Radical scavengers, TEMPO, were employed in the standard reaction, and no desired product was detected. This result suggested that a single electron transfer process (SET) was involved in the reaction.



^aReaction conditions: 1a (0.2 mmol), DTBP (0.4 mmol), TEMPO(0.5 mmol),4 mL of cyclohexane as solvent, 125 ℃, 12h.

2.2 Deuterium-labeling experiment



Mixture of THF-D₈ (1 mL), 1a (0.2mmol) and DTBP (0.4 mmol) was sealed in a Teflon septum screw-capped tube underN₂. The mixture was stirred in an oil bath at 125 °C for 12 h. The corresponding reaction mixture was purified by flash column chromatography on a silica gel to give the desired product in 46% yield. The ¹NMR (in CD₃OD) was described in details as follows.



The addition of t-BuOD (2 eq) to the reaction mixture of **1a** with **8a** on the standard condition did not lead to any deuterium-containing products thus practically ruling out the possibility that the terminal hydrogen abstraction step of the tandem reaction might be involved by t-BuOH formed from homo-cleavage of DTBP.



2.3 Kinetic isotope effect studies

THF (2 mL, 25 mmol), **1a** (0.2 mmol) and DTBP (0.4 mmol) was sealed in a Teflon septum screw-capped tube under N₂. Same procedures was operated with THF-D₁₂. The mixture was stirred in an oil bath at 125 °C for 1.5 h. After completion of the reactions, the mixtures were detected by GC analysis with dibenzofuran as the internal standard. As a result, the desired coupling products were obtained in 82% and 39% yields, respectively, in which $k_H/k_D = 2.1$.



3. Copies of ¹H NMR and ¹³C NMR



kdl-183-H



3ab



3ac

ppm

190 180 170 160 150 140 130 120 110

kdl-177-H



kdl-217-thf-hun

3h

kdl-197-H

5aa

kdl-181-H

I.

ppm

5ab

S15

5d

kdl-212-DMA-H

kd1-245-H

S18

5g

kdl-192-H

kd1-225

S20

E.

7ab

190 180 170 160 150 140 130 120 110 100 90 80 70 60 50 40 30 20 ppm kd1-236-H

7ac

7ae

kd1-237-H

kd1-215-H

Constant of the second se

S26

kd1-240-h

9ac

9bb

10a

kdl-187-h

12b

130 120

190 180 170 160 150 140

ppm

kd1-214-DMA-H

13b

14b

kdl-214-cyh-guanhuan

190 180 170 160 150 140 130 120 110 100 90 80 70 60 50 40 30 20

14a

ppm