

## Supporting Information

Effect of CO and CO<sub>2</sub> on the desulfurization of H<sub>2</sub>S using ZnO sorbent:

### A density functional theory study

Lixia Ling <sup>a</sup>, Zhongbei Zhao <sup>a</sup>, Baojun Wang <sup>\*, b</sup>, Maohong Fan <sup>\*, c</sup>, Riguang Zhang <sup>b</sup>

a) *College of Chemistry and Chemical Engineering, Taiyuan University of Technology, Taiyuan 030024, Shanxi, P.R. China E-mail: [linglixia@tyut.edu.cn](mailto:linglixia@tyut.edu.cn)*

b) *Key Laboratory of Coal Science and Technology (Taiyuan University of Technology), Ministry of Education and Shanxi Province, Taiyuan 030024, Shanxi, People' Republic of China E-mail: [wangbaojun@tyut.edu.cn](mailto:wangbaojun@tyut.edu.cn)*

c) *Department of Chemical and Petroleum Engineering, the University of Wyoming, Laramie, WY 82071, USA E-mail: [mfan@uwyo.edu](mailto:mfan@uwyo.edu)*

Page S1: **Table S1**: The imaginary frequency of every transition state.

Page S1: **Table S2**: Electronic total energy, enthalpy correction ( $H'$ ) and Gibbs free energy correction ( $G'$ ) of every species for R1 and R2 at different temperatures (K).

Page S1: **Table S3**: Electronic total energy, enthalpy correction ( $H'$ ) and entropy (S) for the rate-determined step of forming COS at different temperatures (K).

**Table S1** The imaginary frequency of every transition state.

H <sub>2</sub> S+CO		H <sub>2</sub> S+CO <sub>2</sub>	
Transition state	Imaginary frequency(cm <sup>-1</sup> )	Transition state	Imaginary frequency(cm <sup>-1</sup> )
TS1	-138.33	TS1*	-195.13
TS2	-936.01	TS2*	-235.71
TS3	-112.03	TS3*	-1095.74
TS4	-185.32	TS4*	-134.57
TS5	-298.51	TS5*	-345.24
		TS6*	-1685.37
		TS7*	-71.32
		TS8*	-1671.95

**Table S2** Electronic total energy, enthalpy correction ( $H'$ ) and Gibbs free energy correction ( $G'$ ) of every species for R1 and R2 at different temperatures (K).

	$E_{\text{elec}}(\text{kJ}\cdot\text{mol}^{-1})$	$H'$ (including ZPVE) ( $\text{kJ}\cdot\text{mol}^{-1}$ )				$G'$ (including ZPVE) ( $\text{kJ}\cdot\text{mol}^{-1}$ )			
		298.15	500	650	1000	298.15	500	650	1000
H <sub>2</sub> S	-1048628.5	51.6	60.5	67.5	85.3	-15.8	-64.0	-102.2	-197.6
CO	-297524.7	21.4	27.3	31.9	43.1	-37.6	-79.2	-111.7	-191.6
COS	-1343129.4	31.0	39.9	47.1	64.9	-29.1	-72.3	-106.9	-193.8
H <sub>2</sub>	-3073.2	34.8	40.7	45.0	55.3	-5.8	-35.0	-58.3	-116.3
CO <sub>2</sub>	-495223.1	40.2	49.5	57.2	76.8	-1.0	-31.4	-56.7	-122.4
H <sub>2</sub> O	-200695.2	64.7	71.6	77.0	90.6	6.6	-34.6	-67.2	-147.9

**Table S3** Electronic total energy, enthalpy correction ( $H'$ ) and entropy ( $S$ ) for the rate-determined step of forming COS at different temperatures (K).

	$E_{\text{elec}}(\text{kJ}\cdot\text{mol}^{-1})$	$H'$ (including ZPVE) ( $\text{kJ}\cdot\text{mol}^{-1}$ )			$S$ ( $\text{J}\cdot\text{mol}^{-1}\cdot\text{K}^{-1}$ )		
		298.15	500	1000	298.15	500	1000
IM4	-39478285.8	944.1	1384.5	2574.3	2358.6	3480.9	5124.8
TS4	-39478198.1	939.3	1378.7	2564.9	2396.7	3516.2	5155.2
IM6*	-39675956.4	964.0	1406.7	2606.0	2369.4	3497.3	5154.1
TS6*	-39675814.0	949.9	1391.6	2588.4	2363.7	3488.9	5142.3