

## Electronic Supplementary Information for

# DFT study to design Super- and Hyperacids with 1-(cyclopenta-2,4-dien-1-yl)-4-nitrobenzene and 3-(cyclopenta-2,4-dien-1-ylmethylene)-6-methylenecyclohexa-1,4-diene molecules

Mrinal Kanti Si<sup>1,2</sup>, Bishwajit ganguly\*<sup>1,2</sup>

<sup>1</sup>Computation and Simulation Unit (Analytical Discipline and Centralized Instrument Facility), CSIR-Central Salt & Marine Chemicals Research Institute

<sup>2</sup>Academy of Scientific and Innovative Research, CSIR-CSMCRI, Bhavnagar, Gujarat, India-364 002.

\*Corresponding Author. Fax: (+91)-278-2567562, E-mail: [ganguly@cmcri.org](mailto:ganguly@cmcri.org)

## Theoretical Methods

The gas phase acidity is measured by the calculating the enthalpy change ( $\Delta H_{\text{acid}}$ ) for the proton dissociation reaction given below



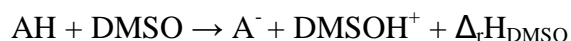
Where  $\Delta H_{\text{acid}}$  can be calculated by the equation 2

$$\Delta H_{\text{acid}} = \Delta E_{\text{acid}} + \Delta(pV) \quad (2)$$

$\Delta E_{\text{acid}}$  is the change of total molecular energy of the species involved in the equation 1 and the electronic energy involves the electronic energy, the repulsion of nuclei, the zero point vibrational energy (ZPVE) and the finite room temperature correction at 298.15 K.

The DFT-B3LYP method gives the satisfactory result for calculating the acidity which is comparable to experimental result <sup>1, 2</sup>. The series of papers published that reported the DFT methods are useful for calculating the electron affinity of molecules <sup>3-7</sup>. Hence we have optimized the compounds in B3LYP/6-31G(d) and further the single point energy calculated in

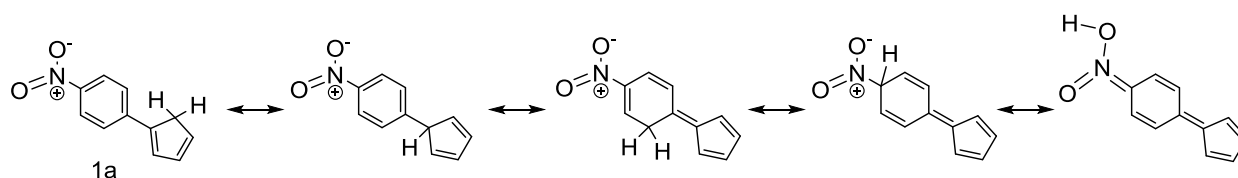
B3LYP/6-311+G(2d,P) method. We have calculated the acidities of the compounds in solvent medium. The PCM solvent model is important for determining the acidity<sup>8,9</sup>. We have calculated the acidity of compounds **4a** in DMSO solvent which based on previously reported the proton transfer reaction between solvated acids and dimethyl sulfoxide molecule<sup>10</sup>.



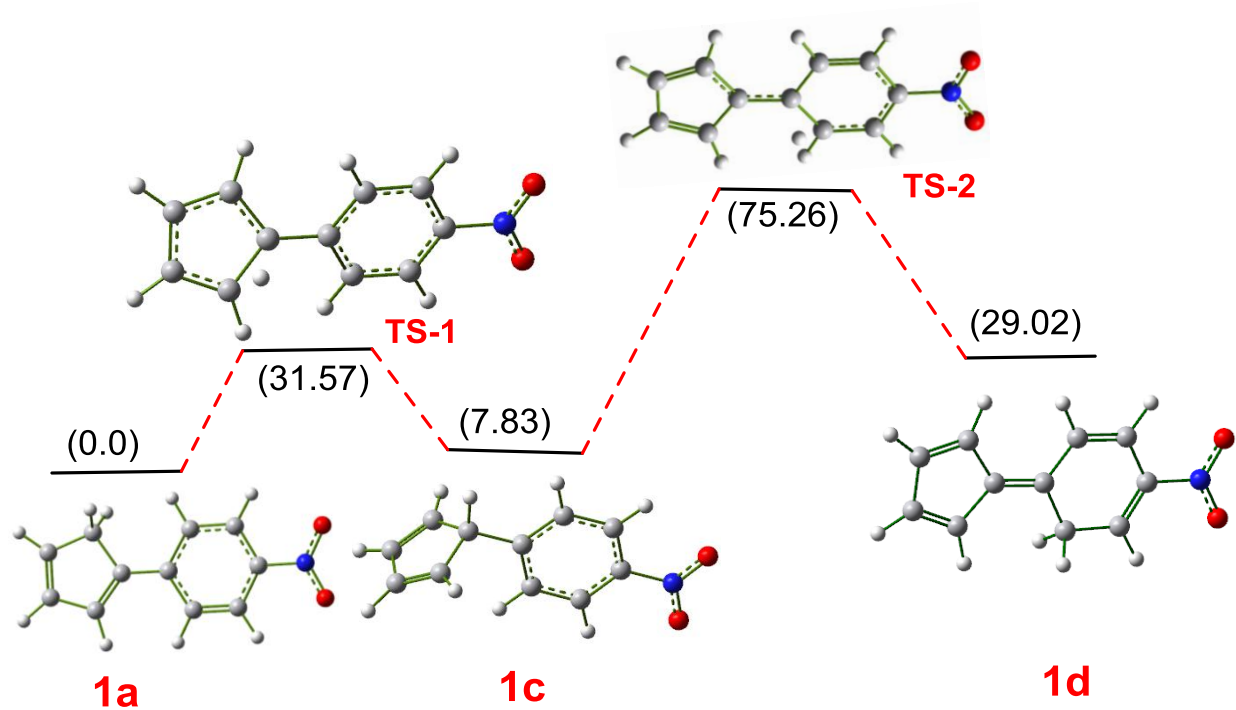
Where  $\text{A}^-$  is the conjugate base and  $\Delta_r H_{\text{DMSO}}$  is enthalpy change in above equation. All calculations are performed at IPCM/B3LYP/6-311+G(2d,p)//B3LYP/6-31G(d) level of theory by using the thermal correction at B3LYP/6-31G(d) level of theory. We have calculated the pKa values for compound **4a** by using equation 3 which is previous reported to calculate the pKa for neutral organic acid<sup>10</sup>.

$$\text{pKa}(\text{exp}) = 0.661. \Delta_r H_{\text{DMSO}} - 7.7 \quad (3)$$

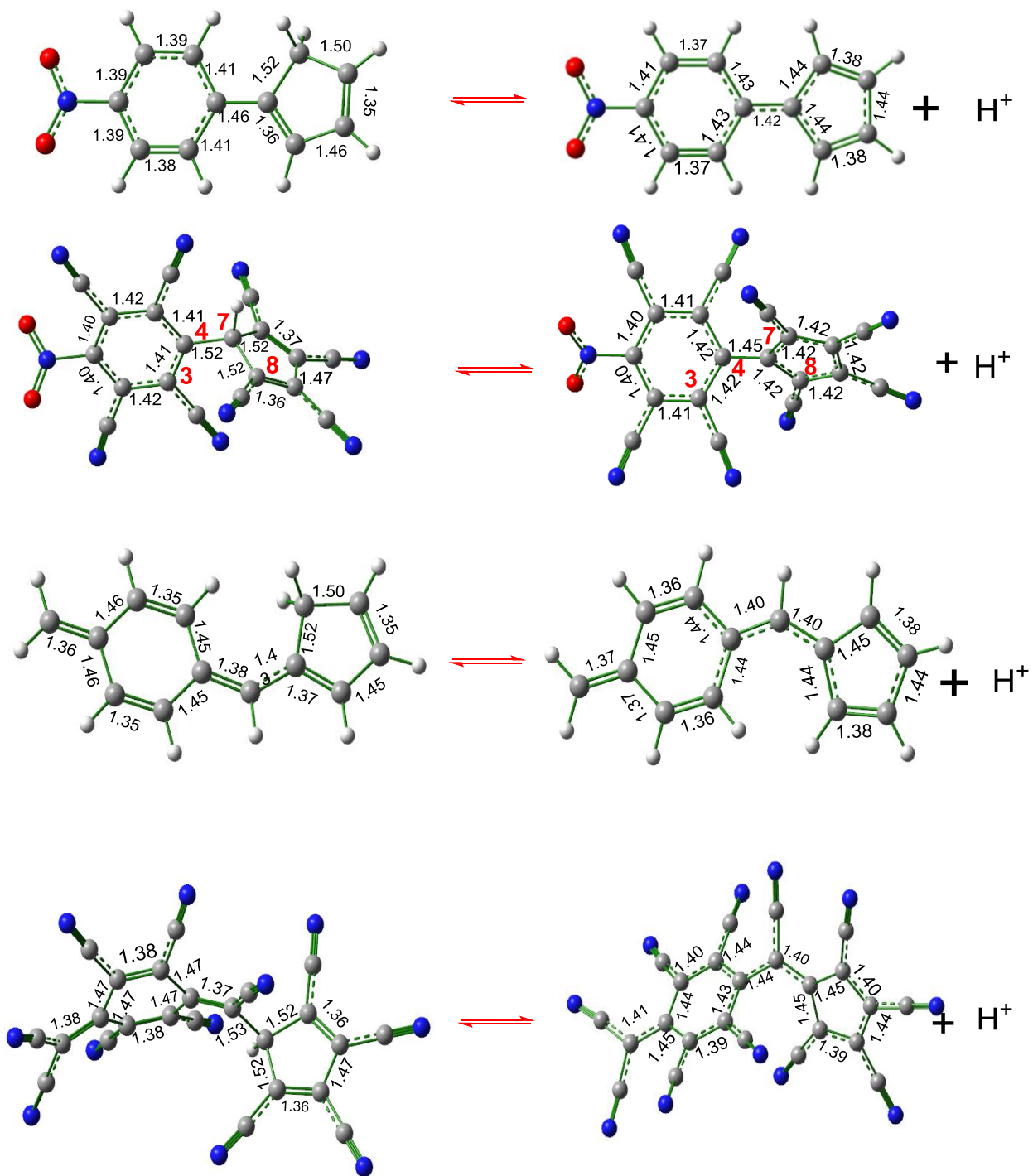
All calculations were performed with the GAUSSIAN09 program package<sup>11</sup>.



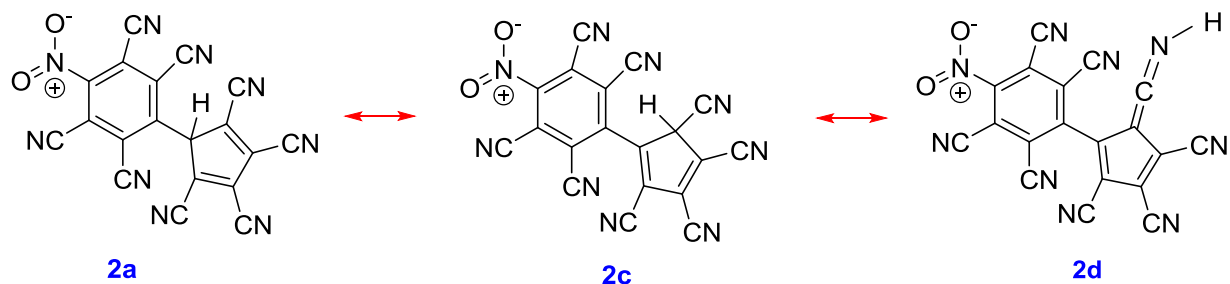
**Scheme S1** The process of tautomerization of hydrogen from  $\text{sp}^3$  carbon atom of **1a** on the nitro group, forming  $-\text{N}(\text{O})\text{OH}$ .



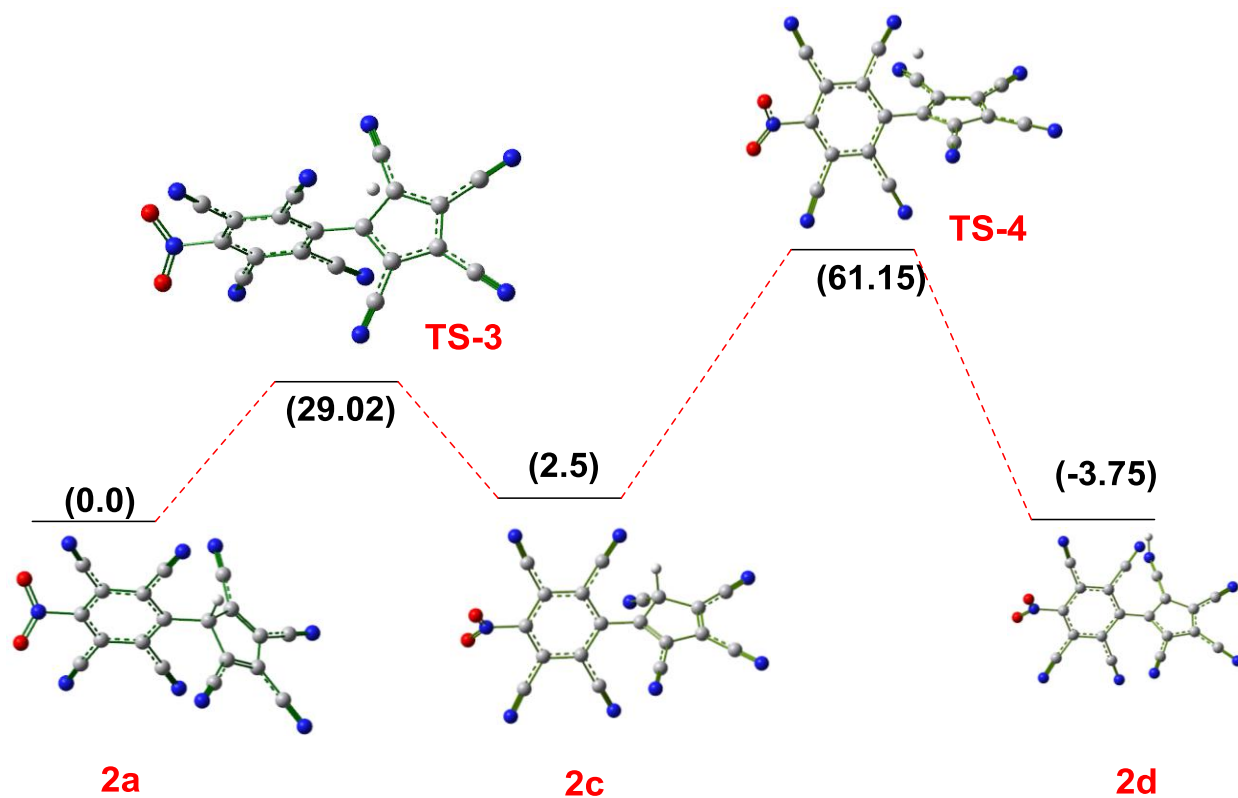
**Figure S1** The potential energy diagram for the transfer of proton from  $sp^3$  carbon center of **1a** to  $sp^2$  carbon center of **1c** and from **1c** to **1d** at B3LYP/6-311+G(2d, p)//B3LYP/6-31G(d) method. (The energies are given in kcal/mol)



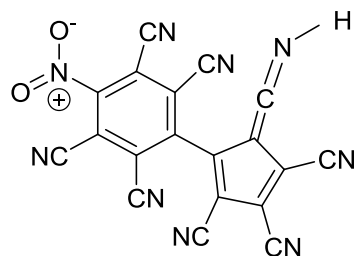
**Figure S2** The geometrical parameters for **1a**, **2a**, **3a** and **4a** molecules and their conjugate bases. (The red color numbers indicate the dihedral angle (-C<sub>3</sub>-C<sub>4</sub>-C<sub>7</sub>-C<sub>8</sub>)



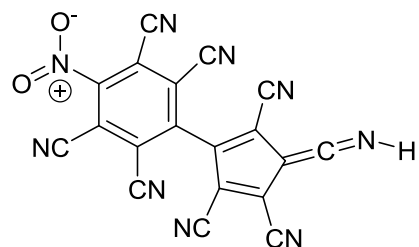
**Scheme S2** The process of tautomerization of hydrogen from  $sp^3$  carbon atom of **2a**, forming ketene imine (**2d**).



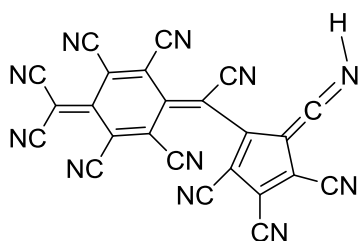
**Figure S3** The potential energy diagram for the transfer of proton from **2a** to **2d** at B3LYP/6-311+G(2d, p)//B3LYP/6-31G(d) method. (The energies are given in kcal/mol).



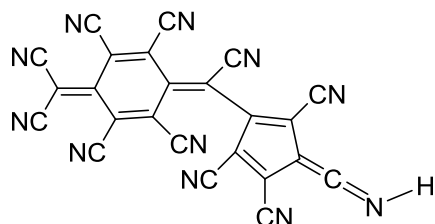
**2d**  
**(-3.7)**



**2e**  
**(-6.3)**



**4d**  
**(-3.0)**

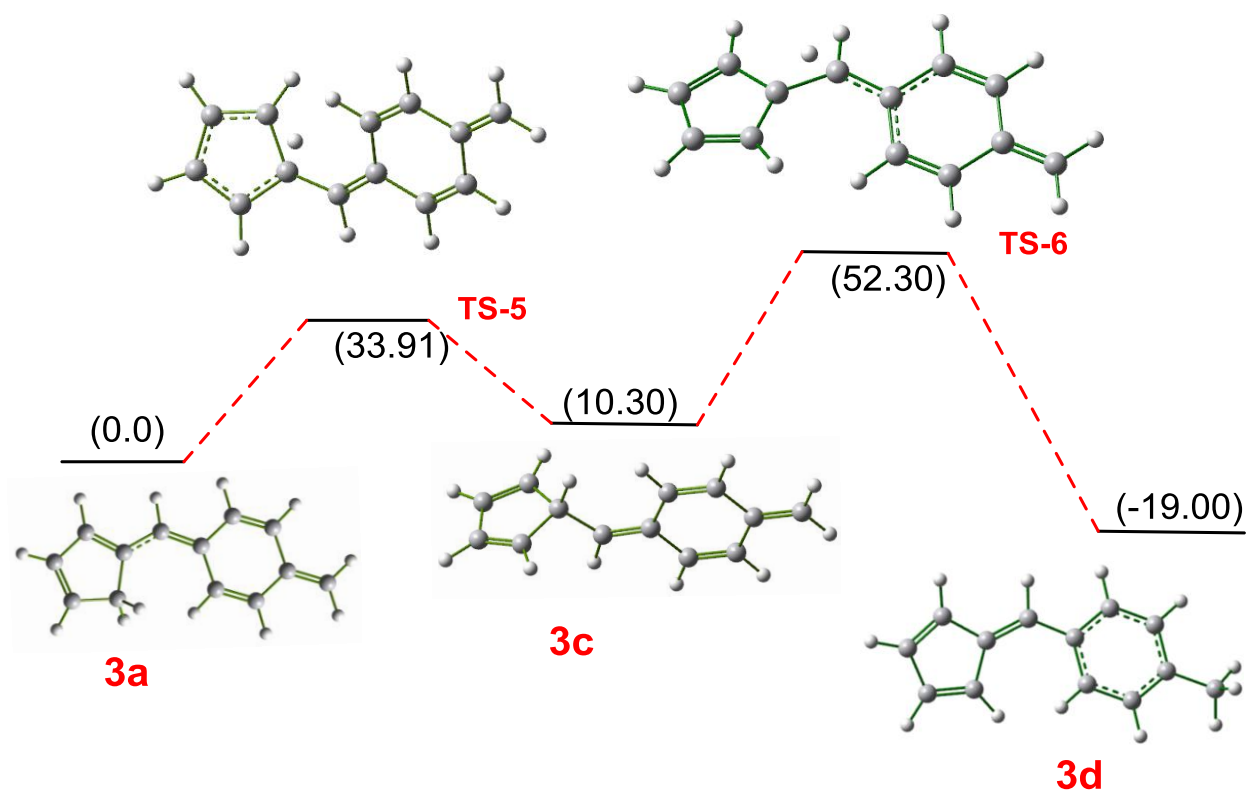


**4e**  
**(-7.9)**

**Figure S4** Structure of ketene imine tautomers of octacyano substituted 1-(cyclopenta-2,4-dien-1-yl)-4-nitrobenzene and cyano group substituted 3-(cyclopenta-2,4-dien-1-ylmethylene)-6-methylenecyclohexa-1,4-diene and their relative energies with respect to tautomers (**2a** and **4a** respectively) obtained with B3LYP/6-311+G(2d,p)//B3LYP/6-31G(d) in gas phase (in kcal/mol)

**Table S1** The  $\Delta H_{\text{acid}}$  values calculated at B3LYP/6-311+G(2d,p)//B3LYP/6-31G(d) level of theory at gas phase for the compounds **2d**, **2e**, **4d**, and **4e**.

Molecules	Energy	$H_{\text{corrected}}$	$\Delta H_{\text{acid}}$
<b>2d</b>	-1367.89780	0.18942	256.5
<b>2e</b>	-1367.90185	0.18950	258.9
<b>2'</b>	-1367.48015	0.17809	-
<b>4d</b>	-1517.49048	0.22001	238.8
<b>4e</b>	-1517.49048	0.21020	246.63
<b>4'</b>	1517.09638	0.2088	-



**Figure S5** The potential energy diagram for the transfer of proton from  $sp^3$  carbon center of **3a** to  $sp^2$  carbon center of **3c** and **3c** to **3d** with B3LYP/6-311+G(2d, p)//B3LYP/6-31G(d) level of theory. (The energies are given in kcal/mol).

## References

1. A. D. Becke, Phys. Rev., 1988, **A38**, 3098-3100.
2. C. Lee, W. Yang, R. G. Parr, Phys. Rev., 1988, **37**, 785-789.
3. R. A. King, J. M. Galbraith, H. F. Schaefer III, J. Chem. Phys., 1996, **100**, 6061-6068.
4. J. M. Galbraith, H. F. J. Schaefer III, Chem. Phys., 1996, **105**, 862-864.
5. Q. -Sh. Li, W. -G. Xu, Y. Xie, H. F. J. Schaefer III, J. Phys. Chem. A, 1999, **103**, 7496-7505.
6. J. C. Rienstra-Kiracofe, G. S. Tschumper, H. F. Schaefer III, N. Sreela, G. B. Ellison, Chem. Rev., 2002, **102**, 231-282.
7. Y. Xie, H. F. Schaefer III, F. A. Cotton, Chem. Commun., 2003, 102-103.
8. S. Miertuš, E. Scrocco, J. Tomasi, Chem phys., 1981, **55**, 117-129.
9. S. Miertuš, J. Tomasi, Chem phys., 1982, **65**, 239-245.
10. R. Vianello, Z. B. Maksić, Eur. J. Org. Chem., 2004, 5003-5010.
11. M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, G. Scalmani, V. Barone, B. Mennucci, G. A. Petersson, et al. Gaussian 09, Revision B01, Gaussian, Inc, Wallingford CT 2010.

**Table S2.** B3LYP/6-31G(d) optimized Cartesian coordinate of compounds **1a**, **2a**, **3a**, and **4a** (The electronic energies (E) are given in atomic unit).

<b>1a</b>				<b>1<sup>+</sup></b>			
<b>E =</b>	<b>-629.87011</b>			<b>E =</b>	<b>-629.40585</b>		
C	0.69107800	0.05654000	-0.00000700	C	-0.76203500	0.00000000	-0.00001700
C	-0.01771600	-1.16414500	-0.00009400	C	0.00541100	1.21146500	-0.00005000
C	-1.40597400	-1.19894700	-0.00008000	C	1.38131600	1.21646400	-0.00006100



C	-2.11144100	0.00310400	0.00002200	C	2.09560000	0.00000000	-0.00003000
C	-1.44944800	1.23291200	0.00011900	C	1.38131600	-1.21646400	0.00001700
C	-0.06372400	1.25151600	0.00010900	C	0.00541100	-1.21146500	0.00003100
H	0.52736200	-2.10250200	-0.00018300	H	-0.52127400	2.16076200	-0.00008800
H	-1.94973400	-2.13564300	-0.00015100	H	1.94061700	2.14514400	-0.00010200
H	-2.02838100	2.14826600	0.00021000	H	1.94061700	-2.14514400	0.00005300
H	0.44530300	2.20975000	0.00021700	H	-0.52127500	-2.16076200	0.00010500
C	2.15114400	0.07234000	-0.00003200	C	-2.18402800	0.00000000	0.00003400
C	2.98463600	1.15161200	-0.00024300	C	-3.05032600	-1.15370500	0.00009200
C	3.01674700	-1.17324400	0.00016300	C	-3.05032600	1.15370400	0.00029200
H	2.68299100	2.19345200	-0.00045200	H	-2.72527100	-2.18896700	0.00003900
C	4.37559400	0.71601700	-0.00015900	C	-4.36320200	-0.71886500	-0.00022700
H	2.82739800	-1.80706500	-0.87878700	H	-2.72527000	2.18896700	0.00053200
C	4.41988700	-0.63388300	0.00007800	C	-4.36320200	0.71886600	0.00012700
H	5.30580800	-1.25797200	0.00018500	H	-5.24711700	-1.35142000	-0.00049100
N	-3.57449900	-0.02379600	0.00003700	H	-5.24711700	1.35142100	0.00018800
O	-4.16895200	1.05620200	0.00011000	N	3.51473800	0.00000000	-0.00004100
O	-4.12941500	-1.12465800	-0.00005700	O	4.12045800	-1.09542400	-0.00004400
H	5.22560600	1.38981200	-0.00028300	O	4.12045800	1.09542300	-0.00010600
H	2.82737900	-1.80680800	0.87929300			<b>2'</b>	
		<b>2a</b>		<b>E = -1367.48015</b>			
<b>E = -1367.89176</b>				C	-0.13386000	-0.00033900	-0.00029500
C	-0.18716600	0.43327400	0.02479200	C	-0.87066800	1.13376800	-0.42979000
C	-1.24285000	1.36429700	0.07208200	C	-2.28447500	1.15156900	-0.40595500
C	-2.59625700	0.94499600	0.03634100	C	-2.97345600	-0.00004900	-0.00029400
C	-2.87446300	-0.42421200	-0.01726900	C	-2.28467900	-1.15180600	0.40532500
C	-1.85078100	-1.37824900	-0.05089800	C	-0.87087900	-1.13426700	0.42921400
C	-0.50461400	-0.93703100	-0.04441700	C	1.32535200	-0.00027700	-0.00005700
C	1.25451500	0.92437600	0.04900700	C	2.16499700	-0.92864500	-0.67444200
C	2.11739800	0.59363300	-1.15733300	C	2.16458700	0.92830800	0.67455300
C	2.12281100	0.46255500	1.20728700	C	3.51254400	-0.56890900	-0.42434000
C	3.33451900	0.13870500	-0.73583200	C	3.51229600	0.56899700	0.42471100
C	3.33780300	0.05705500	0.73265200	N	-4.44433700	0.00019200	-0.00010700
N	-4.28448900	-0.87525000	-0.03866500	O	-5.00170700	-1.02479800	-0.38130600
O	-4.56516000	-1.81302100	0.69509200	O	-5.00125700	1.02537000	0.38123300
O	-5.03692700	-0.26738700	-0.78773300	C	-0.19080300	-2.26664800	0.98454300
C	0.53740500	-1.91697600	-0.10630200	C	-2.97059100	-2.31759800	0.87999100
C	-2.11804700	-2.78191000	-0.13164300	C	-2.97018800	2.31768700	-0.88011400
C	-3.62850800	1.93434200	0.09524800	C	-0.19034100	2.26630800	-0.98450700
C	-0.95637700	2.76591000	0.15454600	C	1.72218800	1.96057700	1.53958600
C	1.68909600	0.80925500	-2.48792100	C	1.72312300	-1.96086500	-1.53979800
C	4.44200800	-0.20199800	-1.55957400	C	4.65833400	-1.21846800	-0.94835200
C	4.44877200	-0.37284700	1.50860200	C	4.65777900	1.21889100	0.94897500
N	-2.25947300	-3.93141800	-0.22486100	N	-3.46303500	-3.28062800	1.30467800
N	1.39506400	-2.69970100	-0.15648100	N	0.31611900	-3.18660300	1.48240500
N	-4.40985700	2.79091300	0.17220700	N	-3.46266900	3.28093400	-1.30426500
N	-0.68227600	3.89394700	0.22119700	N	0.31711900	3.18626900	-1.48181000
N	1.28425000	0.99119900	-3.56446300	N	1.32773600	-2.79395300	-2.25292900
N	5.35527300	-0.47698900	-2.22416300	N	5.60003400	-1.75210400	-1.37965900
N	5.36482500	-0.71998300	2.13444900	N	5.59921700	1.75296200	1.38031400
H	1.22501400	2.02453000	0.10979300	N	1.32640900	2.79379100	2.25235300
C	1.70011300	0.52958100	2.55534700			<b>3'</b>	



N	-0.77134900	4.10461400	-0.20794000	C	1.21802800	2.54817600	0.83759700
C	-0.22813900	-1.40337700	-2.06865800	N	1.34161800	3.62055300	1.27507200
N	0.40010000	-2.02235600	-2.82813900	C	1.11768900	-1.12610300	-1.71206600
C	1.38094300	1.79352300	0.63311800	N	0.23817200	-1.50591700	-2.37491900
N	1.87810200	2.62597400	1.27754200	C	3.97814900	-2.14233600	-1.87601900
C	3.30001700	0.76594200	-2.17935900	N	4.37533600	-2.98561900	-2.57359100
N	3.27416300	1.28863900	-3.21938400	C	5.72833800	-0.48578000	0.03556900
C	5.70115800	0.15141100	-0.35671700	N	6.87456000	-0.63243200	0.17720100
N	6.82314800	0.38024500	-0.55723900	C	3.99975200	1.56022400	1.44184000
C	4.68365100	-1.24987700	2.18152300	N	4.43028800	2.28903700	2.24134300
N	5.37678700	-1.60779800	3.04343600				
C	1.61833900	-1.53725000	1.98631400				
N	0.84506000	-2.00290400	2.72198800				
H	1.63546400	-1.22576900	-0.89394400				