

## Supporting Information

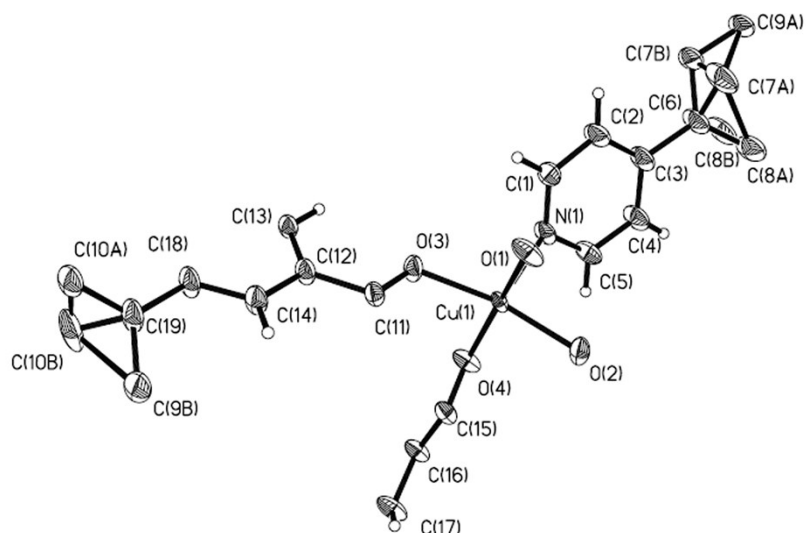
**Synthesis, structure and gas adsorption properties of a stable microporous metal-organic framework assembled from T-shaped pyridyl dicarboxylate ligand**

Di Wang, Libo Sun, Yuchuan Liu, Jianfeng Du, Shun Wang, Xiaowei Song\*  
and Zhiqiang Liang\*

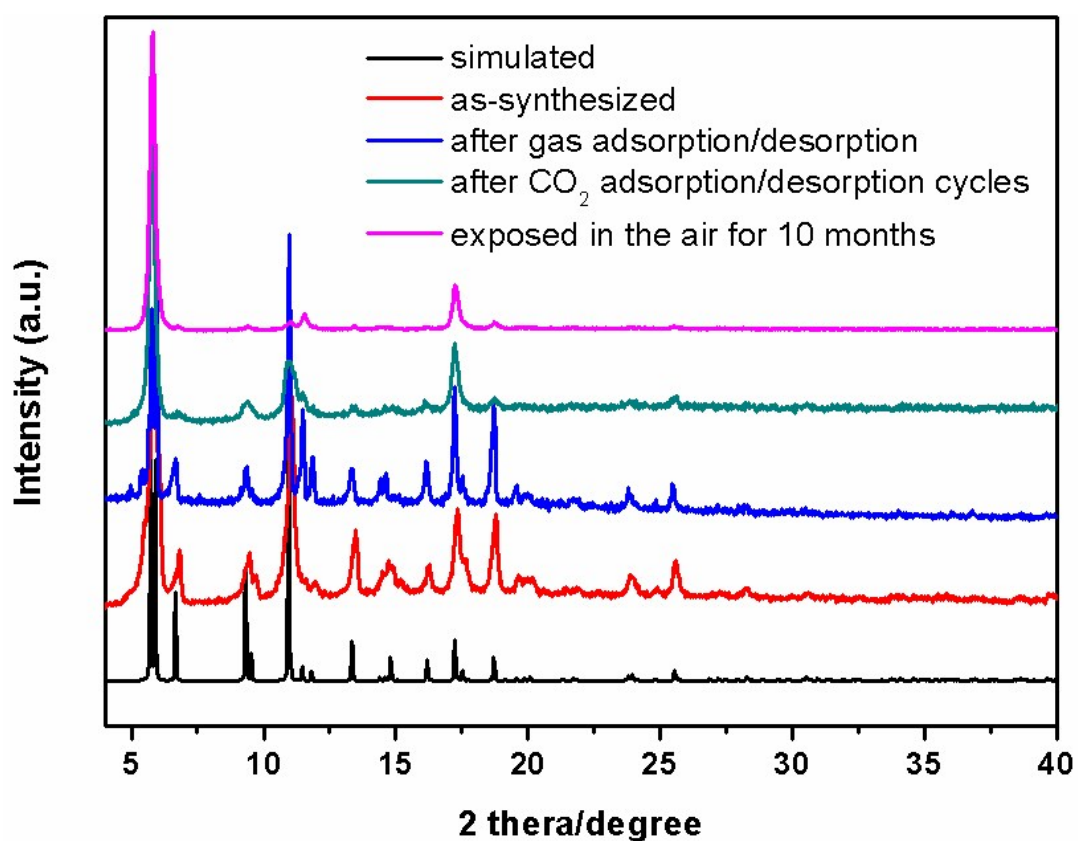
*State Key Laboratory of Inorganic Synthesis and Preparative Chemistry, College of Chemistry, Jilin University, Changchun 130012, P. R. China*

*E-mail: [xiaoweisong@jlu.edu.cn](mailto:xiaoweisong@jlu.edu.cn); [liangzq@jlu.edu.cn](mailto:liangzq@jlu.edu.cn)*

*Fax: +86-431-85168609*



**Fig.S1** Representations of the asymmetric units of compounds **1** showing ellipsoid at the 50% probability level.



**Fig.S2** Powder X-ray diffraction patterns of the simulated, as-synthesized, after gas adsorption/desorption measurements, after CO<sub>2</sub> adsorption/desorption cycles and exposed in the air for 10 months of compound **1**.

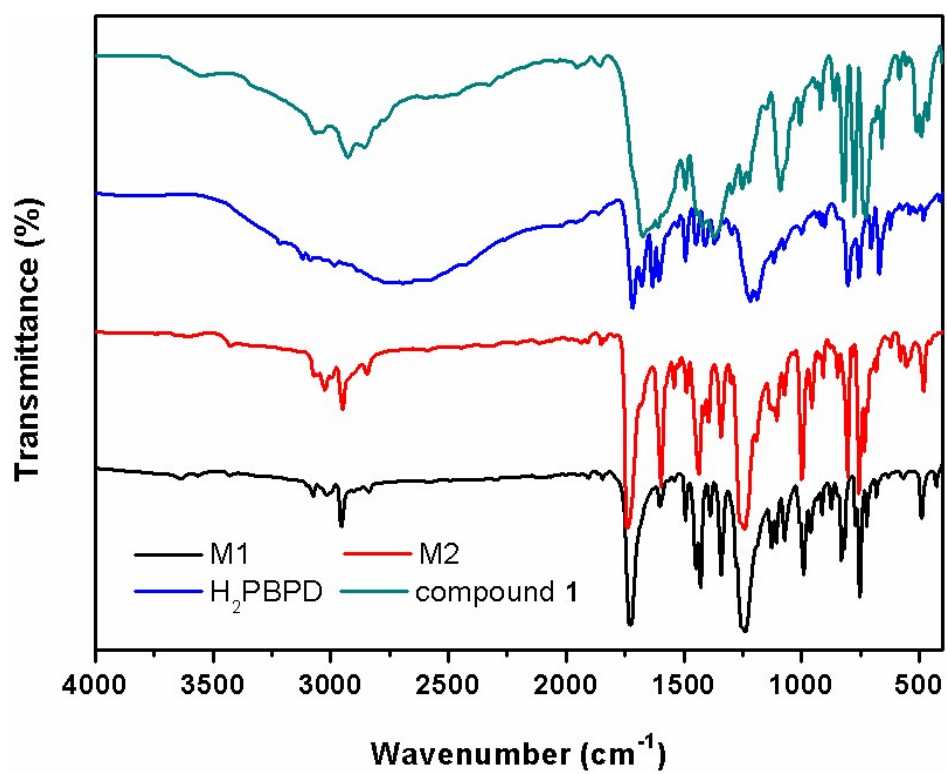


Fig.S3 Infra-red spectra of M1, M2, H<sub>2</sub>PBD and compound 1.

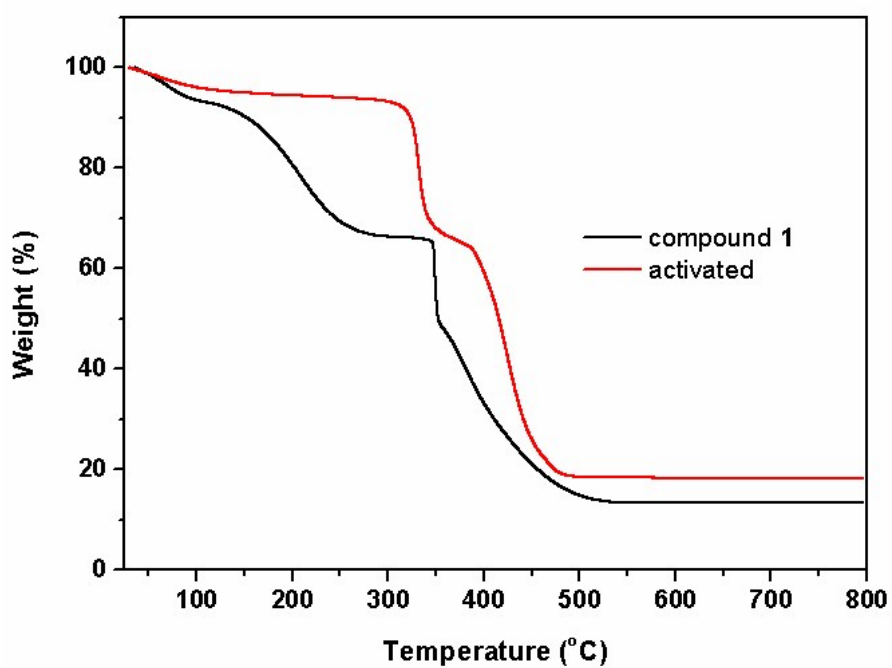
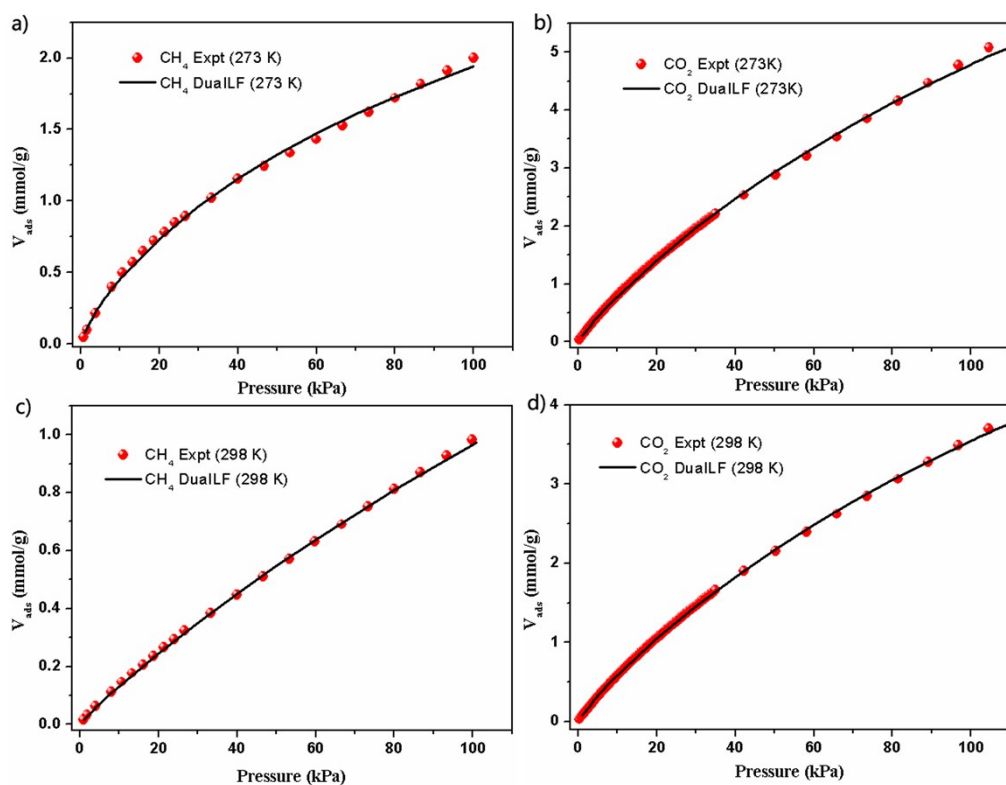
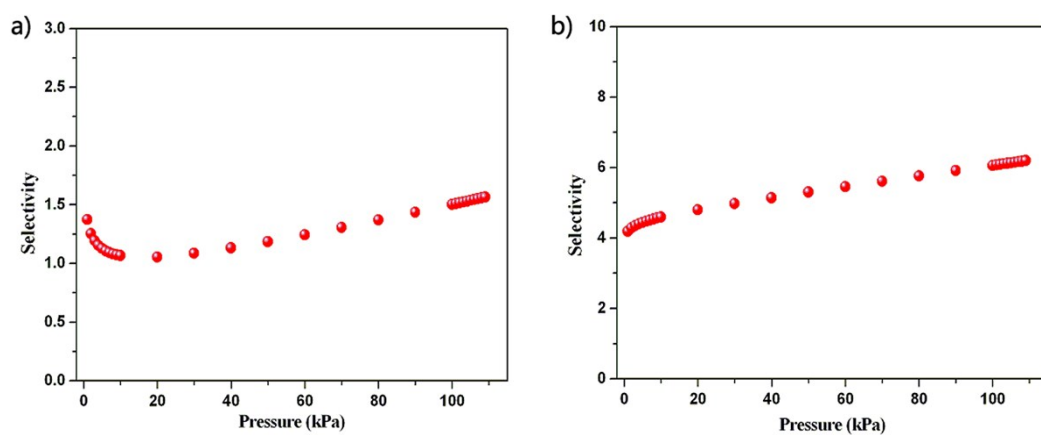


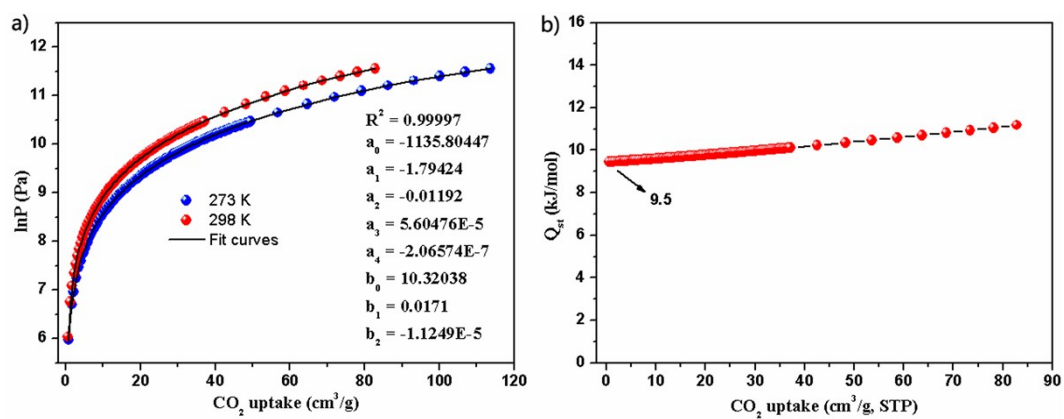
Fig. S4 TGA curves of compound 1 and activated.



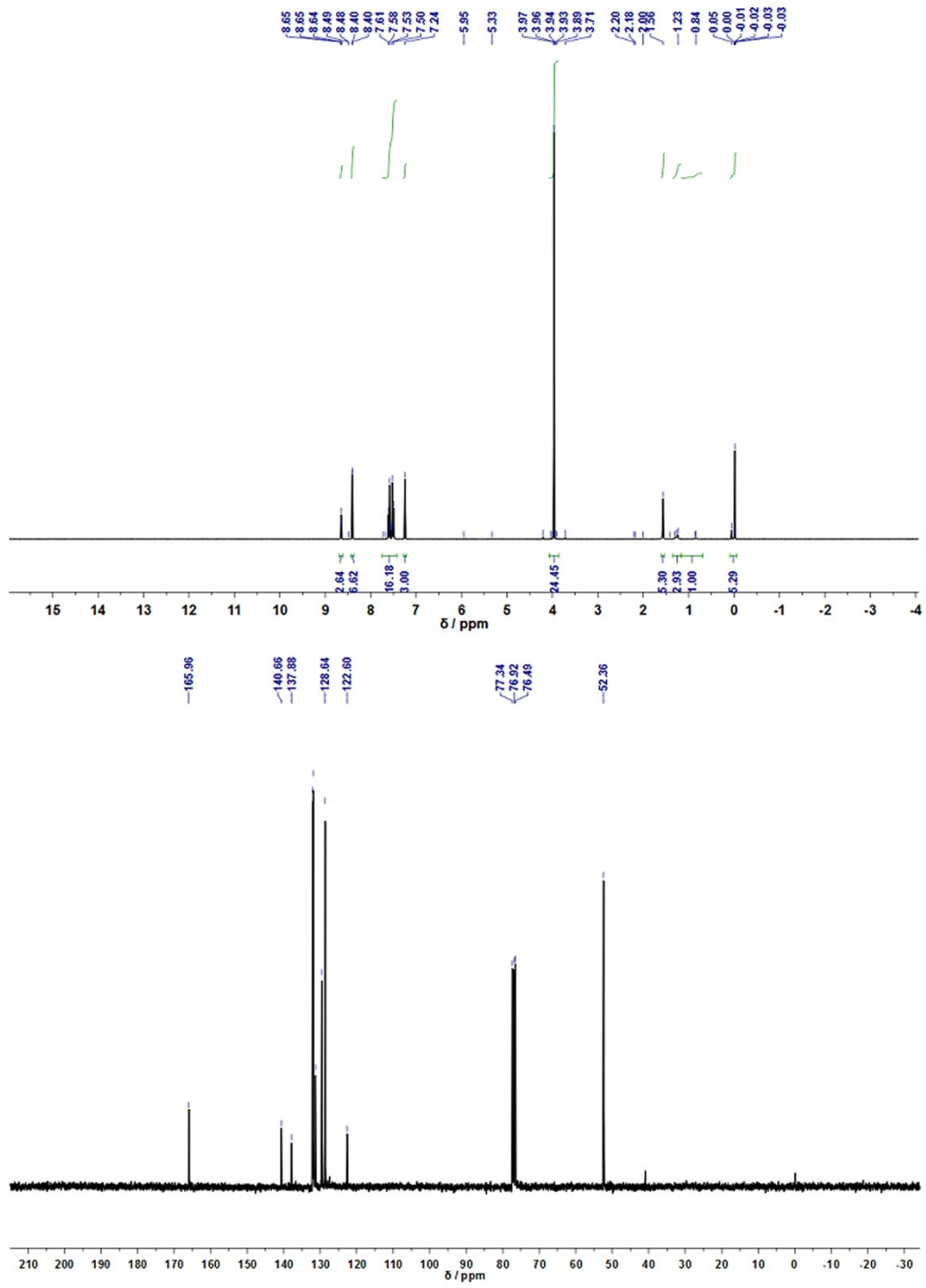
**Fig. S5** Measured  $CH_4$  and  $CO_2$  isotherms at 273 K and 298 K along with the DSLF fits for compound **1**.



**Fig. S6** IAST predicted equimolar gas mixture adsorption selectivities at 273 K (a) and 298 K (b) for compound **1**.



**Fig.S7** (a) Nonlinear curves fitting of  $\text{CO}_2$  for compound **1** at 273 K and 298 K; (b) Isostatic heat of  $\text{CO}_2$  for for compound **1**.



**Fig.S8** The  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of M1.

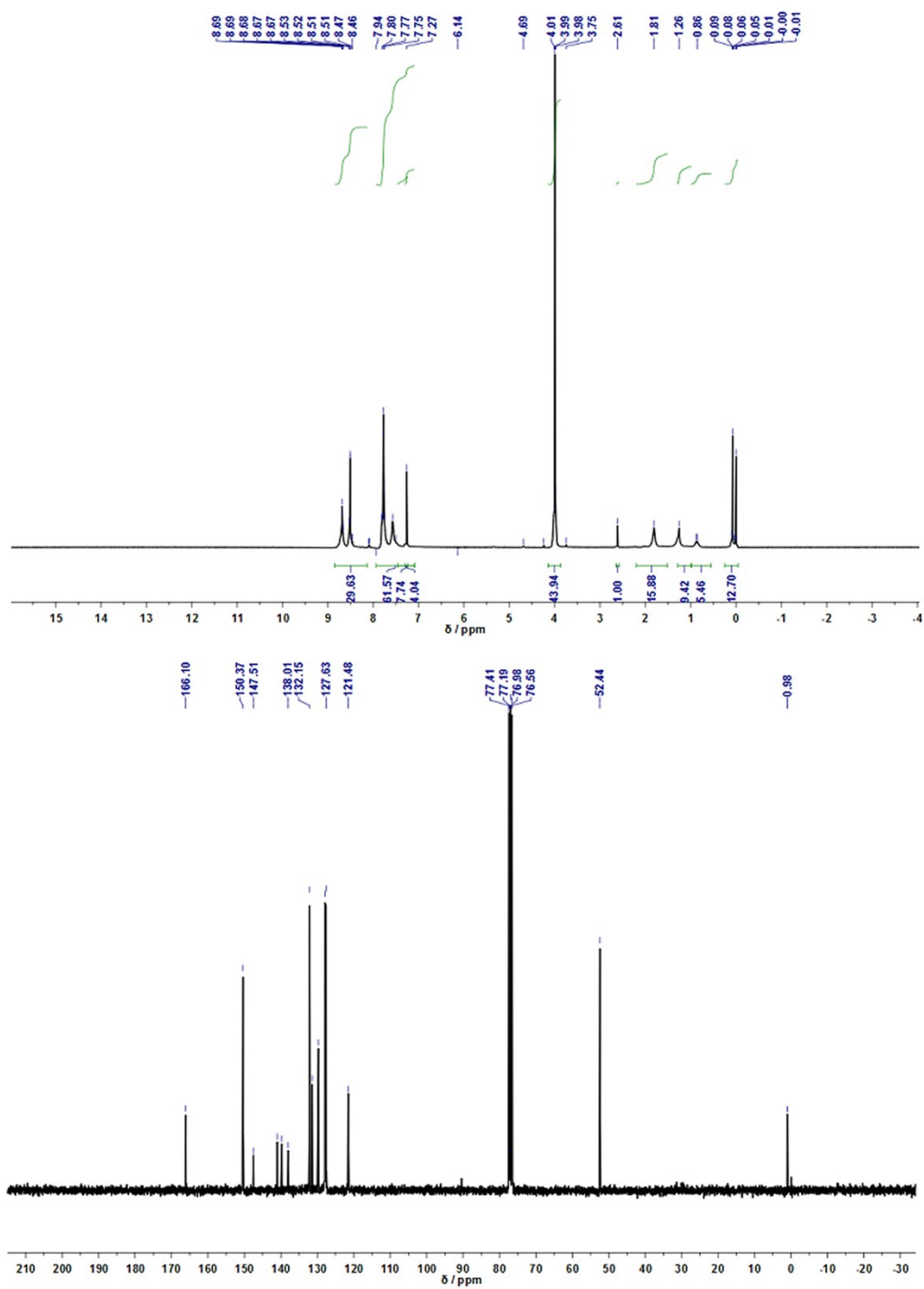
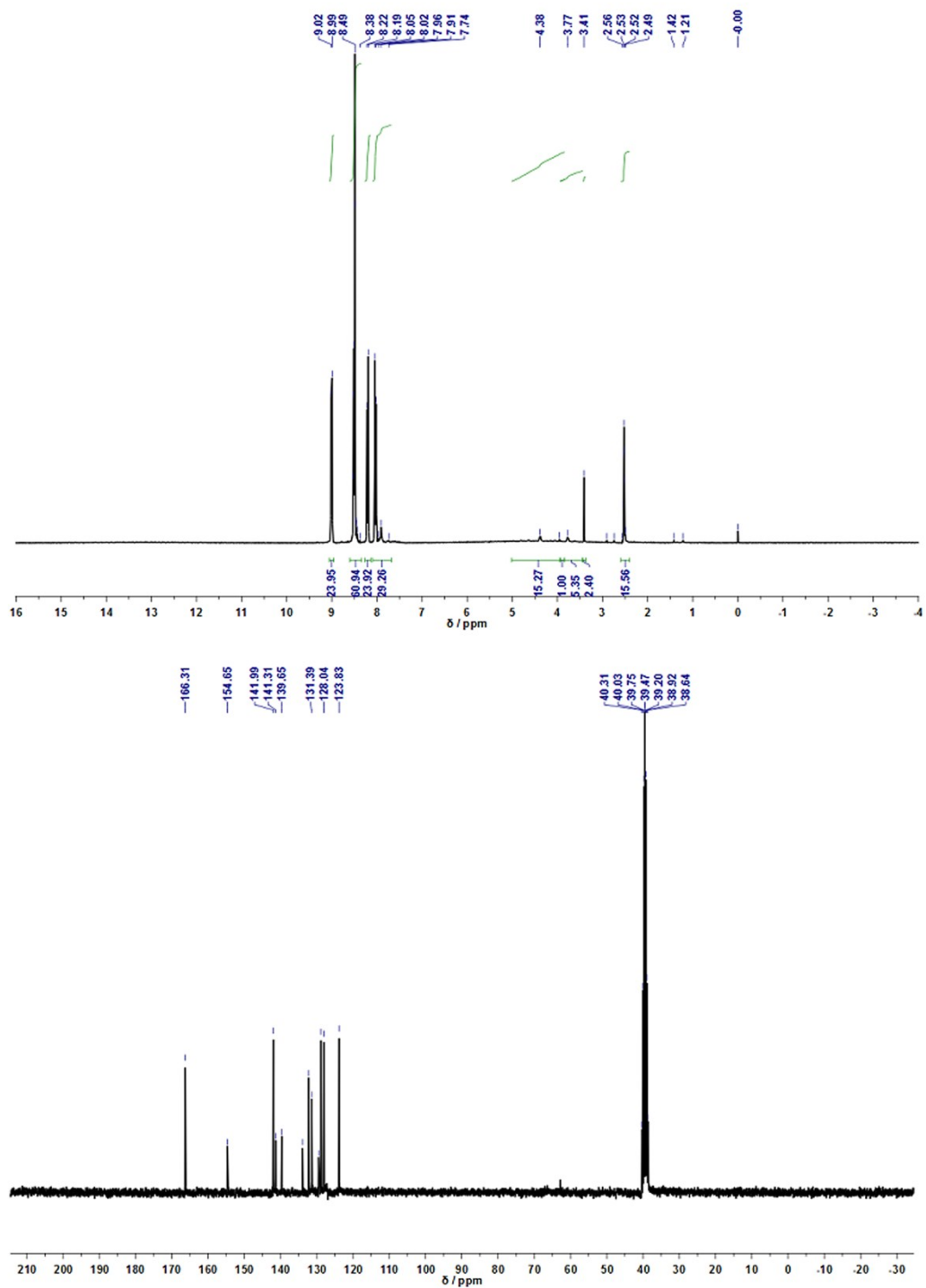


Fig.S9 The  $^1\text{H}$  and  $^{13}\text{C}$  NMR spectra of M2.



**Fig.S10** The <sup>1</sup>H and <sup>13</sup>C NMR spectra of H<sub>2</sub>PBDP.



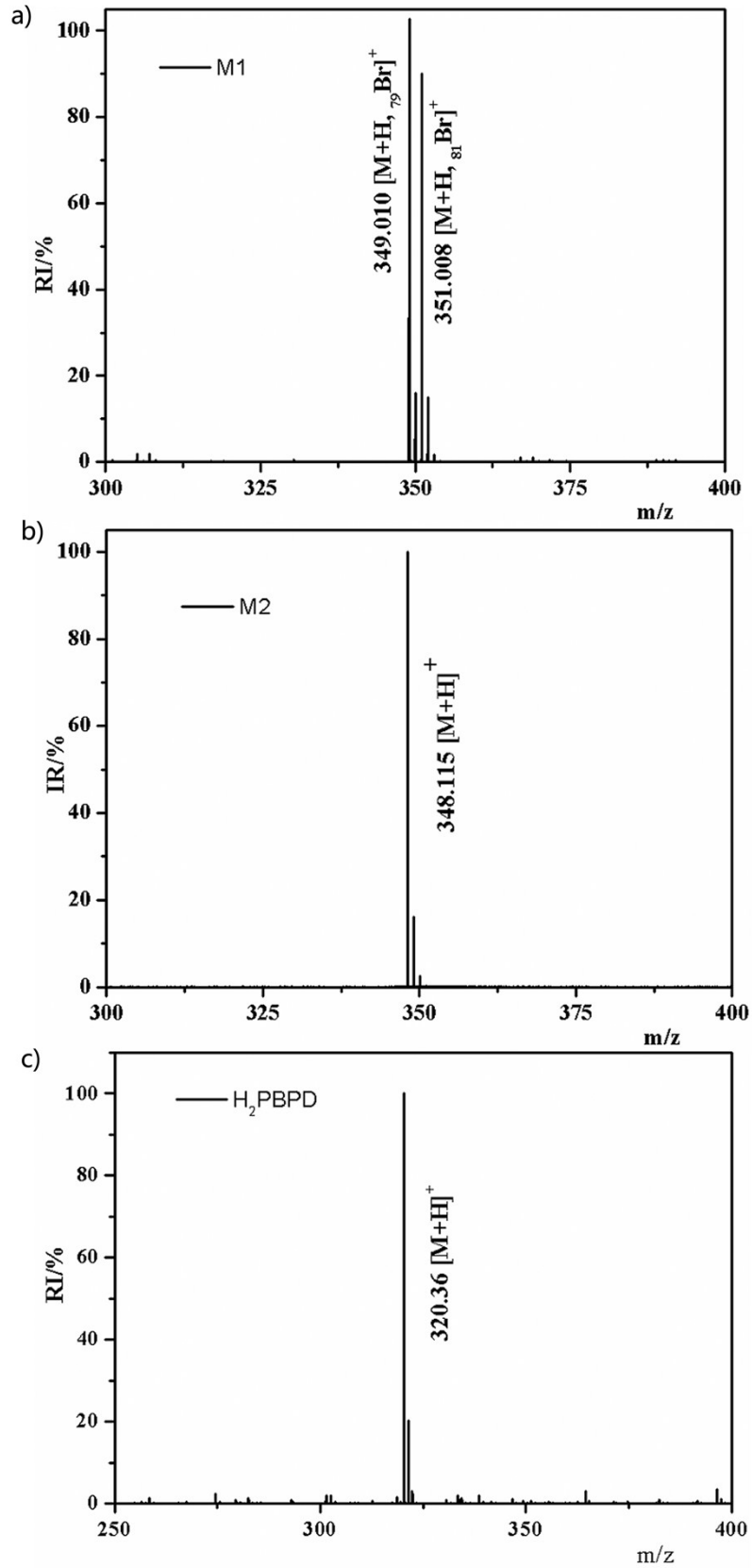


Fig.S11 MS of M1, M2 and H<sub>2</sub>PBPD.

**Table S1** Crystal Data and Structure Refinement for compound **1**.<sup>a</sup>

Formula	C <sub>25</sub> H <sub>31</sub> N <sub>3</sub> O <sub>9</sub> Cu
Formula weight	580.14
Temperature	296(2) K
Wavelength	0.71073 Å
Crystal system, space group	Hexagonal, <i>R</i> -3
Unit cell dimensions	$a = 18.9711(9)$ Å $\alpha = 90^\circ$ $b = 18.9711(9)$ Å $\beta = 90^\circ$ $c = 44.954(2)$ Å $\gamma = 120^\circ$
Volume	14011.3(11) Å <sup>3</sup>
<i>Z</i>	18
Calculated density	0.812 mg/cm <sup>3</sup>
Absorption coefficient	0.714 mm <sup>-1</sup>
<i>F</i> (000)	3474
Theta range for data collection	1.32 to 25.33°
Limiting indices	-20 ≤ <i>h</i> ≤ 22, -22 ≤ <i>k</i> ≤ 22, -49 ≤ <i>l</i> ≤ 53
Reflections collected / unique	30455 / 5645
<i>R</i> <sub>int</sub>	0.0452
Completeness to theta = 25.33	99.1%
Absorption correction	Semi-empirical from equivalents
Refinement method	Full-matrix least-squares on <i>F</i> <sup>2</sup>
Data / restraints / parameters	5645/0/262
Goodness-of-fit on <i>F</i> <sup>2</sup>	1.184
Final <i>R</i> indices [ <i>I</i> > 2σ( <i>I</i> )] <sup>b</sup>	<i>R</i> <sub>1</sub> = 0.0597, <i>wR</i> <sub>2</sub> = 0.2127
<i>R</i> indices (all data)	<i>R</i> <sub>1</sub> = 0.0735, <i>wR</i> <sub>2</sub> = 0.2315

<sup>a</sup>Date based on *PLATON/SQUEEZE* mode.

<sup>b</sup> $R_1 = \sum ||F_o| - |F_c|| / \sum |F_o|$ .  $wR_2 = [\sum [w (F_o^2 - F_c^2)^2] / \sum [w (F_o^2)]]^{1/2}$ .

**Table S2.** Selected bond lengths [Å] and angles [deg] for compound **1**.

Bond lengths	
Cu(1)-O(3)	1.953(3)
Cu(1)-O(4)	1.953(3)
Cu(1)-O(2)	1.971(3)
Cu(1)-O(1)	1.971(3)
Cu(1)-N(1)	2.188(3)
N(1)-C(1)	1.333(6)
N(1)-C(5)	1.336(6)

Bond angles	
O(3)-Cu(1)-O(4)	88.39(15)
O(3)-Cu(1)-O(2)	166.66(12)
O(4)-Cu(1)-O(2)	89.53(16)
O(3)-Cu(1)-O(1)	89.47(16)
O(4)-Cu(1)-O(1)	166.75(12)
O(2)-Cu(1)-O(1)	89.54(17)
O(3)-Cu(1)-N(1)	104.03(13)
O(4)-Cu(1)-N(1)	104.09(13)
O(2)-Cu(1)-N(1)	89.25(13)
O(1)-Cu(1)-N(1)	89.11(13)