Supporting Information

Insight into the synergism between MnO₂ and Acid Sites over Mn-SiO₂@TiO₂ Nano-cups for Low-Temperature Selective Catalytic Reduction of NO with NH₃

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Experimental section

Characterization methods

The NO adsoprtion of the catalysts was detected using the temperatureprogrammed desorption of nitric oxide (NO-TPD) on a Quantachrome Autosorb-IQ-C chemisorption analyser. The sample powders were firstly pre-treated in pure He at 300 °C for 1h and then cooled down to 50 °C. The catalysts were constantly saturated in anhydrous NO (10% in He) at a flow rate of 30ml/min for 30min and the physical adsorption of NO was removed by He for 30min. Finally, the desorption was carried out by heating the samples from 50 to 400 °C at a heat rate of 10 °C/min.

Catalyst test

The oxidation of NO to NO_2 was measured in the same fixed-bed reactor as mentioned in the manuscript. The composition of reactants were as follows: 1000ppm NO, 5% O_2 and balance N_2 , with the total flow rate of 200 ml min⁻¹. And the GHSV was also kept at 30000h⁻¹.

SiO ₂ @TiO ₂ -2.5 sample				
Catalysts	SSA ^[a]	Pore	Pore	
	(m^2g^{-1})	volume ^[a] (cm ³ g ⁻¹)	Diameter ^[a] (nm)	
Used Mn-SiO ₂ @TiO ₂ -2.5	64.49	0.136	11.98	
sample				

Table S1. The physical properties of the 120h-stability-testedMn-

[a] Specific surface area (SSA) and pore volume by N_2 -adsorption method.



Figure S1 XRD patterns of the fresh and stability-tested Mn-SiO₂@TiO₂-2.5 catalysts



Figure S2 SEM image of Mn-SiO₂@TiO₂-2.5 catalyst after 120h-stability test



Fig. S3 SO₂ and H₂0 poisoning effect on the NO conversion of Mn-SiO₂@TiO₂-2.5 catalyst. Reaction conditions: [NO]= [NH₃] =1000ppm, [O₂]=5%, [H₂O]=5 vol% (when used), [SO₂]=200ppm (when used), balance N₂, GHSV=30000h⁻¹

Fig. S4 TCD signal of Mn-SiO₂@TiO₂-2.5 catalyst saturated in He from 50-800 $^{\circ}$ C

Fig. S5 The NO conversion to NO₂ of the Mn-SiO₂@TiO₂-2.5 catalyst. Reaction conditions: [NO]=1000ppm, [O₂]=5%, balance N₂, GHSV=30000h⁻¹

Fig. S6 NO-TPD profiles of the Mn-based nano-cup catalysts