

## Flowery nickel-cobalt hydroxide via a solid-liquid sulphur ions grafting route and its application in hybrid supercapacitive storage

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### Electrochemical Calculations

The loading mass of positive and negative electrodes should meet the requirement of charge balance, as shown in Equation (1):

$$m_+/m_- = C_{s-}V_- / (C_{s+}V_+) \quad (1)$$

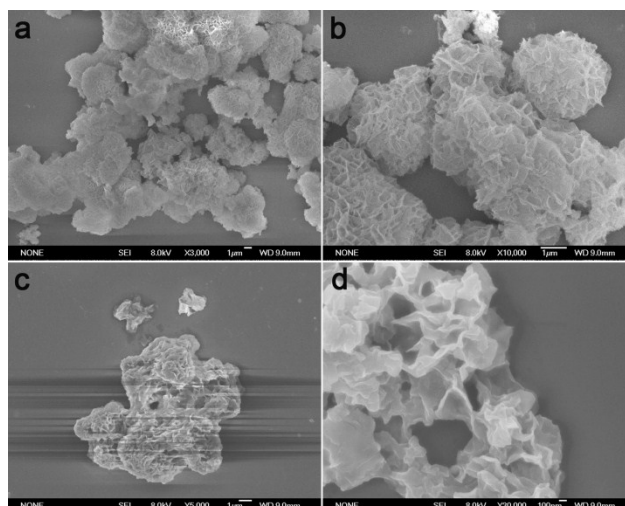
wherein  $m$ ,  $C_s$  and  $V$  are the mass loadings (g), specific capacitance ( $F\ g^{-1}$ ) and potential window (V), and + and – are corresponding to the positive and negative electrodes.

The specific capacitance ( $C_s$ ,  $F\ g^{-1}$ ), energy density ( $E$ ,  $Wh\ kg^{-1}$ ) and power density ( $P$ ,  $W\ kg^{-1}$ ) values of the individual electrode or ASC were calculated according to the Equations as follow:

$$C_s = I\Delta t / (mV) \quad (2)$$

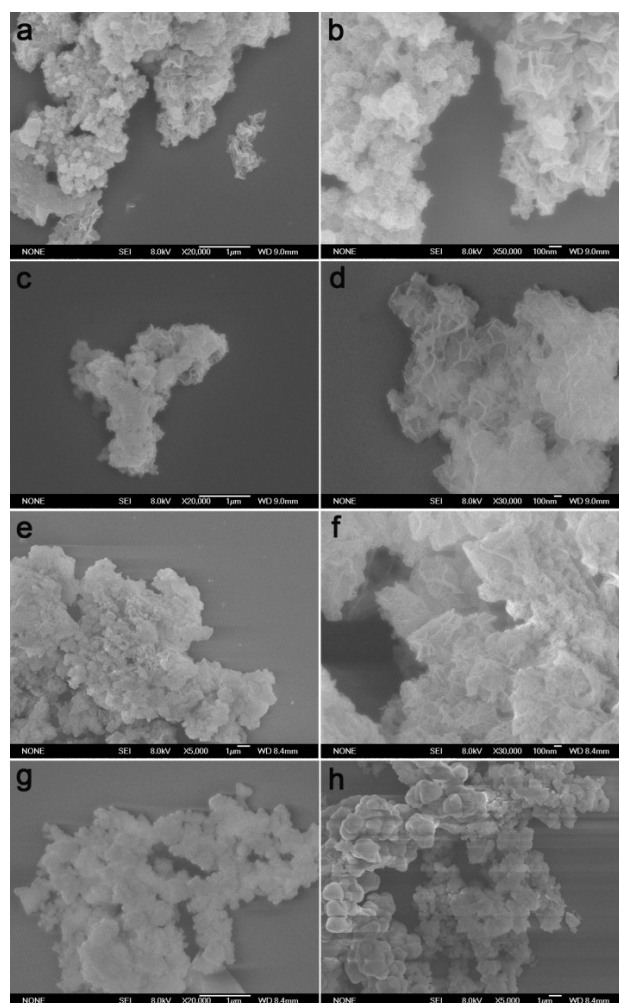
$$E = 1/2 C_s V^2 \quad (3)$$

$$P = E / \Delta t \quad (4)$$



**Fig. S1** FESEM figures for different samples: (a, b) Ni<sub>1</sub>Co<sub>2</sub> precursor; (c, d) Ni<sub>1</sub>Co<sub>2</sub>-

S.

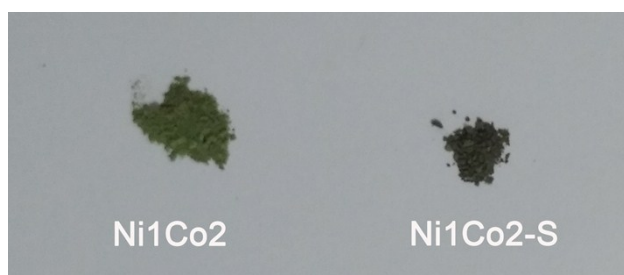


**Fig. S2** FESEM figures for different comparisons: (a, b) Ni<sub>1</sub>Co<sub>2</sub>-AN; (c, d) Ni<sub>2</sub>Co<sub>1</sub>;

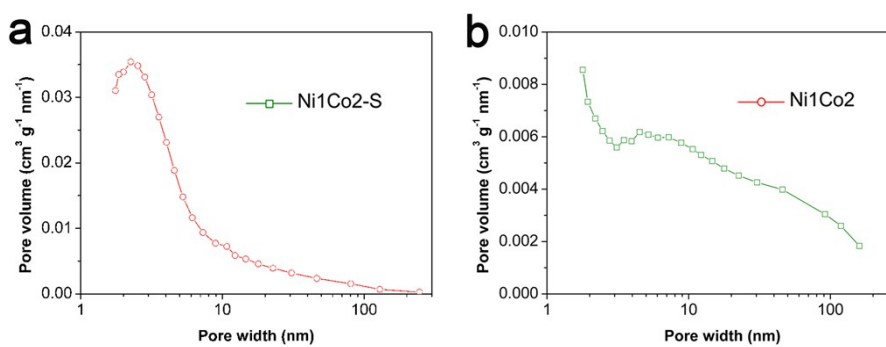
(e, f) Co<sub>3</sub> and (g, h) Ni<sub>3</sub>.

**Table S1.** ICP-OES measurement results of Ni1Co2, Ni1Co2-AN, Ni2Co1, Co3, Ni3 and Ni1Co2-S samples.

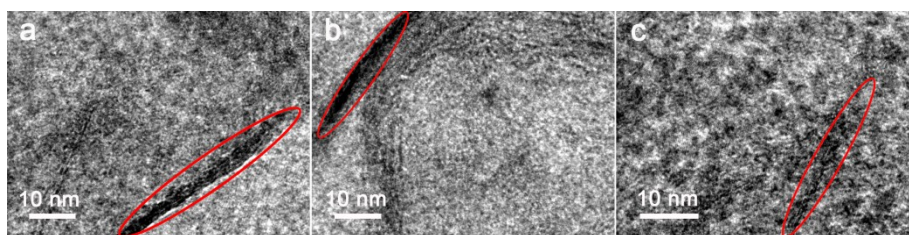
	average Ni concentration (mg L <sup>-1</sup> )	average Co concentration (mg L <sup>-1</sup> )	mass ratio Ni/Co	atomic ratio Ni/Co
Ni1Co2	0.4523	1.0317	1:2.28	1:2.27
Ni1Co2-AN	0.2697	0.6115	1:2.27	1:1.26
Ni2Co1	1.3533	0.7296	1.85:1	1.86:1
Co3	-0.0035 (ignore)	2.249	0:3	0:3
Ni3	2.766	-0.0017 (ignore)	3:0	3:0
Ni1Co2-S	0.1864	0.4179	1:2.24	1:2.23



**Fig. S3** Optical image of Ni1Co2 and Ni1Co2-S

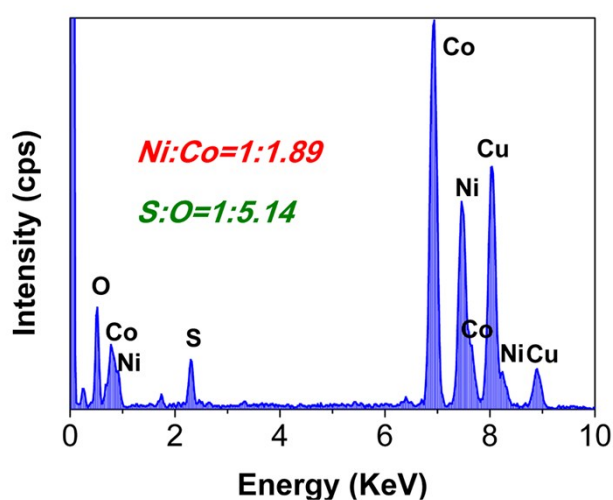


**Fig. S4** Pore width distribution of (a) Ni1Co2 and (b) Ni1Co2-S

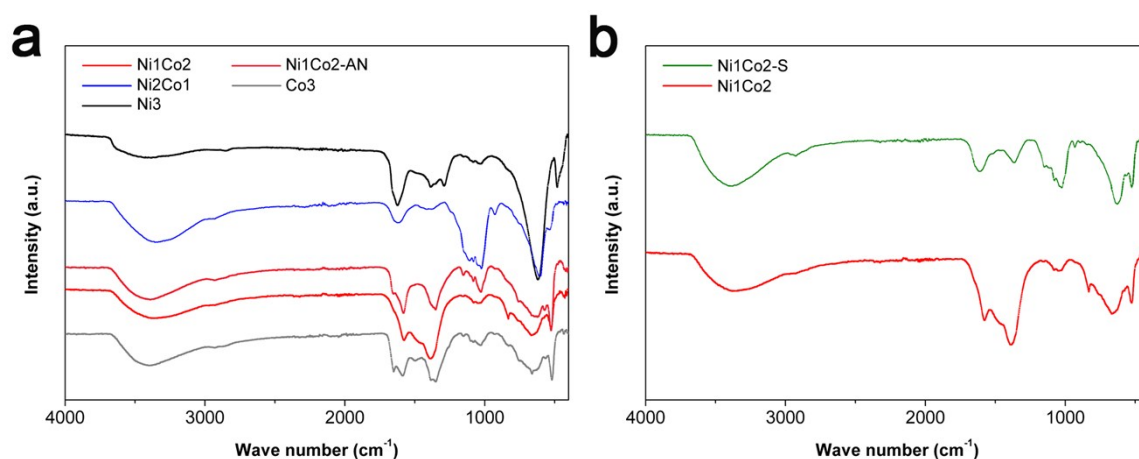


**Fig. S5** HRTEM figures of Ni<sub>1</sub>Co<sub>2</sub>-S at differnt areas

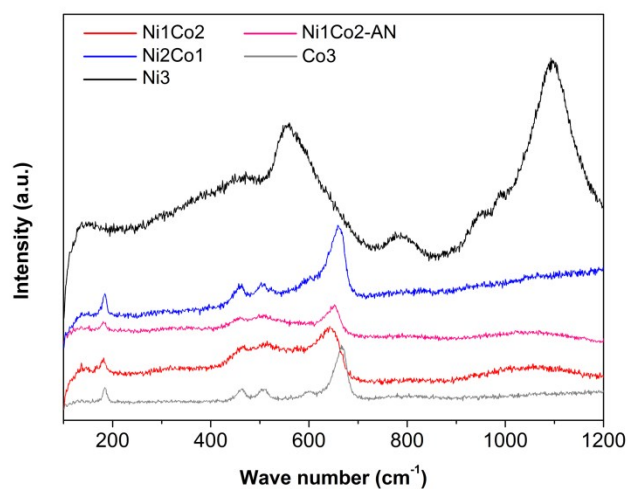
The nanosheet thickness for Fig. S5a-c is ca. 2-5 nm, also indicative of ultrathin structure of Ni<sub>1</sub>Co<sub>2</sub>-S nanosheets.



**Fig. S6** EDS spectrum of Ni<sub>1</sub>Co<sub>2</sub>-S.

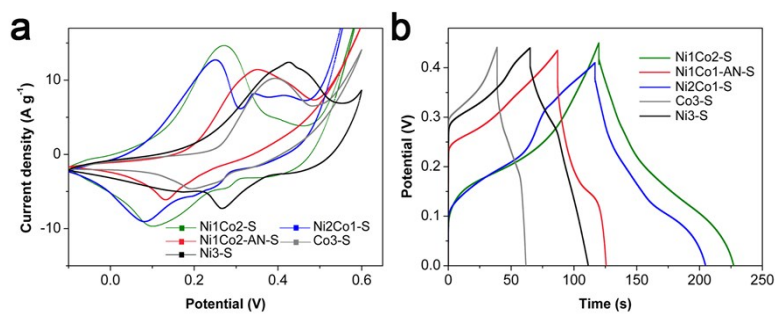


**Fig. S7** (a) FTIR patterns of Ni<sub>1</sub>Co<sub>2</sub>, Ni<sub>1</sub>Co<sub>2</sub>-AN, Ni<sub>2</sub>Co<sub>1</sub>, Co<sub>3</sub> and Ni<sub>3</sub> samples; (b) FTIR patterns of Ni<sub>1</sub>Co<sub>2</sub> and Ni<sub>1</sub>Co<sub>2</sub>-S samples.

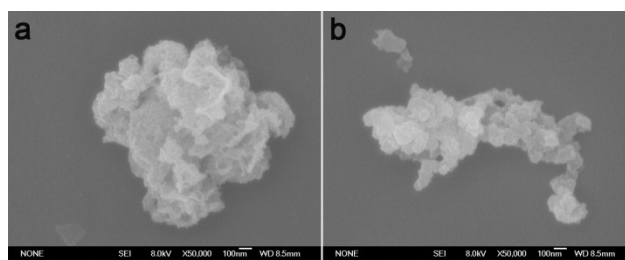


**Fig. S8** Raman patterns of Ni1Co2, Ni1Co2-AN, Ni2Co1, Co3 and Ni3 samples.

Herein, the cumulative number of Raman measurement is 8 times for every samples.



**Fig. S9** (a) CV curves of Ni1Co2-S, Ni1Co2-AN-S, Ni2Co1-S, Co3-S and Ni3-S electrodes at  $5 \text{ mV s}^{-1}$ , (b) GCD curves of Ni1Co2-S, Ni1Co2-AN-S, Ni2Co1-S, Co3-S and Ni3-S electrodes at  $5 \text{ A g}^{-1}$ .



**Fig. S10** (a and b) The FESEM figures of Ni<sub>1</sub>Co<sub>2</sub>-S electrode after 6000 cycles at 10 A g<sup>-1</sup>.