Electronic Supplementary Material (ESI) for Catalysis Science & Technology. This journal is © The Royal Society of Chemistry 2019

Supplementary Information

Oxygen Reduction Reaction on Core@Shell Nanostructured Au-d@Ni_mPt_m:

Effect of Core Size and Shell Thickness

Min Zhang,^a Shu Miao^b and Bo-Qing Xu*a

^aInnovative Catalysis Program, Key Lab of Organic Optoelectronics & Molecular Engineering, Department of Chemistry, Tsinghua University, Beijing 100084, China.

^bState Key Laboratory of Catalysis, Dalian Institute of Chemical Physics, Chinese Academy of Sciences, Dalian 116023, China.

* Corresponding author: Prof. Bo-Qing Xu Department of Chemistry Tsinghua University Beijing, 100084, China Tel.: +86 10 6279 2122 Fax: +86 10 6277 1149 E-mail: bqxu@mail.tsinghua.edu.cn



Figure S1. TEM images of (a) $(Pt_{0.15}^Au-2.2)$ $@Ni_2Pt_2/C$, and (b) $(Pt_{0.25}^Au-2.2)$ $@Ni_2Pt_2/C$.



Figure S2. TEM images and their measured size histograms in diameter of Au-*d* NPs: (a) d = 2.2, (b) d = 3.3, (c) d = 4.5, (d) d = 7.7, (e) d = 10.3.



Figure S3. (a) $Pt_{0.3}$ ^Au-2.2, (b) $Pt_{0.3}$ ^Au-3.3, (c) $Pt_{0.3}$ ^Au-4.5, (d) $Pt_{0.3}$ ^Au-7.7, (e) $Pt_{0.3}$ ^Au-10.3. (f) UV-vis spectra of Au-*d* NPs and $Pt_{0.3}$ ^Au-*d* NPs.



Figure S4. Structure and elemental analysis of as-prepared Au-2.2@Ni_{0.5}Pt_{0.5} NPs. (a) HAADF-STEM images and (b) elemental distribution analysis, elemental mapping of Au (green), Ni (red), Pt (blue) and their overlay.



Figure S5. Structure and elemental analysis of as-prepared Au-10.3@Ni₂Pt₂ NPs. (a) HAADF-STEM images and (b) elemental distribution analysis, elemental mapping of Au (green), Ni (red), Pt (blue) and their overlay. The NPs without Au core were marked in the yellow ellipses.



Figure S6. TEM images of Au-7.7@Ni_mPt_m/C samples: (a) m=1, (b) m=4, (c) m=6.

1. XRD characterizations



Figure S7. XRD patterns for Au-2.2@Ni₁Pt₁/C, Au-2.2@Ni₃Pt₃/C, Au-2.2@Ni₆Pt₆/C, Au-2.2/C and PtNi/C.



2. Activity of Au-7.7@Ni_mPt_m/C samples

Figure S8. (a) MSA_{Pt} and IA_{Pt} data of Au-7.7@Ni_mPt_m/C samples (*m*=1, 2, 4, 6, 9) at 0.9 V.

3. Detailed information for Au-*d*@Ni_mPt_m/C samples

Sample	EAS _{Pt}	Half-wave Potentials	IA _{Pt}	MSA _{Pt}
	$(m^2 \cdot g_{Pt}^{-1})$	(V)	$(mA \cdot cm_{Pt}^{-2})$	$(mA \cdot mg_{Pt}^{-1})$
Au-2.2@Ni _{0.5} Pt _{0.5}	102 (58)*	0.86	0.15 (0.1)*	157 (59)*
Au-2.2@Ni ₁ Pt ₁	86	0.91	0.37	328
Au-2.2@Ni ₂ Pt ₂	83	0.92	0.51	423
Au-2.2@Ni _{2.5} Pt _{2.5}	81	0.92	0.42	346
Au-2.2@Ni ₃ Pt ₃	78 (48)*	0.91	0.43 (0.31)*	330 (153)*
Au-2.2@Ni ₆ Pt ₆	57 (40)*	0.91	0.48 (0.45)*	278 (184)*
Au-3.3@Ni ₂ Pt ₂	71	0.91	0.56	402
Au-4.5@Ni ₂ Pt ₂	61	0.91	0.60	363
Au-7.7@Ni ₂ Pt ₂	47 (33)*	0.93	1.06 (0.98)*	496 (322)*
Au-7.7@Ni ₁ Pt ₁	76	0.93	0.74	562
Au-7.7@Ni ₄ Pt ₄	45	0.93	0.79	360
Au-7.7@Ni ₆ Pt ₆	44	0.94	0.99	432
Au-7.7@Ni ₉ Pt ₉	42	0.93	0.92	388
Au-10.3@ Ni ₂ Pt ₂	36	0.92	1.09	392
Pt/C (E-TEK)	76 (32)*	0.88	0.25 (0.17)*	189 (66)*
PtNi/C	39 (35)*	0.92	1.01 (0.37)*	389 (135)*

Table S1. Electrochemical properties of the Au-*d*@Ni_{*m*}Pt_{*m*}/C samples.

* The numbers in the parentheses were obtained in the measurements after ADT.

Table S2. Detailed experimental parameters/conditions for the synthesis of Au NPs of different sizes

Sample	$d_{\text{seeds}}(\text{nm})^{a}$	V _{seeds} (ml) ^b	$V_{\rm HAuCl4}(ml)^{\rm c}$	$V_{AA}(ml)^d$	D (nm) ^e
Au-2.2	-	-	1.00	-	2.2 ± 0.3
Au-3.3	2.2	10.0	0.25	3.8	3.3 ± 0.5
Au-4.5	2.2	10.0	1.00	15.0	4.5 ± 0.7
Au-7.7	4.7	8.2	0.50	7.5	7.7 ± 1.2
Au-10.3	4.7	5.8	1.00	15.0	10.3 ± 1.9

^a Average size by diameter of Au seeds. ^b The concentration of Au was 1.0×10^{-3} M. ^c The concentration of HAuCl₄ was 5.0×10^{-3} M. ^d The concentration of ascorbic acid was 5.0×10^{-3} M. ^e Average size by diameter of Au NPs.