# Electronic Supplementary Information

# An aggregation-induced emission-based fluorescence turn-on probe for Hg<sup>2+</sup> and its application to detect Hg<sup>2+</sup> in food samples

Lijun Tang,<sup>a,\*</sup> Haili Yu,<sup>a</sup> Keli Zhong,<sup>a</sup> Xue Gao,<sup>b</sup> Jianrong Li<sup>b,\*</sup>

<sup>a</sup> College of Chemistry and Chemical Engineering, Bohai University, Jinzhou, 121013, China. E-mail: ljtang@bhu.edu.cn (L. Tang)

<sup>b</sup> College of Food Science and Technology, Bohai University; National & Local Joint Engineering Research Center of Storage, Processing and Safety Control Technology for Fresh Agricultural and Aquatic Products; The Fresh Food Storage and Processing Technology Research Institute of Liaoning Provincial Universities, Jinzhou, 121013, China. E-mail: lijr6491@163.com (J. Li)

## Sample pretreatment

#### 1. Pretreatment of shrimp and crab meat samples

Shrimp and crab were purchased from the supermarket and the meat was firstly treated with a digestion procedure. Each sample (0.5 g) was soaked overnight in a beaker with  $HNO_3$  (10 mL) at room temperature, then mixture was heated to boil until it was completely dissolved. After cooling, the solution was centrifuged, and the supernatant was adjusted to pH = 7.4 with 1M NaOH solution, and constant the volume to 50 mL a 50 mL volumetric flask.

### 2. Pretreatment of tea samples

0.500 g of mashed tea was placed in a 100 mL beaker, to which 20 mL of concentrated nitric acid was added separately. The beaker was sealed with a plastic wrap and placed overnight. Then it was put in a microwave oven and digested 6h under 400 W of power. It was then placed in a fume hood to cool, and the supernatant was adjusted to pH = 7.4 with 1 M NaOH solution. Then the solution was transferred to a 100 mL volumetric flask and brought up to volume.

# **Supplementary figures**



**Fig. S1**. Photograph of Tyndall phenomena of a CH<sub>3</sub>OH/PBS (20 mM, pH = 7.4) (3:7, v/v) solution with (A) and without (B) compound **4** via illuminating with a laser pointer.



**Fig. S2**. Particle size distributions of compound **4** in CH<sub>3</sub>OH/PBS (20 mM, pH = 7.4) (3:7, v/v) solution.



Fig. S3. The linear relationship between absorbance and TPE-M concentration in CH<sub>3</sub>OH/PBS (20 mM, pH = 7.4) (3:7, v/v) solution.



**Fig. S4**. Particle size distributions of **TPE-M** in CH<sub>3</sub>OH/PBS (20 mM, pH = 7.4) (3:7, v/v) solution before (A) and after (B) addition of Hg<sup>2+</sup>.



Fig. S5. Standard calibration curve of emission intensity of probe TPE-M against  $Hg^{2+}$  concentrations (0 to 15  $\mu$ M) in CH<sub>3</sub>OH/PBS (20 mM, pH = 7.4) (3:7, v/v) solution.



Fig. S6. HRMS (ESI+) spectrum of TPE-M+Hg<sup>2+</sup>.



Fig. S7. <sup>1</sup>H NMR spectrum of Compound 3 in DMSO-*d*<sub>6</sub>.



Fig. S8. <sup>13</sup>C NMR spectrum of compound 3 in DMSO- $d_6$ .



Fig. S9. HRMS (ESI+) spectrum of compound 3.



Fig. S10. <sup>1</sup>H NMR spectrum of compound 4 in DMSO- $d_6$ .



Fig. S11. <sup>13</sup>C NMR spectrum of compound 4 in DMSO- $d_6$ .



Fig. S12. HRMS (ESI-) spectrum of compound 4.



**Fig. S13**. <sup>1</sup>H NMR spectrum of probe **TPE-M** in DMSO- $d_6$ .



Fig. S14. <sup>13</sup>C NMR spectrum of probe TPE-M in DMSO- $d_6$ .



Fig. S15. HRMS (ESI+) spectrum of probe TPE-M.