Supporting Information for

Controlled formation of core shell structures with a uniform AlPO₄ nanoshell

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Experimental section

Materials. The pristine LiCoO₂ powders (with an average particle size of ~10 μ m) and Co₃O₄ powders (with an average particle size of ~5 μ m) were purchased from Pulead Technlogy Industry Co., Ltd. Ag nanoparticles (55 nm) and TiO₂ (Rutile) were purchased from Beijing DK nano technology Co. Ltd. Aluminum nitrate, disodium hydrogen phosphate, urea and nitric acid were of analytical pure and purchased from Sinopharm Chemical Reagent Co. Ltd.. Mili-Q water (resistance > 18 MΩ) was used for all experiments.

Preparation of PS@AIPO₄. The sulfonated polystyrene spheres (d=460nm) were synthesized following the literature procedures.¹ 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, 550 μ L (100 mg/mL) the obtained PS spheres solution were added to the coating solution. Finally, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 90 °C for 3h under vigorously stirring. The obtained product was centrifuged at 8000 rpm for 1min, which was followed by washing with water and ethanol and drying at 80 °C for 10h, respectively. The obtained powder was annealed at 500 °C for 2h in air to remove the core and obtain aluminum phosphate hollow spheres. The concentration of AlPO₄ solution to PS was varied to control the thickness of AlPO₄ coating layer.

Preparation of Ag@AIPO₄. 0.2g Ag, 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 90 °C for 3h under vigorously stirring. The obtained product was separated from the solution by centrifuging at 8000 rpm for 1min, washing with water and ethanol and drying at 80 °C for 10h, respectively.

Preparation of nano-Co₃O₄@AIPO₄. Nano-Co₃O₄, derived from micro-Co₃O₄ with ball milling, was annealed at 900 °C for 5h in air. 0.2g the pretreatment nano-Co₃O₄, 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 90 °C for 3h under vigorously stirring. The obtained product was separated from the solution by centrifuging at 8000 rpm for 1min, washing with water and ethanol and drying at 80 °C for 10h, respectively.

Preparation of TiO₂@AIPO₄. 0.26g TiO₂, 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 90 °C for 3h under vigorously stirring. The obtained product was separated from the solution by centrifuging at 8000 rpm for 1min, washing with water and ethanol and drying at 80 °C for 10h, respectively.

Preparation of nano-LiCoO₂@AIPO₄. Nano-LiCoO₂, derived from micro-LiCoO₂ with ball milling, was annealed at 900 °C for 5h in air. 0.6g the pretreatment nano-LiCoO₂, 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 90 °C for 3h under vigorously stirring. The obtained product was separated from the solution by centrifuging at 8000 rpm for 1min, washing with water and ethanol and drying at 80 °C for 10h, respectively.

Preparation of LiCoO₂(*AIPO*₄. 2.5g LiCoO₂, 0.046g Al(NO₃)₃·9H₂O, trace amount of HNO₃, 0.044g Na₂HPO₄·12H₂O and 0.5g CO(NH₂)₂ were slowly dissolved in 30 mL distilled water under mechanical magnetic stirring. Subsequently, the resultant solution was titrated to pH=2.5 by nitric acid and then incubated at 80 °C for 3h under vigorously stirring. The obtained product was separated from the solution by centrifuging at 8000 rpm for 1min, washing with water and ethanol, drying at 80 °C for 10h. The samples have been annealed at 700 °C for 2 h for their electrochemical test.

Electrochemical measurement. Charge-discharge tests were performed with CR2032 coin cells assembled in an argon-filled glove box, which consisted of a lithium metal as the counter and reference electrode, a cathode as the working electrode, 1M LiPF₆ in ethylene carbonate (EC)/dimethyl carbonate (DMC)/diethyl carbonate (DEC) (1:1:1, in wt %) as the electrolyte, and polypropylene membrane (Celgard) as a separator. The cathode was prepared as a mixed slurry of 80 wt % active materials, 10 wt% of Super P carbon black, and 10 wt % of polyvinylidene fluoride (PVDF, Aldrich) binder in a n-methyl pyrrolidone (NMP) solvent, casting onto Al foil (99%, Goodfellow) and cutting into circular electrodes of 1.13 cm² area after drying under a vacuum at 80 °C for 10h. Galvanostatic tests of the assembled cells were performed with a Land CT2001A battery test system with a voltage window of 2.75-4.5V at room temperature.

Characterization. Scanning electron microscopy (SEM) was collected on Filed Emission Scanning Microscopy (FE-SEM, Hitachi S-4800). Transmission electron microscopy (TEM) and high resolution transmission electron microscopy (HRTEM) images were recorded on a JEOL-

2100F and FEI Tecnai F20 operated at 200kV. X-ray diffraction (XRD) data was collected on a Rigaku D/MAX-2500 with Cu K α radiation at 50 kV and 250 mA. X-ray photoelectron spectroscopy (XPS) was performed on the Thermo Scientific ESCALab250Xi using 200 W monochromated Al K α radiation.

References.

1. P. M. P, Google Patents, 1967.



Figure S1. (a-c) SEM images of pristine PS (a), PS@AlPO₄ (b), and PS@AlPO₄ after annealed at 500 °C in air for 2 h (c). Low magnification TEM image of PS@AlPO₄ after annealed at 500 °C in air for 2 h (d). The broken sphere particles and the brighter edge clearly identified the hollow character of AlPO₄spheres.



Figure S2. (a) The dispersive spectroscopy (EDS) pattern of the randomly picked hollow spheres. (b) XRD pattern of the PS@AlPO₄ after annealed at 500 °C in air for 2 h.



Figure S3. TEM images of core shell structures synthesized with different seeds: (a) Ag@ AlPO₄, (b) nano-Co₃O₄@AlPO₄, (c) TiO₂@AlPO₄, (d) nano-LiCoO₂@AlPO₄.



Figure S4. Low magnification TEM image of $PS@AlPO_4$, without controlling the precipitation process of AlPO₄, revealed phase separation.



Figure S5. (a-b): SEM images of pristine $LiCoO_2$ and $LiCoO_2@AIPO_4$ samples. The insets are high magnification SEM images. (c-d): TEM images of surface modified $LiCoO_2$ samples: c) uniform AIPO₄ coating and d) partial AIPO₄ coating.



Figure S6. (a) XRD patterns of the pristine $LiCoO_2$ and uniformly coated one. (b) Chargedischarge profiles of the 100^{th} cycle of the pristine $LiCoO_2$, partially coated one, uniformly coated one, which are operate at 0.1C with a voltage window between 2.75V and 4.5V at room temperature.