

Electric Supplemental Information for:

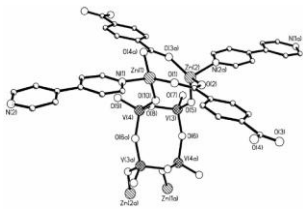
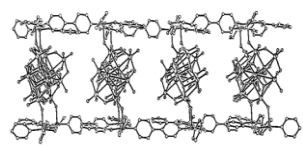
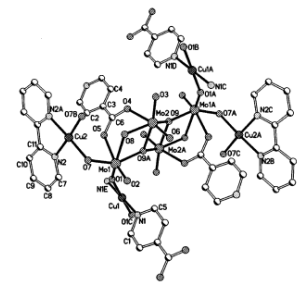
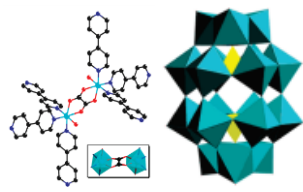
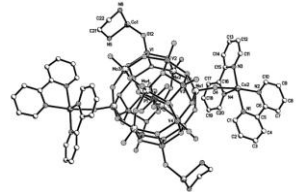
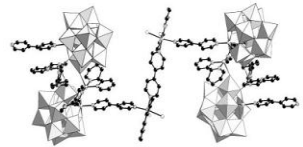
Tuning the Dimension of the POM-based Inorganic-organic Hybrids from 3D Self-penetrating Framework to 1D Poly-pendant Chain *via* Changing POM Clusters and Introducing Secondary Spacers†

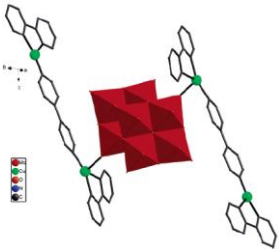
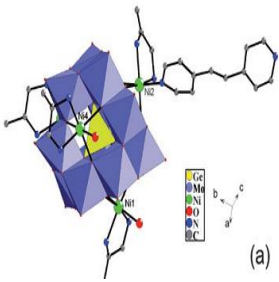
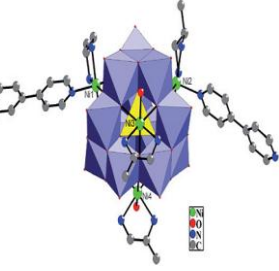
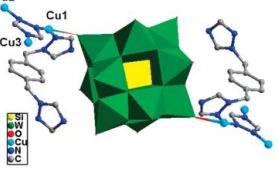
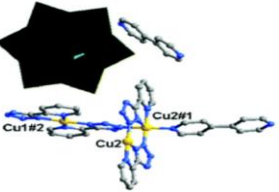
Shaobin Li, Huiyuan Ma*, Haijun Pang*, Zhuanfang Zhang, Yan Yu, Heng Liu and Tingting Yu

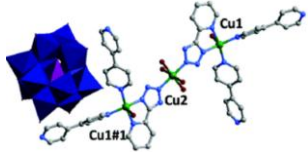
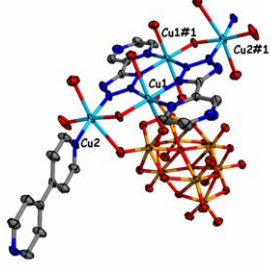
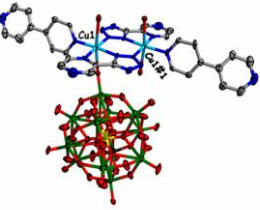
Table of contents:

1. **Table S1.** Summarization of typical POM-based hybrids constructed by mixed-ligands.
2. **Fig. S1** The two kinds of TMCs in **1**.
3. **Fig. S2** The process of 2D layers is formed in **1**.
4. **Fig. S3** The infinite organic-inorganic hybrid chain formed by β -Mo₈ anions and subunit C in **1**.
5. **Fig. S4** The detailed view of the self-penetrating structure in **1**.
6. **Fig. S5** View of the interdigitated structure in **3**.
7. **Fig. S6** Combined polyhedral/ball/stick representation of the infinite 1D Poly-pendant chain in **4**.
8. **Fig. S7** The IR spectra of compounds **1-4**.
9. **Fig. S8** The simulative (black) and experimental (red) powder X-ray diffraction patterns for **1-4**.
10. **Fig. S9** Solid-state emission spectra of pzta ligand at room temperature.
11. **Fig. S10** The CVs of **1-CPE** (left) and **2-CPE** (right) in the 1 M H₂SO₄ solution at the scan rate of 50 mV·s⁻¹.
12. **Fig. S11** (a) Reduction of H₂O₂ and (b) oxidation of AA for **2-CPE** in 1M H₂SO₄ solution at the scan rate of 50 mV·s⁻¹.

Table S1 Summarization of typical POM-based hybrids constructed by mixed-ligands

Compounds	Ligands 1 and 2	Types of mixed-ligands	References
 <p>[Zn₂(tp)(4,4'-bpy)V₂O₆]</p>	tp 4,4'-bpy	I	<i>Tong and Chen et al., J. Chem. Soc., Dalton Trans.</i> 2001, 770
 <p>[Cu(4,4'-bpy)(nic)(H₂O)]₂Mo₈O₂₆</p>	nic 4,4'-bpy	I	<i>Lu et al., Chem. Mater.</i> 2002, 14, 2649
 <p>[{Cu^{II}(2,2'-bpy)}{Cu^{II}(IN)₂}{Mo₄O₁₂(OH)₂}]</p>	IN 2,2'-bpy	I	<i>Wang et al., Inorg. Chem.</i> 2003, 42, 6956
 <p>[Ni₂(4,4'-bpy)₃(H₂O)₂(ox)][P₂W₁₈O₆₂]₂ (H₂4,4'-bpy)·H₂O</p>	ox 4,4'-bpy	I	<i>Liu and Su et al., Inorg. Chem.</i> 2008, 47, 7133
 <p>[Co(en)₂][Co(2,2'-bipy)₂] -[PMo^{VI}₅Mo^V₃V^{IV}₈O₄₄]-4.5H₂O</p>	en 2,2'-bipy	II	<i>Zhang and Zhu et al. Chem. Commun.</i> 2002, 1416
 <p>[4,4'-Hbipy][{Cu₂(2,2'-bipy)₂(4,4'-bpy)_{2.5}}</p>	2,2'-bipy 4,4'-bipy	II	<i>Dolbecq et al. Dalton Trans.</i> 2005, 3919

<p>$-\text{PW}_{11}\text{CuO}_{39}] \cdot 16\text{H}_2\text{O}$</p>  <p>$\text{K}[\{\text{Cu}^{\text{I}}(2,2'\text{-bipy})\}(4,4'\text{-bipy})$ $-\{\text{Cu}^{\text{I}}(2,2'\text{-bipy})\}_{0.5}\}_2[\text{Mo}_8\text{O}_{26}]$</p>	<p>2,2'-bipy 4,4'-bipy</p>	<p>II</p>	<p><i>Wang et al.,</i> <i>Cryst. Growth</i> <i>Des.</i> 2006, 6, 2693</p>
 <p>$[\text{GeMo}^{\text{V}}_8\text{Mo}^{\text{VI}}_4\text{O}_{36}(\mu_2\text{-OH})_4\{\text{Ni}(\text{pda})(\text{H}_2\text{O})\}_2$ $-\{\text{Ni}(\text{pda})\}\{\text{Ni}(\text{pda})(\text{bpe})\}(\text{bpe})_{0.5}]$ (1)</p>  <p>$[\text{GeMo}^{\text{V}}_8\text{Mo}^{\text{VI}}_4\text{O}_{36}(\mu_2\text{-OH})_4\{\text{Ni}(\text{pda})\}_2$ $-\{\text{Ni}(\text{pda})(4,4'\text{-bpy})_{0.5}\}\{\text{Ni}(\text{pda})(4,4'\text{-bpy})\}]$ $\cdot 7\text{H}_2\text{O}$ (2)</p>	<p>pda bpe</p>	<p>II</p>	<p><i>Xu et al.,</i> <i>CrystEngComm.</i> 2009, 11, 2488</p>
 <p>$[\text{Cu}_6(\text{trz})_2(\text{bbtz})_2(\text{SiW}_{12}\text{O}_{40})]$</p>	<p>trz bbtz</p>	<p>III</p>	<p><i>Peng and Su et</i> <i>al.,</i> <i>Cryst. Growth</i> <i>Des.</i> 2010, 10, 1104</p>
 <p>$[\text{Cu}_3(2\text{-pytz})_4(4,4'\text{-bpy})][\text{H}_2\text{SiMo}_{12}\text{O}_{40}]$</p>	<p>2-pytz 4,4'-bipy</p>	<p>III</p>	<p><i>Yan et al.,</i> <i>Dalton Trans.,</i> 2013, 42, 1667</p>

 <p>[Cu₃(2-pytz)₂(4,4'-bpy)₄(H₂O)₆] -[H₄SiW₁₂O₄₀]₂·6H₂O</p>			
 <p>Cu₄(pzta)₂(4,4'-bipy)(H₂O)₄(OH)₂ -[Mo₈O₂₆]₂·2H₂O</p>  <p>[Cu₂(pzta)₂(4,4'-bipy)₂(H₂O)₂] -[HPW₁₂O₄₀]₂·8H₂O</p>	<p>pzta 4,4'-bipy</p>	<p>III</p>	<p>This work</p>

The abbreviations in the table: tp = terephthalate, 2,2'-bipy = 2,2'-bipyridine, 4,4'-bipy = 4,4'-bipyridine, nic = nicotinic acid, IN = isonicotinateion, ox = oxalic acid, en = ethylenediamine, pda = 1,2-propanediamine, bpe = 1,2-bis(4-pyridine)-ethane, trz = 1-H-1,2,4-triazole, bbtz = 1,4-bis(1,2,4-triazol-1-ylmethyl)benzene, 2-pytz = 5-(2-pyridyl)tetrazolate, pzta = 5-(2-pyrazinyl) tetrazole.

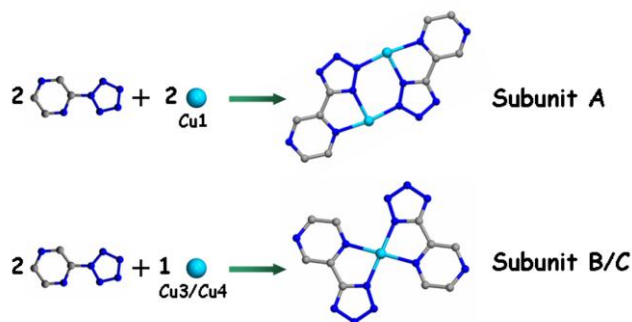


Fig. S1 The two kinds of TMCs in **1**.

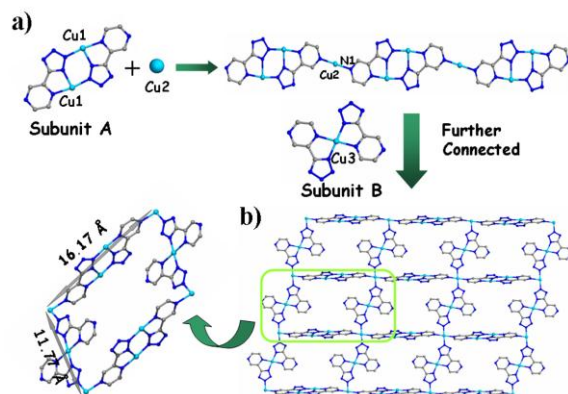


Fig. S2 (a) Ball/stick representation of subunit A, subunit B, infinite organic-inorganic hybrid chain formed by subunit A and Cu2, (b) 2D layers formed by subunit B further connected in **1**.

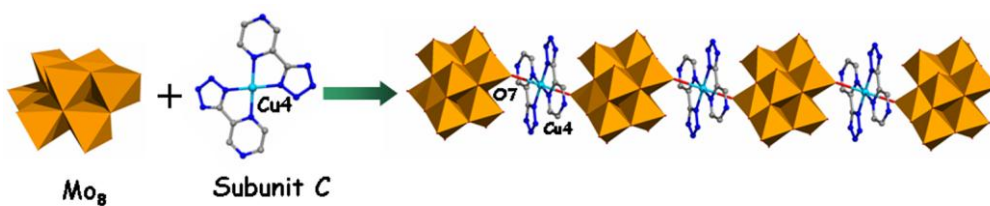


Fig. S3 Ball/stick representation of subunit C, infinite organic-inorganic hybrid chain formed by β -Mo₈ anions and subunit C.

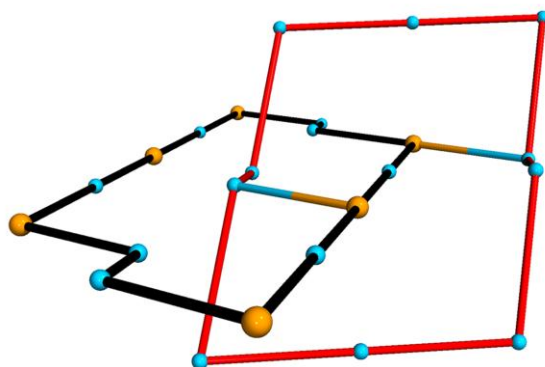


Fig. S4 The detailed view of the self-penetrating structure in **1**.

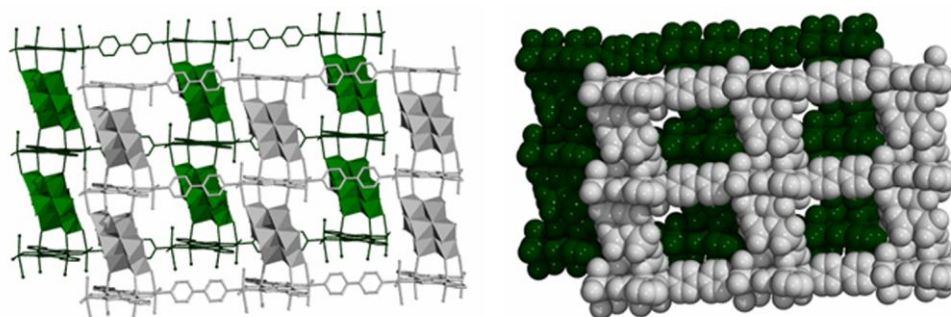


Fig. S5 View of the interdigitated structure in **3**.

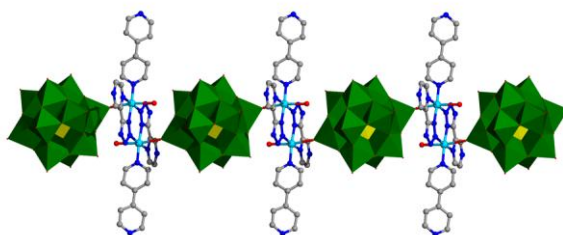


Fig. S6 Combined polyhedral/ball/stick representation of the infinite 1D Poly-pendant chain.

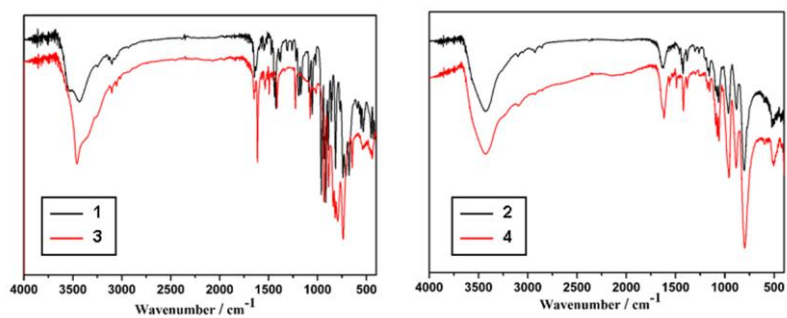


Fig. S7 The IR spectra of compounds **1-4**.

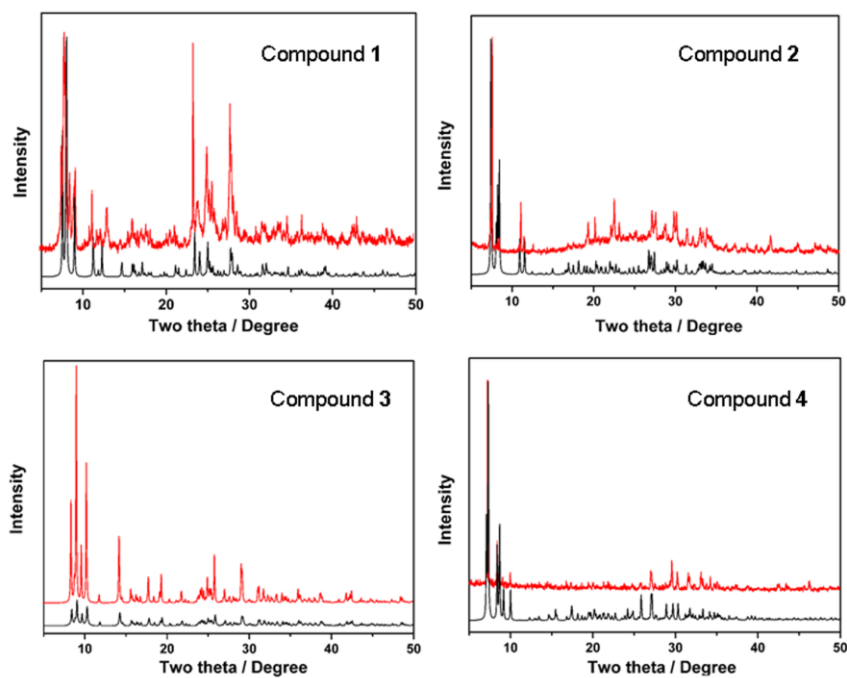


Fig. S8 The simulative (black) and experimental (red) powder X-ray diffraction patterns for **1-4**.

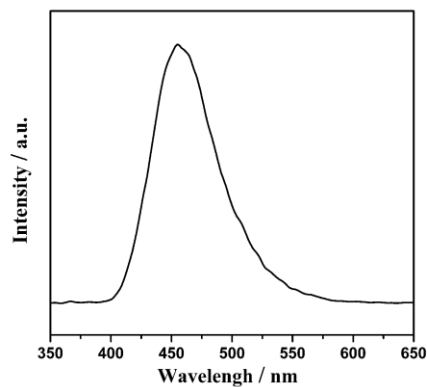


Fig. S9 Solid-state emission spectra of pzta ligand at room temperature.

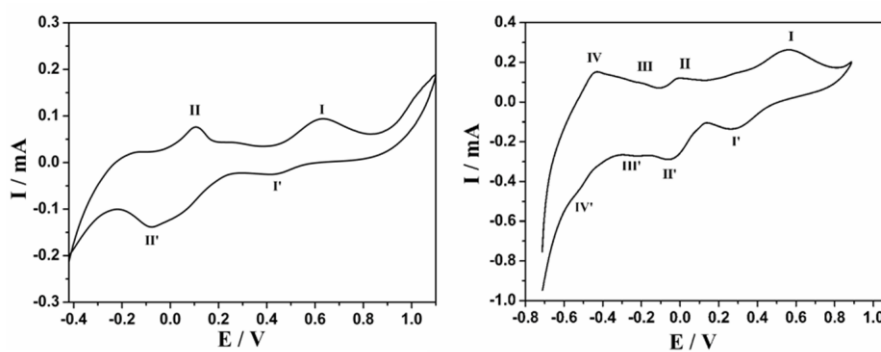


Fig. S10 The CVs of **1**-CPE (left) and **2**-CPE (right) in the 1 M H₂SO₄ solution at the scan rate of 50 mV·s⁻¹.

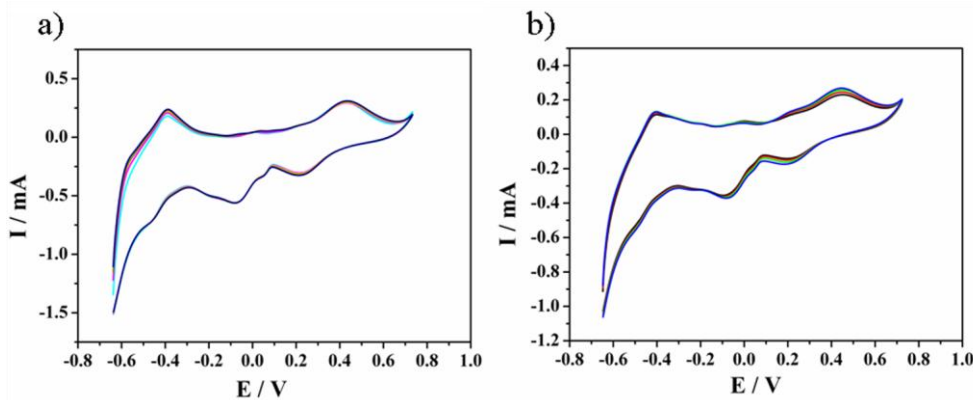


Fig. S11 (a) Reduction of H₂O₂ and (b) oxidation of AA for **2**-CPE in 1M H₂SO₄ solution (scan rate: 50 mV·s⁻¹). The concentrations (from inner to outer) are 0.0, 10, 20, 30 mM for H₂O₂ and 0.0, 0.1, 0.2, 0.3 mM for AA.