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A facile approach to construct organic D- π -A dyes *via* sequential condensation reactions for dye-sensitized solar cells

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General

All chemicals and reagents were used as received from chemical company (Aldrich) without further purification. Column chromatography was performed using with silica gel (200 ~300 mesh) as a stationary phase. Cyclic voltammetry (CV) was performed on a CHI760E Electrochemical Workstation (CHI Instruments Co., Shanghai, China). All CV measurements were carried out in anhydrous CH₂Cl₂ containing 0.1 M TBAHFP as a supporting electrolyte, purging with argon prior to conduct the experiment. Platinum electrode was used as a working electrode, Ag/AgCl in saturated KCl (aq.) as a reference electrode, and a platinum wire as a counter electrode. UV-Vis spectra were measured in CH₂Cl₂ solution or TiO₂ film using UV-3600 Spectrophotometer (SHIMADZU). Fluorescence spectra were measured using FL6500 spectrophotometer (PerkinElmer). The ¹H- and ¹³C-NMR measurements were performed by a DRX spectrometer (Bruker BioSpin). Mass spectra were measured on a Shimadzu Biotech matrix-assisted laser desorption ionization (MALDI) mass

spectrometer. Geometry optimization and Molecular orbital distributions of three dyes were performed using B3LYP functional and 6-31G (d,p) basis set implemented in the Gaussian 09 program package.¹ To estimate the dye-loading amount of each dye in DSSCs, dyes were desorbed from the nanocrystalline TiO₂ film by dipping in 0.1 M TBAOH solution of 1: 1 mixture of H₂O and ethanol. The content of uptake dye was estimated from the absorption peak of each resulting solution.

Cell fabrication and characterization

The device fabrication was performed as follows. TiO₂ paste was received from dyesol Ltd. A double-layered TiO₂ film as photoanode containing 11 μ m main transparent layer with ca. 20 nm sized titania particles and a 5 μ m scattering layer with ca. 400 nm sized titania particles were screen printed on the fluorine-doped tin oxide (FTO) conducting glass substrate (~10 Ohm⁻², Solaronix). The double-layered films were heated to 520 °C and sintered for 1 hour and then cooled to 80 °C, following additional treatment with 0.1 M HCl aqueous solution. The resulted TiO₂ films were washed, dried and then dipped into a 3 × 10⁻⁴ M acetonitrile/n-BuOH (1/1, v/v) solution of the corresponding sensitizer in CH₃CN/n-BuOH (1/1, v/v) for 40 h. Afterwards, the dye-loaded TiO₂ film and a platinum coated conducting glass were assembled into a DSSC of a sandwich type and sealed by heating the Surlyn spacer (40 mm thick). An electrolyte consisting of 0.6 M dimethylpropylimidazolium iodide, 0.05 M I₂, 0.1 M LiI and 0.7 M TBP in acetonitrile was injected into the spacer from the counterelectrode side through a pre-drilled hole, and then the hole was sealed with a Bynel sheet and a thin-glass-slide cover by heating.

The *J-V* characteristics were carried out by using a black metal mask with an aperture area of 0.2304 cm⁻² under standard AM 1.5 sunlight, 100 mW·cm⁻² (WXS-155S-10: Wacom Denso Co. Japan). Monochromatic IPCE spectra were determined with monochromatic incident light of 1×10^{16} photons per cm² under 100 mW·cm⁻² in director current mode (CEP-2000BX, Bunko-Keiki). The IMVS were characterized with a potentiostat (Solartron1287) equipped with a frequency response analyzer (Solartron1255B) at an open-circuit condition based on a monochromatic illumination (420 nm) controlled by a Labview system to obtain the photovoltaic response induced by the modulated light.



Fig. S1 Normalized UV-Vis and emission spectra of dyes LJ-10, LJ-11, LJ-12 and LJ-13 in $\rm CH_2Cl_2.$



Fig. S2 Absorption spectra of dyes LJ-10, LJ-11, LJ-12 and LJ-13 anchored on a transparent TiO_2 films.

DFT Calculation



Fig. S3 Optimized ground-state geometries of dyes LJ-10, LJ-11, LJ-12 and LJ-13.



Fig. S4 The HOMO and LUMO of dyes LJ-10, LJ-11, LJ-12 and LJ-13 calculated at B3LYP/6-31G** level.



Fig. S5 Cyclic voltammograms of ferrocene in CH₂Cl₂/TBAHFP (0.1 M), $[c] = 1 \times 10^{-4} \text{ mol} \cdot \text{L}^{-1}$, 293 K, scan rate = 100 mV·s⁻¹.



Fig. S6 Differential pulse voltammetry of dyes LJ-10, LJ-11, LJ-12 and LJ-13.



Fig. S7 J-V curve (a) and IPCE (b) of DSSC based on dye N719.



Fig. S8 V_{OC} as a function of extracted electron density for DSSCs based on dyes LJ-10, LJ-11, LJ-12 and LJ-13, respectively.





Fig. S9 ¹H NMR spectra of compound 3a.



Fig. S10 ¹³C NMR spectra of compound 3a.



Fig. S11 ¹H NMR spectra of compound 3b.



Fig. S12 ¹³C NMR spectra of compound 3b.

8.94 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 8.13 7.49 7.14 7.15 7.15 7.15 7.15 7.15 7.15 7.15 7.15 7.15 7.15 7.16 7.05 7.05 7.05 7.05



Fig. S13 ¹H NMR spectra of dye **LJ-10**.



Fig. S14 ¹³C NMR spectra of dye LJ-10.



Fig. S15 ¹H NMR spectra of dye LJ-11.



Fig. S16 ¹³C NMR spectra of dye LJ-11.



Fig. S17 ¹H NMR spectra of dye LJ-12.



Fig. S18 ¹³C NMR spectra of dye LJ-12.



Fig. S19 ¹H NMR spectra of dye LJ-13.



Fig. S20 ¹³C NMR spectra of dye LJ-13.

Reference

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