Electronic Supplementary Information

for

Controlling the unpaired electron by electrostatic attraction in the solid state

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Table of Contents

1. Materials and General

All operations were carried out under an atmosphere of dry argon or nitrogen by using modified Schlenk line and glovebox techniques. All solvents were freshly distilled from Na/benzophenone and degassed prior to use. Elemental analyses were performed on an Elementar Vario EL III instrument at Shanghai Institute of Organic Chemistry, the Chinese Academy of Sciences. EPR spectra were obtained using Bruker plus-6/1 X-band variable-temperature apparatus. Commercially available reagents were purchased from Aldrich, Acros or Alfa-Assar, and used as received. Compound 2,7-*t*Bu$_2$-PTO (**1**) was synthesized according to the reported literatures.[1]

Procedure for X-ray diffraction measurement

In glove box under N$_2$ atmosphere, the crystal sample of **1K(cryp)** or **1K(18c6)** was coated with polybutylene. This sample was stable for several weeks in the glove box and suitable for X-ray diffraction measurements. The crystals were each mounted on a glass capillary in perfluorinated oil and measured in a cold N$_2$ flow. The data were collected on a Bruker APEX-II CCD area detector with a radiation source of Ga(Kα) (1.34139 Å) or Mo(Kα) (0.71073 Å) at setting temperatures. Multiscan or empirical absorption corrections (SADABS) were applied. The structures were solved by direct methods and all refined on F^2 with the SHELX-2018/3 software package.[2] The positions of the H atoms were calculated and considered isotropically according to a riding model.

Procedure for EPR measurement

In glove box under N$_2$ atmosphere, the crystal sample of **1K(cryp)** or **1K(18c6)** was dissolved in THF and loaded into sealed paramagnetic tube for EPR measurements.

Procedure for UV/Vis spectral measurement

All the sample solutions were prepared in glove box. The crystal sample of **1K(cryp)** or **1K(18c6)** was dissolved in THF. The prepared sample solution was injected into a sealed cuvette, the absorption was recorded at room temperature.

2. Synthesis

Synthesis of 1K(cryp)

A mixture of **1** (58.3 mg, 0.16 mmol), K (6.6 mg, 0.17 mmol) and cryptand[2.2.2] (64.1 mg, 0.17 mmol) in THF (≈ 20 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to brown green. After filtration, the filtrate was concentrated to ca. 2 ml and stored at -20 °C for 1 day to afford black crystals of **1K(cryp)** (48.5 mg, 39.4%). M.p. 186-189 °C. Elemental analysis (%) Calcd for C$_{42}$H$_{58}$KN$_2$O$_{10}$•C$_4$H$_{10}$O$_2$: C, 62.77; H, 7.79; N, 3.18. Found: C, 62.59; H, 7.55; N, 3.55.

Synthesis of 1K(18c6)

A mixture of **1** (98.5 mg, 0.26 mmol), K (11.1 mg, 0.28 mmol) and 18-c-6 (69.5 mg, 0.26 mmol) in THF (≈ 30 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to green. After filtration, the filtrate was concentrated to ca. 5 ml and stored at 5 °C for 1 day to afford green crystals of **1K(18c6)** (77.1 mg, 43.2%). M.p. 184-186 °C. Elemental analysis (%) Calcd for $C_{36}H_{46}KO_{10} \cdot C_4H_8O$: C, 63.79; H, 6.84. Found: C,63.23; H, 6.98.

Synthesis of [1K(15c5)]₄

A mixture of **1** (81.5 mg, 0.22 mmol), **K** (9.0 mg, 0.23 mmol) and 15-c-5 (71.2 mg, 0.32 mmol) in DME (≈ 30 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to dark green. After filtration, the filtrate was concentrated to ca. 5 mL and stored at 5 °C for 2 day to afford dark red solid of **[1K(15c5)]₄** (25.6 mg, 17.4%). M.p. 163-165 °C. Elemental analysis (%) Calcd for $C_{136}H_{168}K_4O_{36}$: C, 64.43; H, 6.68; Found: C, 64.91; H, 7.32.

Synthesis of (1K)₂ₙ

A mixture of **1** (63.2 mg, 0.17 mmol), **K** (8.2 mg, 0.21 mmol) in DME (≈ 30 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to green. After filtration, the filtrate was concentrated to ca. 5 mL and stored at 5 °C for 1 day to afford black crystals of **(1K)₂ₙ** (20.3 mg, 23.8%). M.p. >300 °C. Elemental analysis (%) Calcd for $C_{28}H_{32}KO_6$: C, 66.77; H, 6.40. Found: C, 65.89; H, 7.63.

3. Reciprocal transformation

From 1K(cryp) to 1K(18c6)

A mixture of **1** (60.1 mg, 0.16 mmol), **K** (6.6 mg, 0.17 mmol) and cryptand[2.2.2] (63.6 mg, 0.17 mmol) in THF (≈ 30 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to brown green. After filtration, KI (36.3 mg, 0.22 mmol) and 18-c-6 (45.3 mg, 0.17 mmol) were added into the filtrate. The color soon changed to from brown green to green. The mixture was allowed to stir overnight and then concentrated to ca. 3 ml and stored at -20 °C for 1 day to afford green crystals of **1K(18c6).** Cell parameters: a = 8.03 Å, b = 22.37 Å, c = 24.55 Å, V = 4408 Å³, Orthorhombic P.

From 1K(18c6) to 1K(cryp)

A mixture of **1** (64.8 mg, 0.17 mmol), **K** (7.1 mg, 0.18 mmol) and 18-c-6 (48.0 mg, 0.18 mmol) in THF (≈ 30 ml) was stirred at room temperature overnight, whereupon the color of the solution was changed from yellow to green. After filtration, cryptand[2.2.2] (68.4 mg, 0.18 mmol) was added into the filtrate. The color changed to from green to brown green 5 min later. The mixture was allowed to stir overnight and then concentrated to ca. 3 mL and stored at -20 °C for 1 day to afford black crystals of **1K(cryp)**. Cell parameters: a = 13.01 Å, b = 16.42 Å, c = 20.60 Å, β = 98.77°, V = 4349 Å³, Monoclinic C.

4. Crystal structures

Table S1. Crystal data and structure refinement. Crystals of **1K(cryp)** and **1K(18c6)** suitable for X-ray analysis were grown from THF. And crystals of **[1K(15c5)]₄** and **(1K)₂ₙ** suitable for X-ray analysis were grown from DME. Two seriously disordered solvent molecules (THF) in **1K(18c6)** have been squeezed.

|	1K(cryp)	1K(18c6)	[1K(15c5)]₄	(1K)₂ₙ
CCDC No.	2061470	2061471	2061472	2061473
Formula	$C_{42}H_{58}KN_2O_{10}$	$C_{36}H_{46}KO_{10}$	$C_{144}H_{188}K_4O_{40}$	$C_{28}H_{32}KO_6$
Formula weight	790.00	677.83	2715.33	503.63
Temp. (K)	193(2) K	296(2)	296(2)	193(2) K
Crystal system	Monoclinic	Orthorhombic	Triclinic	Triclinic
Space group	$C2/c$	$P2_12_12_1$	$P\text{-}1$	$P\text{-}1$
a (Å)	12.8968(4)	8.0229(5)	15.9555(13)	11.7079(6)
b (Å)	16.1529(6)	22.4041(14)	17.9505(14)	11.8131(6)
c (Å)	20.4741(8)	24.5984(16)	27.485(2)	12.0431(10)
α (°)	90	90	88.330(2)	109.702(4)
β (°)	98.9680(10)	90	74.113(2)	98.730(3)
γ (°)	90	90	86.647(3)	119.522(2)
V [Å³]	4213.0(3)	4421.5(5)	7557.7(11)	1257.61(14)
Z	4	4	2	2
ρ_{calcd} (g·cm⁻³)	1.245	1.018	1.193	1.330
μ (mm⁻¹)	0.184	0.164	0.192	1.452
$F(000)$	1692	1444	2896.0	534
Collected data	18968	33362	65907	4067
Unique data	4831 [R(int) = 0.0516]	7772 [R(int) =0.0490]	34259 [R(int) = 0.0369]	4067 [R(int) = 0]
GOF on F^2	0.872	1.025	0.943	1.165
Final R indexes [I > 2σ(I)]	R_1 = 0.0528 ωR_2 = 0.1449	R_1 = 0.0551 ωR_2 = 0.1542	R_1 = 0.0828 ωR_2 = 0.1973	R_1 = 0.0967 ωR_2 = 0.2112
R indexes (all data	R_1 = 0.0643 ωR_2 = 0.1571	R_1 = 0.0664 ωR_2 = 0.1649	R_1 = 0.1421 ωR_2 = 0.2325	R_1 = 0.1373 ωR_2 = 0.2297
Completeness	0.996	0.995	0.984	0.980
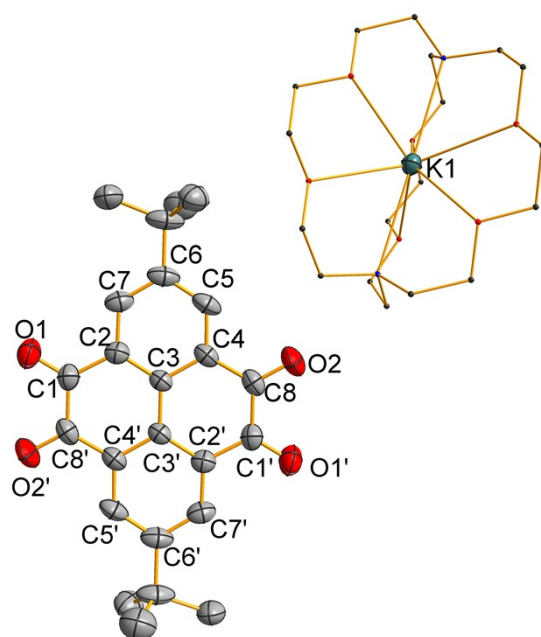
Fig. S1 Thermal ellipsoid drawing of the molecular structure of **1K(cryp)** at 50% probability. Hydrogen atoms and K-cryptand cation are omitted for clarity. Selected bond lengths (Å) and angles (°): O1–C1 1.238(3), O2–C8 1.229(3), C1–C2 1.469(3), C1–C8' 1.490(3), C3–C3' 1.443(3), C2–C7 1.395(3), C2–C3 1.410(3), C3–C4 1.411(3), C4–C5 1.395(3), C4–C8 1.470(3), C5–C6 1.396(4), C6–C7 1.384(3); O1-C1-C2 121.9(2), O1-C1-C8' 119.9(2), O2-C8-C4 121.7(2), O2-C8-C1' 119.8(2).
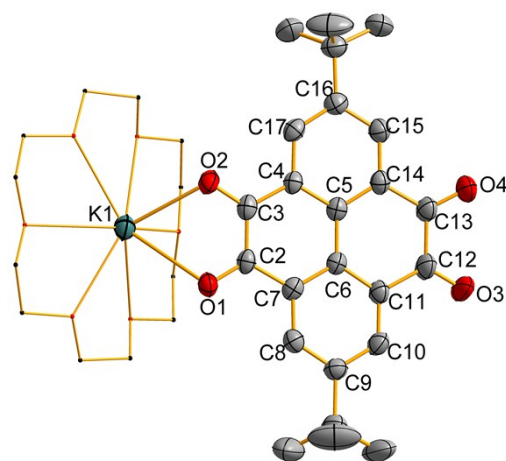
Fig. S2 Thermal ellipsoid drawing of the molecular structures of **1K(18c6)** at 50% probability. Hydrogen atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): C2–O1 1.253(5), C3–O2 1.263(5), C12–O3 1.202(5), C13–O4 1.218(5), C2–C3 1.458(6), C2–C7 1.478(5), C3–C4 1.459(5), C4–C5 1.406(5), C5–C6 1.455(5), C6–C7 1.414(5), C4–C17 1.410(5), C5–C14 1.406(5), C6–C11 1.408(5), C7–C8 1.409(5), C8–C9 1.388(6), C9–C10 1.388(6), C10–C11 1.386(6), C11–C12 1.491(6), C12–C13 1.534(6), C13–C14 1.474(5), C14–C15 1.392(6), C15–C16 1.394(6), C16–C17 1.381(6), K1–O1 2.847(3), K1–O2 2.668(3); O1-C2-C3 121.6(3), O1-C2-C7 120.7(3), C3-C2-C7 117.8(3), O2-C3-C2 120.6(3), O2-C3-C4 120.4(4), O3-C12-C11 123.2(4), O3-C12-C13 118.7(4), O4-C13-C14 123.6(4), O4-C13-C12 117.8(4).
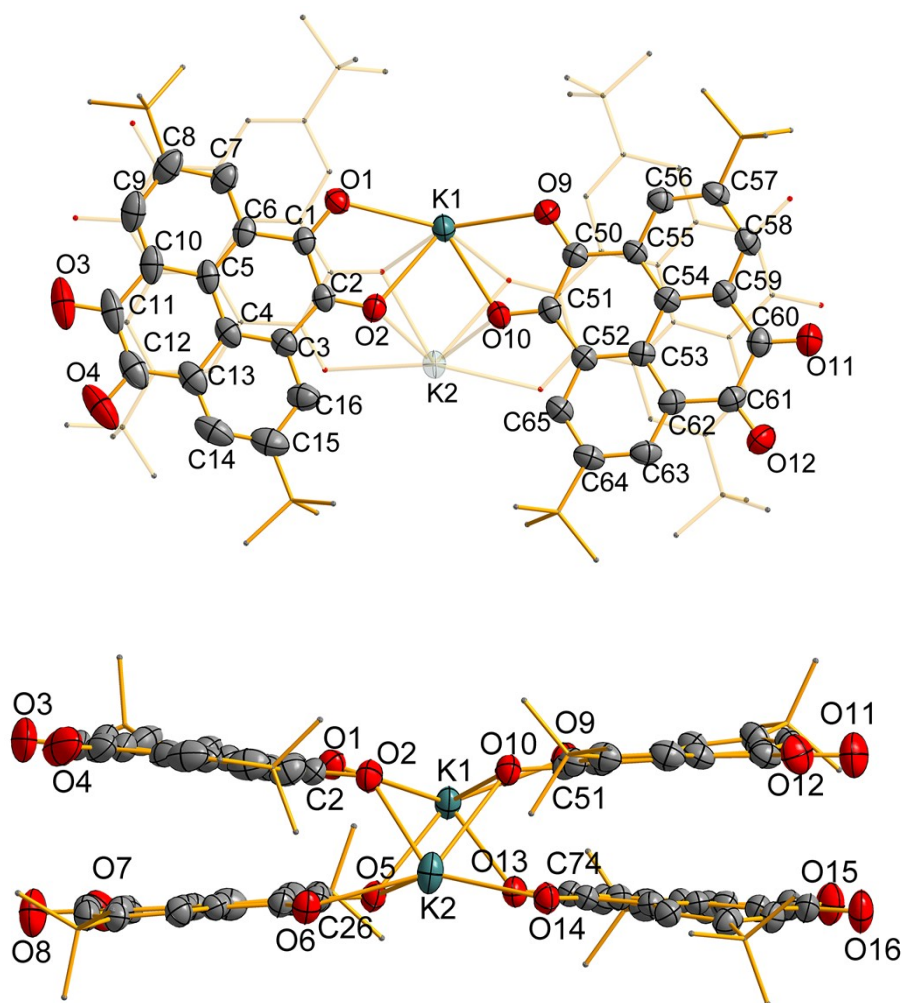
Fig. S3 Thermal ellipsoid drawing of the molecular structures of $(1_4K_2)^{2-}$ dianion in **[1K(15c5)]₄** at 50% probability (up: topview; down: side view). Two $(15c5)_2K^+$ cations, solvent molecules and hydrogen atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): C1–O1 1.256(4), C2–O2 1.260(4), C11–O3 1.219(4), C12–O4 1.223(5), C1–C2 1.457(4), C11–C12 1.515(7), C4–C5 1.438(5), C50–O9 1.255(4), C51–O10 1.256(3), C60–O11 1.218(4), C61–O12 1.220(4), C50–C51 1.464(4), C60–C61 1.534(5), C53–C54 1.449(4), K1–O1 2.760(2), K1–O2 2.670(2), K1–O9 2.749(2), K1–O10 2.692(2), C1–C6 1.480(4), C2–C3 1.462(4), C3–C4 1.391(5), C3–C16 1.404(5), C4–C13 1.428(5), C5–C6 1.409(5), C5–C10 1.409(5), C10–C11 1.482(6), C12–C13 1.451(6); C50–C55 1.479(4), C51–C52 1.460(4), C52–C65 1.395(4) , C52–C53 1.415(4), C53–C62 1.402(4), C54–C55 1.402(4), C54–C59 1.412(4), C59–C60 1.458(5), C61–C62 1.464(5); O1-C1-C2 121.3(3), O1-C1-C6 121.2(3), O2-C2-C1 120.3(3), O2-C2-C3 120.1(3), O9-C50-C51 121.3(3), O9-C50-C55 120.9(3).
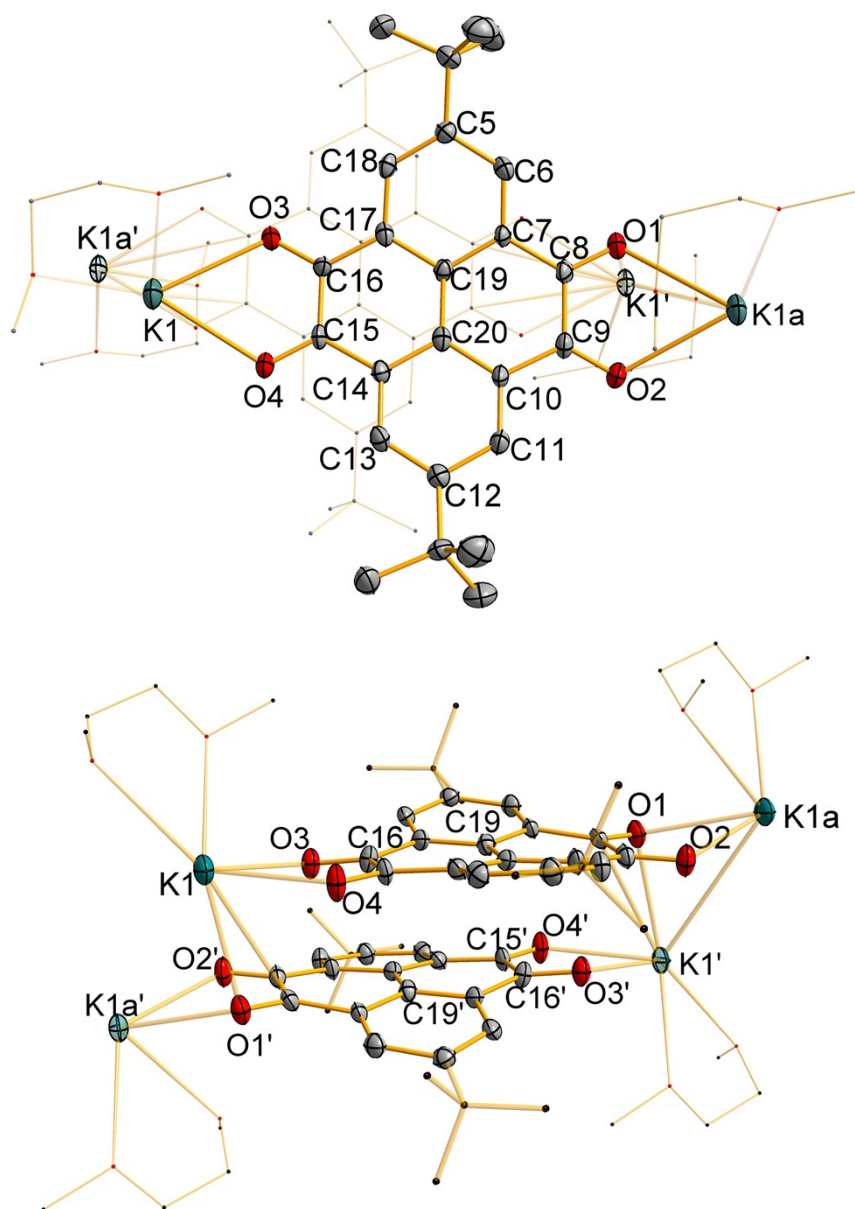
Fig. S4 Thermal ellipsoid drawing of the molecular structures of one unit in **(1K)$_{2n}$** at 50% probability (up: topview; down: side view). Two neighboring K atoms and two DME molecules were added, and hydrogen atoms are omitted for clarity. Selected bond lengths (Å) and angles (°): C8-O1 1.240(6), C9-O2 1.221(6), O3-C16 1.260(6), O4-C15 1.247(6), K1–O3 2.698(4), K1–O4 2.728(4), K1''–O2 2.751(4), K1''–O1 2.788(4), C8-C9 1.531(8), C15-C16 1.480(7), C14-C15 1.468(7), C5–C18 1.375(7), C5–C6 1.407(7), C6–C7 1.384(7), C7–C19 1.412(7), C7–C8 1.473(7), C9–C10 1.480(7), C10–C11 1.382(7), C10–C20 1.414(7), C11–C12 1.419(8), C12–C13 1.398(8), C12–C21 1.521(7), C13–C14 1.400(7), C14–C20 1.410(7), C16–C17 1.450(7), C17–C18 1.415(7), C17–C19 1.423(7), C19–C20 1.433(7); O1-C8-C7 123.5(5),O1-C8-C9 118.2(5), O2-C9-C10 123.5(5), O2-C9-C8 119.0(5), O4-C15-C14 121.6(5), O4-C15-C16 119.6(5), O3-C16-C17 122.8(5), O3-C16-C15 119.2(4).

5. EPR and SQUID measurements
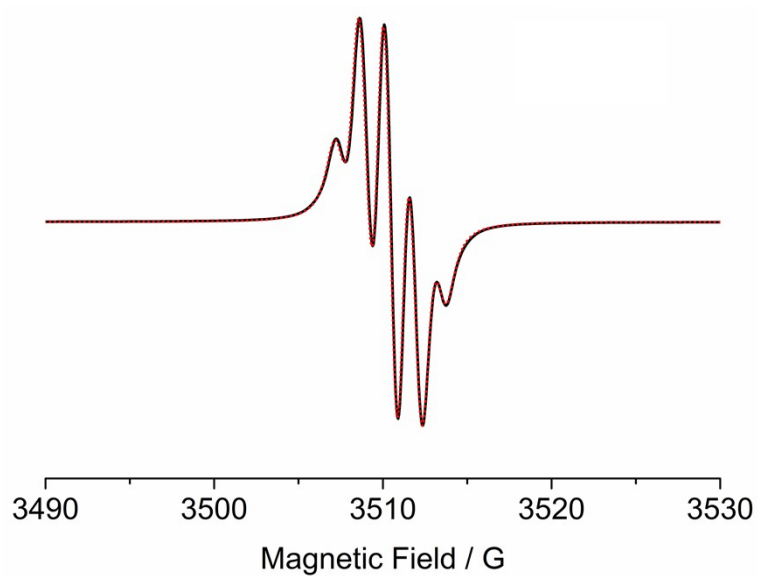
Fig. S5 Experimental (black line) and simulated (red dotted line) EPR spectra of **1K(cryp)** in THF solution at room temperature.
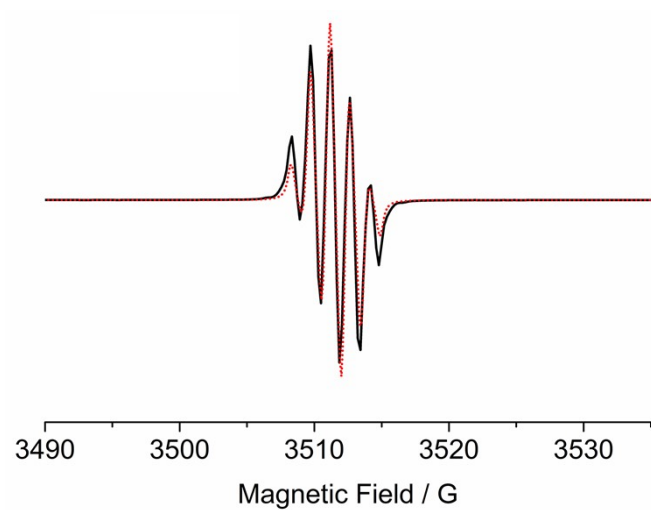
Fig. S6 Experimental (black line) and simulated (red dotted line) EPR spectra of **1K(18c6)** in THF solution at room temperature.
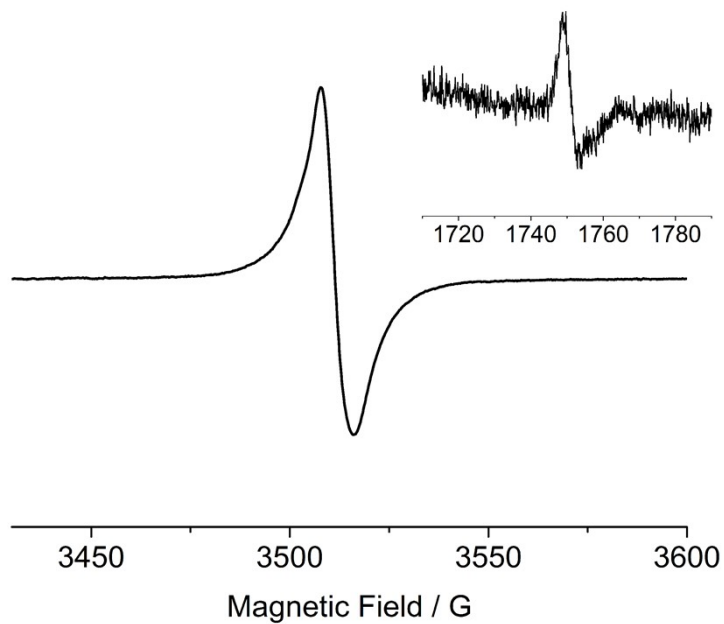
Fig. S7 EPR spectrum of **[1K(15c5)]₄** in power at room temperature with inset showing the half-field signal.
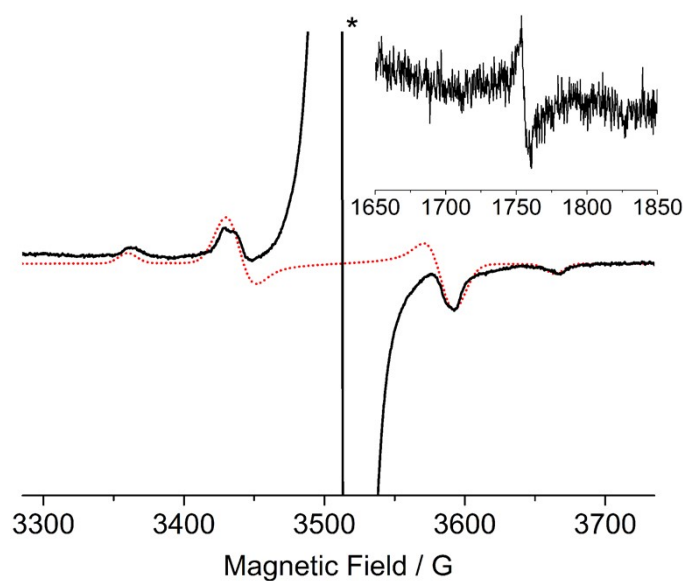
Fig. S8 EPR spectrum (black line) of **(1K)₂ₙ** in powder at room temperature with simulation (red dotted line) with inset showing the half-field signal. The central signals labelled with star belongs to the monoradical.
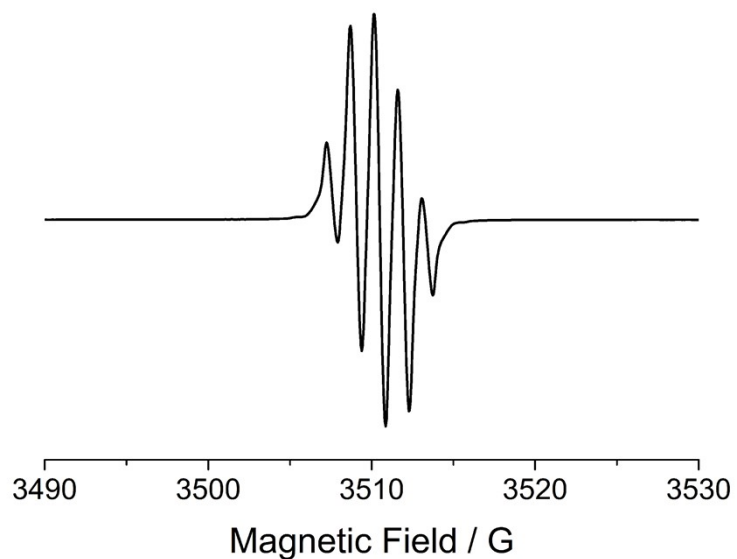
Fig. S9 EPR spectrum of **[1K(15c5)]₄** in THF solution at room temperature.
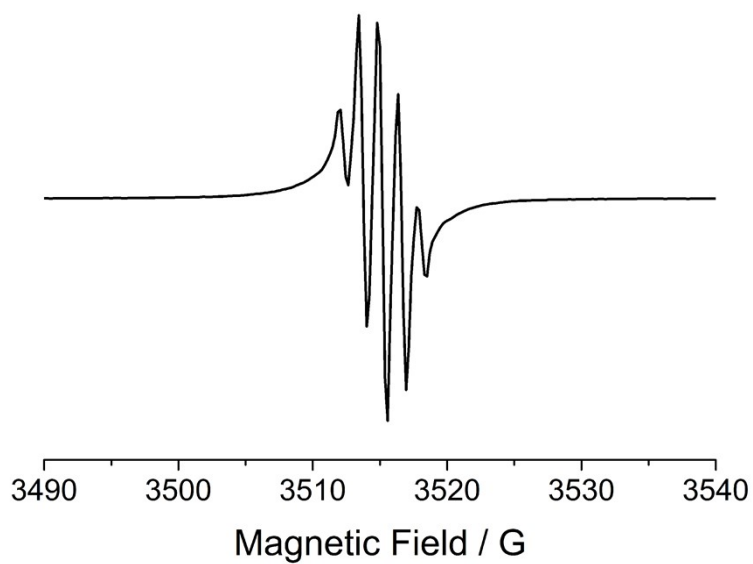
Fig. S10 EPR spectrum of **(1K)₂ₙ** in THF solution at room temperature.

6. UV Spectra

Fig. S13 UV-vis spectra of **[1K(15c5)]₄** in THF solution at room temperature.
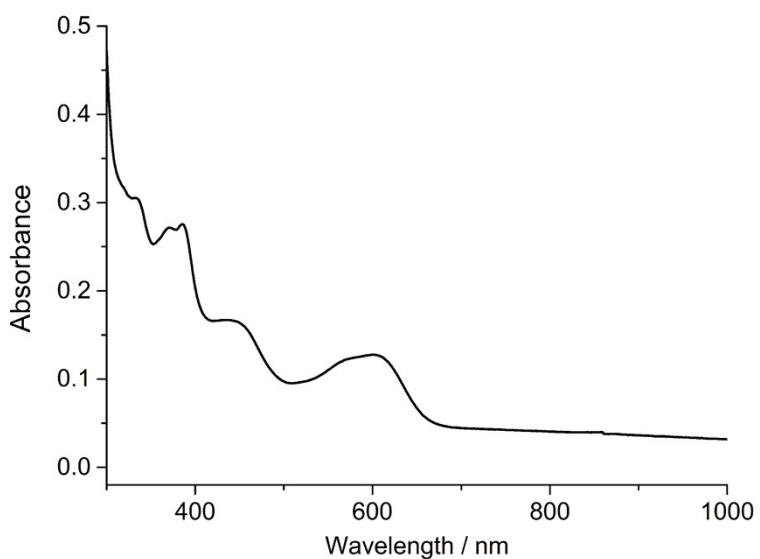
Fig. S14 UV-vis spectra of **(1K)₂n** in THF solution at room temperature.

The EPR spectra of **[1K(15c5)]₄** and **(1K)₂n** in THF solution at ambient temperature (Fig. S9 and S10) both show monoradical signals (g = 2.00491 G, $a(^1H)$ = 1.44 G for **[1K(15c5)]₄**; g = 2.00486, $a(^1H)$ = 1.44 G for **(1K)₂n**), and diradical signals were not observed. Consequently, we believe that both **[1K(15c5)]₄** and **(1K)₂n** depolymerize in THF solution at room temperature.

Then, we speculate that PTO moieties of **(1K)$_{2n}$** remain coordinate to K$^+$ cation in THF solution, just like **1K(18c6)**. So its UV-vis spectrum in THF solution is similar to that of **1K(18c6)**.

While, in the THF solution of **[1K(15c5)]$_4$**, two K atoms are surrounded by coordinated four 15-crown-5, and there are not enough naked K$^+$ cations to be coordinated by all the four PTO moieties. Then part of the PTO moieties exist either in the form of **1$^{\cdot-}$**, or in a coordinated form between **1$^{\cdot-}$** and **1K(18c6)**. That may be the reason why both **1K(cryp)** and **[1K(15c5)]$_4$** exhibit absorption bands of 700-800 nm.

7. Computation details

All calculations were performed with the Gaussian 09 program suite.[3,4] The geometry optimizations of **1K(cryp)** and **1K(18c6)** were carried out at the UB3LYP/6-31G* levels of theory. The obtained stationary points were characterized by frequency calculations. The frontier molecular orbitals and spin densities were calculated at the level of UB3LYP/6-31G* on the optimized geometries and no imaginary frequency was found, which confirmed the local minimum of the optimized structures. The UV-vis absorption spectra were calculated using the time-dependent DFT (TD-DFT) method at UCAM-B3LYP/6-31+G(d) level on **1$^{\cdot-}$** and **1K(18c6)**.

Table S2. Comparison of selected experimental and calculated bond lengths for **1$^{\cdot-}$** in **1K(cryp)**.
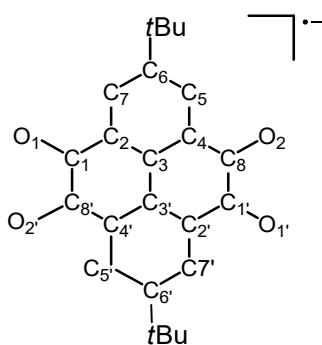
Bond	X-ray	Cal.	Bond	X-ray	Cal.
O1–C1	1.238(3)	1.238	C4–C8	1.470(3)	1.480
O2–C8	1.229(3)	1.238	C3–C3'	1.443(3)	1.429
C1–C8'	1.490(3)	1.521	C4–C5	1.395(3)	1.396
C1–C2	1.469(3)	1.477	C5–C6	1.396(4)	1.405
C2–C3	1.410(3)	1.423	C6–C7	1.384(3)	1.399
C3–C4	1.411(3)	1.425	C2–C7	1.395(3)	1.402

Table S3. Comparison of selected experimental and calculated bond lengths for **1K(18c6)**.
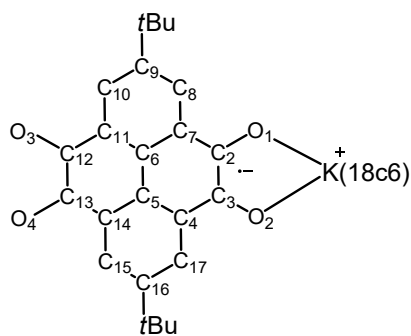
Bond	X-ray	Cal.	Bond	X-ray	Cal.
C2–O1	1.253(5)	1.269	C5–C14	1.406(5)	1.414
C3–O2	1.263(5)	1.269	C6–C11	1.408(5)	1.414
C2–C3	1.458(6)	1.468	C11–C12	1.491(6)	1.481
C12–O3	1.202(5)	1.221	C13–C14	1.474(5)	1.481
C13–O4	1.218(5)	1.221	C4–C17	1.410(5)	1.406
C12–C13	1.534(6)	1.553	K1–O1	2.847(3)	2.691
C2–C7	1.478(5)	1.470	K1–O2	2.668(3)	2.691
C3–C4	1.459(5)	1.470	C7–C8	1.409(5)	1.406
C4–C5	1.406(5)	1.417	C8–C9	1.388(6)	1.397
C5–C6	1.455(5)	1.450	C9–C10	1.388(6)	1.403
C6–C7	1.414(5)	1.417	C10–C11	1.386(6)	1.397

Table S4. Calculated excited wavelengths (λ) and oscillator strengths (f) of selected transitions of **1$^{\cdot-}$** (in **1K(cryp))** and **1K(18c6)**.

Compound	Excited state	Wavelength/nm	f	Transition nature	
1$^{\cdot-}$	S_1	801.9	0.2205	HOMO(α)→LUMO+1(α)	97%
	S_2	727.0	0.0054	HOMO-1(α)→LUMO(α)	74%
				HOMO(β)→LUMO(β)	16%
1K(18c6)	S_1	594.2	0.0211	HOMO-1(α)→LUMO(α)	27%
				HOMO(α)→LUMO+1(α)	15%
				HOMO-2(β) →LUMO(β)	12%
				HOMO(β)→LUMO(β)	22%
	S_2	487.8	0.0478	HOMO(α)→LUMO+1(α)	64%

| | HOMO(β)→LUMO(β) | 20% |
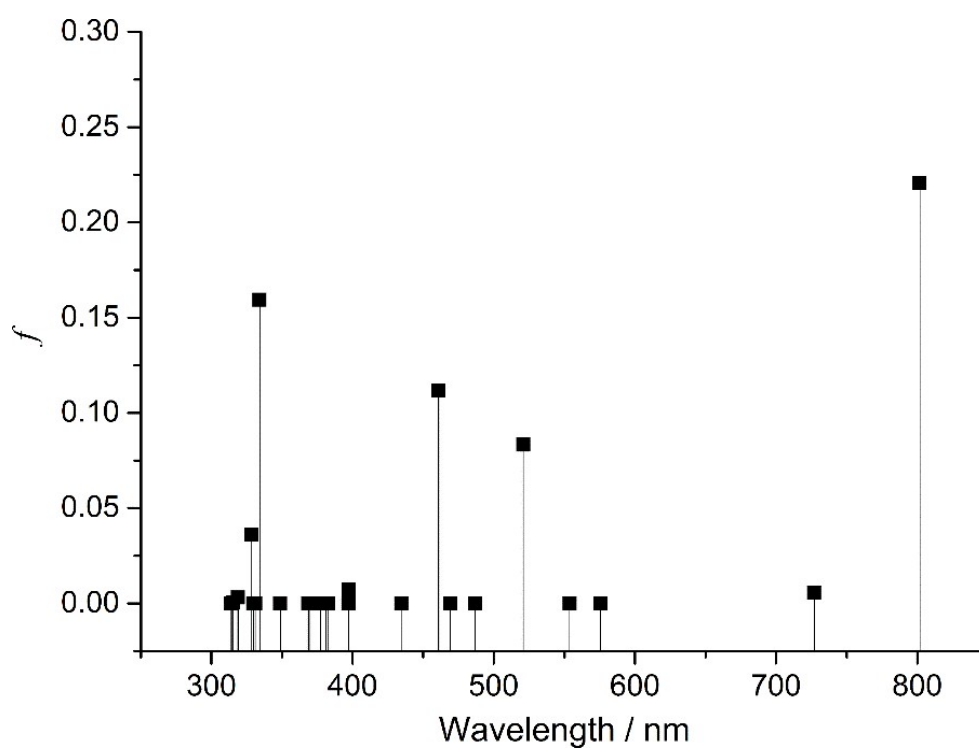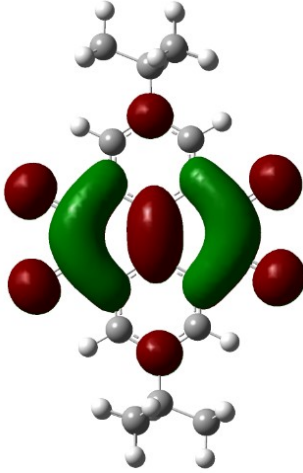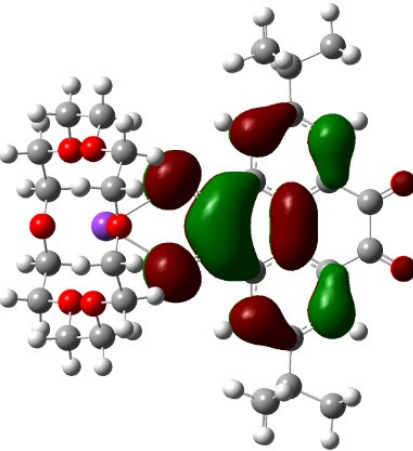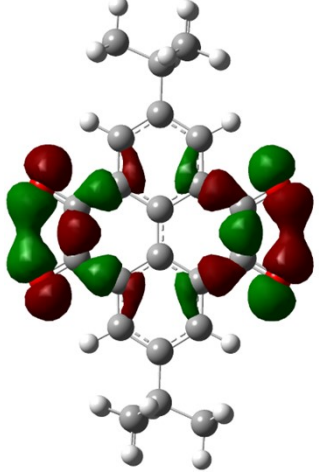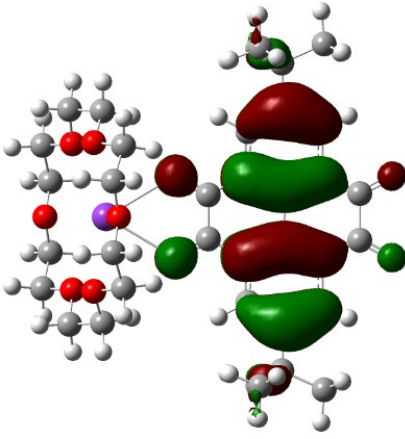
Fig. S15 Calculated UV/vis absorption spectra of **1**$^{\bullet-}$.
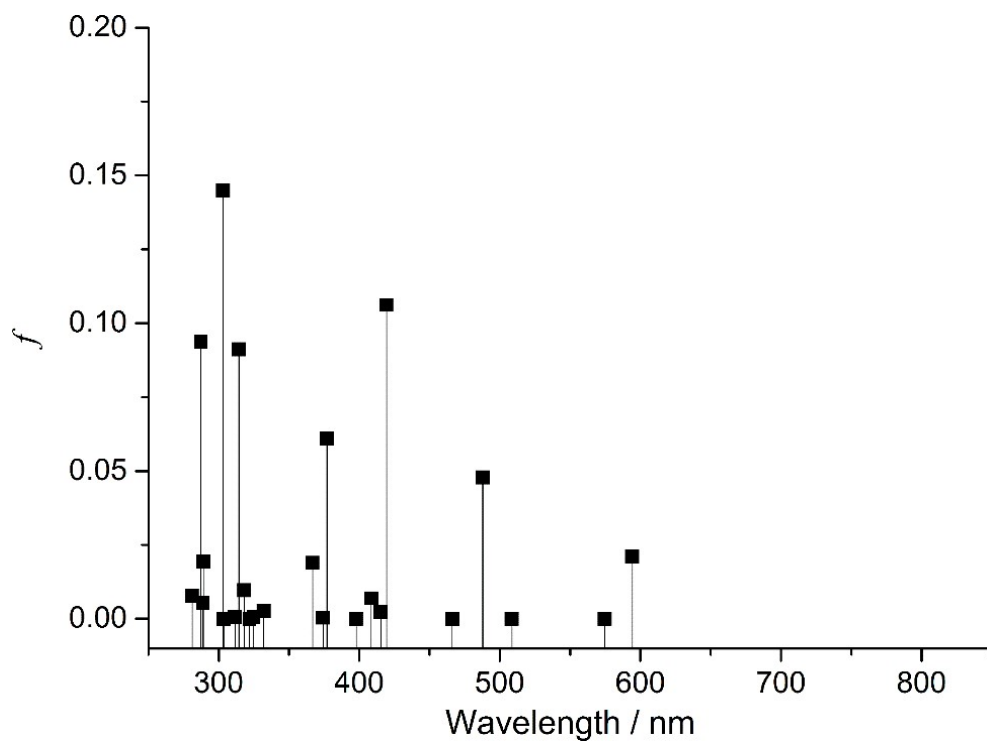
Fig. S16 Calculated UV/vis absorption spectra of **1K(18c6)**.

Table S5. Main frontier molecular orbitals related to UV/vis absorption transitions of **1•−** and **1K(18c6)**.

	1•−	**1K(18c6)**
HOMO(*α*)		
HOMO(*β*)		

HOMO-1(α)		
HOMO-2(β)		
LUMO(α)		

LUMO(β)		
LUMO+1(α)		
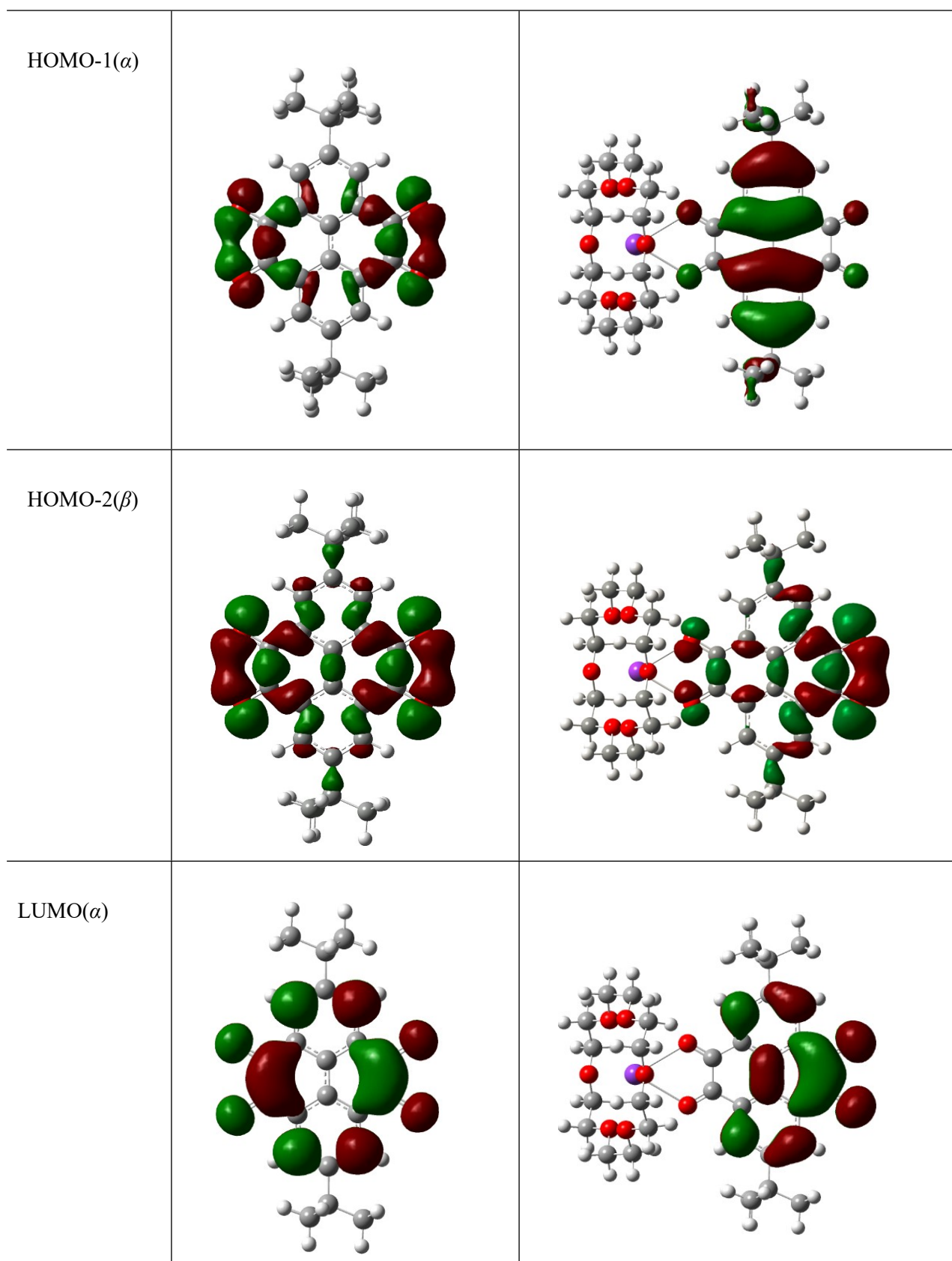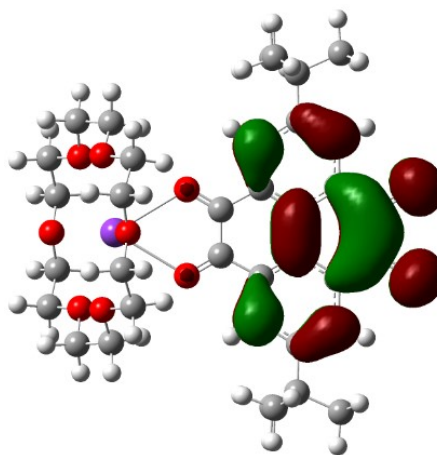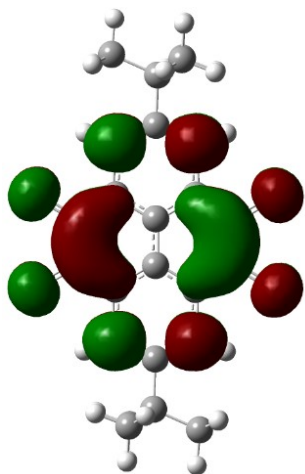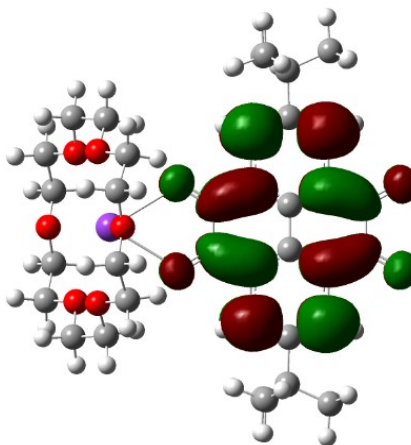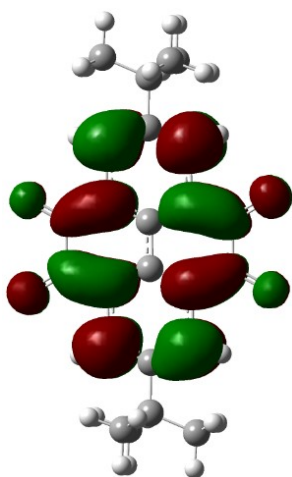
Coordinates of the studied molecules

1·⁻

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Center     Atomic     Atomic          Coordinates (Angstroms)
Number     Number      Type        X          Y          Z
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   1         8         0       -1.417970   3.591586   -0.000302
   2         8         0       -1.343085  -3.621486   -0.000296
   3         6         0       -0.786248   2.527280   -0.000037
   4         6         0       -1.456423   1.208032    0.000028
   5         6         0       -2.851911   1.164300    0.000166
   6         1         0       -3.358766   2.125627    0.000004
   7         6         0       -3.573358  -0.041045    0.000308
   8         6         0       -2.831352  -1.227457    0.000120
   9         1         0       -3.311628  -2.199801   -0.000021
  10         6         0       -1.429888  -1.239061   -0.000007
  11         6         0       -0.714353  -0.008902    0.000006
  12         6         0       -0.734839  -2.542672   -0.000172
  13         6         0       -5.112966  -0.014170    0.000076
  14         6         0       -5.722351  -1.429002    0.000484
  15         1         0       -5.422782  -2.000599    0.886008
  16         1         0       -5.423597  -2.000761   -0.885234
  17         1         0       -6.817322  -1.360545    0.000978
  18         6         0       -5.617503   0.726445   -1.260904
  19         1         0       -5.285051   0.214321   -2.171353
  20         1         0       -5.243254   1.754557   -1.302356
  21         1         0       -6.715121   0.766929   -1.273441
  22         6         0       -5.618098   0.727529    1.260196
  23         1         0       -6.715511   0.773148    1.269221
  24         1         0       -5.239100   1.753781    1.303975
  25         1         0       -5.290972   0.212944    2.171196
  26         8         0        1.417968  -3.591584    0.000258
  27         8         0        1.343083   3.621490    0.000331
  28         6         0        0.786246  -2.527278    0.000265
  29         6         0        1.456422  -1.208030    0.000053
  30         6         0        2.851910  -1.164299   -0.000087
  31         1         0        3.358765  -2.125626    0.000166
  32         6         0        3.573358   0.041045   -0.000314
  33         6         0        2.831353   1.227458   -0.000210
  34         1         0        3.311631   2.199802   -0.000141
  35         6         0        1.429890   1.239064   -0.000099
  36         6         0        0.714354   0.008906   -0.000023
  37         6         0        0.734840   2.542675   -0.000083
  38         6         0        5.112967   0.014168   -0.000064
  39         6         0        5.722355   1.428997   -0.000558
  40         1         0        5.422784   2.000543   -0.886114
  41         1         0        5.423607   2.000810    0.885128
  42         1         0        6.817325   1.360537   -0.001054
  43         6         0        5.617489  -0.726367    1.260969
  44         1         0        5.285025  -0.214186    2.171381
  45         1         0        5.243242  -1.754477    1.302481
  46         1         0        6.715107  -0.766848    1.273520
  47         6         0        5.618109  -0.727615   -1.260131
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1K(18c6)

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Center    Atomic    Atomic          Coordinates (Angstroms)
Number    Number     Type        X          Y          Z
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   5         6        0       2.997865    0.725267   -0.199866
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   9         6        0       2.973308    3.582620   -0.217357
  10         6        0       4.179050    2.865483   -0.248443
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  15         6        0       4.199774   -1.468341   -0.240466
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  17         1        0       5.139017   -3.366614   -0.281806
  18         6        0       2.973552   -3.582333   -0.217237
  19         6        0       1.791642   -2.838365   -0.173294
  20         1        0       0.824605   -3.329988   -0.149442
  21         6        0       2.918165    5.121596   -0.231813
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  24         1        0       1.190158    5.233727    1.127754
  25         1        0       2.152984    6.718892    1.051038
  26         6        0       2.125024    5.595789   -1.472557
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  29         1        0       2.611377    5.265321   -2.397416
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  33         1        0       4.927403    5.481318    0.584461
  34         6        0       2.918501   -5.121320   -0.231745
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  37         1        0       4.927477   -5.480926    0.585224
  38         1        0       4.868667   -5.459768   -1.187105
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  40         1        0       1.103282   -5.203059   -1.475809
  41         1        0       2.065487   -6.690866   -1.493041
  42         1        0       2.612612   -5.265084   -2.397472
  43         6        0       2.210694   -5.623015    1.049068
  44         1        0       1.189965   -5.233727    1.127131
  45         1        0       2.757479   -5.310548    1.946118
  46         1        0       2.153027   -6.718740    1.050748
  47         6        0      -3.043779    3.526649    0.772535
  48         1        0      -2.070122    3.355915    0.293760
  49         1        0      -3.018821    4.494485    1.301162
  50         6        0      -4.137516    3.575988   -0.273109
  51         1        0      -5.123941    3.609152    0.216285
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  53         6        0      -5.033527    2.383509   -2.113046
  54         1        0      -5.004577    3.296225   -2.731103
  55         1        0      -6.036677    2.308188   -1.663263
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59	6	0	-4.776506	-1.183184	-3.002128
60	1	0	-5.482819	-1.203462	-3.848310
61	1	0	-3.753909	-1.233221	-3.408463
62	6	0	-5.033362	-2.382804	-2.114256
63	1	0	-6.036578	-2.308131	-1.664505
64	1	0	-5.003978	-3.295225	-2.732732
65	6	0	-4.136777	-3.576048	-0.275143
66	1	0	-4.020469	-4.492486	-0.876547
67	1	0	-5.123187	-3.610121	0.214201
68	6	0	-2.343226	-2.372144	2.708987
69	1	0	-1.350969	-2.229448	2.260039
70	1	0	-2.330980	-3.290196	3.319989
71	6	0	-2.678730	-1.192063	3.597894
72	1	0	-3.715587	-1.272302	3.962758
73	1	0	-2.006589	-1.200216	4.471681
74	6	0	-2.678602	1.190044	3.598404
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76	1	0	-3.715394	1.270135	3.963482
77	6	0	-2.343252	2.370496	2.709933
78	1	0	-2.330872	3.288282	3.321331
79	1	0	-1.351079	2.227991	2.260729
80	6	0	-3.043081	-3.526990	0.770554
81	1	0	-3.017713	-4.495236	1.298424
82	1	0	-2.069481	-3.355464	0.291952
83	19	0	-2.916346	-0.000134	0.057302
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86	8	0	6.578269	1.364931	-0.308419
87	8	0	6.578371	-1.364406	-0.307856
88	8	0	-3.327472	2.475025	1.687673
89	8	0	-4.039636	2.431290	-1.107675
90	8	0	-4.950466	0.000385	-2.239262
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92	8	0	-3.327255	-2.476204	1.686500
93	8	0	-2.503016	-0.000859	2.852319

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