

Dimethyltin(IV)-4,6-dimethyl-2-pyridylselenolate: An Efficient Single Source Precursor for the Preparation of SnSe Nanosheets as Anode Material for Lithium Ion Batteries

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Figure captions

Figure S1. ⁷⁷Se{¹H} NMR spectra of bis(4,6-dimethyl-2-pyridyl)diselenide

Figure S2. ⁷⁷Se{¹H} NMR spectra of [Me₂Sn{2-SeC₅H₂(Me-4,6)₂N₂}₂] (**1**)

Figure S3. XRD pattern of TG residue of [Me₂Sn{2-SeC₅H₂(Me-4,6)₂N₂}₂] (**1**) obtained at 800°C .

Figure S4. Deconvolution of pXRD peaks (2θ; 30-31.6°) of tin selenide material synthesized in DDT

Figure S5. EDS spectra of tin selenide nanostructures obtained by thermolysis of [Me₂Sn{2-SeC₅H₂(Me-4,6)₂N₂}₂] (**1**) in (a) DDT, (b) OAm+DDT (1:1, v/v) respectively.

Figure S6. EDS spectra of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

Figure S7. 2-D Elemental mapping of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

Figure S8. SAED pattern of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

Figure S9. (a) pXRD pattern, (b) FT-IR spectra, (c) SEM micrograph and (d) EDS spectra of electrode after 100 cycles.

Figure S10. (a) Cycling, coulombic efficiency and normalized capacity of full cell LiNi_{0.6}Mn_{0.2}Co_{0.2}O₂ // SnSe at current density 100 mA g⁻¹ and (b) Cycling and coulombic efficiency of LiNi_{0.6}Mn_{0.2}Co_{0.2}O₂ at current density 100 mA g⁻¹.

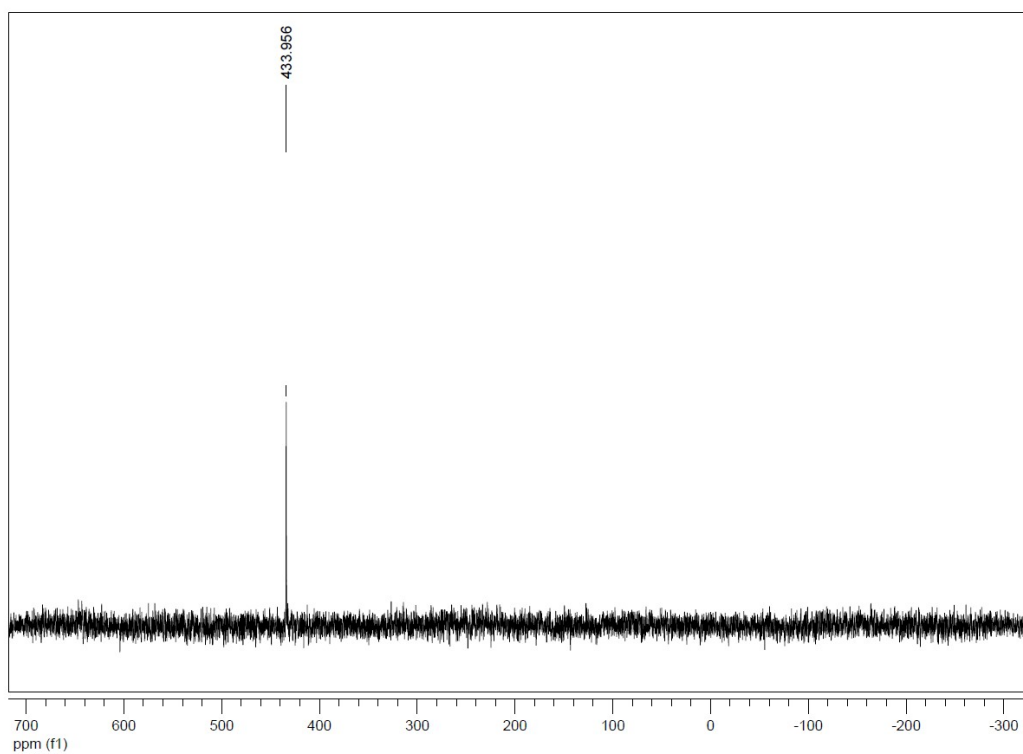


Figure S1. $^{77}\text{Se}\{^1\text{H}\}$ NMR spectra of bis(4,6-dimethyl-2-pyridyl)diselenide

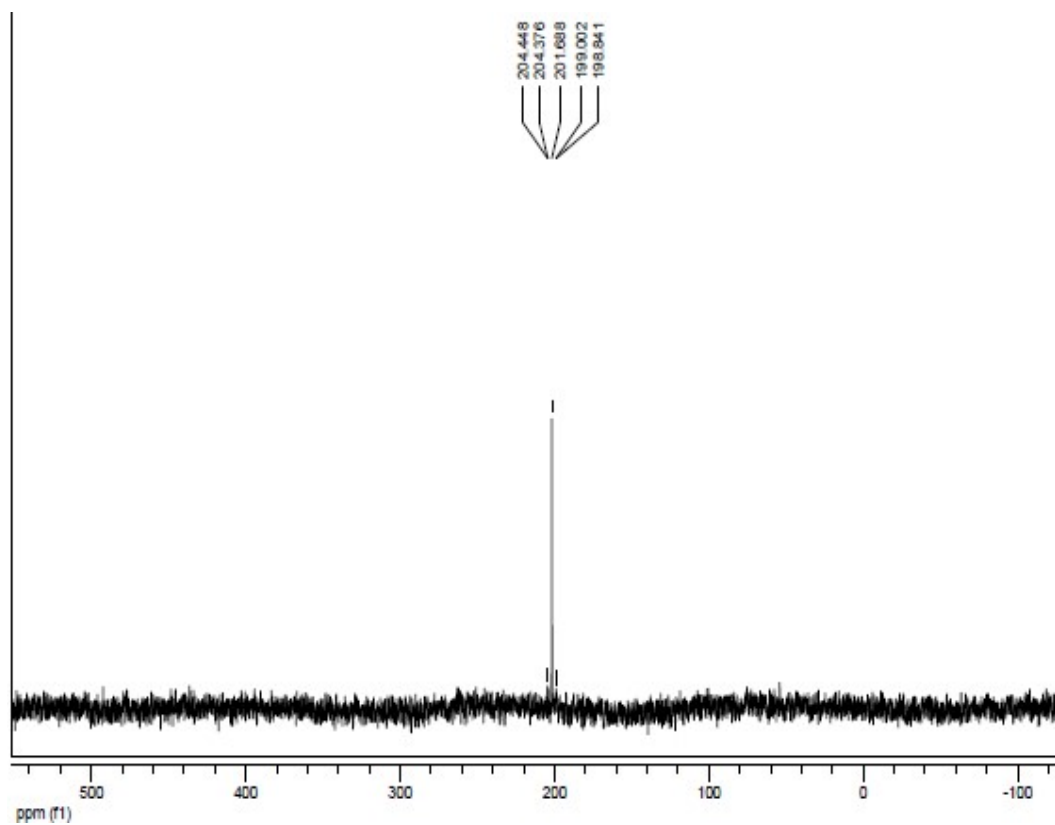


Figure S2. $^{77}\text{Se}\{^1\text{H}\}$ NMR spectra of $[\text{Me}_2\text{Sn}\{2\text{-SeC}_5\text{H}_2(\text{Me-4,6})_2\text{N}\}_2]$ (**1**)

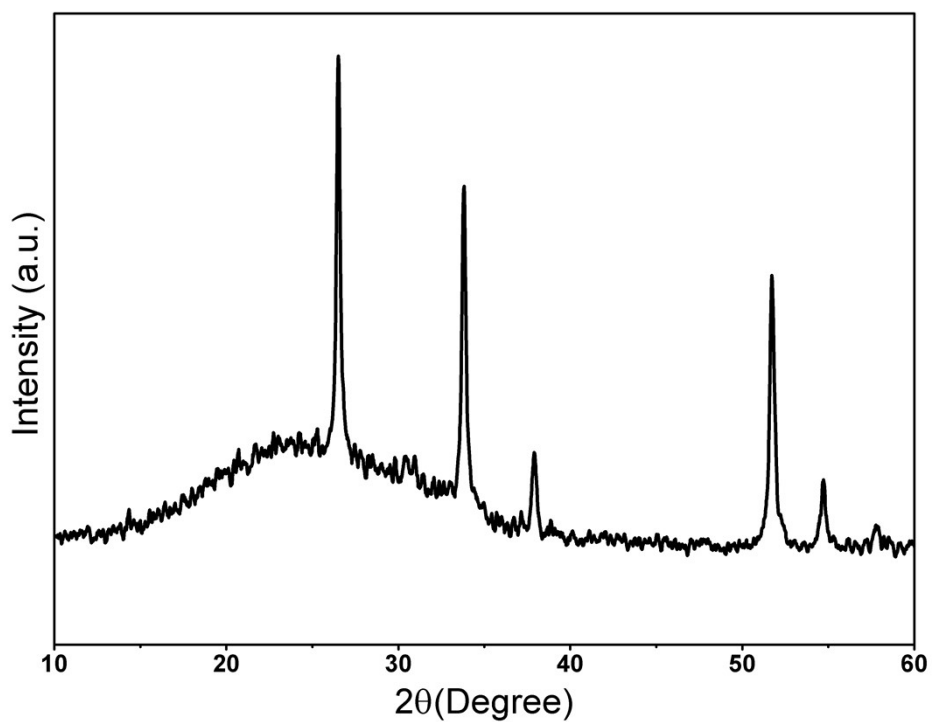


Figure S3. XRD pattern of TG residue of $[\text{Me}_2\text{Sn}\{2\text{-SeC}_5\text{H}_2(\text{Me-4,6})_2\text{N}\}_2]$ (**1**) obtained at 800°C.

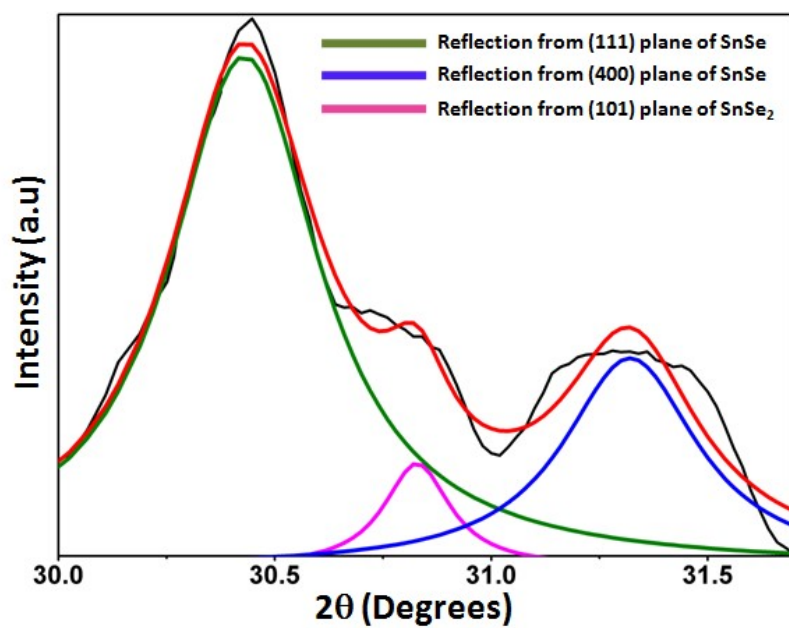


Figure S4. Deconvolution of pXRD peaks (2θ ; 30-31.6°) of tin selenide material synthesized in DDT

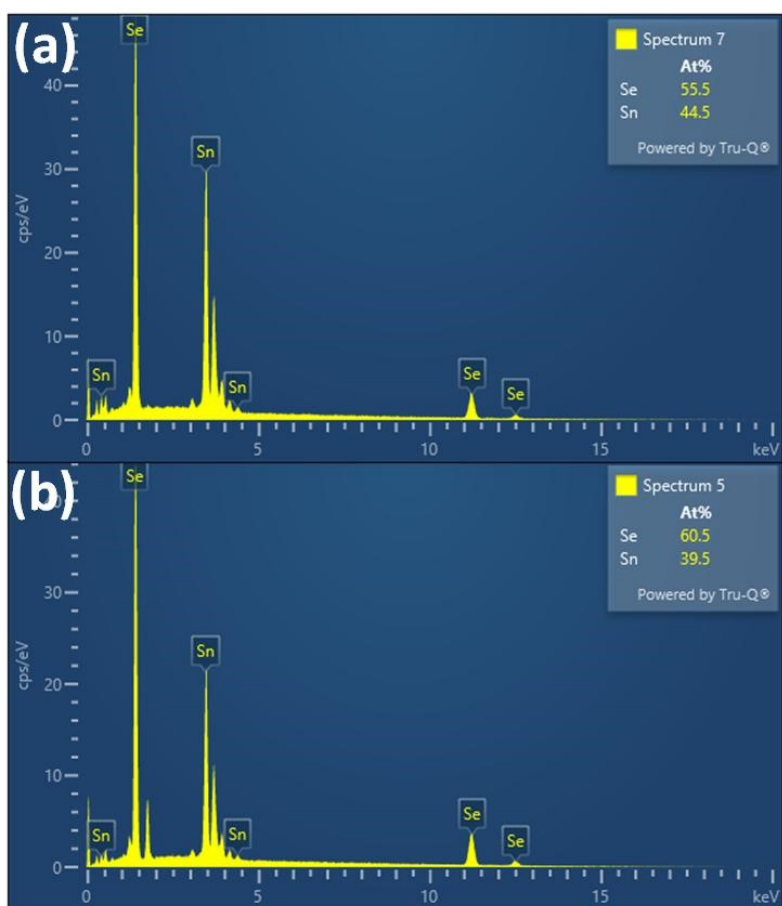


Figure S5. EDS spectra of tin selenide nanostructures obtained by thermolysis of $[\text{Me}_2\text{Sn}\{2\text{-SeC}_5\text{H}_2(\text{Me-4,6})_2\text{N}\}_2]$ (**1**) in (a) DDT, (b) OAm+DDT (1:1, v/v) respectively.

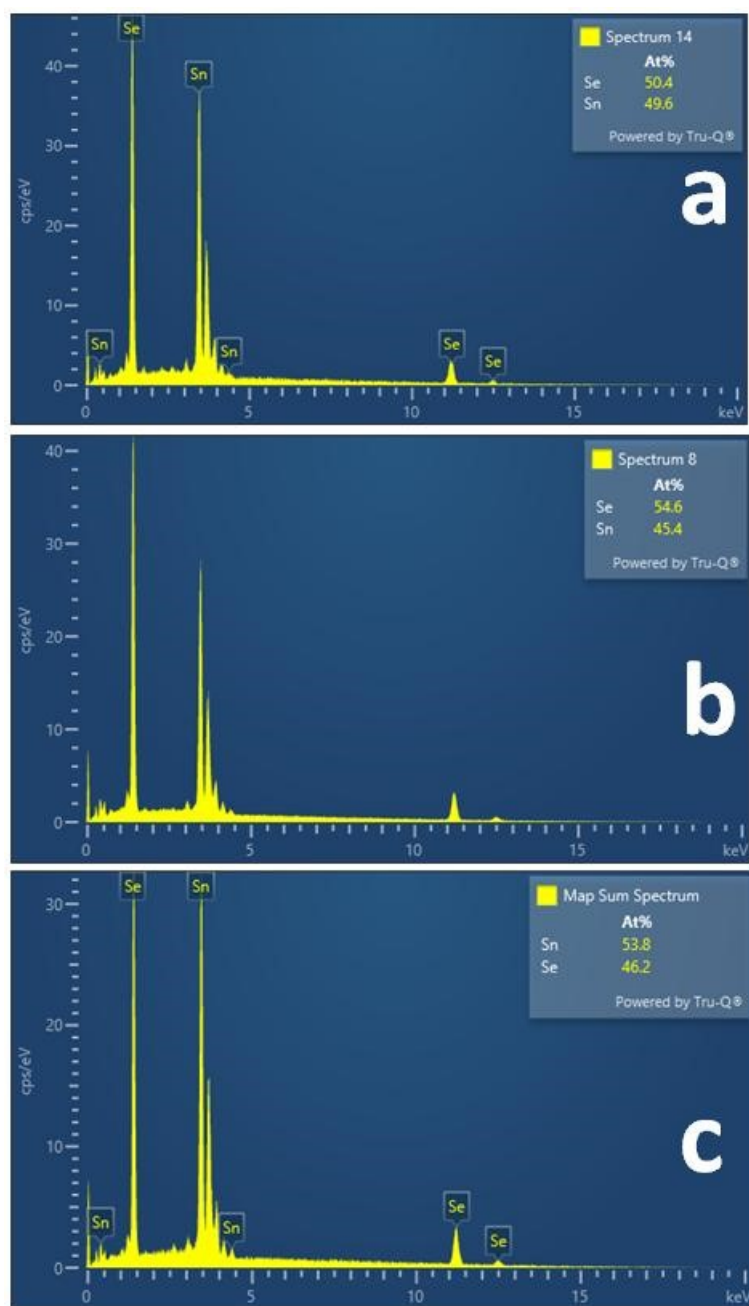


Figure S6. EDS spectra of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

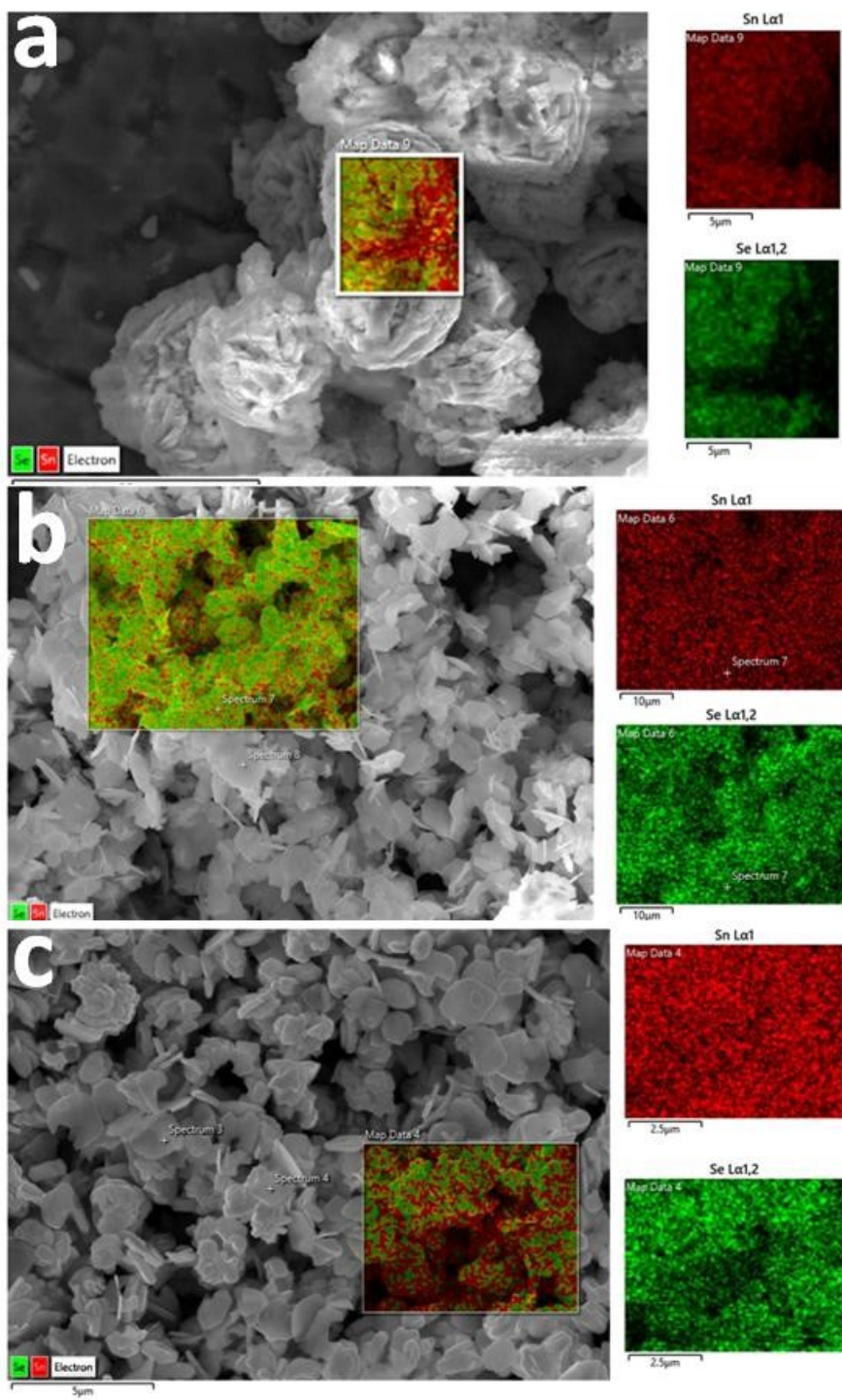


Figure S7. 2-D Elemental mapping of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

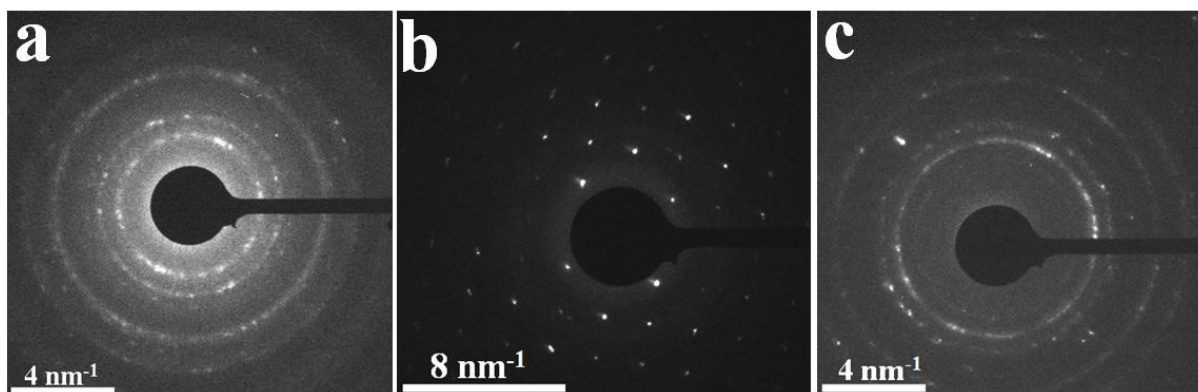


Figure S8. SAED pattern of SnSe nanostructures obtained by (a) pyrolysis of **1** and thermolysis of **1** in (b) OAm and (c) ODE respectively.

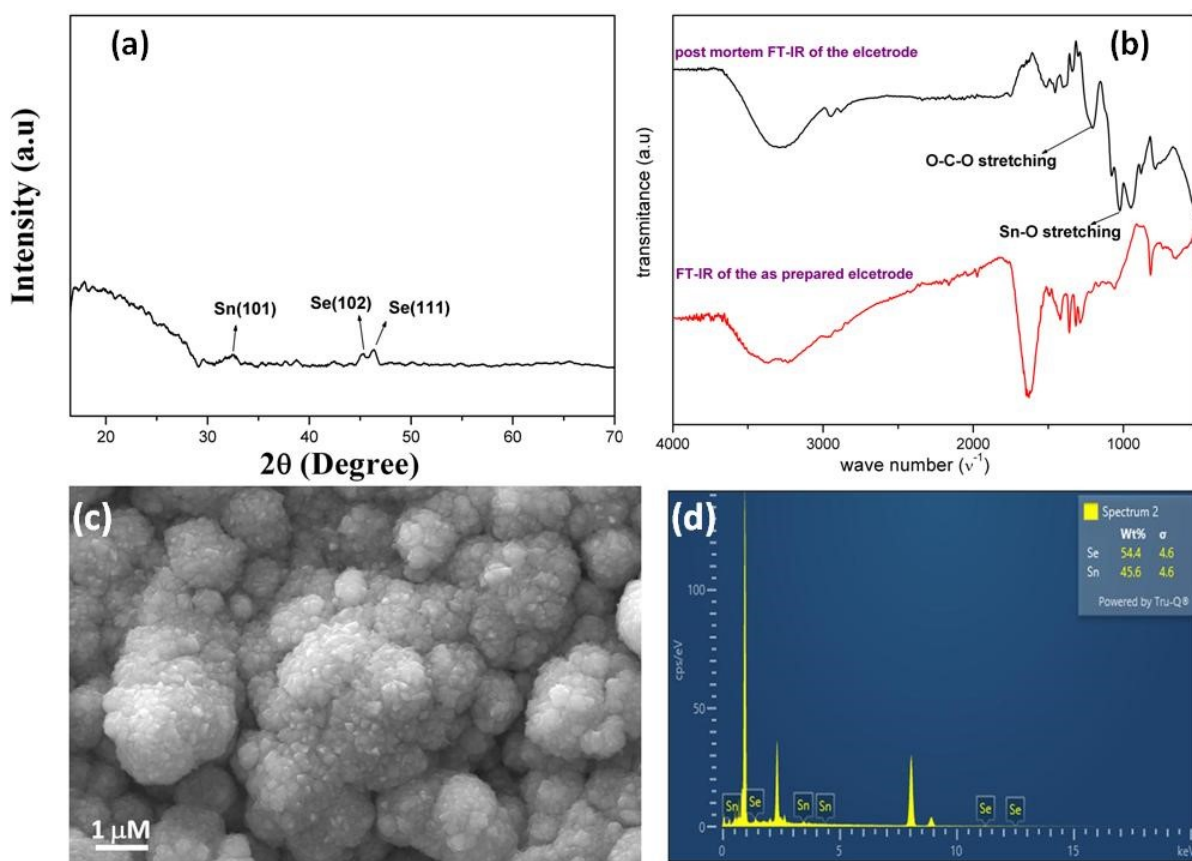


Figure S9. (a) pXRD pattern, (b) FT-IR spectra, (c) SEM micrograph and (d) EDS spectra of electrode after 100 cycles.

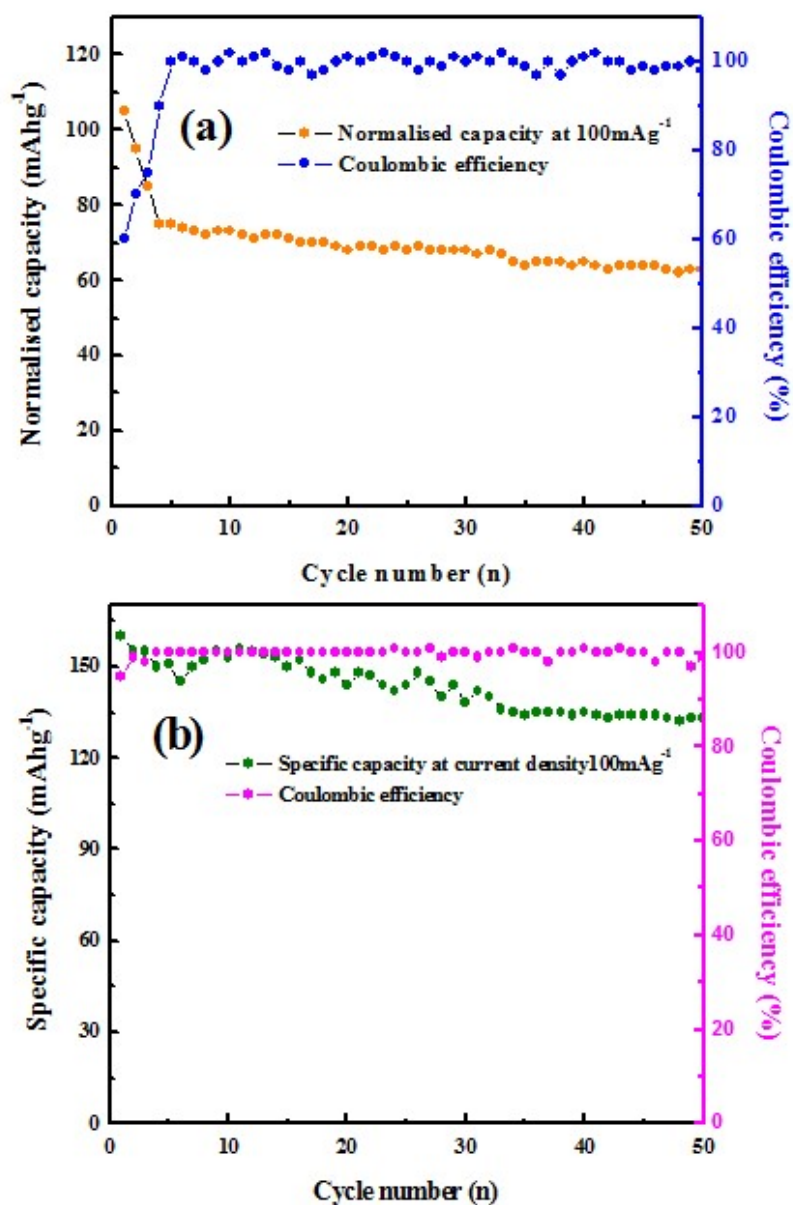


Figure S10. (a) Cycling, coulombic efficiency and normalized capacity of full cell $\text{LiNi}_{0.6}\text{Mn}_{0.2}\text{Co}_{0.2}\text{O}_2 // \text{SnSe}$ at current density 100 mA g^{-1} and (b) Cycling and coulombic efficiency of $\text{LiNi}_{0.6}\text{Mn}_{0.2}\text{Co}_{0.2}\text{O}_2$ at current density 100 mA g^{-1} .