

Electronic Supplementary Information

**Electrocatalytic synthesis of α,α -gem-dihalide ketones from
 α -mono-halide ketones and unexpected dimer condensation**

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1. General experiment

1.1 Experimental details

Electro-catalysis experiments were performed in an argon atmosphere. Chemical reagents were obtained from commercial suppliers (Sigma-Aldrich, Adamas, Shanghai Titan Technology Co., LTD, et al.) and used without further purification. Purified water was purchased from the local market. Reactions were monitored by thin-layer chromatography (TLC) on 0.25 mm silica gel GF254 plate (Merck Silica gel 60-F254) using UV light as a visualizing agent and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. Silica gel (300-400 mesh) was used for flash column chromatography. Yields refer to chromatographically and spectroscopically homogeneous materials unless otherwise noted. Switching DC voltage regulator for electrolysis is a triple display potentiostat (KORAD-KA3305D, made in China). The anodic electrode was carbon plate (10 mm×10 mm×2 mm), and the cathodic electrode was platinum plate (10 mm×10 mm×0.2 mm).

1.2 Characterizations

¹H-NMR, **¹³C-NMR**, and **¹⁹F-NMR** were recorded with Bruker Advance (300, 400, 500 & 600 MHz) spectrometers. All chemical shifts were reported as δ values in parts per million (ppm) and coupling constants (J) in Hz. Tetramethylsilane (TMS) was used as the internal standard for CDCl₃ (7.26 ppm for ¹H, 77.00 ppm for ¹³C), respectively.

IR spectra were recorded on a Bruker Bio-Rad FTS-135 spectrometer with KBr pellets.

Mass spectra: (1) **HR-MS (ESI)** was taken on AB QSTAR Pulsar mass spectrometer or Agilent LC/MSD TOF mass spectrometer. HR-MS data were recorded via electron impact mass spectrometry using a time-of-flight analyzer. (2) **GC-MS** was performed on a Hewlett-Packard 6890 N gas chromatograph (equipped with the same HP-5MS capillary column) under identical operating conditions used in Mass Spectral Library by R. P. Adams (Adams, 2007).

Electrochemistry (CVs): cyclic voltammograms were obtained on a Metrohm PGSTAT302N potentiostat.

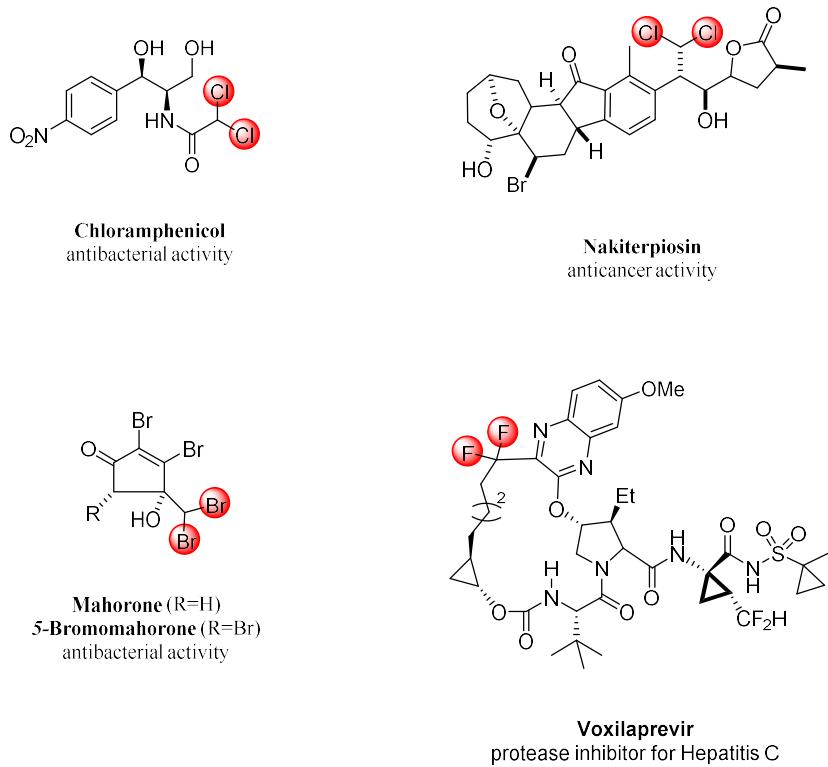
1.3 Electrocatalysis apparatus

Switching DC voltage regulator for electrolysis is a triple display potentiostat (KORAD-KA3305D, made in China). In an oven-dried undivided two-necked (or three-necked) flask (50 mL) equipped with a stir bar. The flask was equipped with a platinum plate (10 mm×10 mm×0.2 mm) as the cathodic electrode and a carbon plate (10 mm×10 mm×2 mm) as the anodic electrode. The reaction was in an argon atmosphere. The reaction was conducted at room temperature.



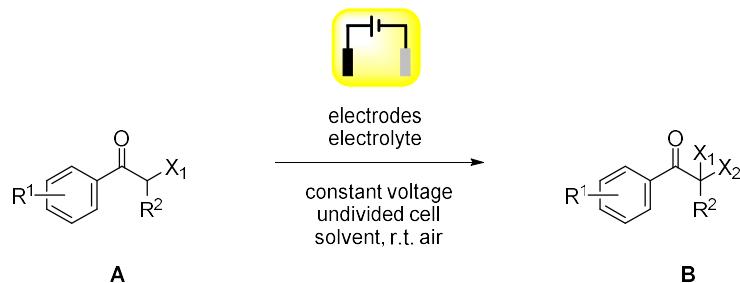
Figure S1. Electrochemical reaction cell apparatus and electrodes

1.4 Bio-active α,α -gem-dihalide natural products



Scheme S1. α,α -Gem-dihalide monoterpenes among marine bio-active natural products.

2. Reaction conditions optimization

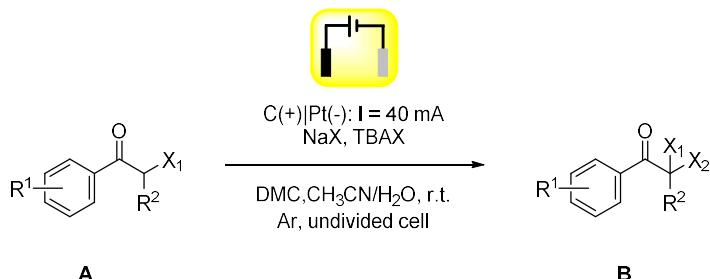


Entry	Variations of standard conditions	Yield (%) ^a
1	none	82
2	no CH_3CN	13
3	toluene, THF, or dioxane replace CH_3CN	13-25
4	H_2O (0.5 or 2.0 $\text{mol}\cdot\text{L}^{-1}$)	35-35
5	no DMC	57
6	without TBAC	37
7	" Bu_4NBF_4 or " Bu_4NPF_6 replace TBAC	35-38
8	TBAF	16
9	Nickel as cathode	18
10	5 mA or 100 mA current	12-35
11	Ar or O_2 atmosphere	81-80
12	6h reaction time	20
13	without electric current	0
14	K_2CO_3 (0.2 $\text{mol}\cdot\text{L}^{-1}$)	12
15	AcOH (0.2 $\text{mol}\cdot\text{L}^{-1}$)	11
16	NaOH (1.0 $\text{mol}\cdot\text{L}^{-1}$)	dimer ^b

Table S1. Typical electro halogenation. Standard reaction conditions: glassy carbon plate (10 mm × 10 mm × 2 mm) as anode, Pt plate (10 mm × 10 mm × 0.2 mm) as cathode, constant current = 40 mA, reactant **A** (200 mg, 1.0 mmol), TBAC (55.6 mg, 0.2 mmol), NaCl (725 mg, 1.0 $\text{mol}\cdot\text{L}^{-1}$), CH_3CN (5.0 mL), H_2O (5.0 mL), DMC (2.5 mL), room temperature, ambient atmosphere, 30 min, undivided cell. ^a Isolated yield by silica gel column chromatography, the same for other yields. ^b A dimer condensation product.

3. General electrocatalysis reactions

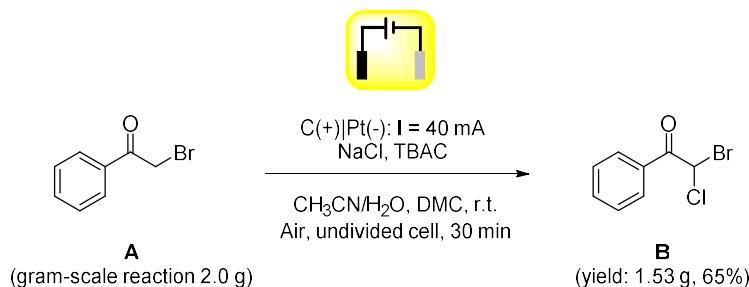
3.1 Typical electro-halogenation



Scheme S2. Electrocatalytic halogenation

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halide acetophenone (200 mg, 1.0 mmol), tetrabutylammonium halides (TBAC or TBAB, 50 mg, 0.2 mmol), brine (e.g., NaCl, KCl, NaBr, KBr, or NaF, 1.0 mol·L⁻¹, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were added. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathode and a carbon plate (10 mm × 10 mm × 2 mm) as the anode. The reaction was conducted under an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA at room temperature for 30 min. The reaction was monitored by TLC, using UV light as a visualizing agent and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. When the material no longer decreased, the solution was extracted with ethyl acetate (EtOAc) (3 × 10 mL) and H₂O (3 × 10 mL). The combined organic layer was dried with Na₂SO₄ and filtered. The solvent was removed with a rotary evaporator. The pure product was obtained by flash chromatography on silica gel using EtOAc and petroleum ether as the eluent. The products were further identified by FTIR spectroscopy, NMR spectroscopy, and HRMS analysis.

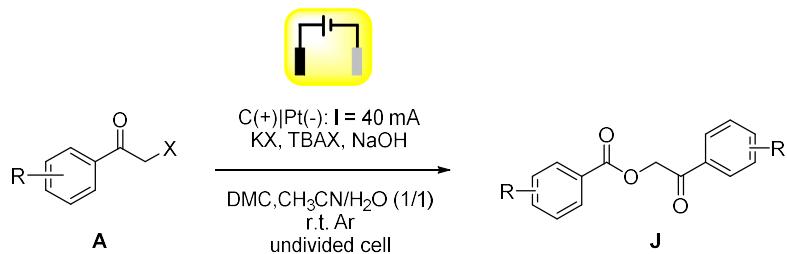
3.2 Gram-scale electro-halogenation



Scheme S3. Gram-scale electro-halogenation

In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 2-bromo-acetophenone (2.0 g, 1.0 mmol), TBAC (564.3 mg, 0.2 mmol), NaCl (3.8 g, 6.5 mmol), CH₃CN (50.0 mL), H₂O (50.0 mL), and DMC (25.0 mL) were combined and added. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathode and a carbon plate (10 mm × 10 mm × 2 mm) as the anode. The reaction was conducted under an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA at room temperature for 30 min. The reaction was monitored by TLC, using UV light as a visualizing agent and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. When the material no longer decreased, the solution was extracted with ethyl acetate (EtOAc) (3 × 10 mL) and H₂O (3 × 10 mL). The combined organic layer was dried with Na₂SO₄ and filtered. The solvent was removed with a rotary evaporator. The pure product was obtained by flash chromatography on silica gel using EtOAc and petroleum ether as the eluent. Product α -bromo- α -chloro-acetophenone was obtained with a yield of 1.53 g, 65%.

3.3 Electrocatalytic dimer condensation



Scheme S4. Electrocatalysis for dimer condensation

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (200 mg, 1.0 mmol), tetrabutylammonium halide (TBAC or TBAB, ~50 mg, 0.2 mmol), brine (e.g., NaCl, KCl, NaBr, KBr, or NaF, 1.0 mol·L⁻¹, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were added. NaOH (500 mg, 1.0 mol·L⁻¹) was added to the reaction solutions. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathode and a carbon plate (10 mm × 10 mm × 2.0 mm) as the anode. The reaction was conducted under an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA at room temperature for 30 min. The reaction was monitored by TLC, using UV light as visualizing agent and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. When the material no longer decreased, the solution was extracted with EtOAc (3 × 10 mL) and H₂O (3 × 10 mL). The combined organic layer was dried with Na₂SO₄ and filtered. The solvent was removed with a rotary evaporator. The pure product was obtained by flash chromatography on silica gel using EtOAc and petroleum ether as the eluent. The products were further identified by FTIR spectroscopy, NMR spectroscopy, and HRMS analysis.

3.4 X-ray of the dimer condensation product

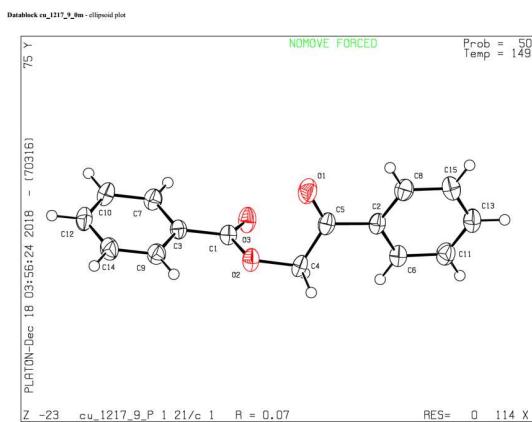
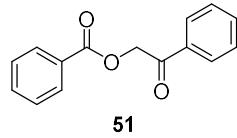


Figure S2. X-ray of dimer condensation product **51** (CCDC No. 2106444)

Table 1. Crystal data and structure refinement for C:_1217_9.

Identification code	cu_1217_9_0m	
Empirical formula	C15 H12 O3	
Formula weight	240.25	
Temperature	149.44 K	
Wavelength	1.54178 Å	
Crystal system	Monoclinic	
Space group	P 1 21/c 1	
Unit cell dimensions	a = 8.9535(3) Å	α= 90°.
	b = 14.0575(5) Å	β= 90.799(2)°.
	c = 9.4542(3) Å	γ = 90°.
Volume	1189.83(7) Å ³	
Z	4	
Density (calculated)	1.341 Mg/m ³	
Absorption coefficient	0.763 mm ⁻¹	
F(000)	504	
Crystal size	? x ? x ? mm ³	
Theta range for data collection	4.940 to 72.084°.	
Index ranges	-11<=h<=11, -17<=k<=16, -11<=l<=11	
Reflections collected	19949	
Independent reflections	2335 [R(int) = 0.1052]	
Completeness to theta = 67.679°	99.8 %	

Absorption correction	Semi-empirical from equivalents
Max. and min. transmission	0.7538 and 0.3884
Refinement method	Full-matrix least-squares on F^2
Data / restraints / parameters	2335 / 0 / 164
Goodness-of-fit on F^2	1.097
Final R indices [I>2sigma(I)]	R1 = 0.0669, wR2 = 0.1616
R indices (all data)	R1 = 0.0826, wR2 = 0.1745
Extinction coefficient	0.0119(18)
Largest diff. peak and hole	0.316 and -0.333 e. \AA^{-3}

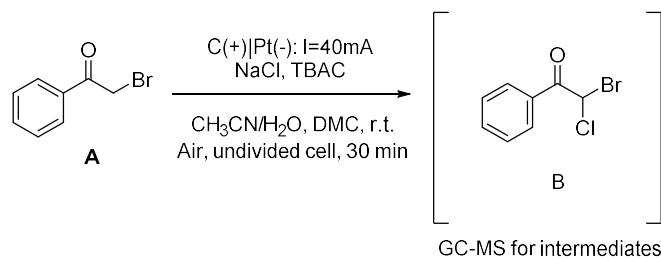
Table 2. Atomic coordinates ($\times 10^4$) and equivalent isotropic displacement parameters ($\text{\AA}^2 \times 10^3$) for C:_1217_9. U(eq) is defined as one-third of the trace of the orthogonalized U^{ij} tensor.

	x	y	z	U(eq)
O(1)	8939(2)	4904(1)	7079(2)	48(1)
O(2)	7566(2)	3929(1)	5032(2)	35(1)
O(3)	5606(2)	4143(1)	6457(2)	38(1)
C(1)	6470(2)	3614(2)	5884(2)	27(1)
C(2)	8489(2)	6454(2)	6204(2)	28(1)
C(3)	6425(2)	2564(2)	6000(2)	27(1)
C(4)	7644(2)	4951(2)	4893(2)	31(1)
C(5)	8393(2)	5399(2)	6151(2)	31(1)
C(6)	7624(2)	7036(2)	5318(2)	32(1)
C(7)	5478(2)	2164(2)	6991(2)	32(1)
C(8)	9473(2)	6866(2)	7176(2)	35(1)
C(9)	7266(3)	1976(2)	5133(2)	34(1)
C(10)	5376(3)	1183(2)	7111(3)	37(1)
C(11)	7748(3)	8020(2)	5424(3)	38(1)
C(12)	6198(3)	602(2)	6236(3)	37(1)
C(13)	8745(3)	8423(2)	6382(3)	37(1)
C(14)	7147(3)	996(2)	5247(3)	38(1)
C(15)	9599(3)	7845(2)	7265(3)	37(1)

3.5 Control experiments

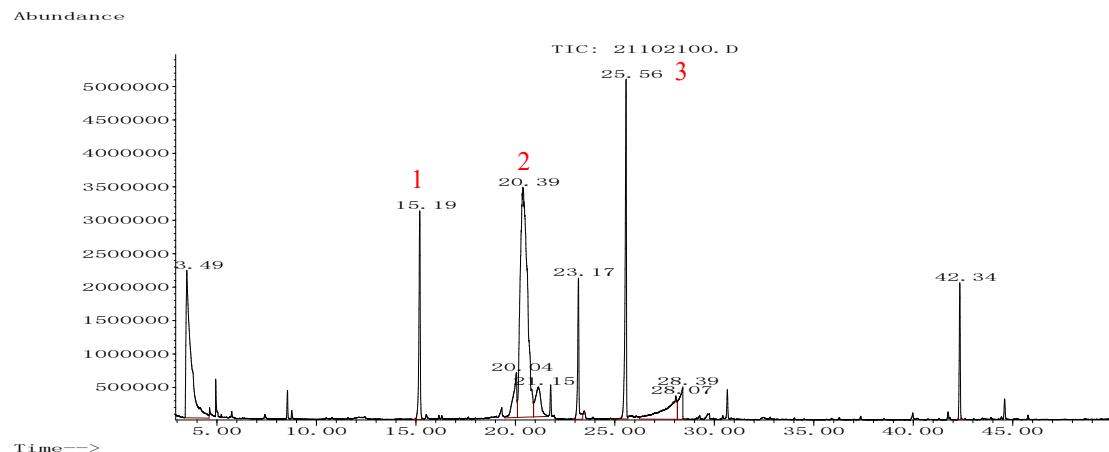
GC-MS analysis of menthyl and isomenthyl acetates was also performed on a Hewlett-Packard 6890 N gas chromatograph (equipped with the same HP-5MS capillary column) under identical operating conditions used in Mass Spectral Library by R. P. Adams (Adams, 2007). In this experiment, the GC was operated under the following conditions: injector and interface temperatures were 220 and 240 °C, respectively; oven temperature was raised from 60 °C to 246 °C at a heating rate of 3 °C/min; the flow of helium carrier gas was 1.0 mL/min.

3.5.1 GC-MS for electrocatalytic halogenation

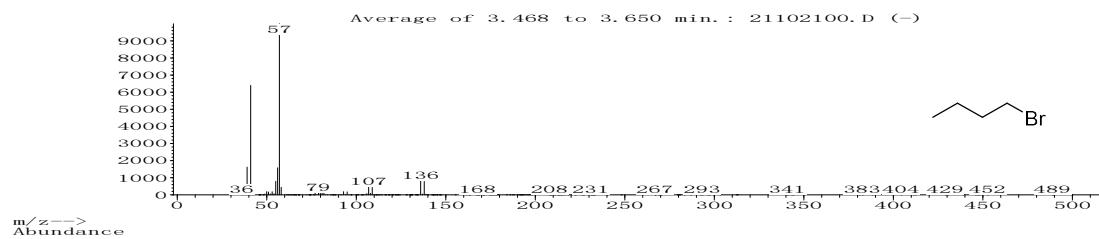


Scheme S5. Electrocatalysis for halogenation

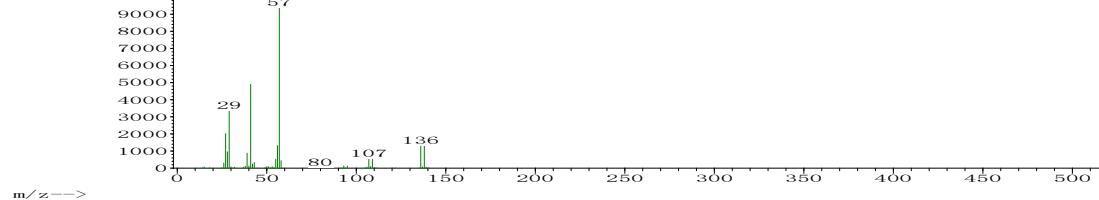
In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 2-bromo-acetophenone (200 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaCl (380 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (50 mm×50 mm×0.2 mm) as the cathodic electrode and a carbon plate (50 mm×50 mm×2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 30 min.



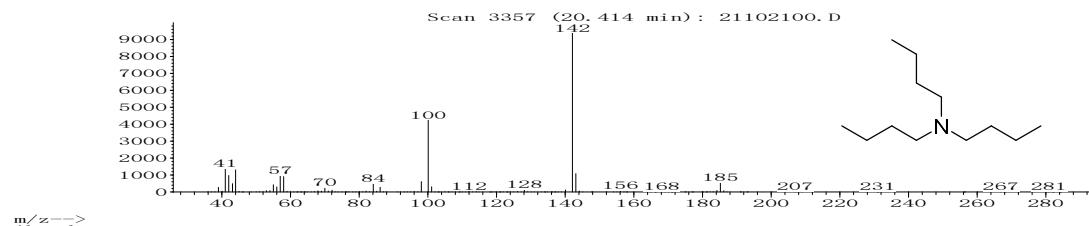
1
Abundance



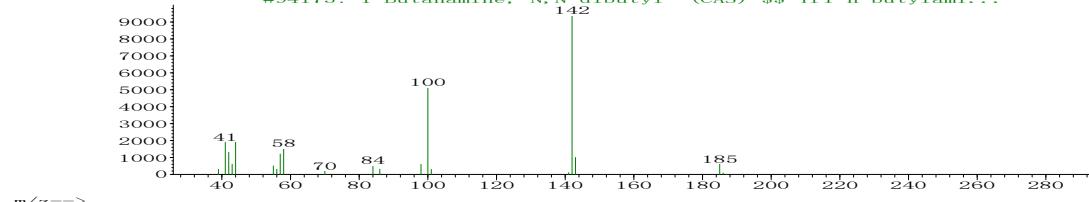
#30566: Butane, 1-bromo- (CAS) \$\$ 1-Bromobutane \$\$ n-Butyl...



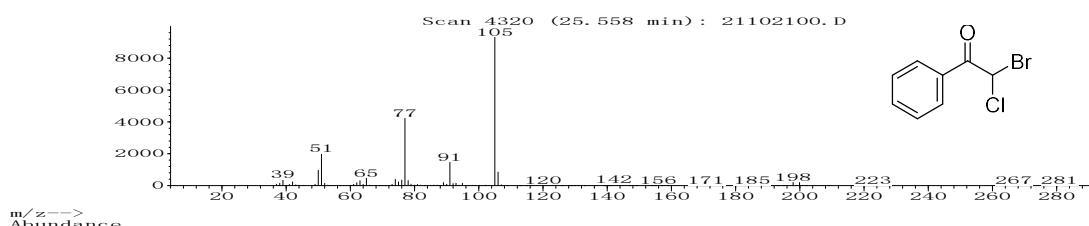
2
Abundance



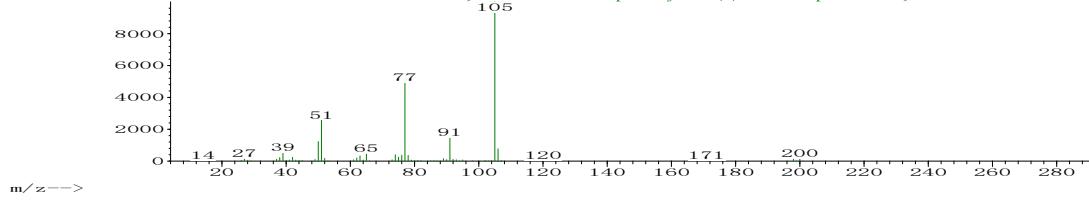
#94175: 1-Butanamine, N,N-dibutyl- (CAS) \$\$ Tri-n-butylami...



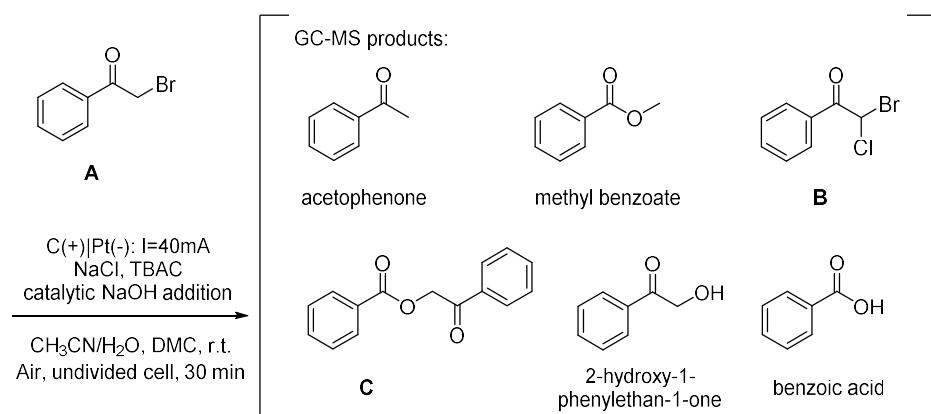
3
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#111480: Ethanone, 2-bromo-1-phenyl- \$\$ Acetophenone, 2-br...

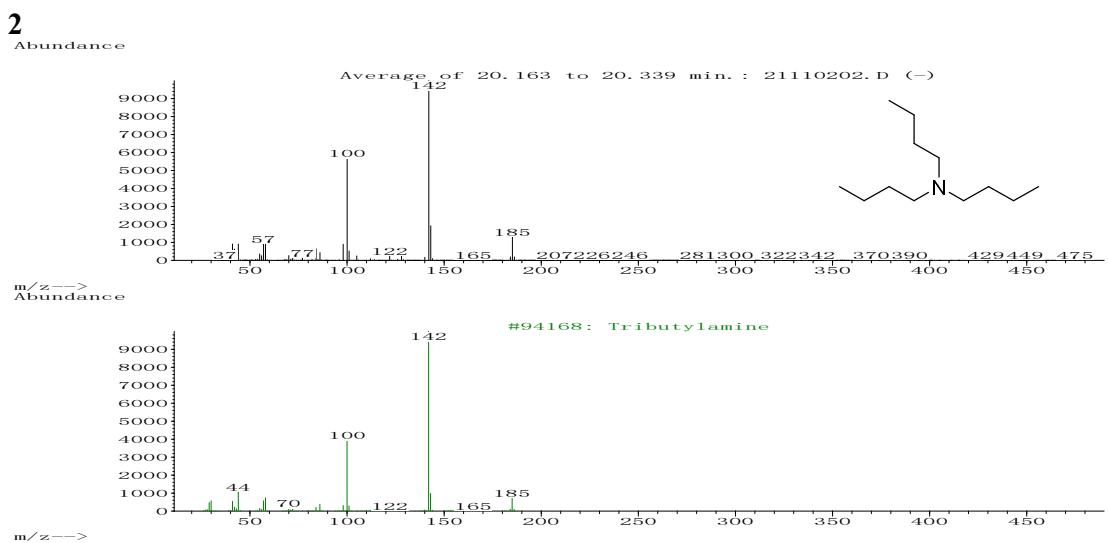
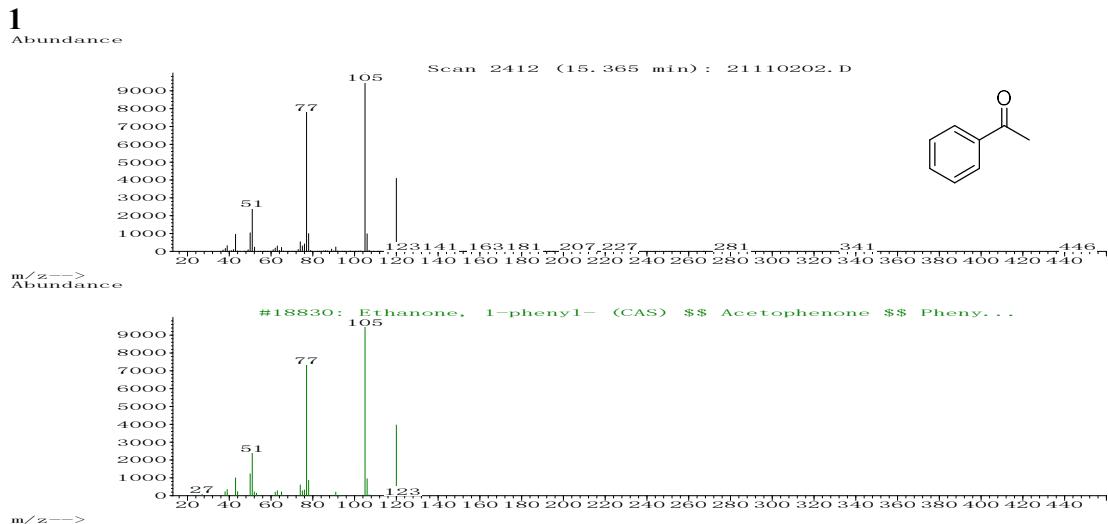
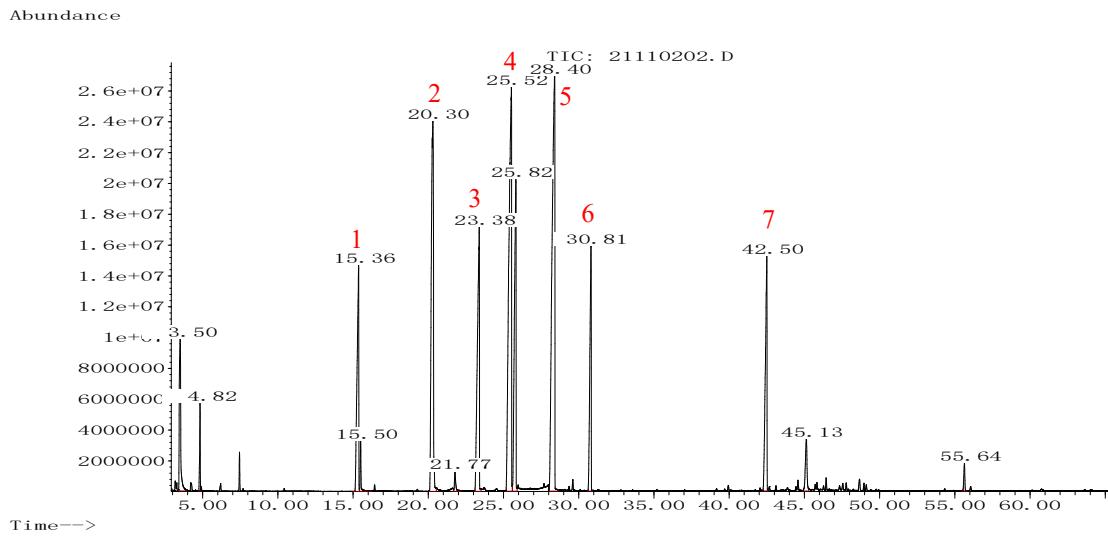


3.5.2 GC–MS for electrocatalytic dimer condensation



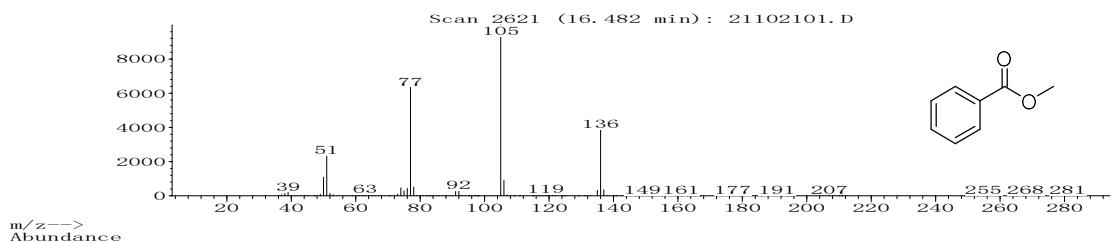
Scheme S6. Electrocatalysis for dimer condensation

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (~ 200 mg, 1.0 mmol), tetrabutylammonium halide (TBAC or TBAB, ~ 50 mg, 0.2 mmol), brine (eg: NaCl, KBr, NaF, 1.0 mol·L⁻¹, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. NaOH (1.0 mmol(L-1, 500 mg)) was added to the reaction solutions. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathodic electrode and a carbon plate (10 mm × 10 mm × 2.0 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 10 mA under room temperature for 10 min.

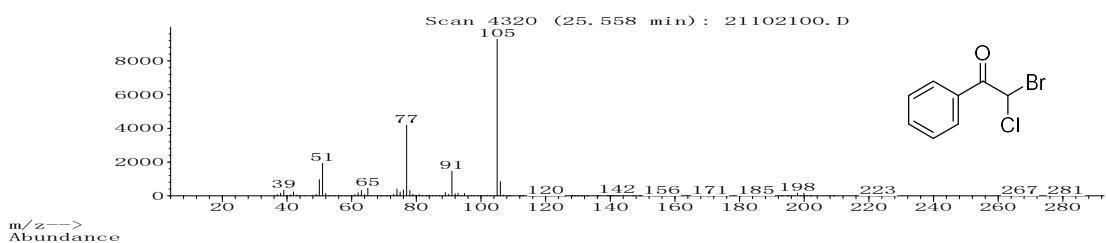


3

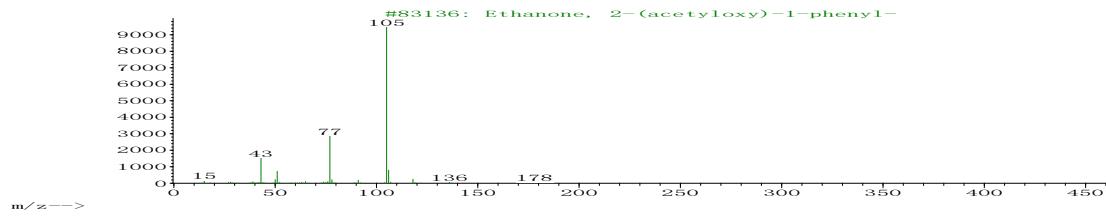
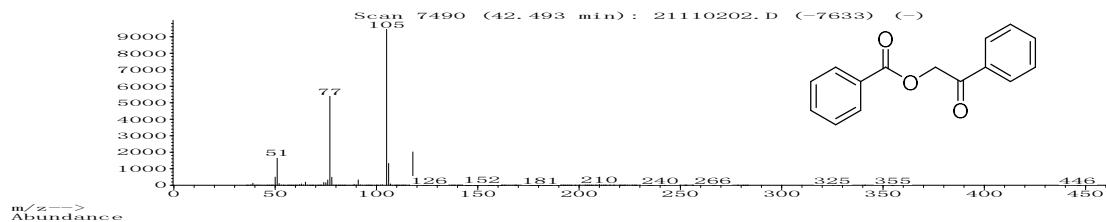
Abundance

**4**

Abundance

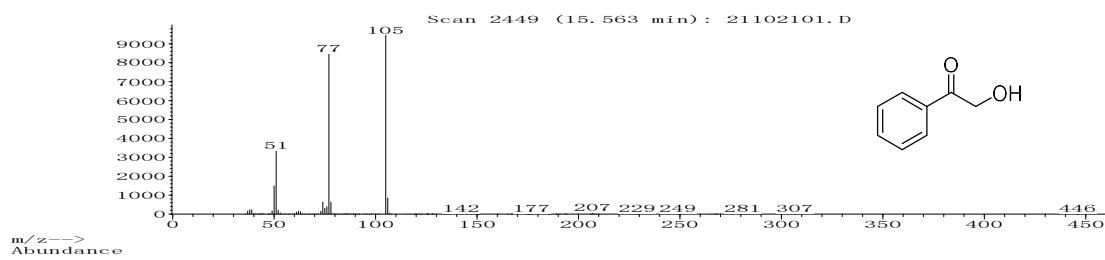
**5**

Abundance

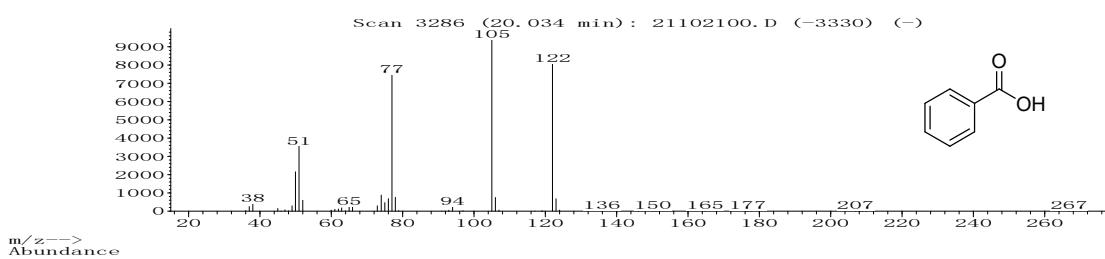


6

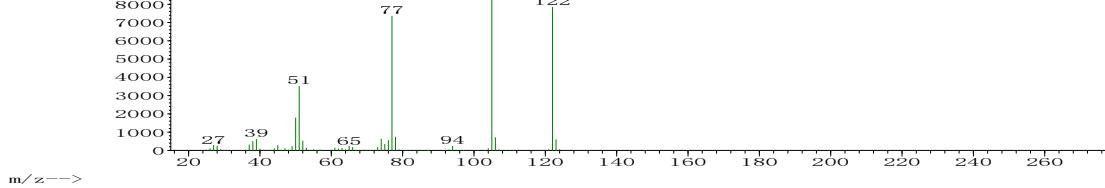
Abundance

**7**

Abundance

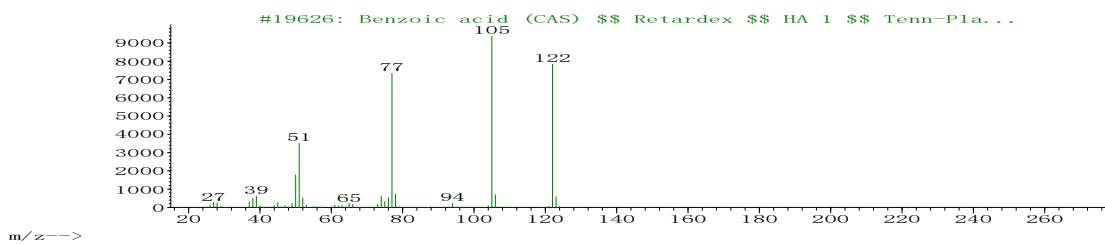
**7**

Abundance

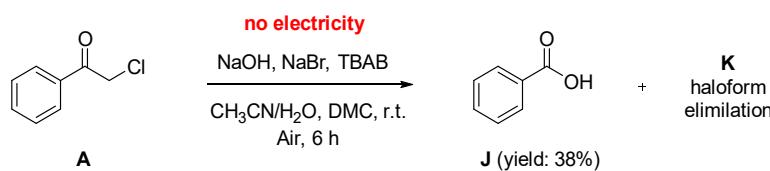


m/z-->

Abundance

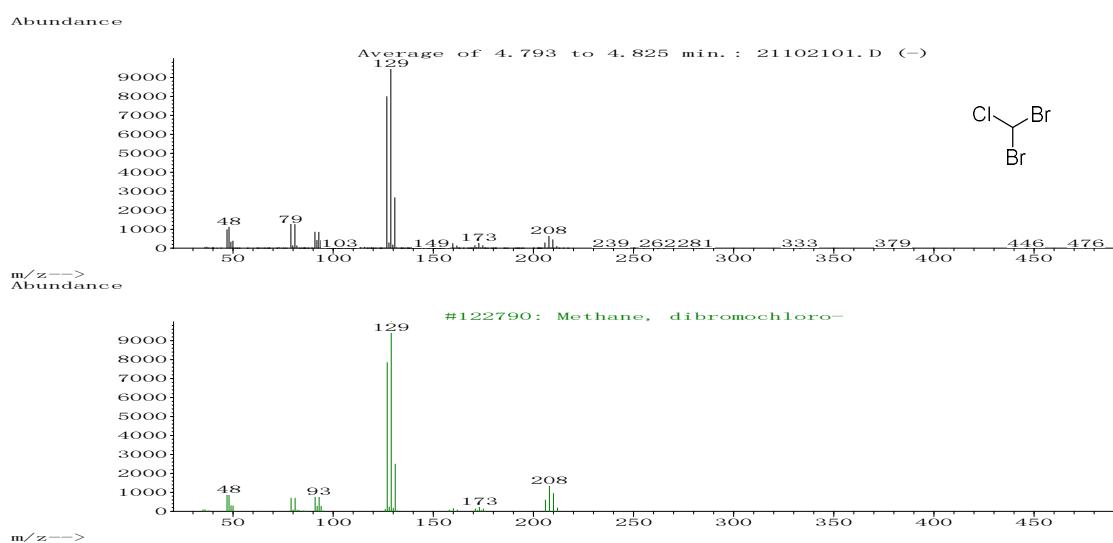


3.5.3 NaOH addition experiment without electricity

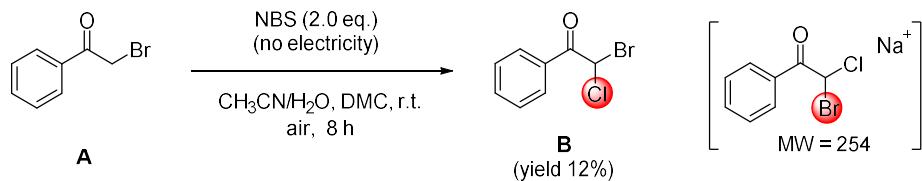


Scheme S7. NaOH addition experiment without electricity

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (200 mg, 1.0 mmol), TBAB(0.2 mmol, 55.6 mg), NaBr (1.0 mol·L⁻¹, 725 mg), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. NaOH (1.0 mol(L-1, 500 mg) was added to the reaction solutions. No electricity was used. The reaction was in an ambient atmosphere. The reaction mixture was stirred at room temperature for 6 h. TLC, UV light as a visualizing agent, the reaction was monitored, and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. When the material was no longer decreased, the solution was extracted with EtOAc (3 × 10 mL) and H₂O (3 × 10 mL). The combined organic layer was dried with Na₂SO₄ and filtered. The solvent was removed with a rotary evaporator. The pure product was obtained by flash chromatography on silica gel using EtOAc and petroleum ether as the eluent solvents. Then benzoic acid was isolated with a yield of 46.4 mg, 38%.



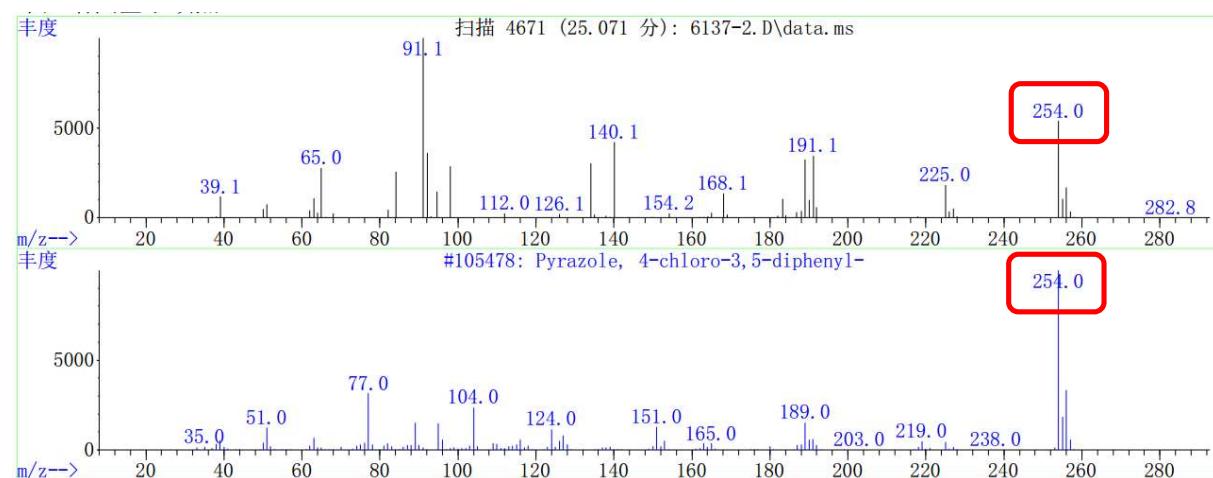
3.5.4 NBS experiment



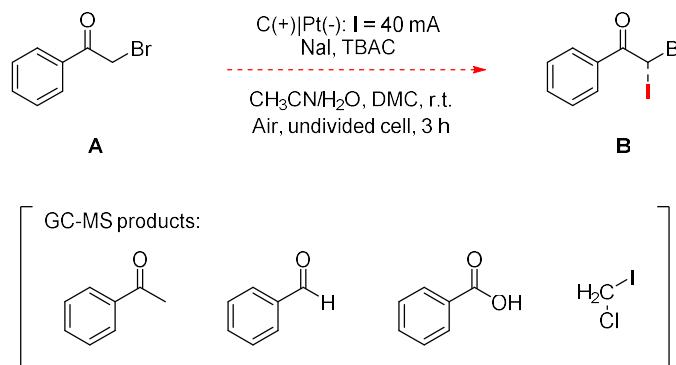
Scheme S8. NBS experiment

In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 2-bromo-acetophenone (200 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaCl (380 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The reaction was in an ambient atmosphere, without electricity for 8 h.

GC-MS result for NBS experiment.



3..5.5 NaI experiment

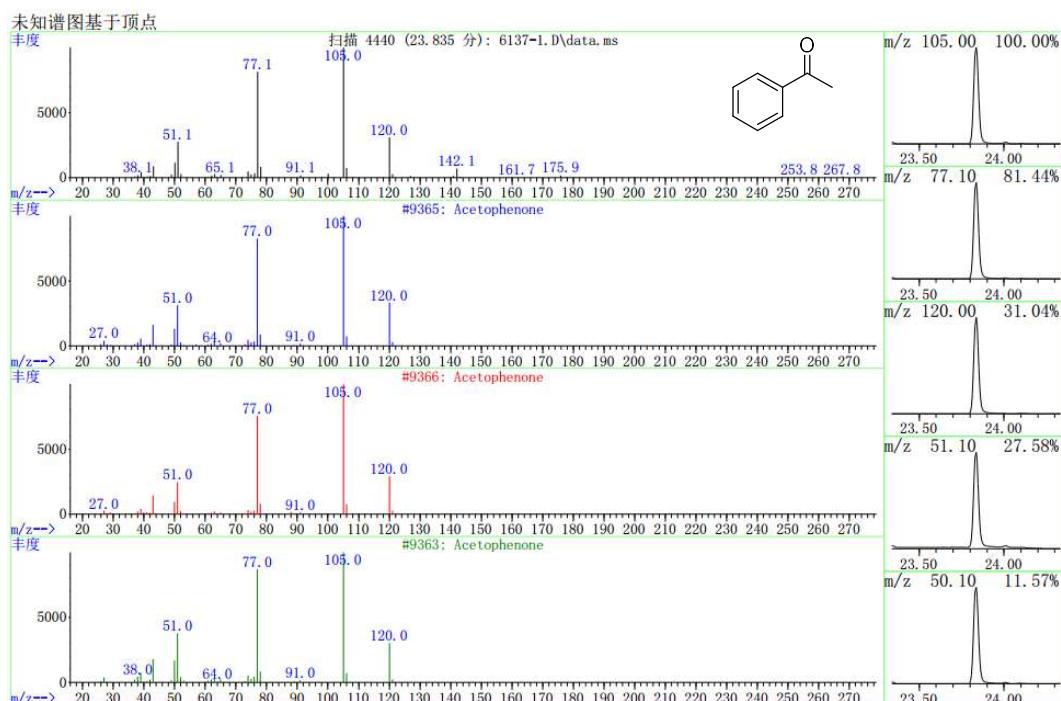


Scheme S9. NaI experiment

In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 2-bromo-acetophenone (200 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaI (715 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (50 mm×50 mm×0.2 mm) as the cathodic electrode and a carbon plate (50 mm×50 mm×2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 3 h.

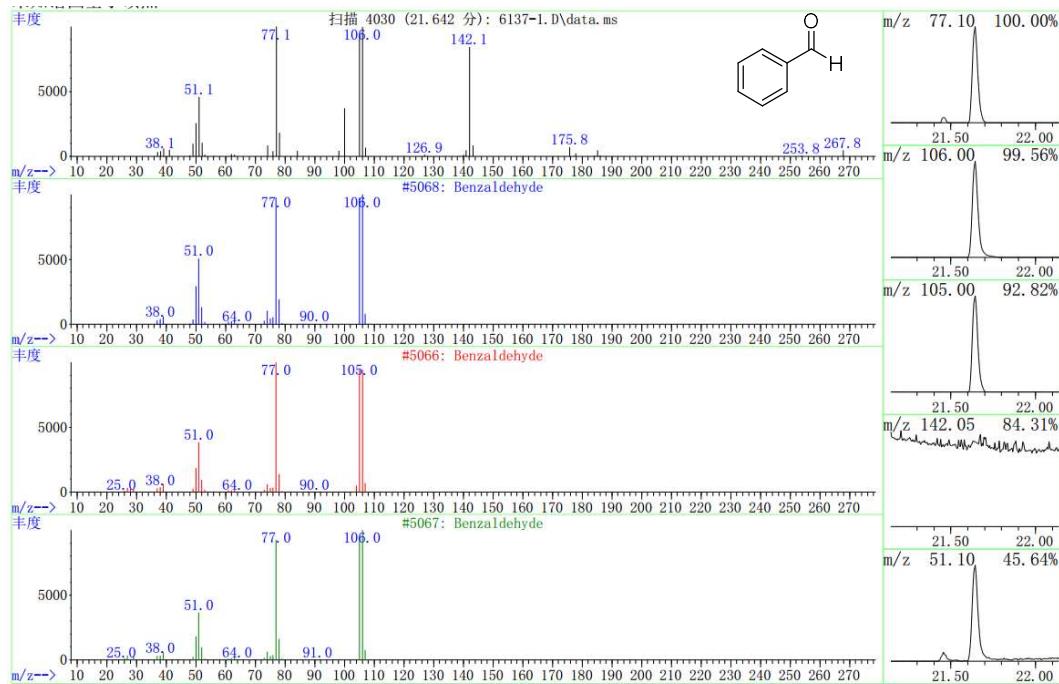
GC-MS result for NaI experiment.

1) acetophenone



	Ref\#	CAS\#	匹配度
C:\Database\NIST11.L			
1 Acetophenone	9365	000098-86-2	94
2 Acetophenone	9366	000098-86-2	91
3 Acetophenone	9363	000098-86-2	91

2) benzaldehyde

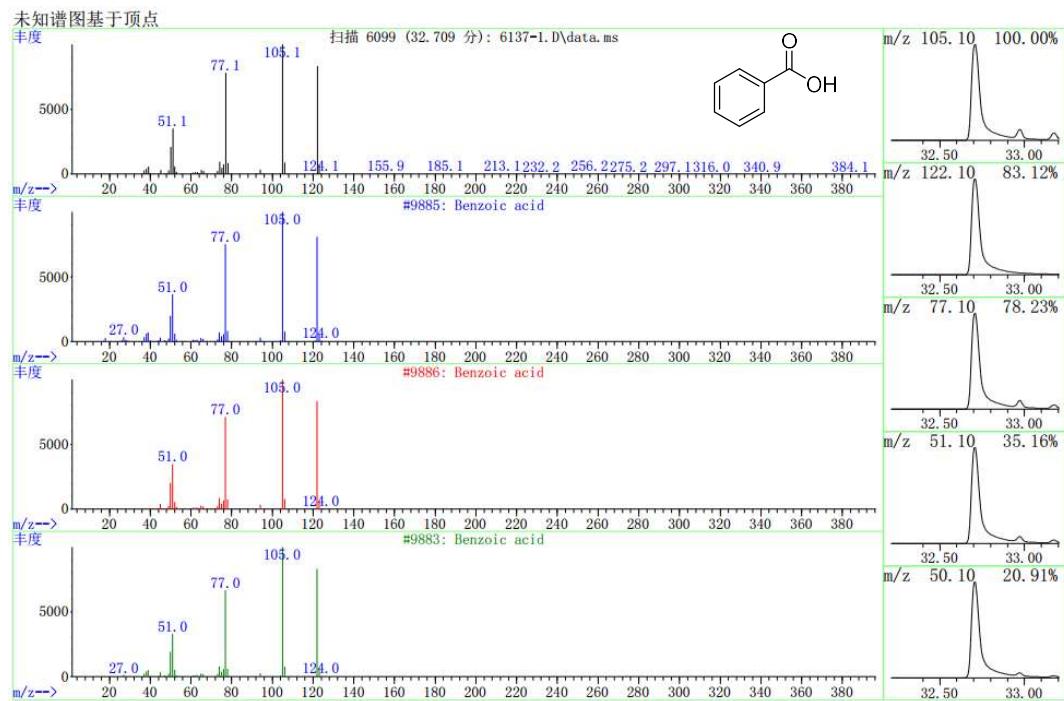


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L			
1 Benzaldehyde	5068	000100-52-7	91
2 Benzaldehyde	5066	000100-52-7	91
3 Benzaldehyde	5067	000100-52-7	64

3) benzoic acid

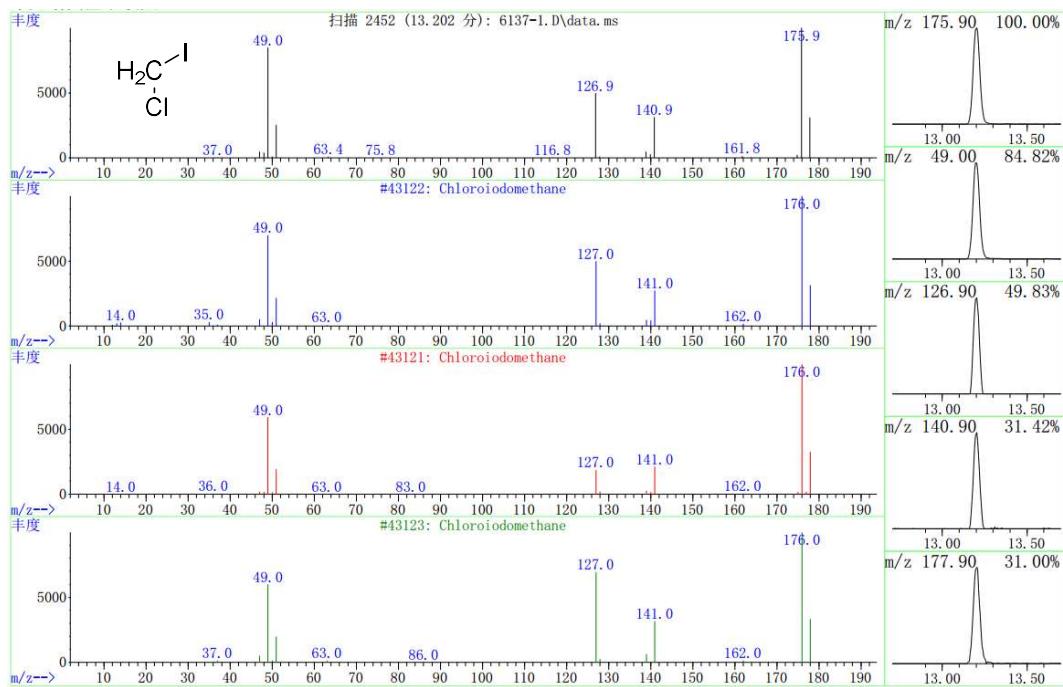


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L	Ref\#	CAS\#	匹配度
1 Benzoic acid	9885	000065-85-0	97
2 Benzoic acid	9886	000065-85-0	96
3 Benzoic acid	9883	000065-85-0	94

4) chloroiodomethane

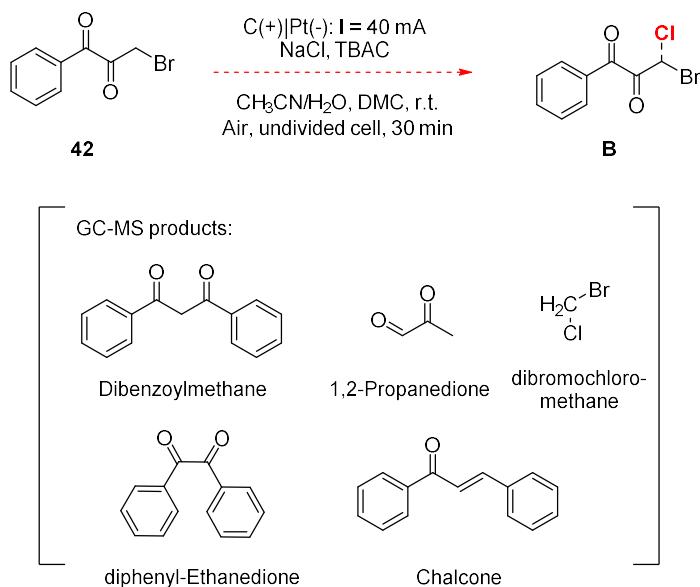


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L			
1 Chloroiodomethane	43122	000593-71-5	91
2 Chloroiodomethane	43121	000593-71-5	49
3 Chloroiodomethane	43123	000593-71-5	43

3.5.6 1,2-dione experiment

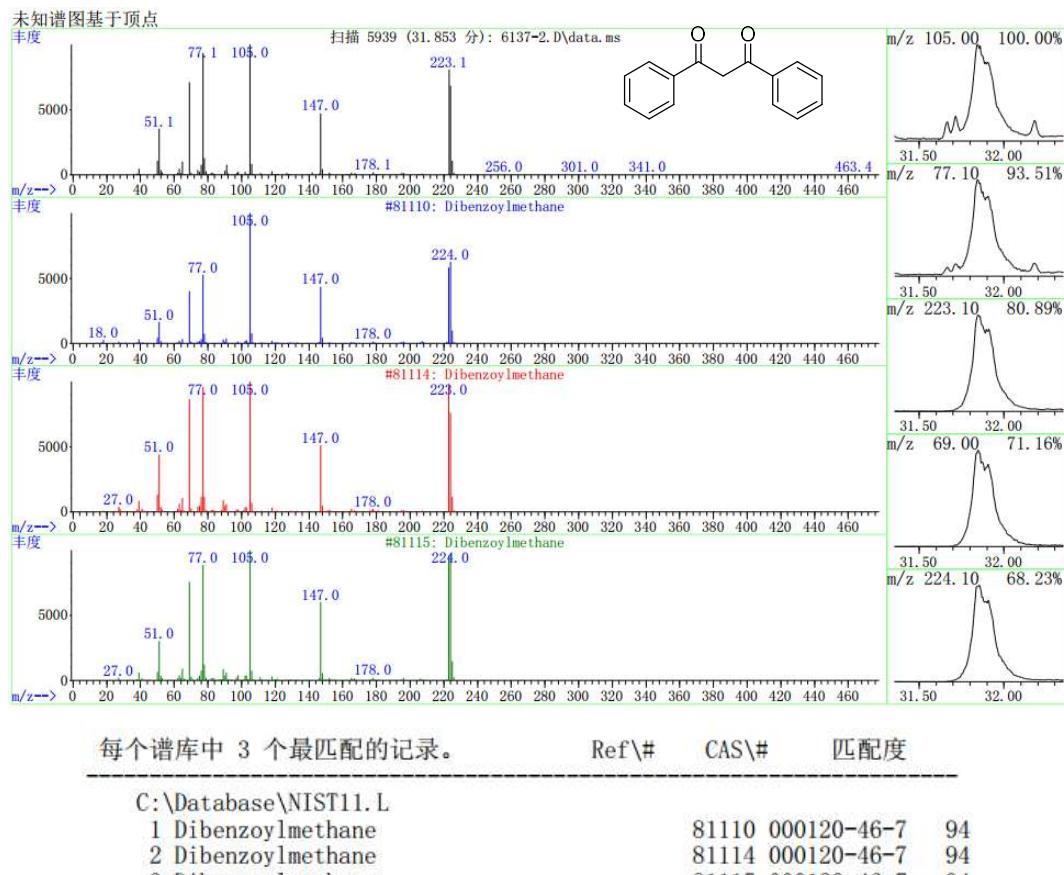


Scheme S10. 3-bromo-1-phenylpropane-1,2-dione experiment

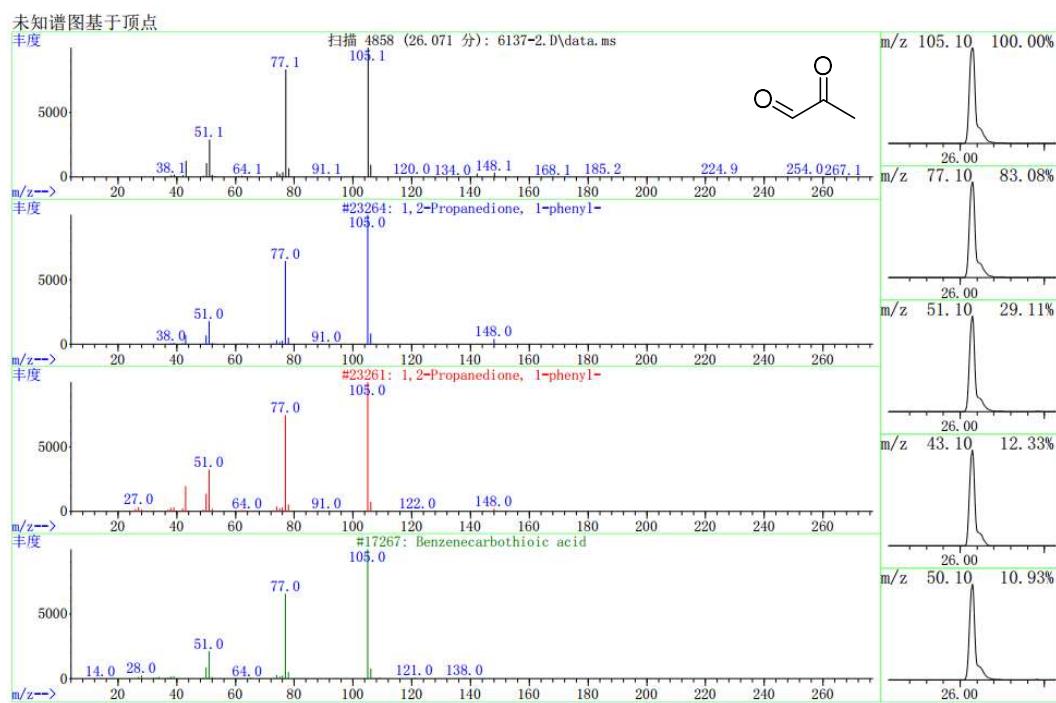
In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 3-bromo-1-phenylpropane-1,2-dione (227 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaI (715 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (50 mm×50 mm×0.2 mm) as the cathodic electrode and a carbon plate (50 mm×50 mm×2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 30 min.

GC-MS result for 1,2-dione experiment.

1) Dibenzoylmethane



2) 1,2-Propanedione

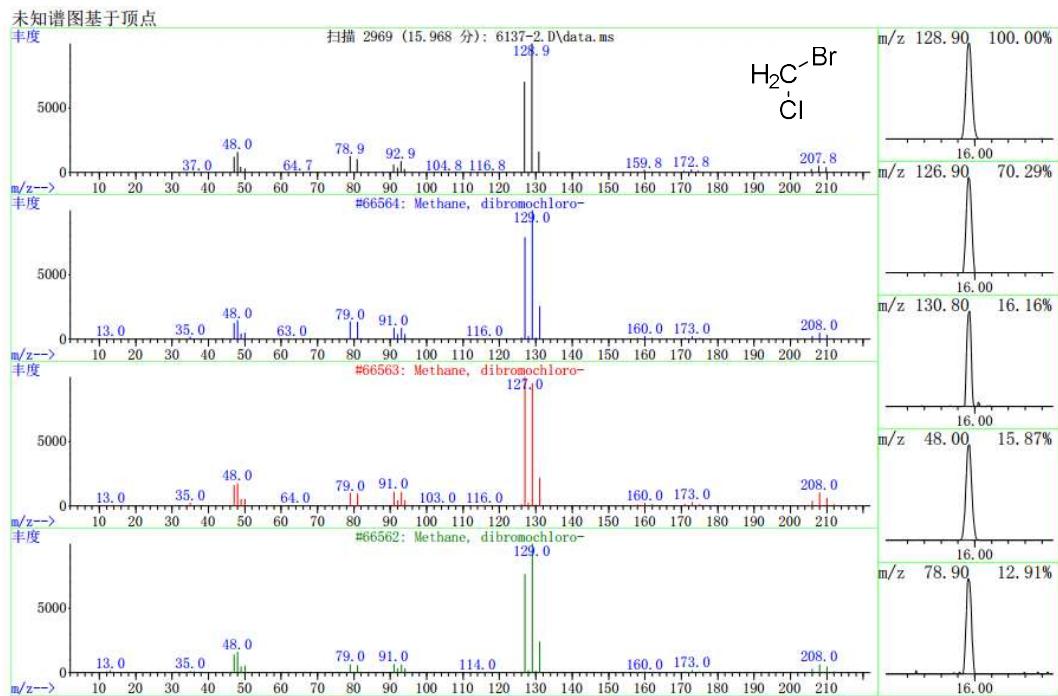


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L	Ref\#	CAS\#	匹配度
1 1, 2-Propanedione, 1-phenyl-	23264	000579-07-7	91
2 1, 2-Propanedione, 1-phenyl-	23261	000579-07-7	91
3 Benzenecarbothioic acid	17267	000098-91-9	90

3) dibromochloro-methane

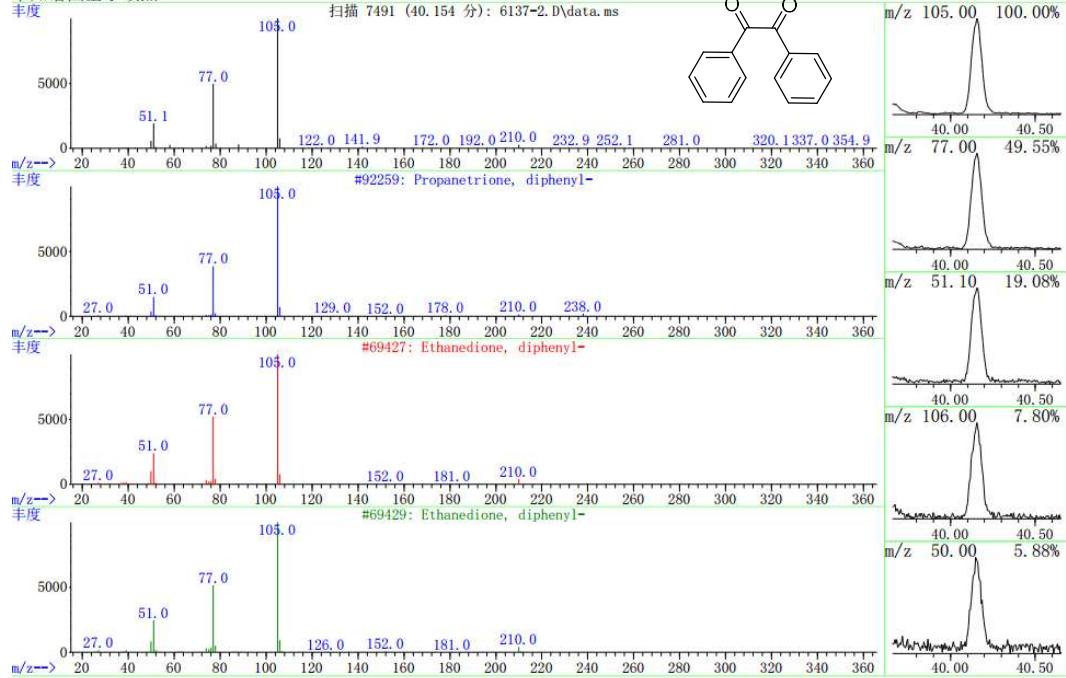


每个谱库中 3 个最匹配的记录。

Ref\#	CAS\#	匹配度
C:\Database\NIST11.L		
1 Methane, dibromochloro-	66564 000124-48-1	95
2 Methane, dibromochloro-	66563 000124-48-1	89
3 Methane, dibromochloro-	66562 000124-48-1	76

4) diphenyl-Ethanedione

未知谱图基于顶点

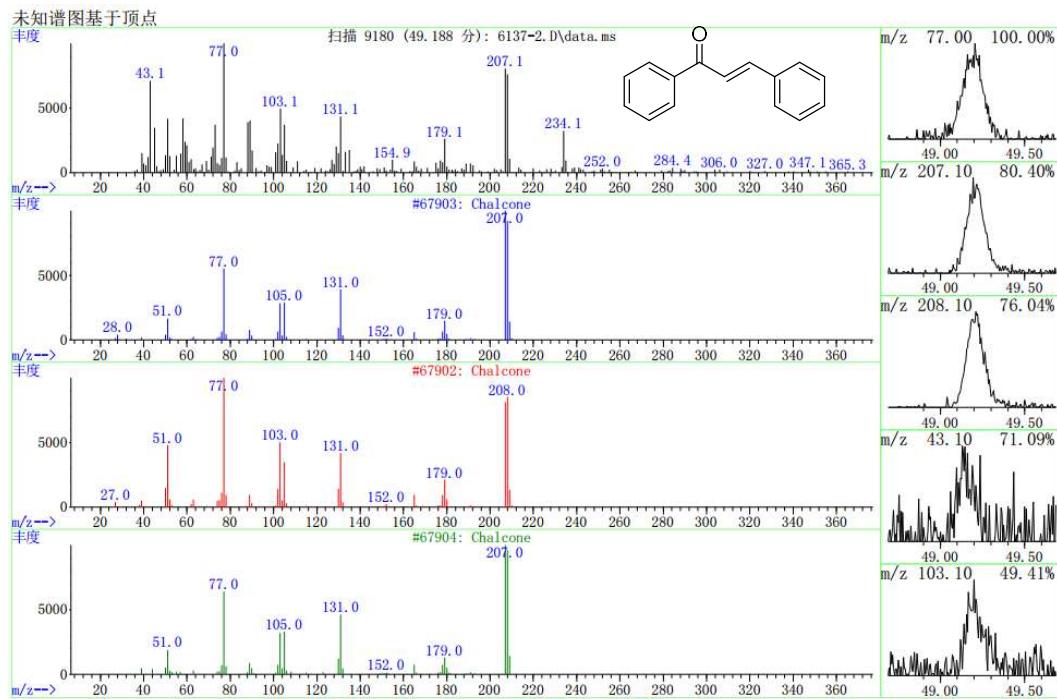


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L	Ref\#	CAS\#	匹配度
1 Propanetrione, diphenyl-	92259	000643-75-4	78
2 Ethanedione, diphenyl-	69427	000134-81-6	78
3 Ethanedione, diphenyl-	69429	000134-81-6	74

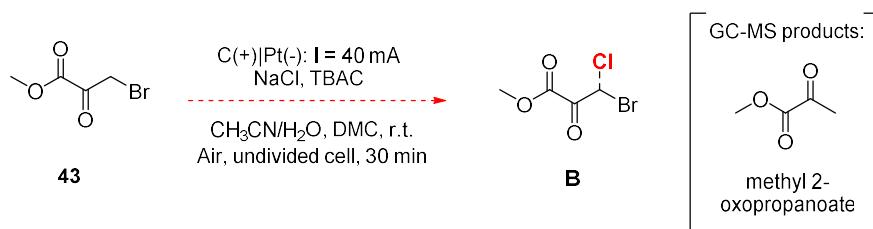
5) Chalcone



每个谱库中 3 个最匹配的记录。

Ref\#	CAS\#	匹配度
C:\Database\NIST11.L		
1 Chalcone	67903 000094-41-7	93
2 Chalcone	67902 000094-41-7	93
3 Chalcone	67904 000094-41-7	86

3.5.7 α -keto ester experiment

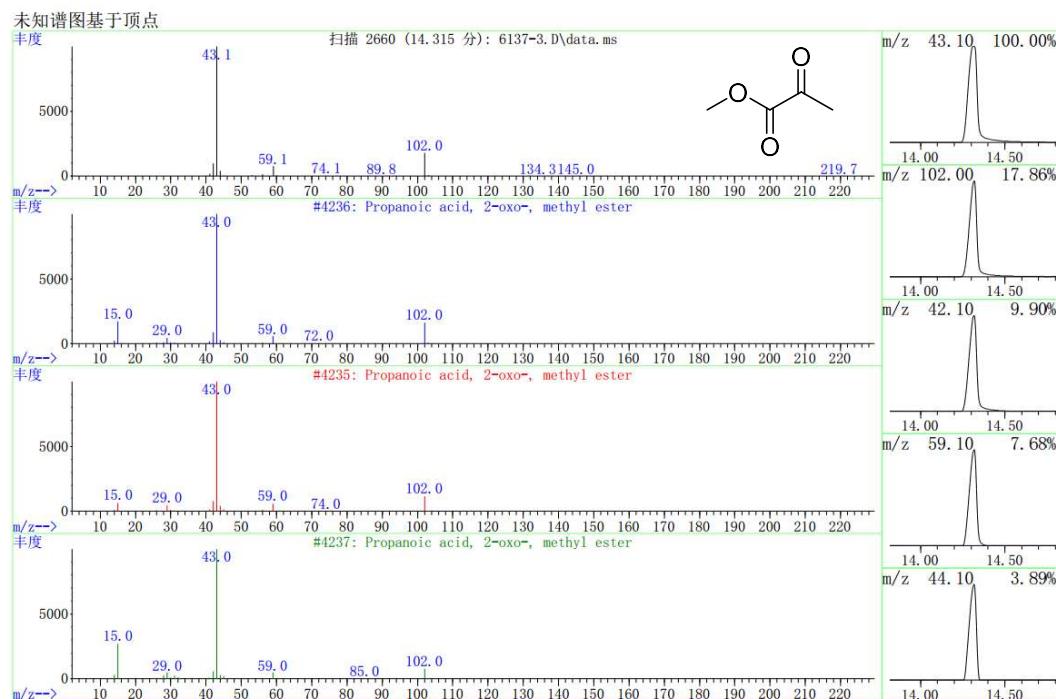


Scheme S11. methyl 3-bromo-2-oxopropanoate experiment

In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, methyl 3-bromo-2-oxopropanoate (181 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaI (715 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (50 mm×50 mm×0.2 mm) as the cathodic electrode and a carbon plate (50 mm×50 mm×2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 30 min.

GC-MS result for α -keto ester experiment.

1) Methyl pyruvate

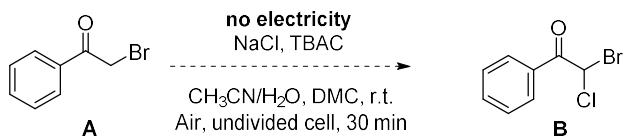


每个谱库中 3 个最匹配的记录。

Ref\# CAS\# 匹配度

C:\Database\NIST11.L	Ref\#	CAS\#	匹配度
1 Propanoic acid, 2-oxo-, methyl e...	4236	000600-22-6	80
2 Propanoic acid, 2-oxo-, methyl e...	4235	000600-22-6	72
3 Propanoic acid, 2-oxo-, methyl e...	4237	000600-22-6	7

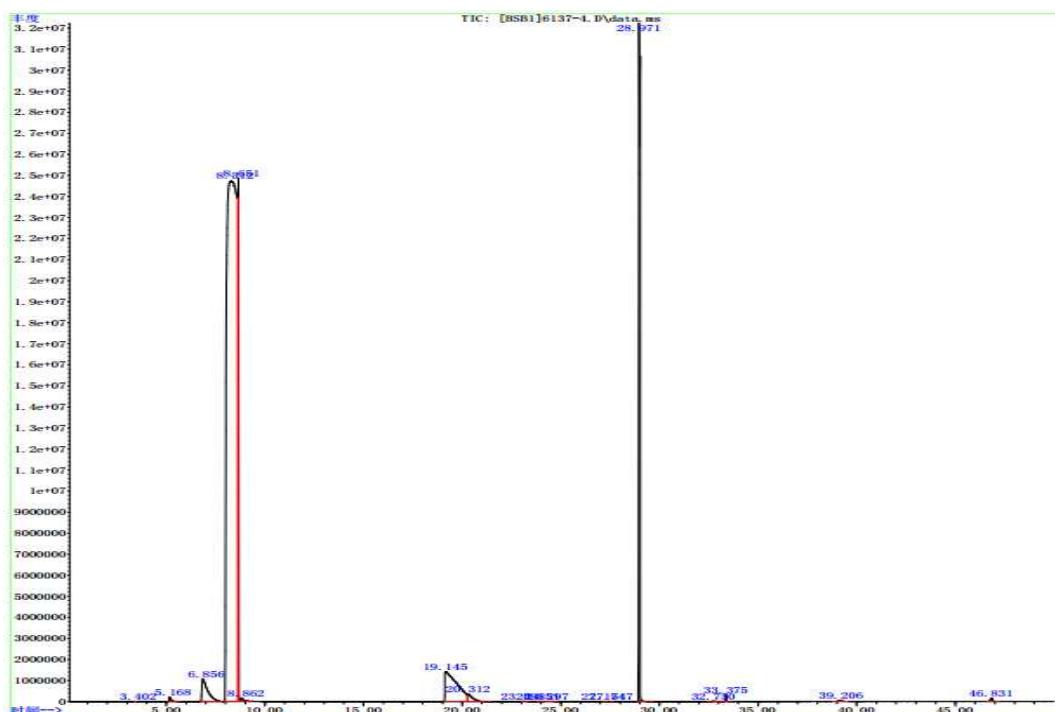
3.5.8 NaCl, TBAC (without electricity) experiment



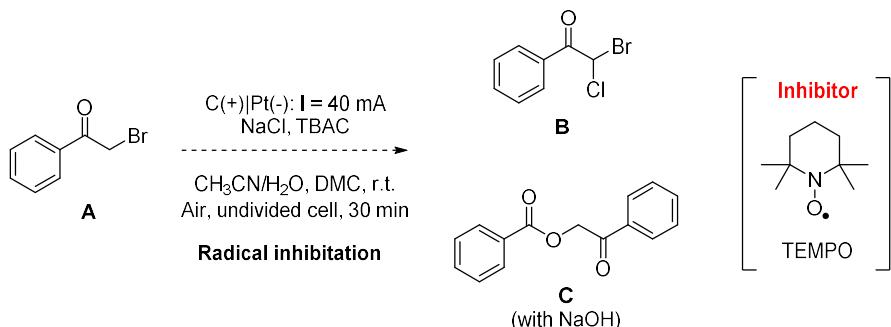
Scheme S12. NaOH addition experiment without electricity.

In an oven-dried undivided two-necked flask (250 mL) equipped with a stir bar, 2-bromo-acetophenone (200 mg, 1.0 mmol), TBAC (56.4 mg, 0.2 mmol), NaCl (380 mg, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The reaction was in an ambient atmosphere, without electricity for 30 min.

GC-MS map for NaCl, TBAC (without electricity) experiment



3.6 Radical inhibition experiment



Scheme S13. Radical inhibition experiments.

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (200 mg, 1.0 mmol), TBAC (0.2 mmol, 55.6 mg), NaCl (1.0 mol·L⁻¹, 725 mg), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. To test the radical inhabitation of this electrocatalysis experiment, a chemical equivalent of TEMPO was added to the reaction solutions. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathodic electrode and a carbon plate (10 mm × 10 mm × 2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 30 min. TLC, UV light as a visualizing agent, the reaction was monitored, and an ethanolic solution of phosphomolybdic acid and cerium sulfate and heat as a developing agent. Another experiment with an addition of a buffer of NaOH (1.0 mol·L⁻¹, 500 mg) was also conducted for the radical inhibition experiment, and no dimer condensation product was observed.

3.7 EPR experiments

EPR spectra were recorded at room temperature on a Bruker EMX micro A300 spectrometer operated at 9.8543 GHz. Typical spectrometer parameters are shown as follows, sweep width: 6000 G; center field set: 3000 G; time constant: 81.92 ms; sweep time: 327.68 s modulation amplitude: 1.00 G; modulation frequency: 100 kHz; receiver gain: 1.00×10^3 ; microwave power: 18.53 mW.

EPR experiment for halogenation:

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (200 mg, 1.0 mmol), TBAC (0.2 mmol, 55.6 mg), NaCl (1.0 mol·L⁻¹, 725 mg), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathodic electrode and a carbon plate (10 mm × 10 mm × 2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 10 min.

EPR experiments were performed to detect the radical intermediates with the addition of free radical spin trapping agent 5,5-dimethyl-1-pyrroline N-oxide (DMPO). The control experiment was conducted without electricity. Results showed that no distinct free radical signal was detected under reaction conditions without electricity with any substrates; however, EPR signals were identified after the electrocatalysis. DMPO could quickly trap the free radicals to form a relatively stable radical (DMPO·R, R represents alkoxy radical) ($g=2.00662$, $a_N=14.28$, $aH^\beta=13.27$).

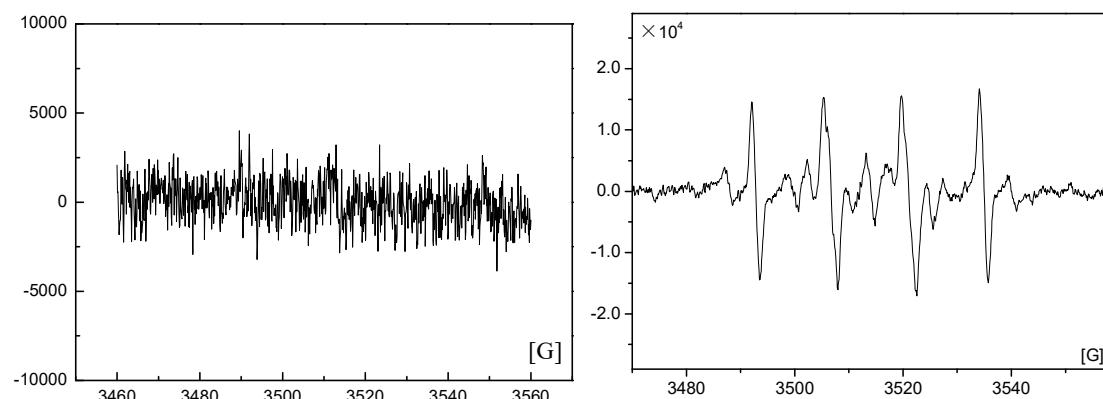


Figure S3. EPR of reactant A for electrohalogenation

EPR experiment for dimer condensation:

In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-halo-acetophenone (200 mg, 1.0 mmol), TBAC (0.2 mmol, 55.6 mg), NaCl (1.0 mol·L⁻¹, 725 mg), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. NaOH (1.0 mol(L-1, 500 mg) was added to the reaction solutions. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathodic electrode and a carbon plate (10 mm × 10 mm × 2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 10 min.

EPR experiments were performed to detect the radical intermediates with the addition of free radical spin trapping agent 5,5-dimethyl-1-pyrroline N-oxide (DMPO). The control experiment was conducted without electricity. Results stated that no distinct free radical signal was detected under reaction conditions without electricity with any substrates; however, EPR signals were identified after the electrocatalysis. DMPO can quickly trap the free radicals to form a typical hydroxyl radical (DMPO·OH, $g=2.00634$, $a_N=14.98$ G, $aH^{\beta}=14.86$ G).

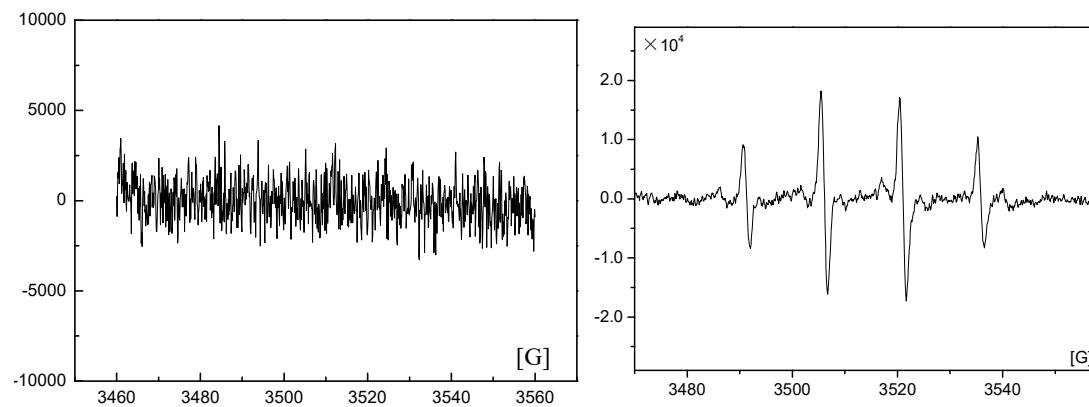


Figure S4. EPR of reactant A for dimer condensation

3.8 Microcoulometric experiments

Microcoulometric experiments was conducted to investigate the electro-catalysis reversibility and electron transfer number (*n*-value). For the halogenation experiments, $Q_{\text{exp}} = it = 40 \times 0.001 \times 30 \times 60 = 72 \text{ C}$, $Q_{\text{theor}} = 0.001 \times 1 \times 96500 = 96.5 \text{ C}$. The Columbic yield (current efficiency) is $96.5/72 \times 0.8 \times 100 = 107\% \approx 1$. This means that the reaction involves in conversion of one haloketone molecule with one electron.

CV experiments for oxidation peak of A:

Using CH₃CN (20 mL), and H₂O (20 mL) as co-solvent. Without NaOH, cyclic voltammetry (CVs) experiments of reactant 2-bromo-acetophenone (A, 199.4 mg) were conducted at different scan rates. At a scan rate of 0.010 V/s, from -0.5 V to 2.0 V, the oxidation peak of A was reversible ($i_{\text{pa1}} = 0.32 \text{ mA}$, $i_{\text{pc1}} = 0.28 \text{ mA}$, $i_{\text{pa1}}/i_{\text{pc1}} = 1.1 \approx 1$). The relationship of E_{pa} and i_{pa} at different scan rate are $E_{\text{pa}} = 0.0728 \log i_{\text{pa1}} + 0.094$ ($R^2 = 0.8397$). Results indicated that, in the electrohalogenation process the transferred electrons for the oxidation peak of A is one ($n = 0.0712/0.0728 = 0.98 \approx 1$).

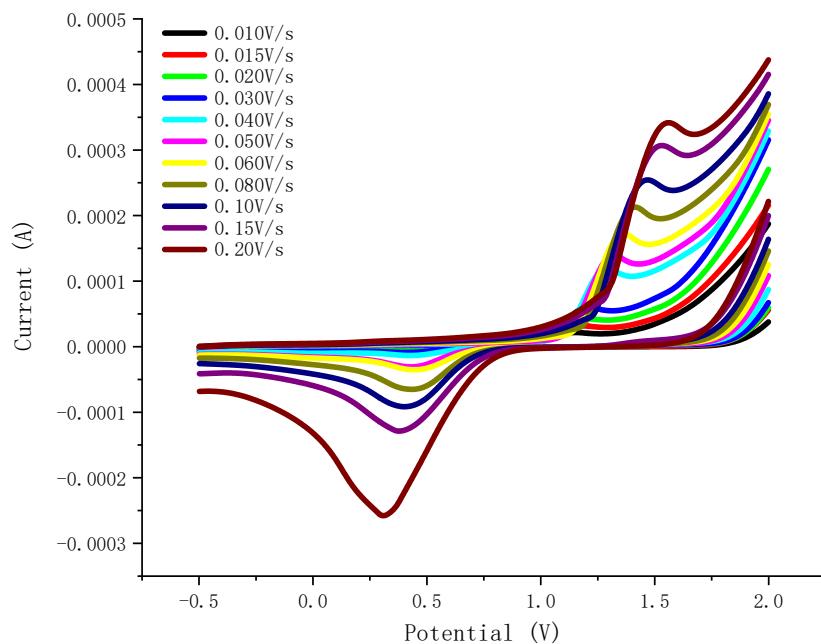


Figure S5. CVs of A at different scan rates

E _{pa1}	i _{pa1}	Lg i _{pa1}
0.50	-2.04*10 ⁻⁶	-5.69
0.496	-4.07*10 ⁻⁶	-5.39
0.473	-4.49*10 ⁻⁶	-5.35
0.44	-8.78*10 ⁻⁶	-5.06
0.447	-1.36*10 ⁻⁵	-4.87
0.43	-3.10*10 ⁻⁵	-4.51
0.447	-3.50*10 ⁻⁵	-4.46
0.429	-6.46*10 ⁻⁵	-4.19
0.40	-9.12*10 ⁻⁵	-4.04
0.38	-1.29*10 ⁻⁴	-3.89
0.31	-2.58*10 ⁻⁴	-3.59

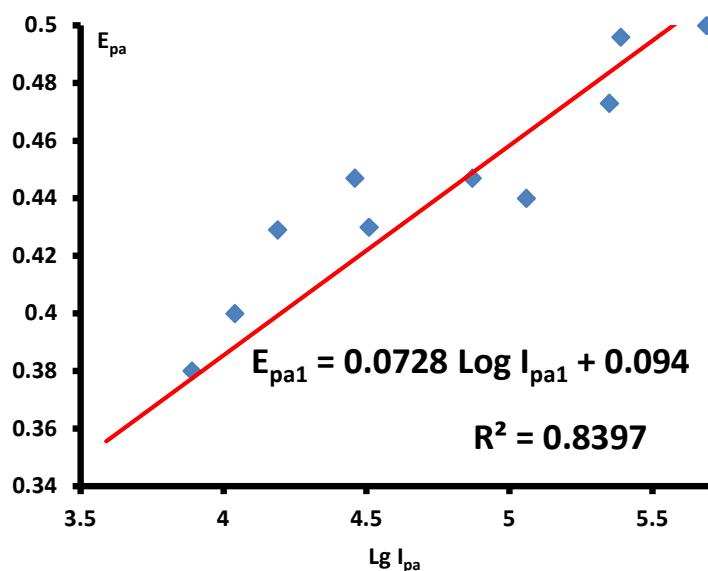


Figure S6. Line fitting of E_{pa1} and $\log i_{pa1}$ for the oxidation peak of A

CV experiments for oxidation peak of A, with NaOH addition:

CH₃CN (20 mL) and H₂O (20 mL) as co-solvent. NaOH (200 mg) was added to the reaction solution, cyclic voltammetry (CVs) experiments of reactant 2-bromo-acetophenone (**A**, 199.4 mg) were conducted at different scan rates. At a scan rate of 0.010 V/s, from – 0.5 V to 2.5 V, the oxidation peak of **A** was irreversible. The relationship of E_{pa} and i_{pa} at different scan rate are E_{pc} = -0.6491 Log i_{pc} + 3.6361 (R² = 0.808).

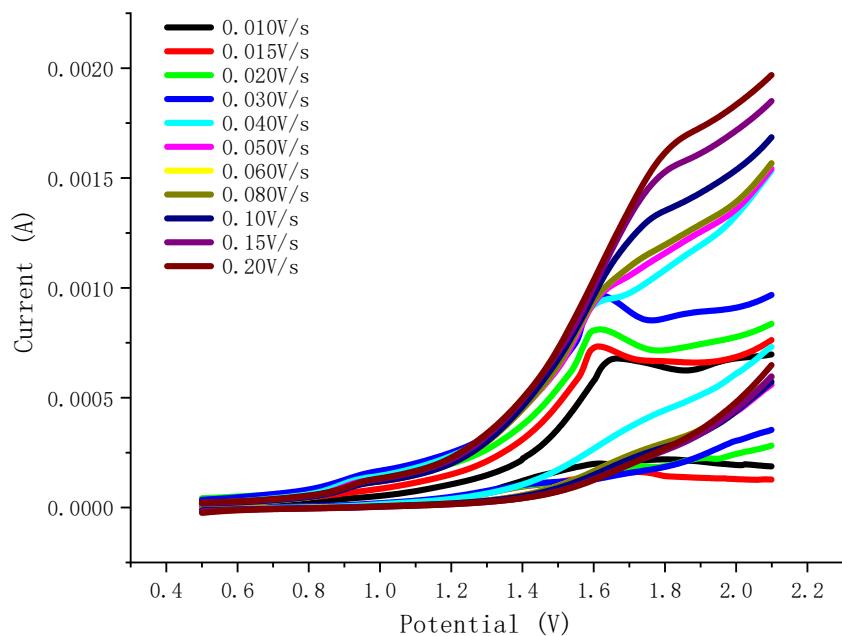


Figure S7. CVs of A at different scan rates with NaOH

E _{pc}	i _{pc}	Lg i _{pc}
1.66	6.78*10 ⁻⁴	-3.17
1.61	7.33*10 ⁻⁴	-3.13
1.62	8.11*10 ⁻⁴	-3.09
1.622	9.64*10 ⁻⁴	-3.02
1.622	9.44*10 ⁻⁴	-3.025
1.64	9.93*10 ⁻⁴	-3.003
1.71	1.10*10 ⁻³	-2.96
1.75	1.16*10 ⁻³	-2.94
1.80	1.35*10 ⁻³	-2.87
1.82	1.55*10 ⁻³	-2.81
1.86	1.70*10 ⁻³	-2.77

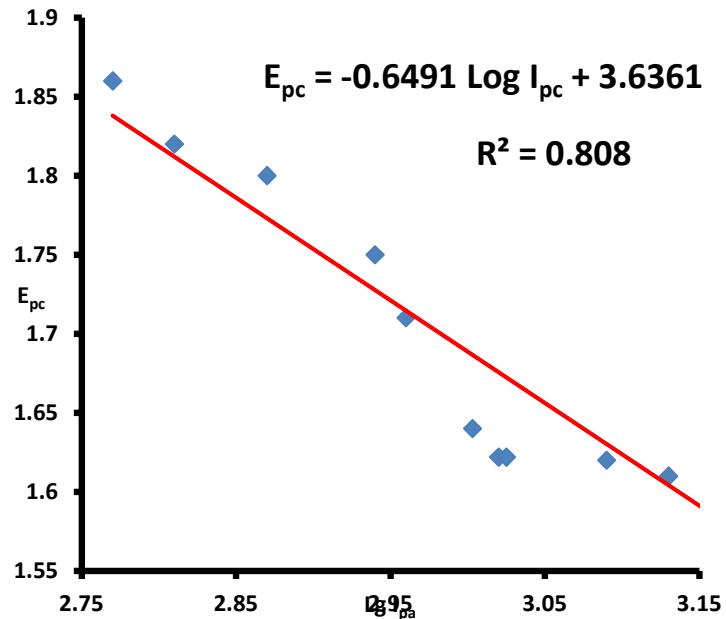


Figure S8. Line fitting of E_{pa1} and $\log i_{pa1}$ for the oxidation peak of A (with NaOH addition)

CV experiments for oxidation peak of Ferrocene:

Using CH₃CN (20 mL) and H₂O (20 mL) as co-solvent. Cyclic voltammetry (CVs) experiments of Ferrocene (186.0 mg) were conducted at different scan rates. At a scan rate of 0.010 V/s, from 0 V to 0.8 V, the oxidation peak of Ferrocene was irreversible. The relationship of E_{pa} and i_{pa} at different scan rate are E_{pc} = 0.0712 Log i_{pa} - 0.1033 ($R^2 = 0.9599$).

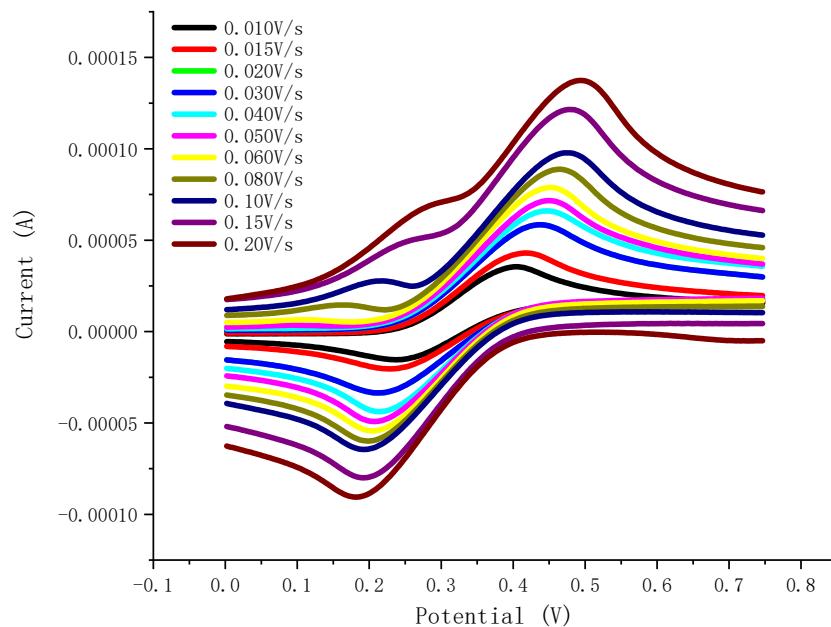


Figure S9. CVs of Ferrocene at different scan rates

E_{pa}	i_{pa}	$\log i_{pa}$
0.24	-1.54×10^{-5}	-4.81
0.23	-2.02×10^{-5}	-4.69
0.211	-3.342×10^{-5}	-4.475
0.212	-3.339×10^{-5}	-4.476
0.213	-4.37×10^{-5}	-4.36
0.206	-4.92×10^{-5}	-4.31
0.204	-5.43×10^{-5}	-4.27
0.20	-5.97×10^{-5}	-4.22
0.192	-6.44×10^{-5}	-4.19
0.19	-7.99×10^{-5}	-4.10
0.18	-9.07×10^{-5}	-4.04

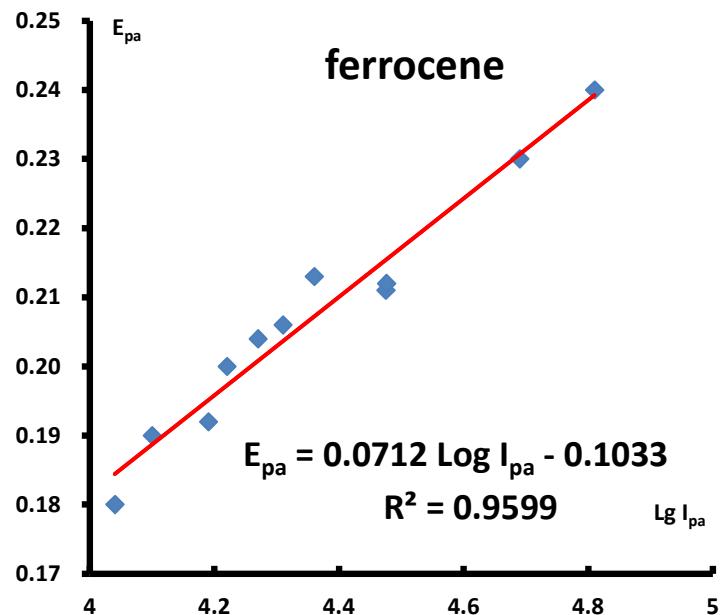


Figure S10. Line fitting of E_{pa1} and $\log i_{pa1}$ for the oxidation peak of Ferrocene

3.9 Cyclic voltammetry experiments

Cyclic voltammograms were conducted on a Metrohm PGSTAT302N potentiostat and performed in a three-electrode cell connected to a Schlenk line under nitrogen at room temperature. The working electrode was a carbon plate electrode. The counter electrode was a platinum wire. The reference was a Hg/Hg₂Cl₂ electrode submerged in saturated aqueous KBr solution and separated from the reaction by a salt bridge.

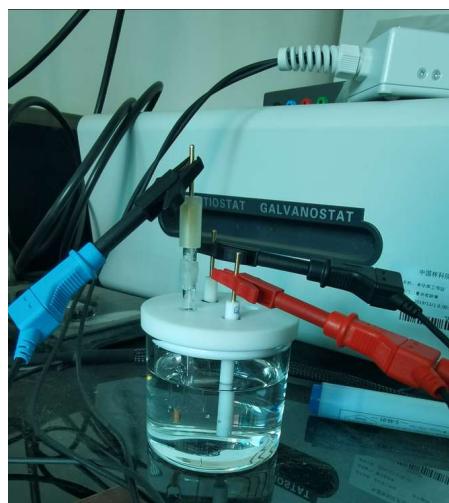


Figure S11. Cyclic voltammograms experiments

(A) CV Procedure for NaCl. A mixed solvent (CH₃CN/H₂O = 15 mL /15 mL) containing NaCl (1.19g, 0.01 mol·L⁻¹, 5.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -2.0 V to 3.0 V.

(B) CV Procedure for NaCl and 2-bromo-acetophenone. A mixed solvent (CH₃CN/H₂O = 15 mL /15 mL) containing NaCl (1.19g, 0.01 mol·L⁻¹, 5.0 chemical equivalent) and 2-bromo-acetophenone (1.5459g, 0.01 mol·L⁻¹, 1.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -2.0 V to 3.0 V.

(C) CV Procedure for NaBr. A mixed solvent (CH₃CN/H₂O = 15 mL /15 mL) containing NaBr (0.52 g, 0.01 mol·L⁻¹, 5.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s,

ranging from -2.0 V to 3.0 V.

(D) CV Procedure for NaF. A mixed solvent ($\text{CH}_3\text{CN}/\text{H}_2\text{O} = 15 \text{ mL /}15 \text{ mL}$) containing NaF (0.21 g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 5.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -1.5 V to 2.5 V.

(E) CV Procedure for NaOH. A mixed solvent ($\text{CH}_3\text{CN}/\text{H}_2\text{O} = 15 \text{ mL /}15 \text{ mL}$) containing NaOH (0.20g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 5.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -2.0 V to 2.5 V.

(F) CV Procedure for NaOH and 2-bromo-acetophenone. A mixed solvent ($\text{CH}_3\text{CN}/\text{H}_2\text{O} = 15 \text{ mL /}15 \text{ mL}$) containing NaOH (0.20g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 5.0 chemical equivalent) and 2-bromo-acetophenone (1.5459g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 1.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 10 mA. The scan rate was 0.10 V/s, ranging from -2.0 V to 2.5 V.

(G) CV Procedure for NaCl, NaOH and 2-bromo-acetophenone. A mixed solvent ($\text{CH}_3\text{CN}/\text{H}_2\text{O} = 15 \text{ mL /}15 \text{ mL}$) containing NaCl (1.19g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 5.0 chemical equivalent), NaOH (0.20g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 5.0 chemical equivalent) and 2-bromo-acetophenone (1.5459g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 1.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -1.0 V to 3.0 V.

(H) CV Procedure for TBAC. A mixed solvent ($\text{CH}_3\text{CN}/\text{H}_2\text{O} = 15 \text{ mL /}15 \text{ mL}$) containing TBAC (0.3230g, $0.01 \text{ mol}\cdot\text{L}^{-1}$, 1.0 chemical equivalent), was added into the electrochemical cell in cyclic voltammetry experiments. A constant current of 1 mA. The scan rate was 0.10 V/s, ranging from -0.7 V to 2.5 V.

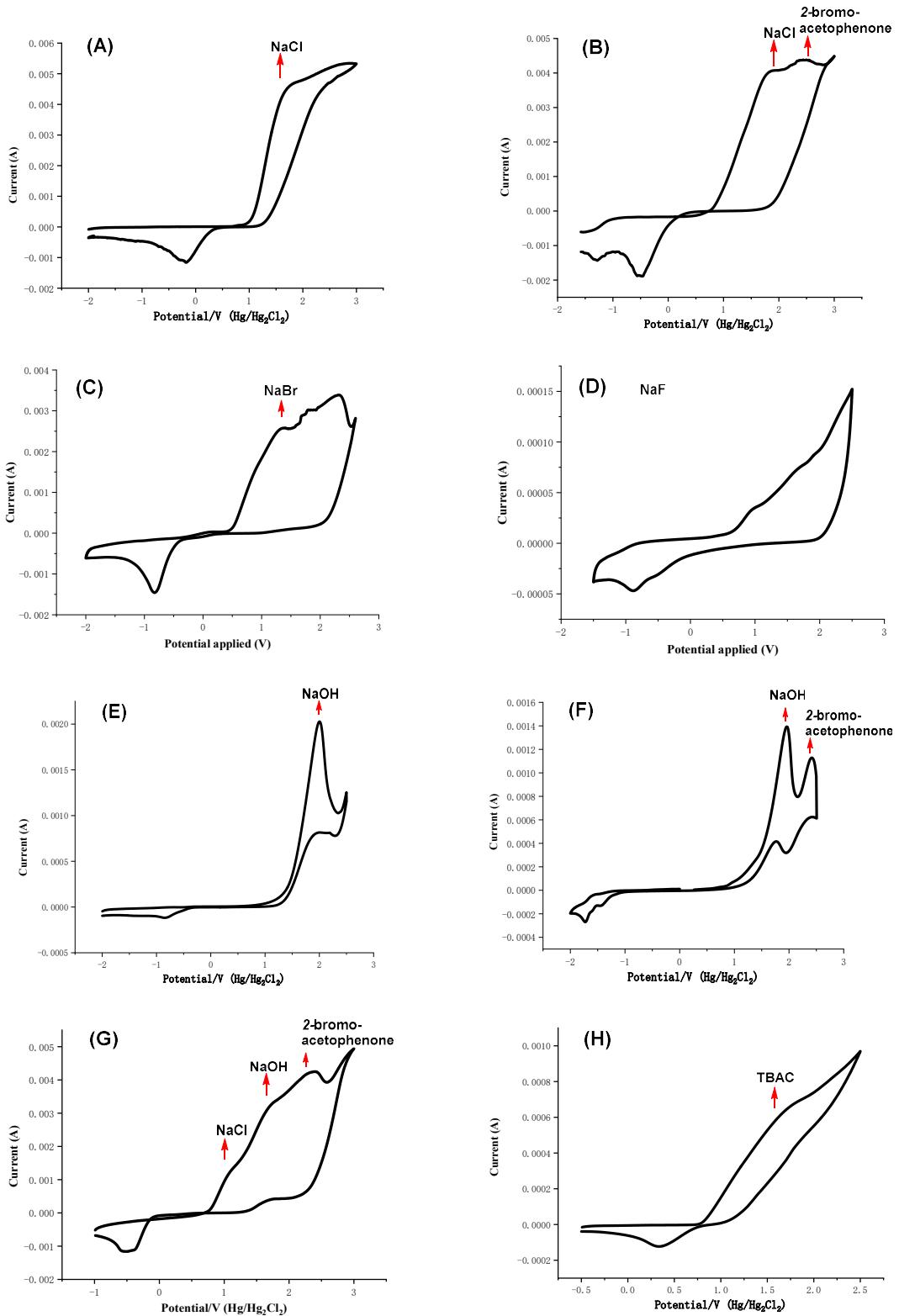
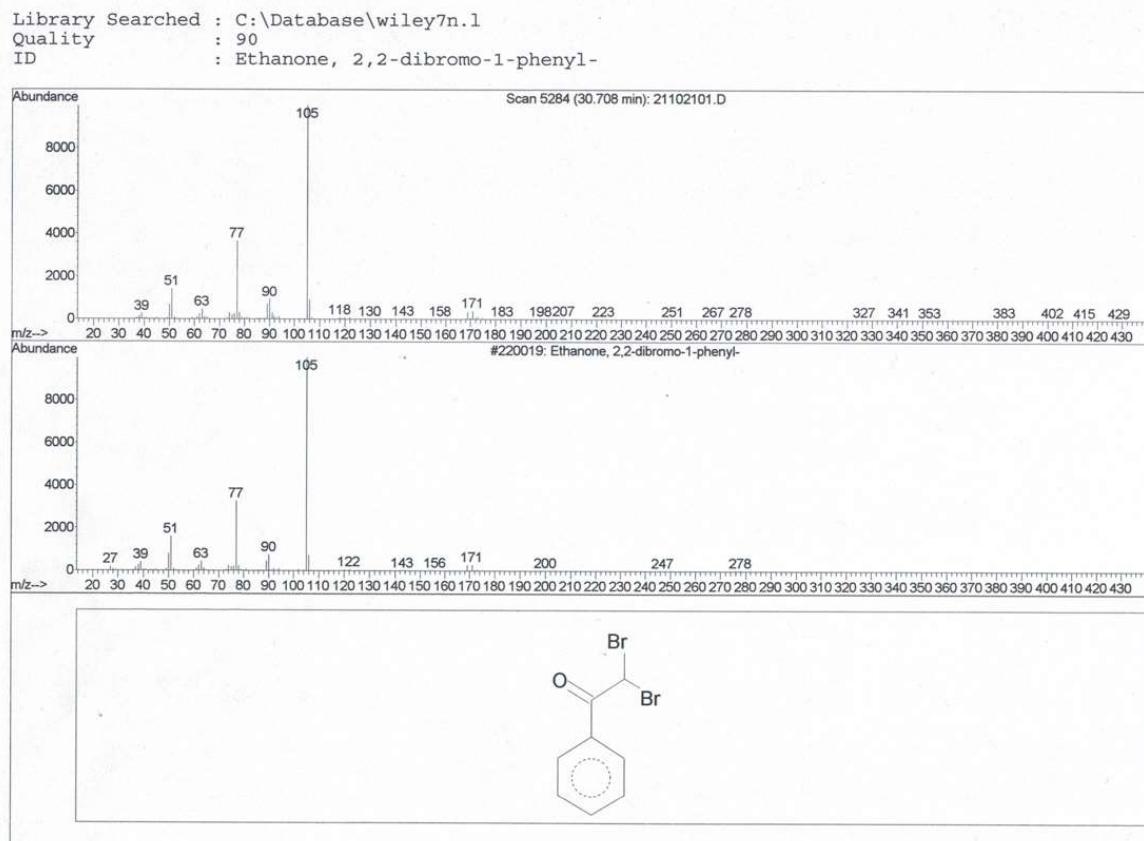
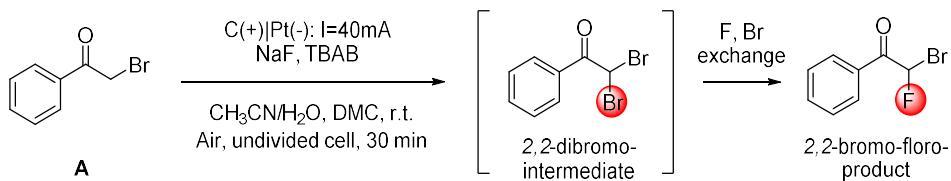


Figure S12. Cyclic voltammograms maps

3.10 GC-MS experiment for electro-fluorination

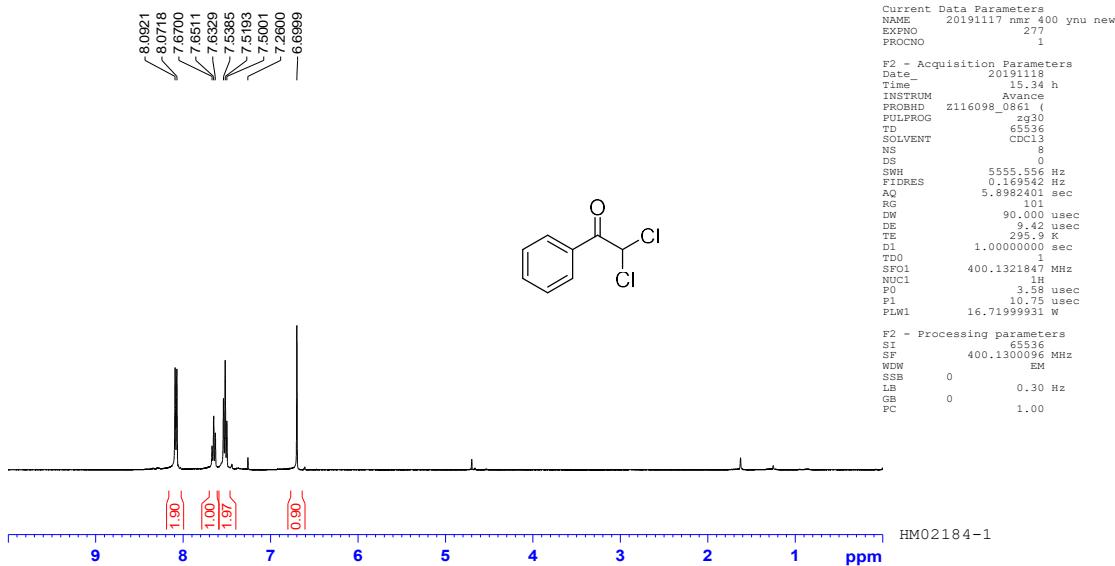
In an oven-dried undivided two-necked flask (50 mL) equipped with a stir bar, 2-bromo-acetophenone (200 mg, 1.0 mmol), TBAB (50 mg, 0.2 mmol), KBr (1.0 mol·L⁻¹, 6.5 mmol), CH₃CN (5.0 mL), H₂O (5.0 mL), and DMC (2.5 mL) were combined and added. The flask was equipped with a platinum plate (10 mm × 10 mm × 0.2 mm) as the cathodic electrode and a carbon plate (10 mm × 10 mm × 2 mm) as the anodic electrode. The reaction was in an ambient atmosphere. The reaction mixture was stirred and electrolyzed at a constant current of 40 mA under room temperature for 30 min. The reaction was conducted with the GC-MS.



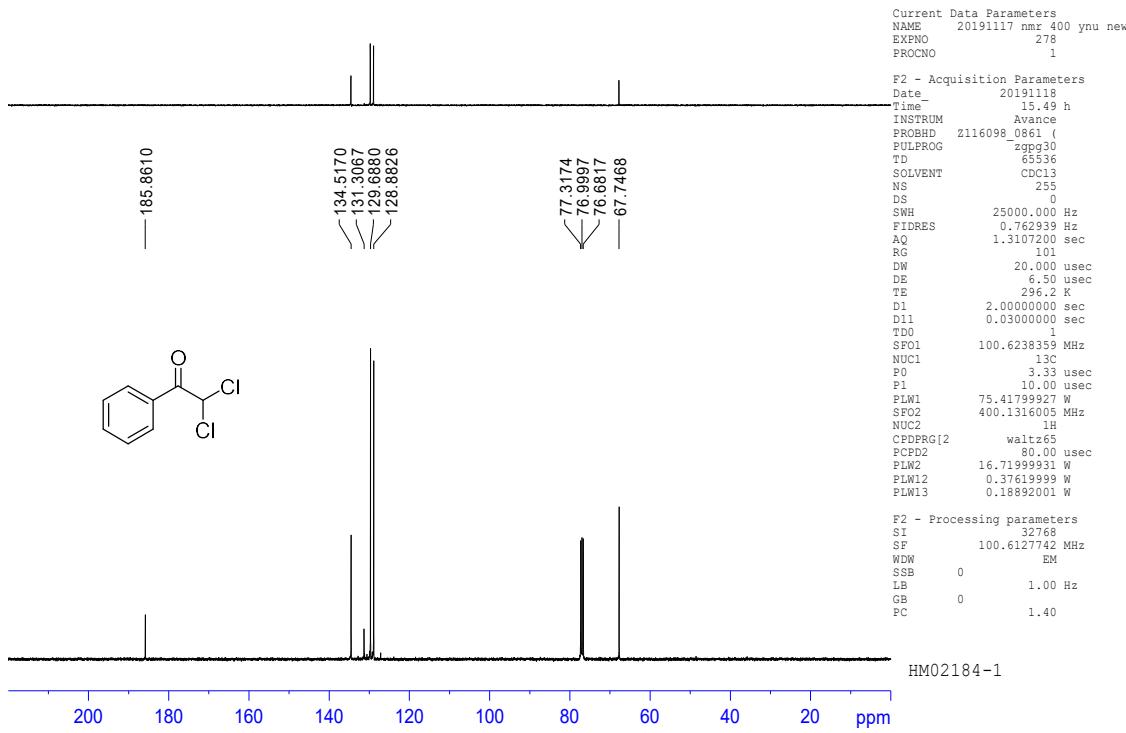
Scheme S14. GC-MS analysis of 2,2-dibromo-intermediate in TBAB involved fluorination

4. ^1H -NMR, ^{13}C -NMR, and ^{19}F -NMR spectra

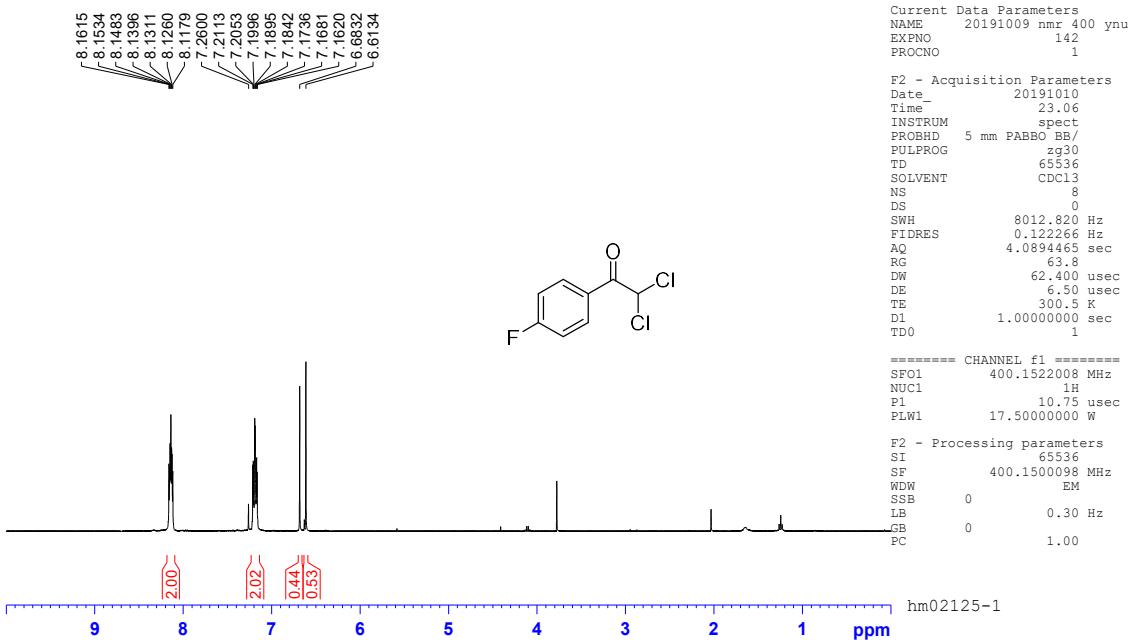
^1H -NMR (400 MHz, CDCl_3)



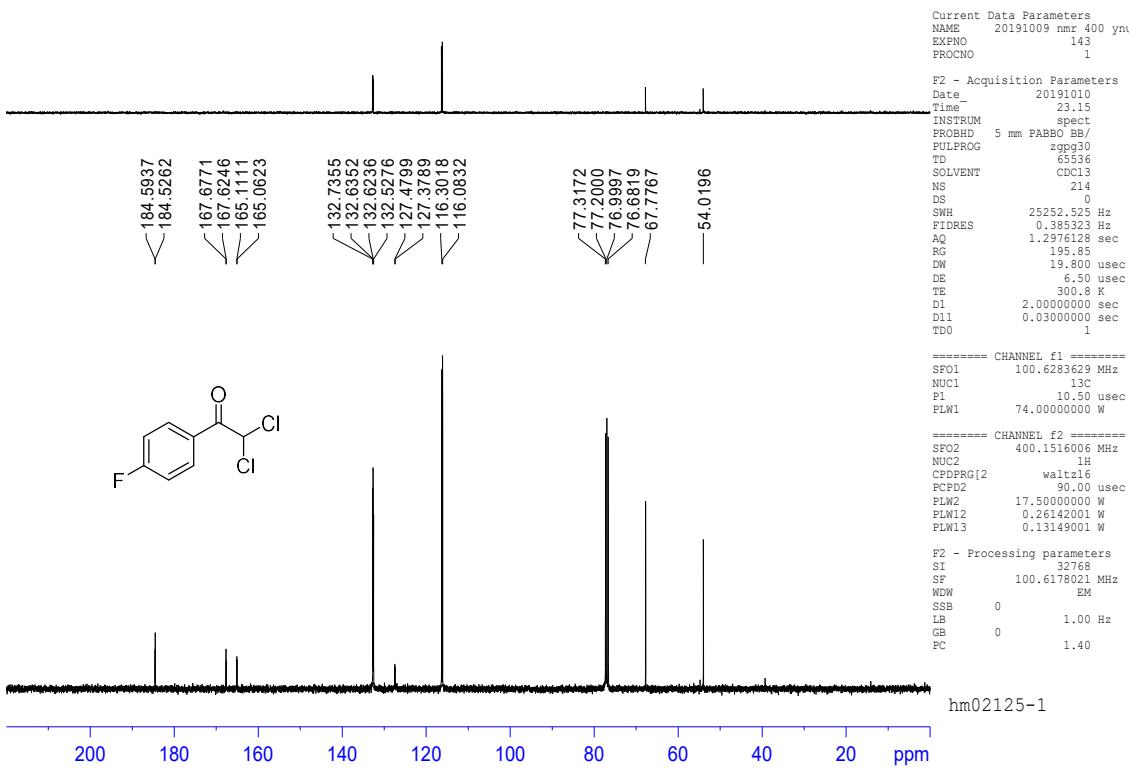
^{13}C -NMR (100 MHz, CDCl_3)



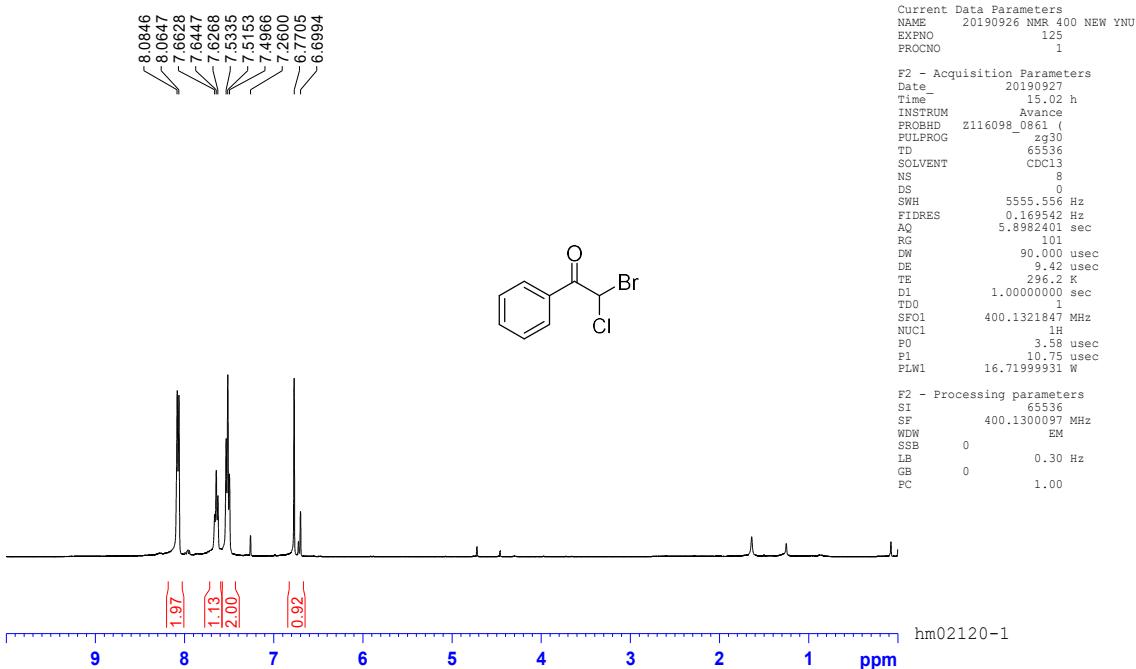
¹H-NMR (400 MHz, CDCl₃)



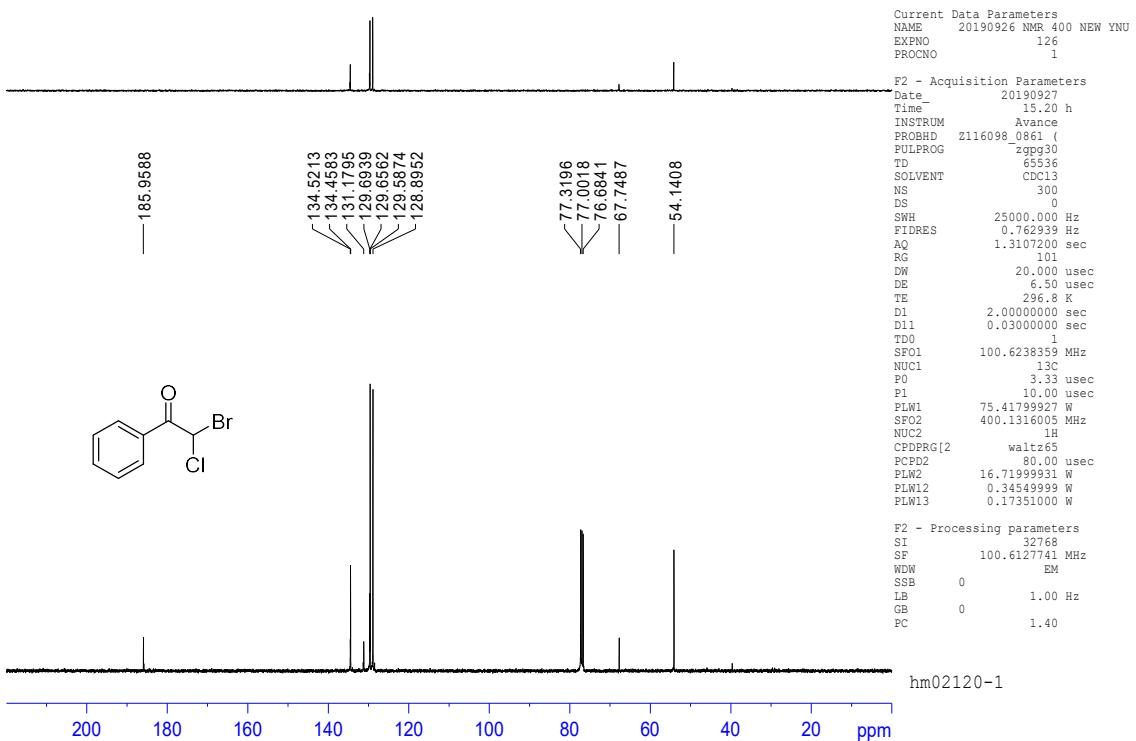
¹³C-NMR (100 MHz, CDCl₃)



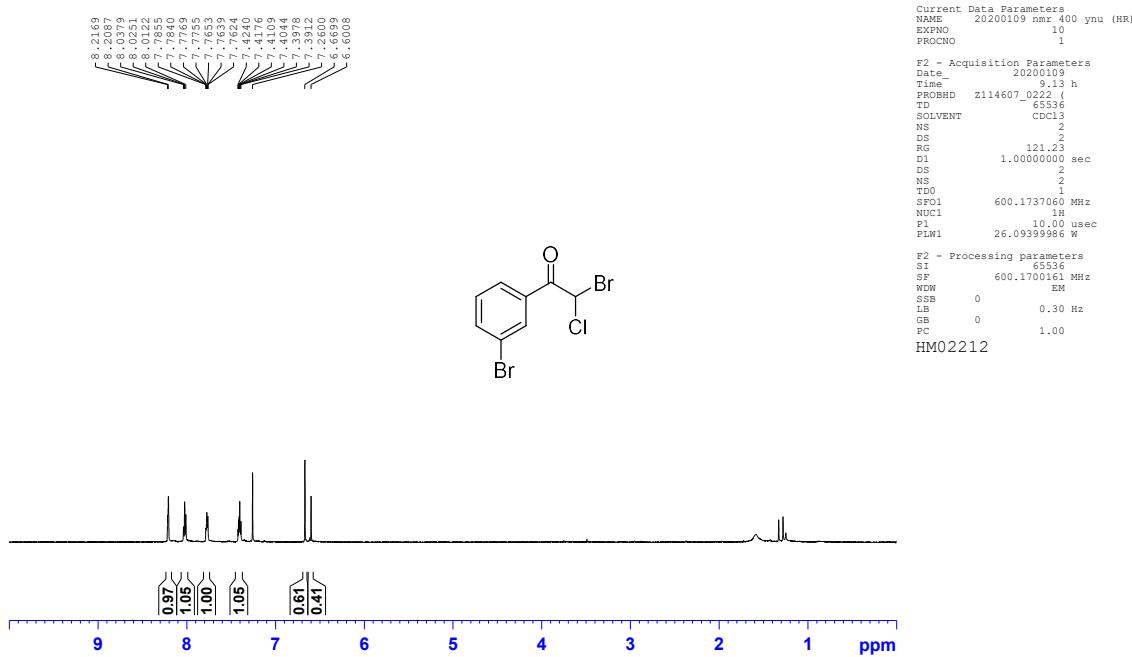
¹H-NMR (400 MHz, CDCl₃)



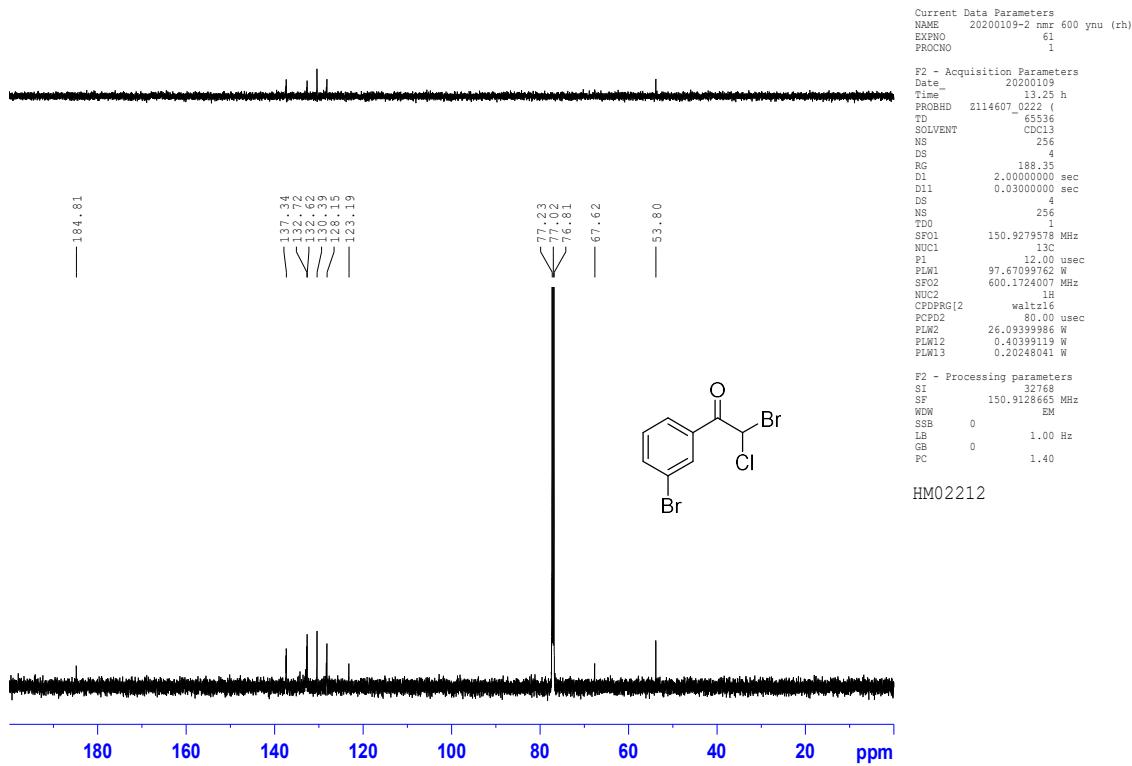
¹³C-NMR (100 MHz, CDCl₃)



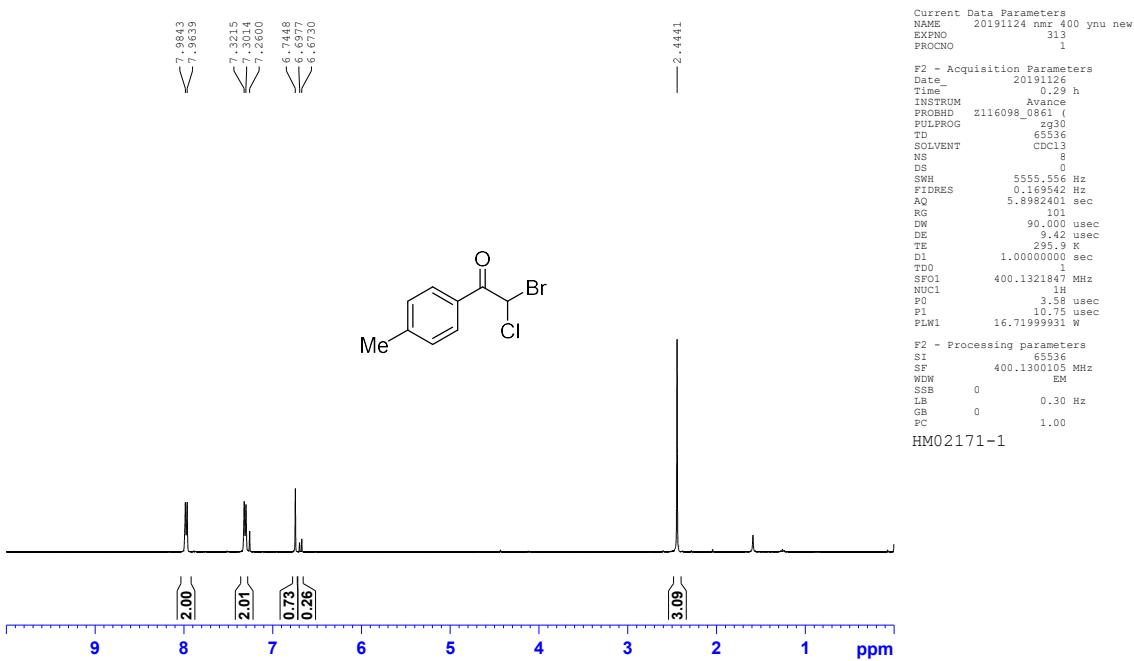
¹H-NMR (400 MHz, CDCl₃)



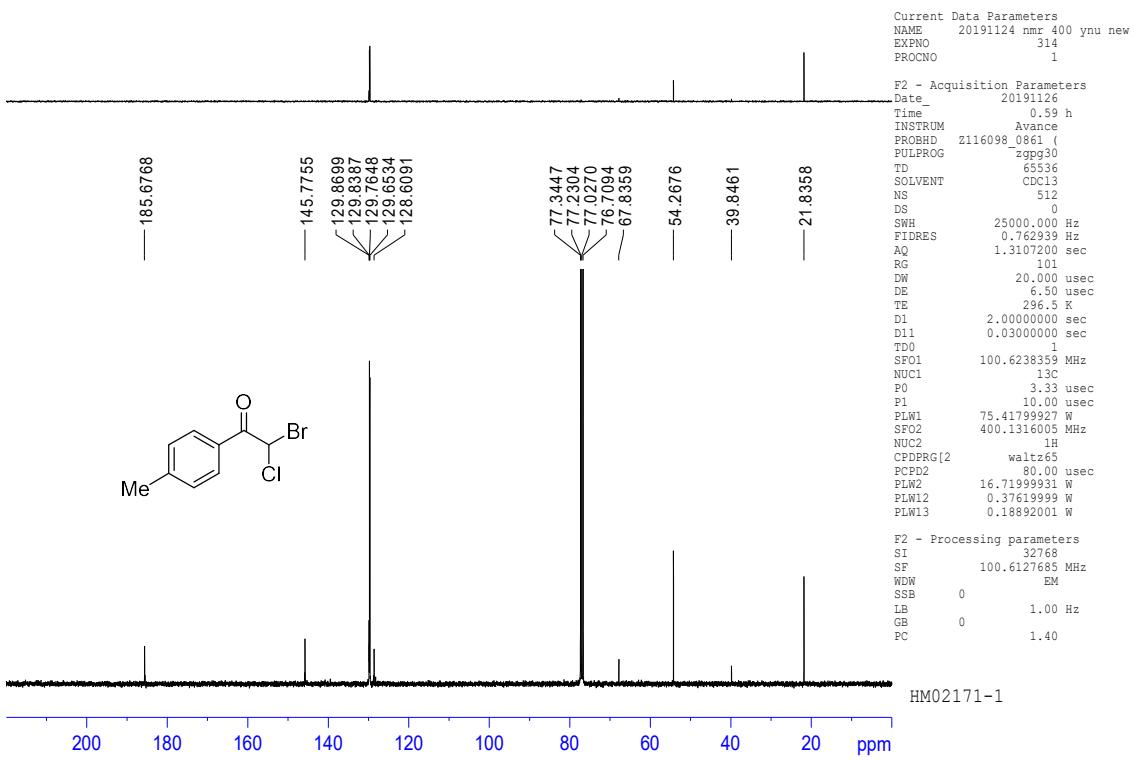
¹³C-NMR (150 MHz, CDCl₃)



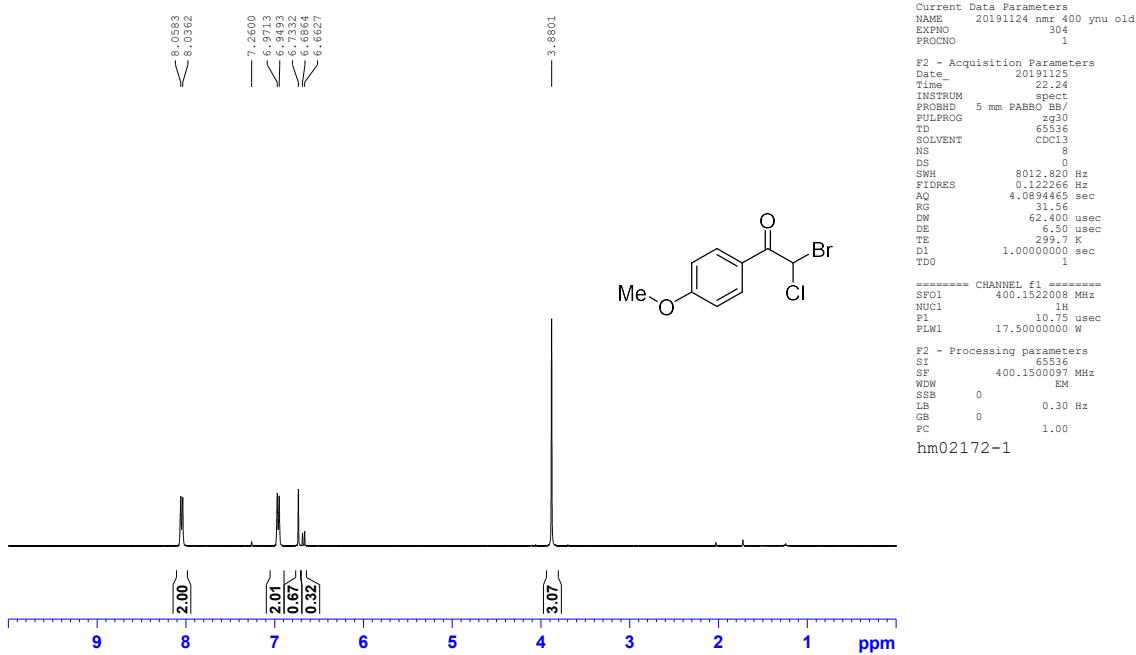
¹H-NMR (400 MHz, CDCl₃)



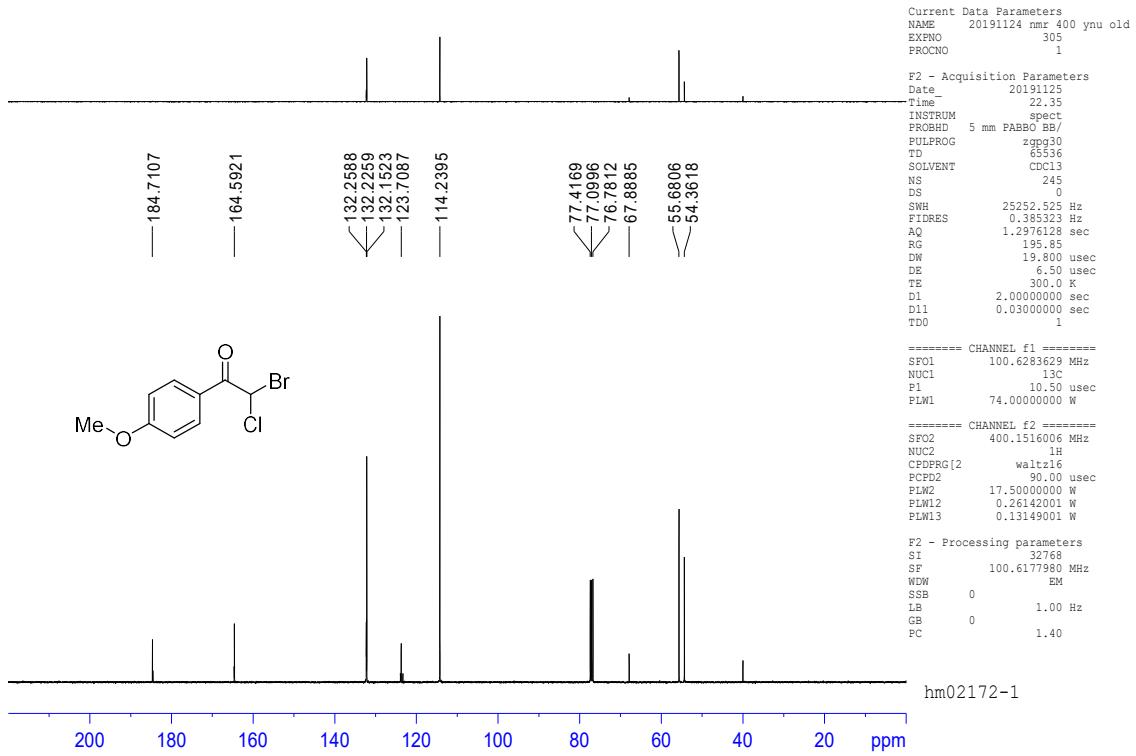
¹³C-NMR (100 MHz, CDCl₃)



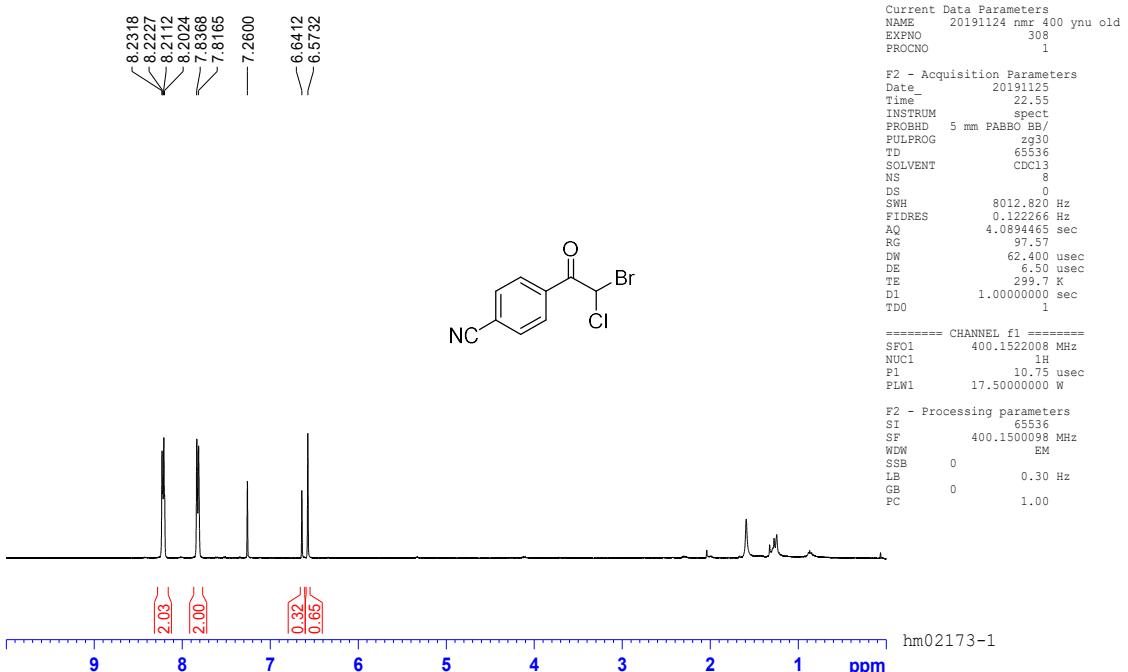
¹H-NMR (400 MHz, CDCl₃)



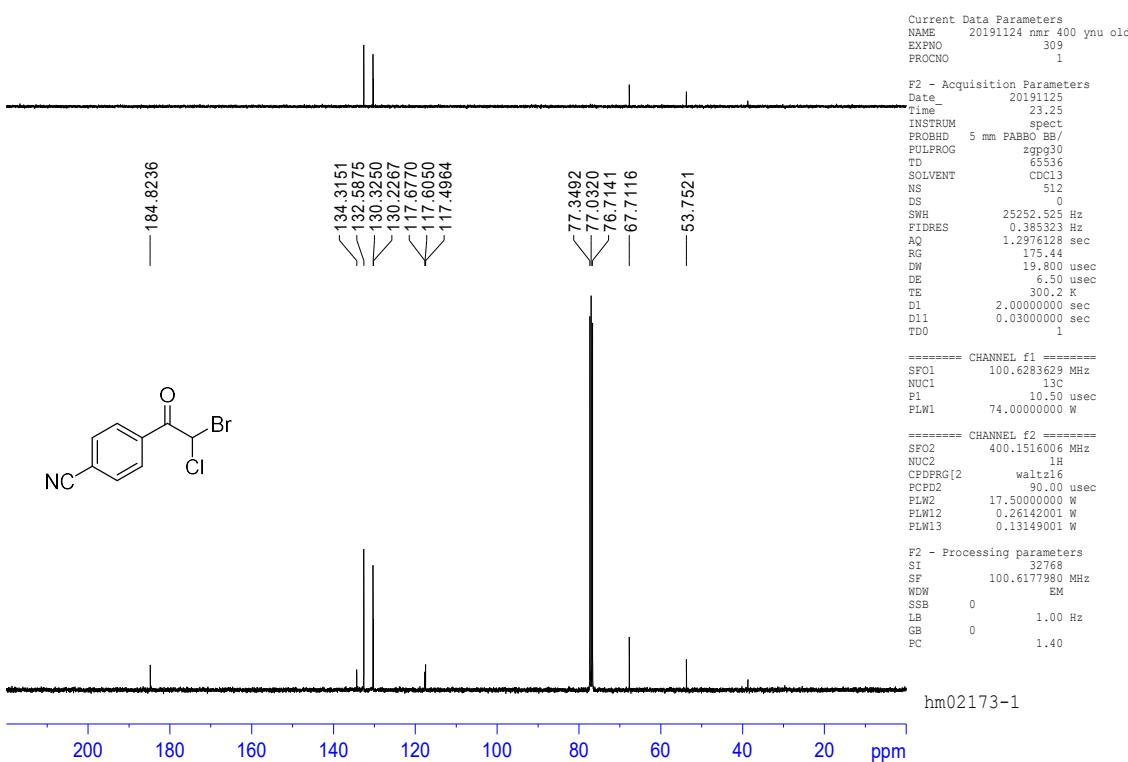
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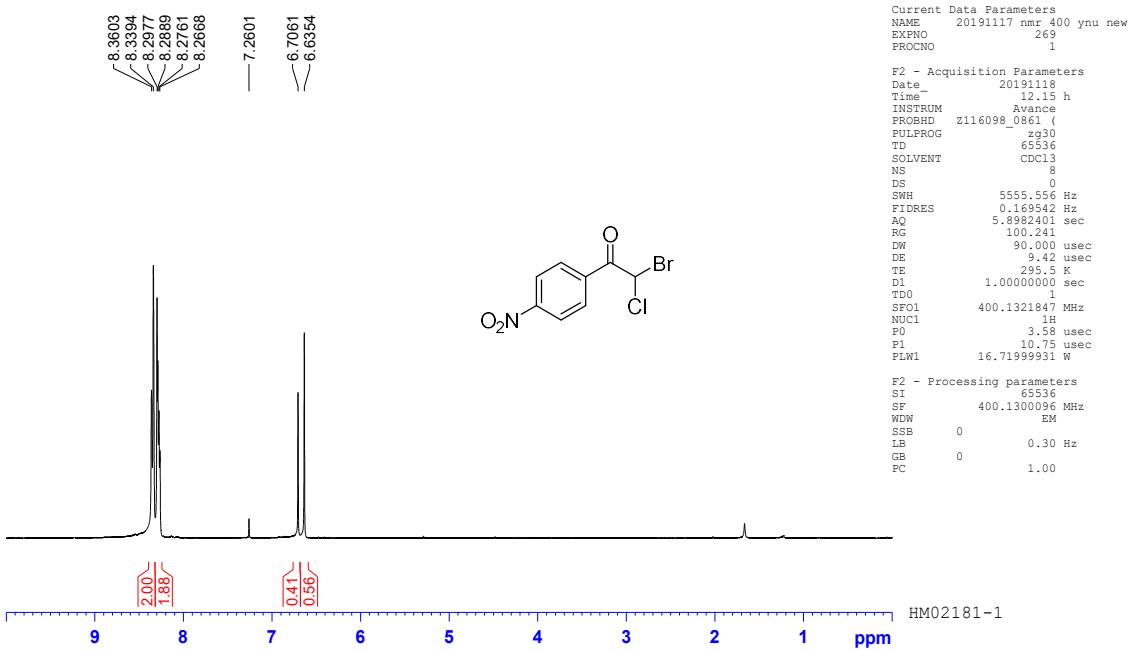
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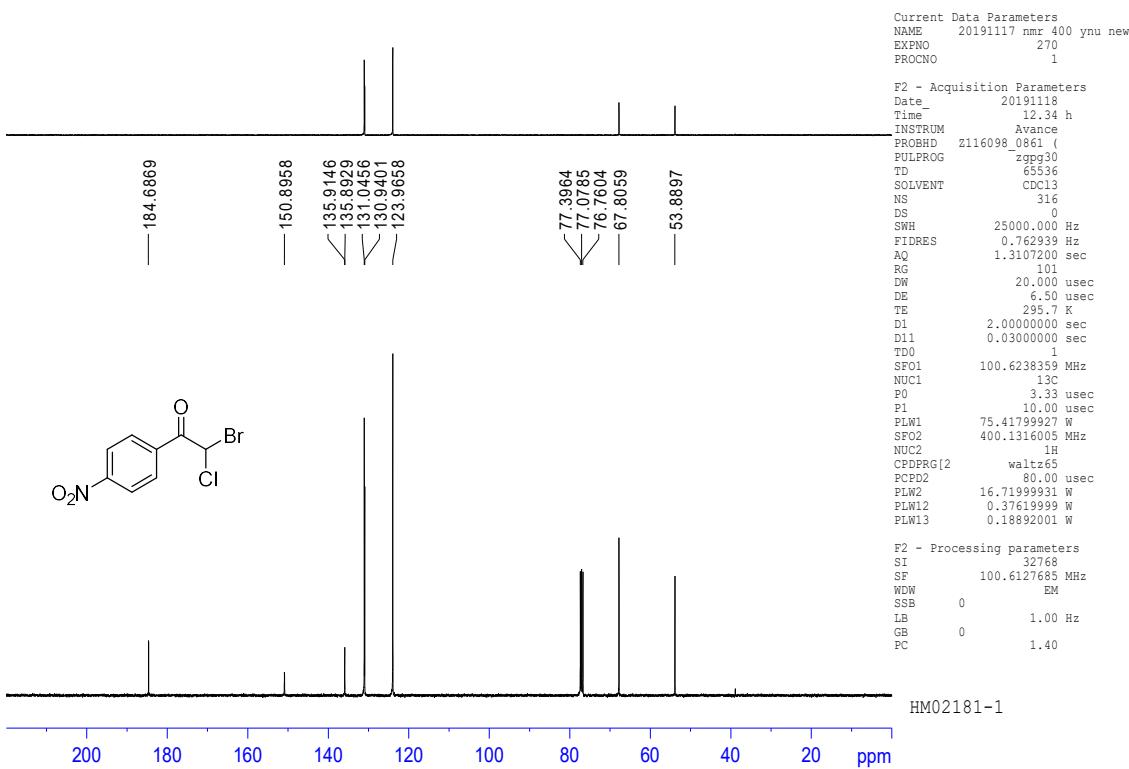
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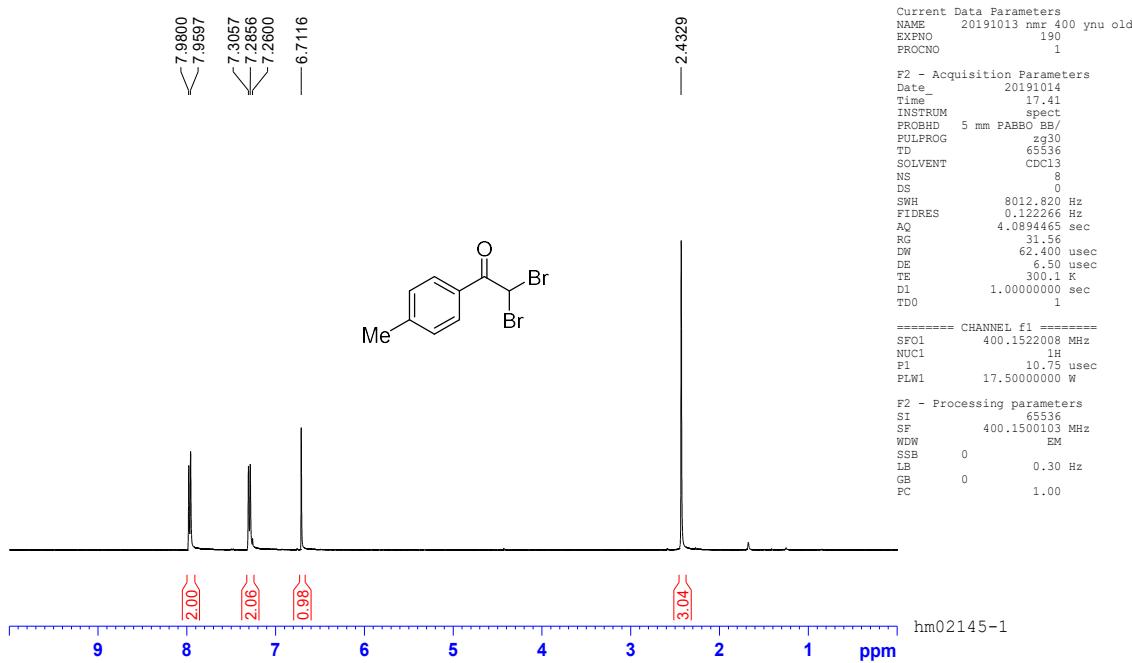
¹H-NMR (400 MHz, CDCl₃)



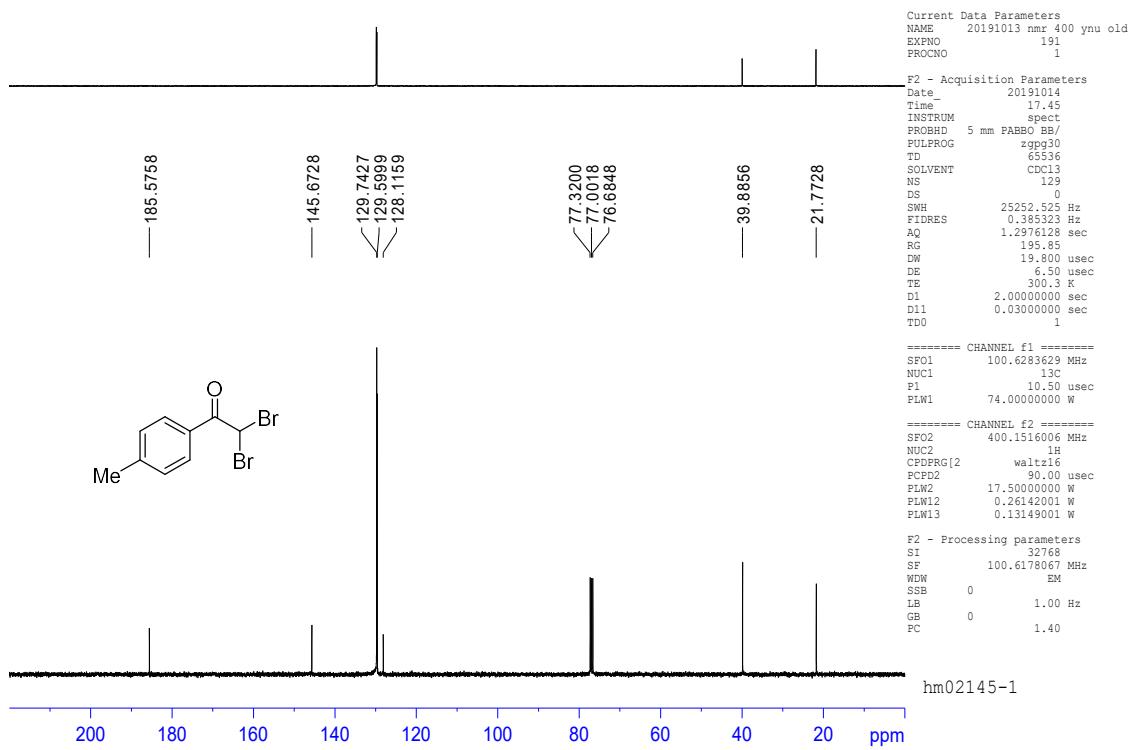
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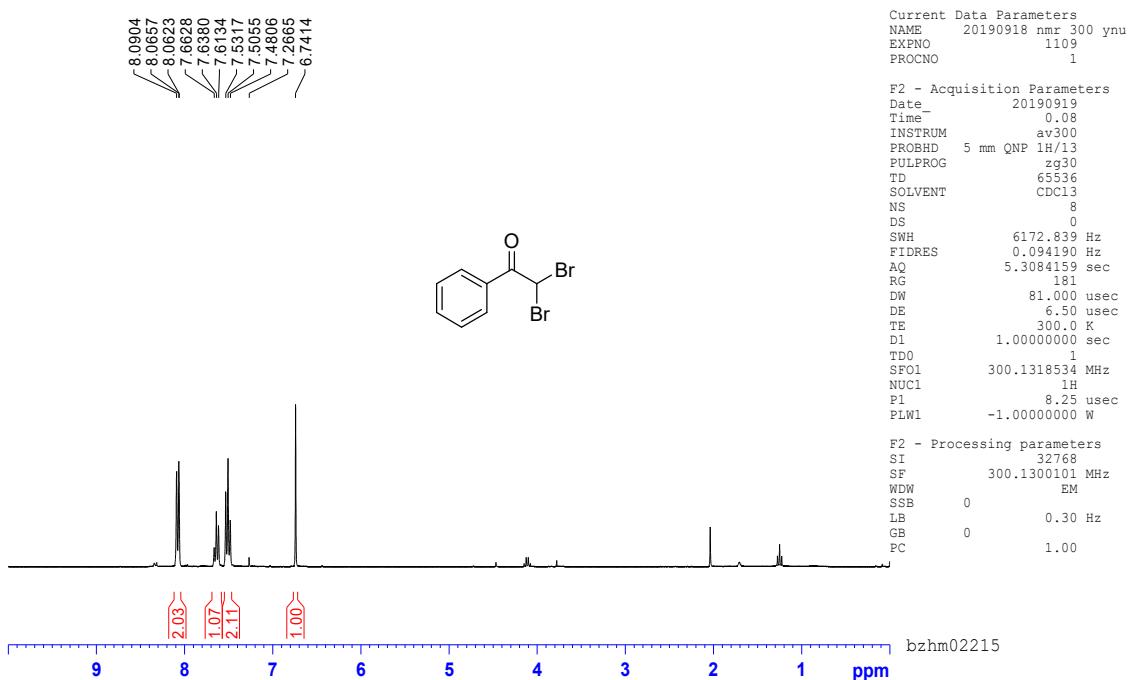
¹H-NMR (400 MHz, CDCl₃)



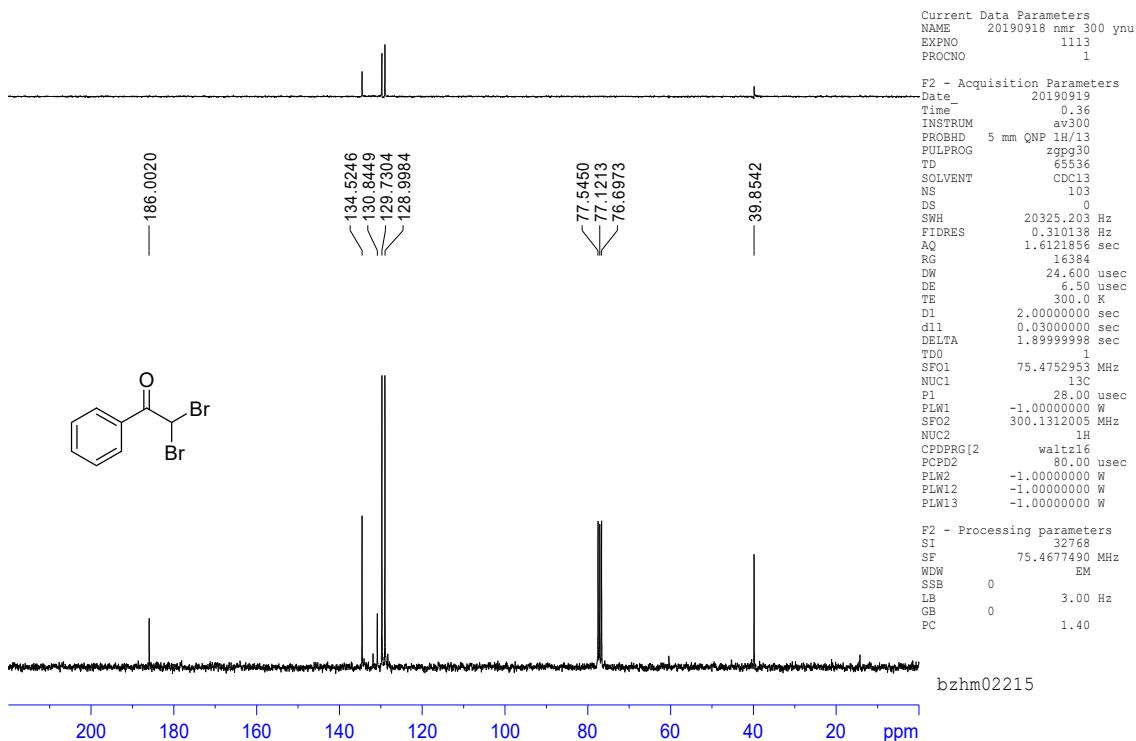
¹³C-NMR (100 MHz, CDCl₃)



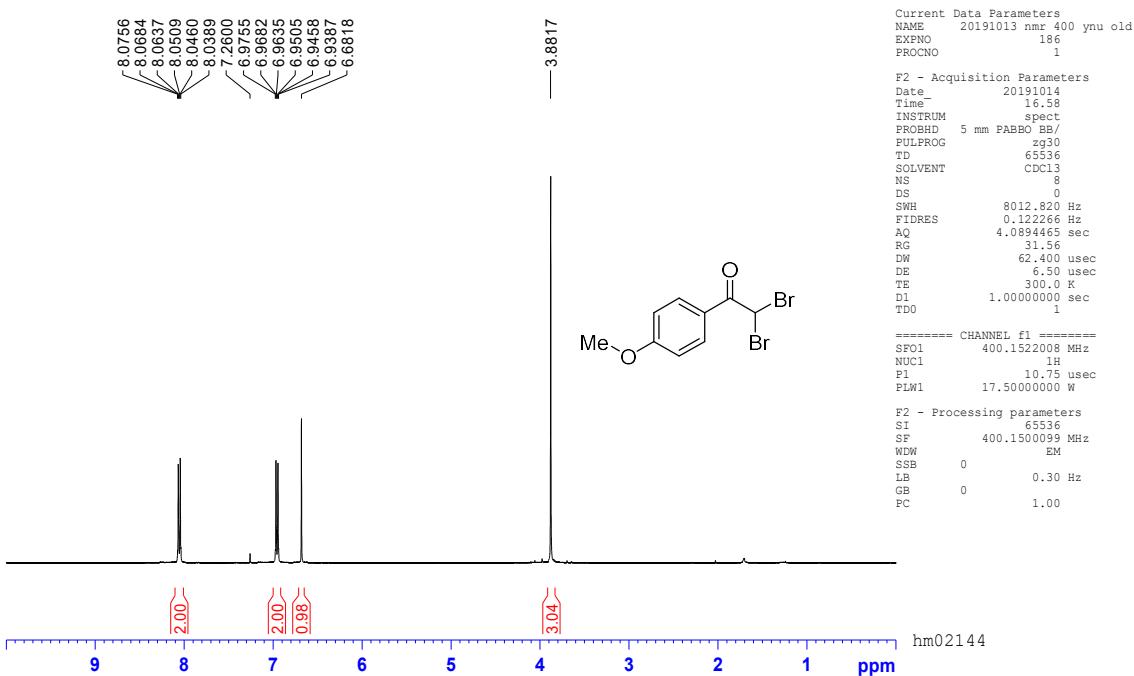
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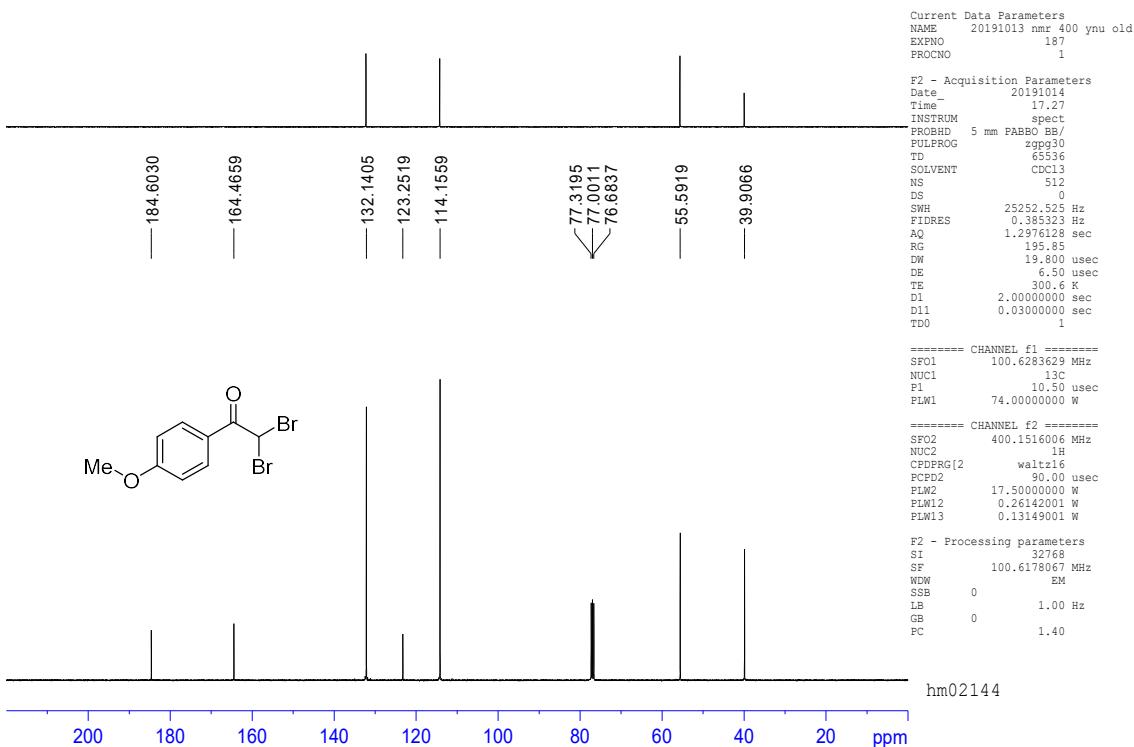
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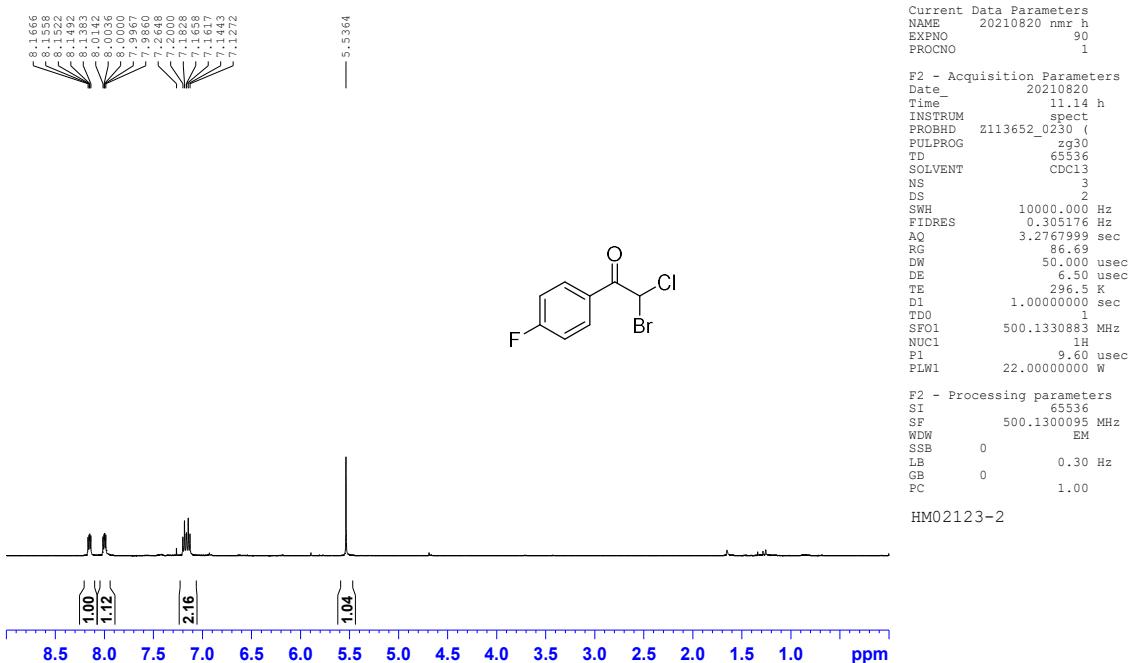
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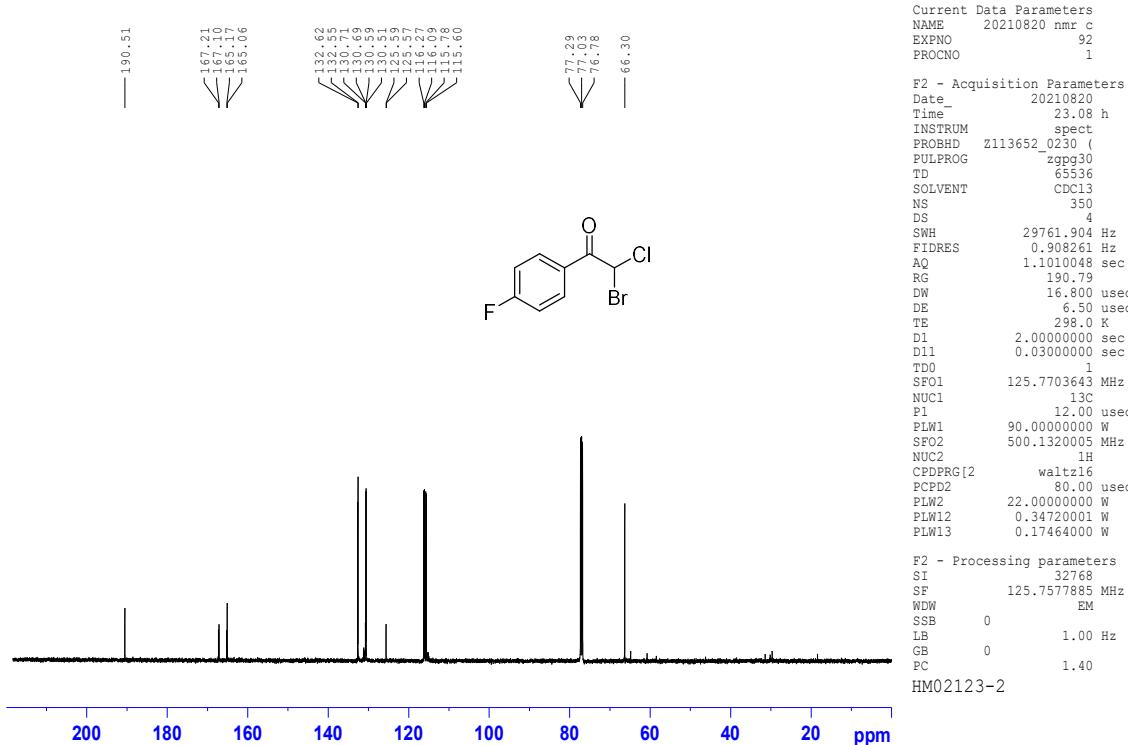
¹³C-NMR (100 MHz, CDCl₃)



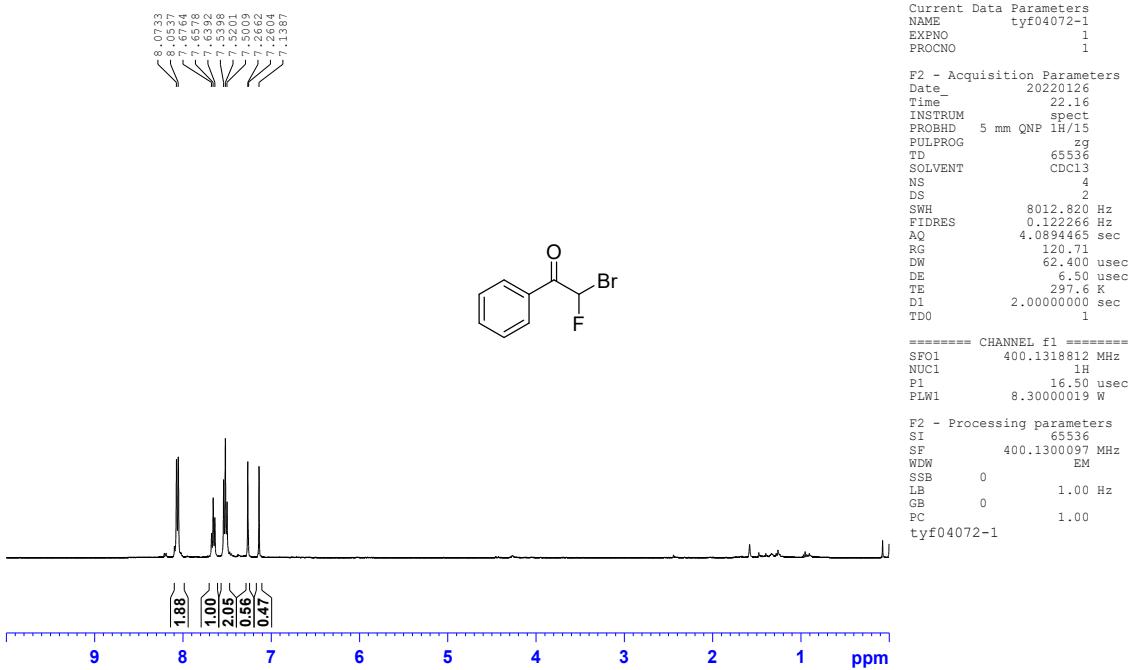
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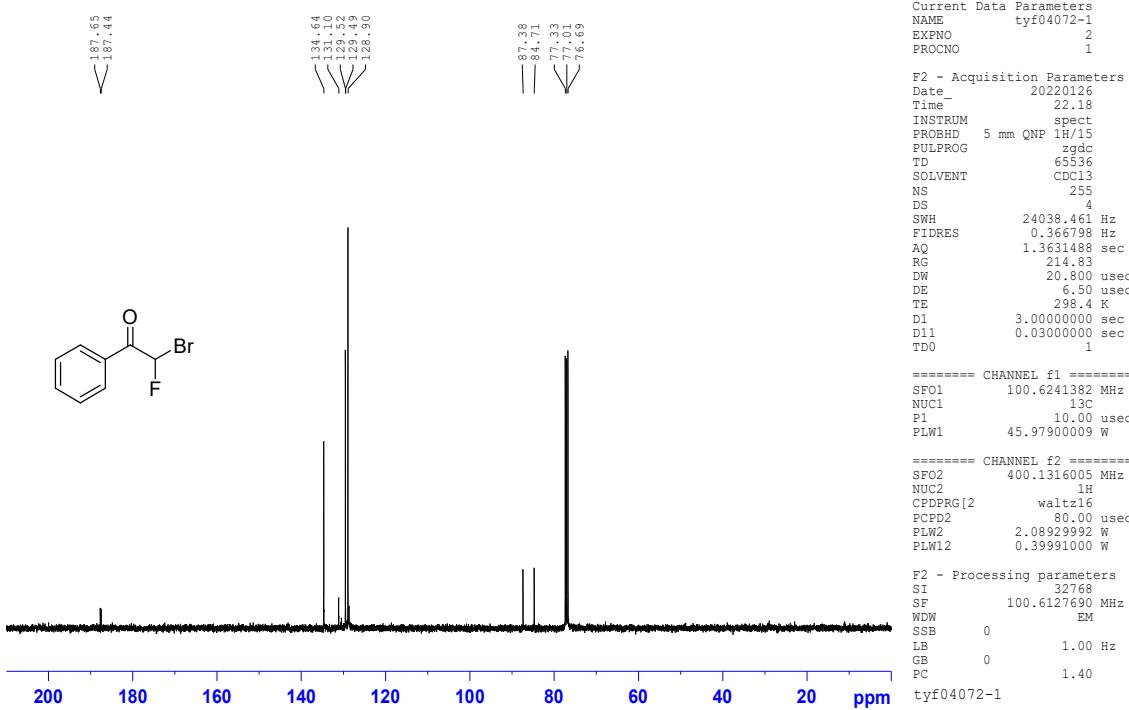
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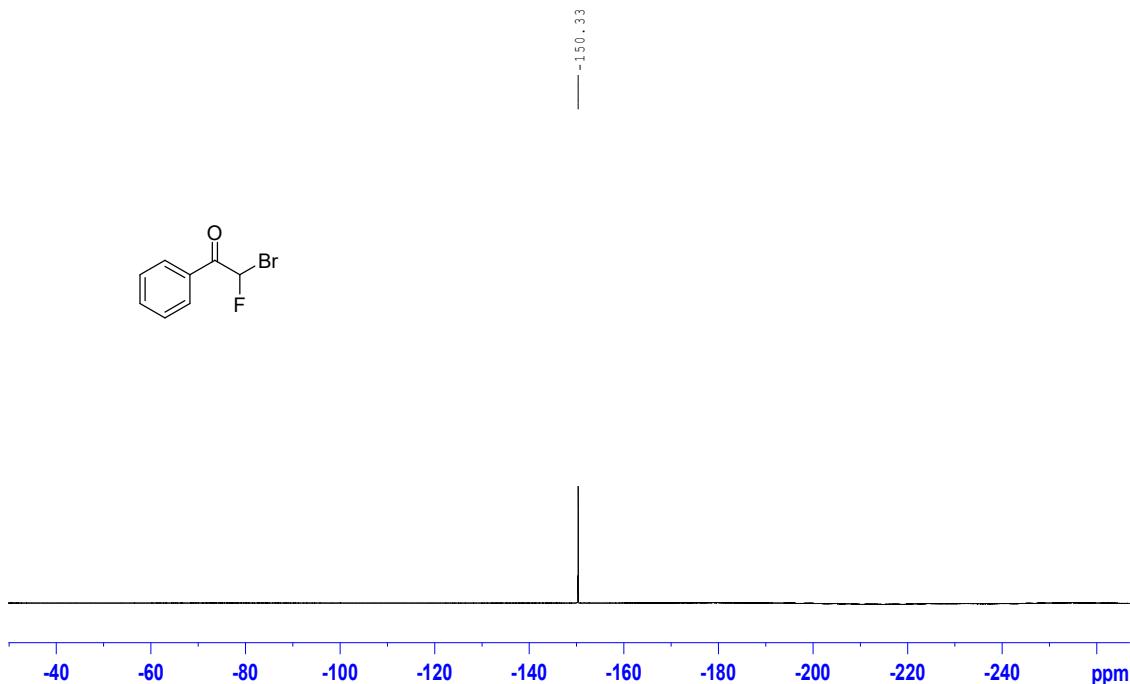
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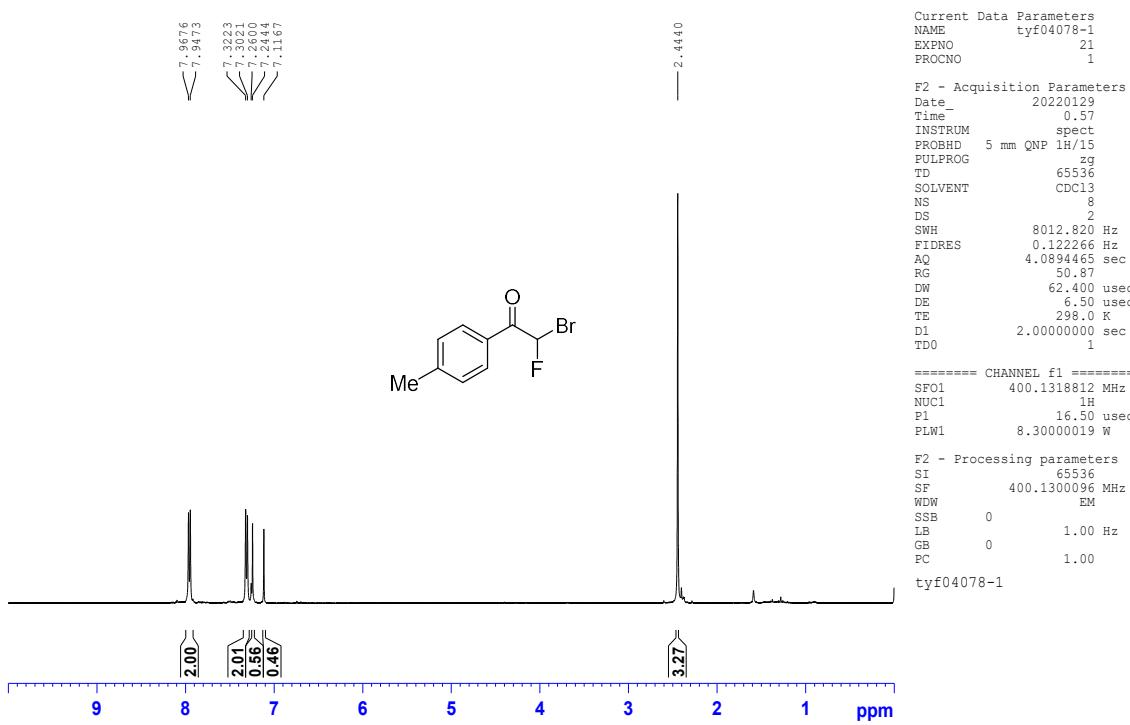
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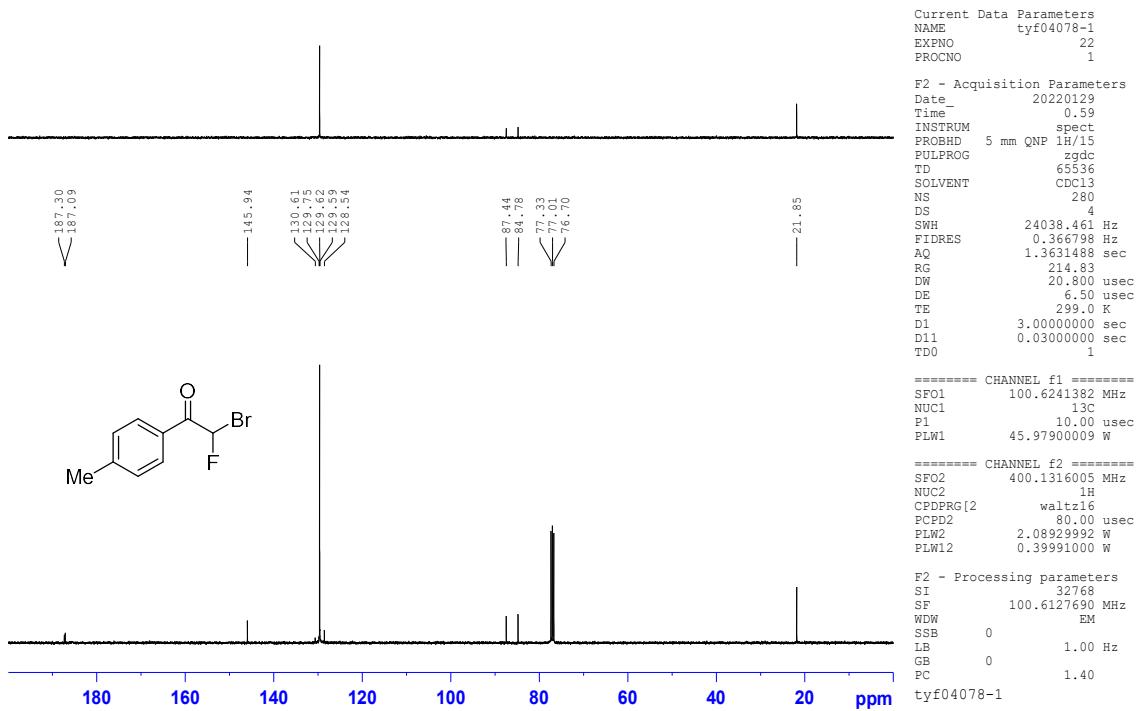
¹⁹F NMR (376 MHz, CDCl₃)



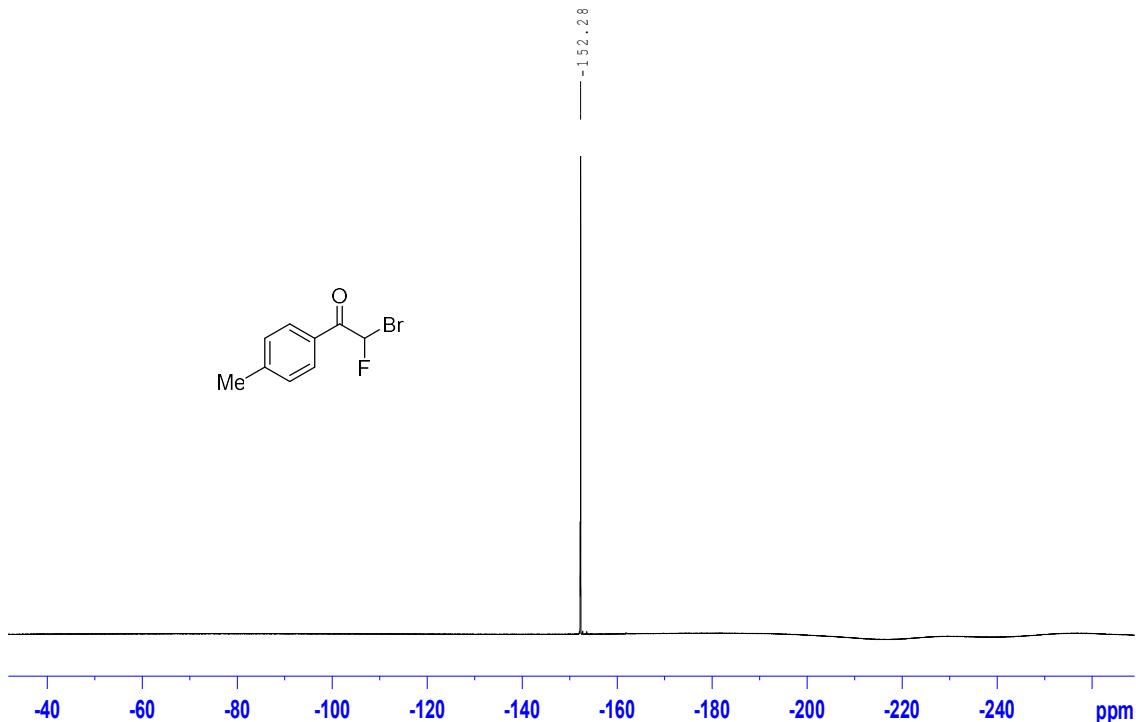
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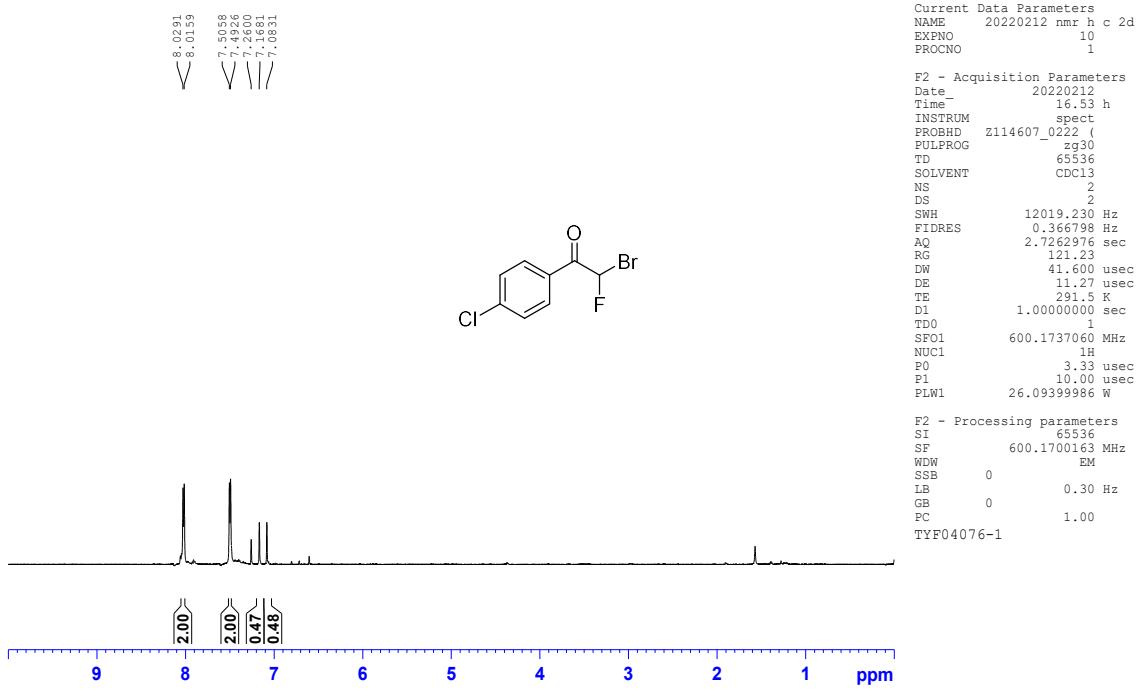
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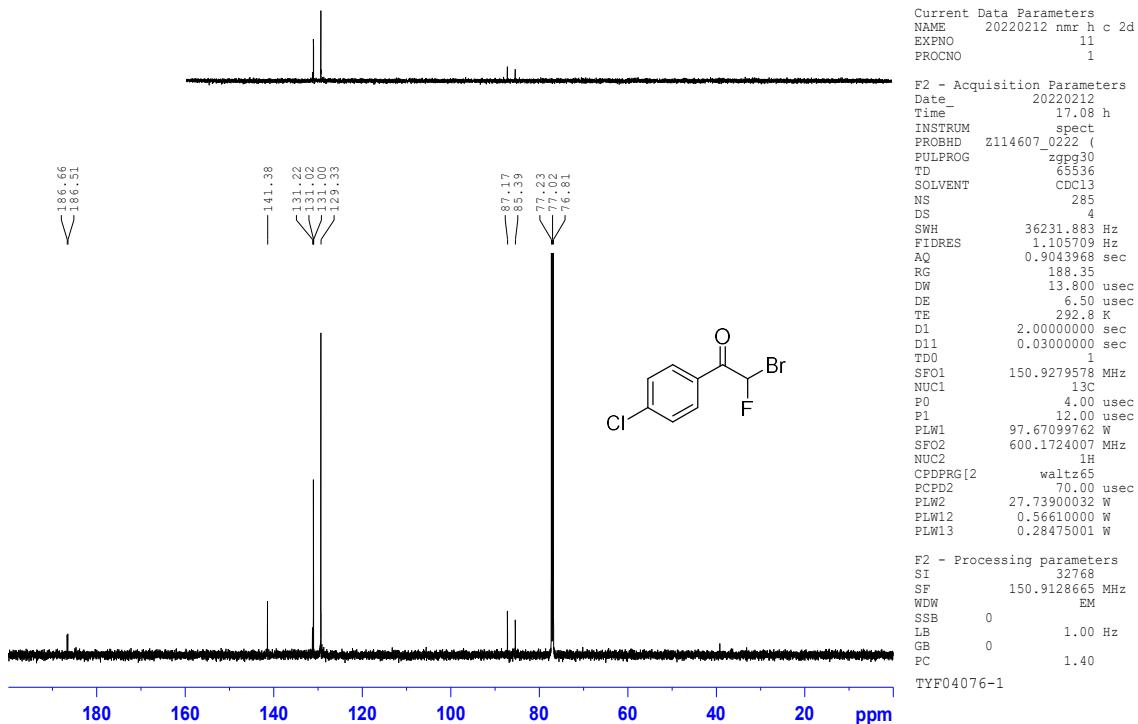
¹⁹F NMR (376 MHz, CDCl₃)



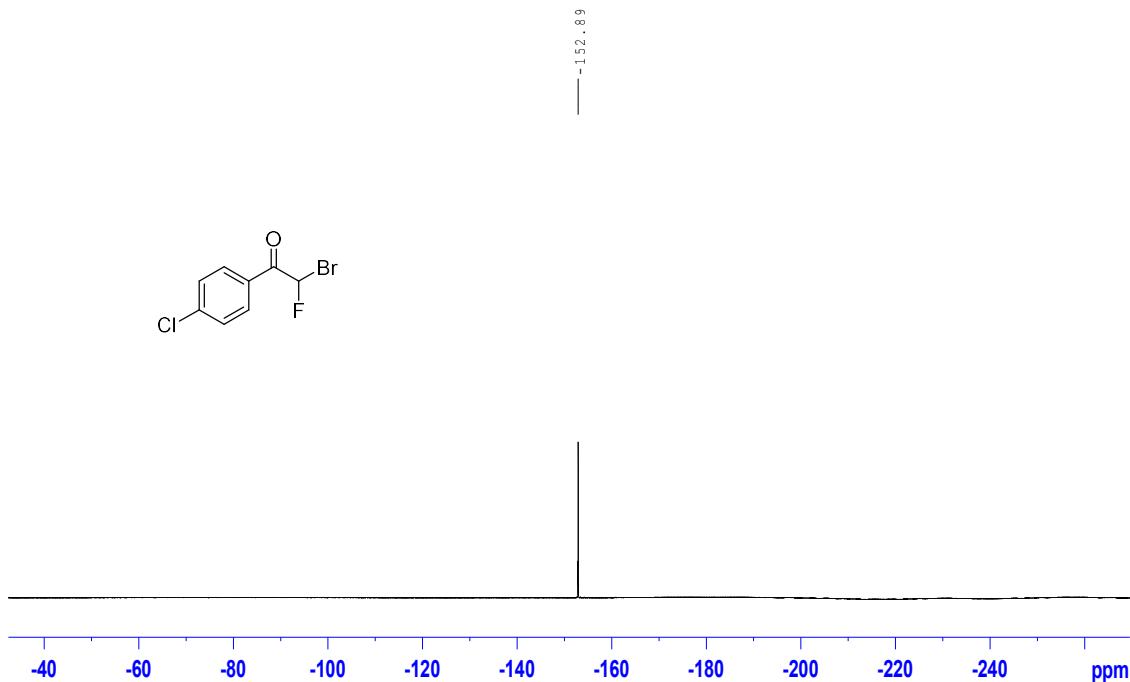
¹H-NMR (600 MHz, CDCl₃)



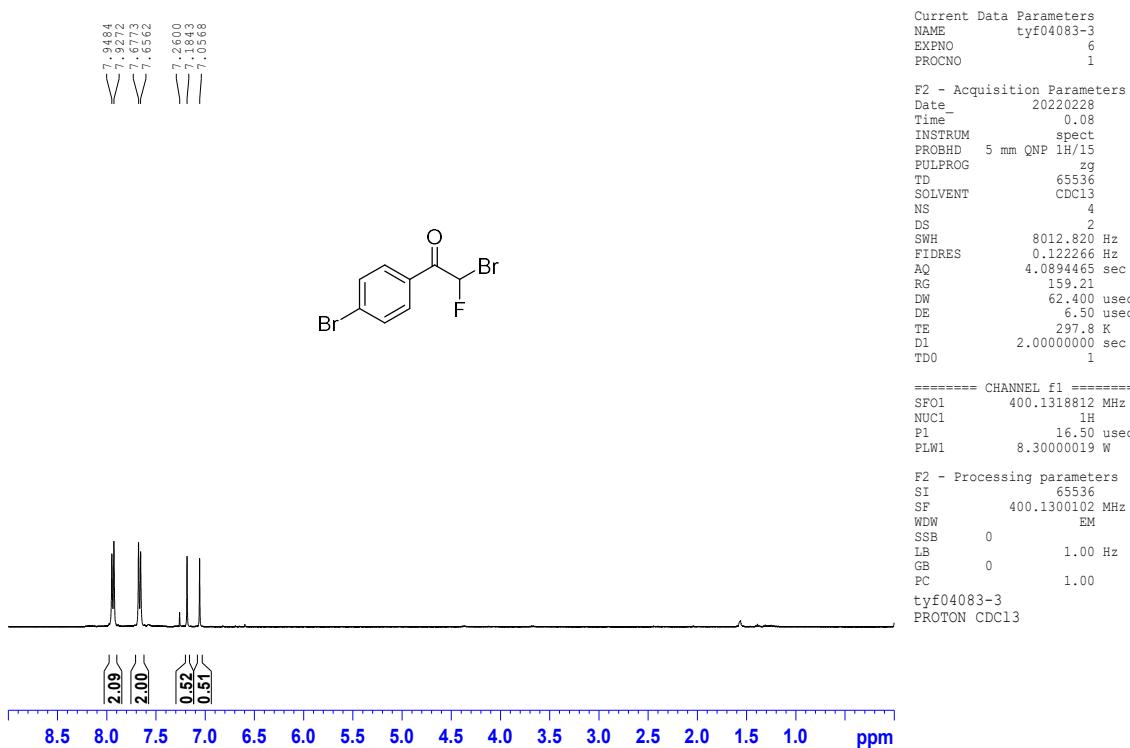
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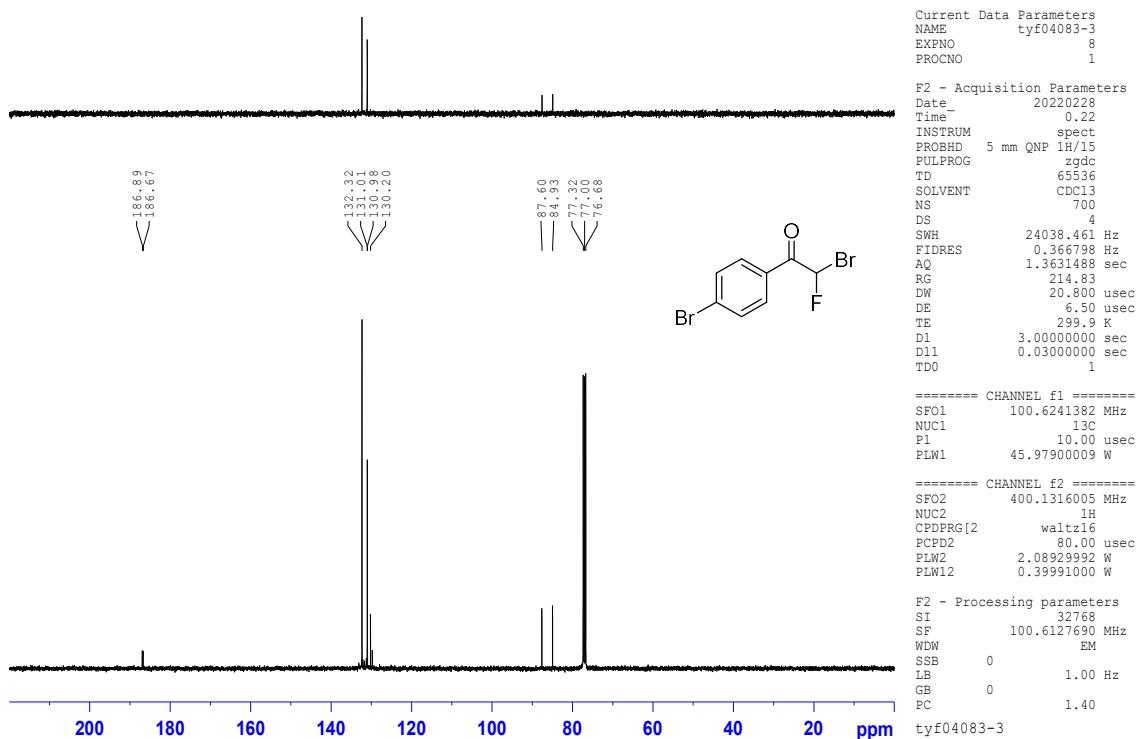
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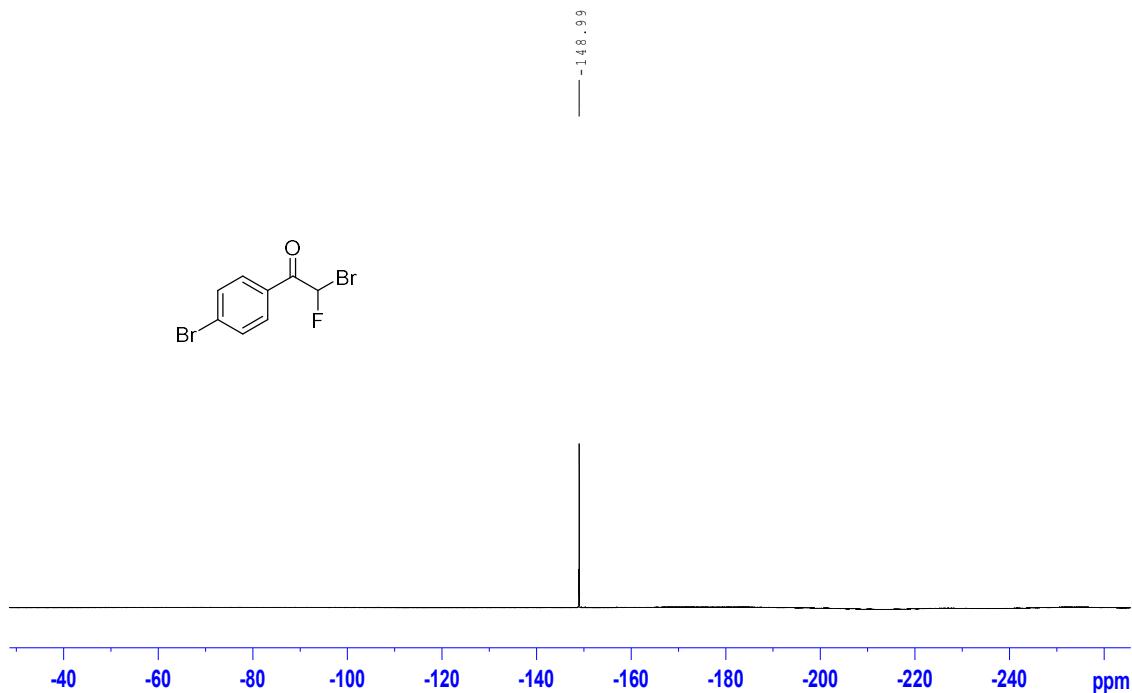
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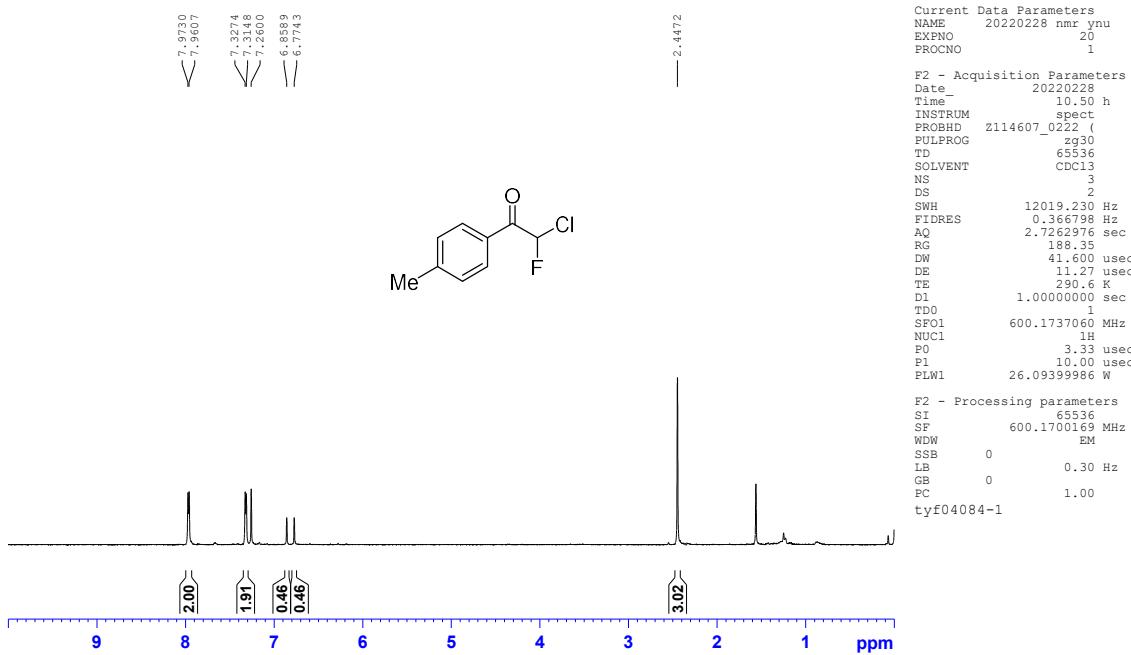
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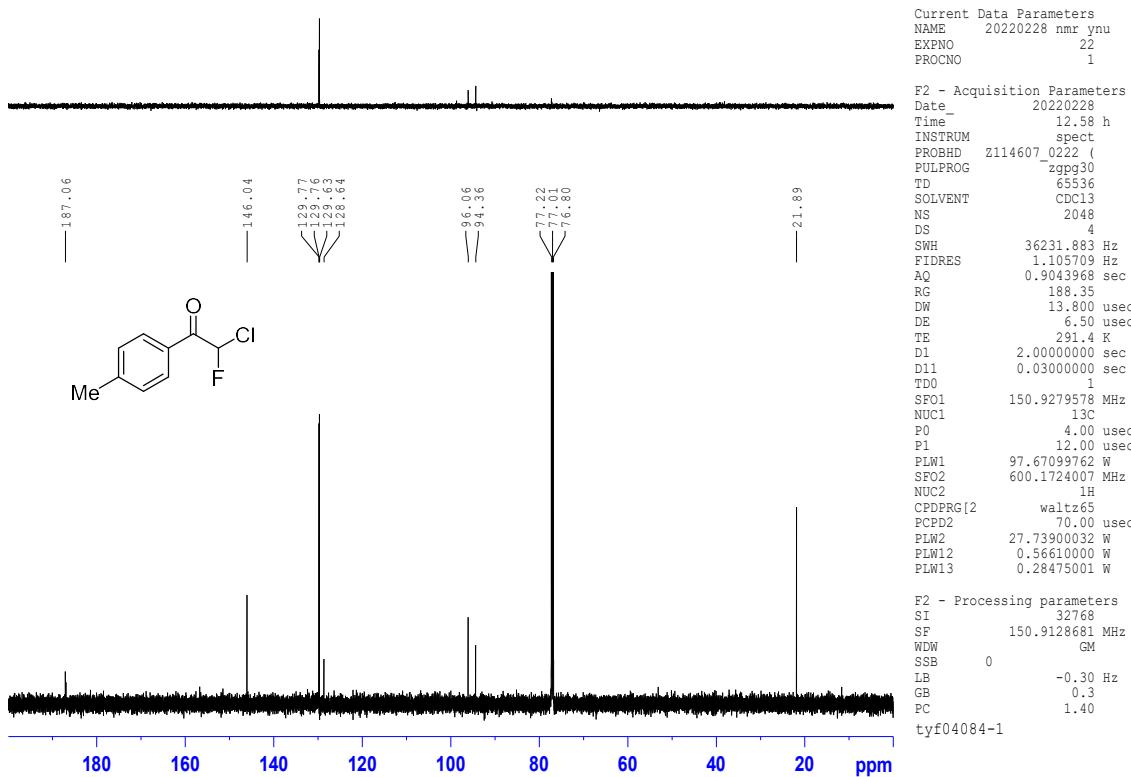
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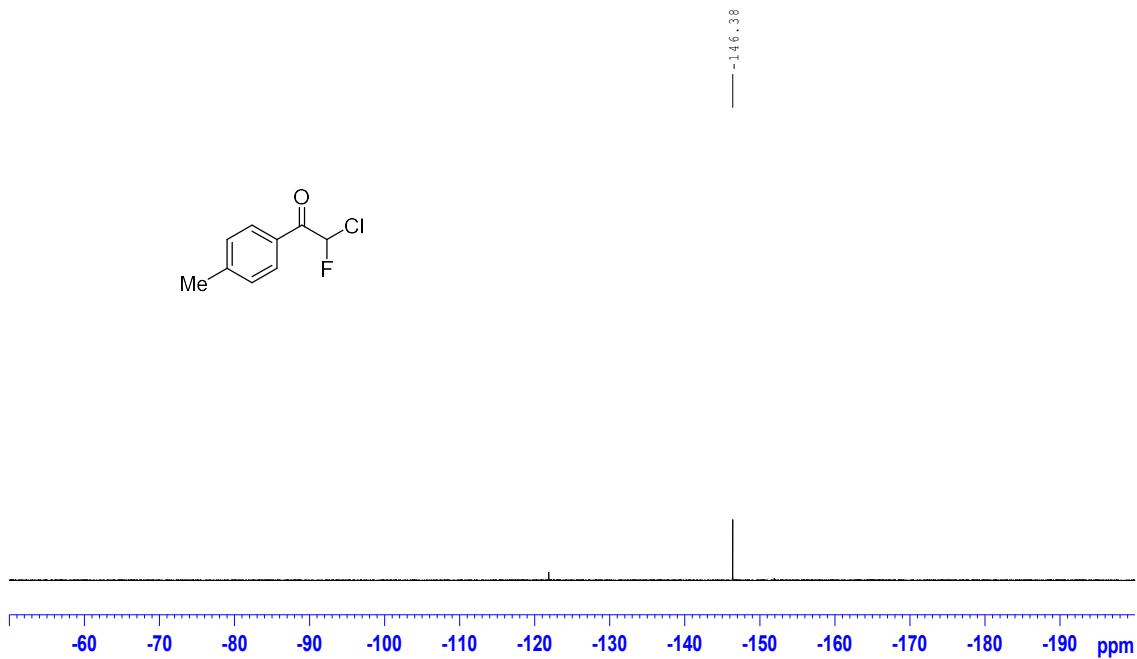
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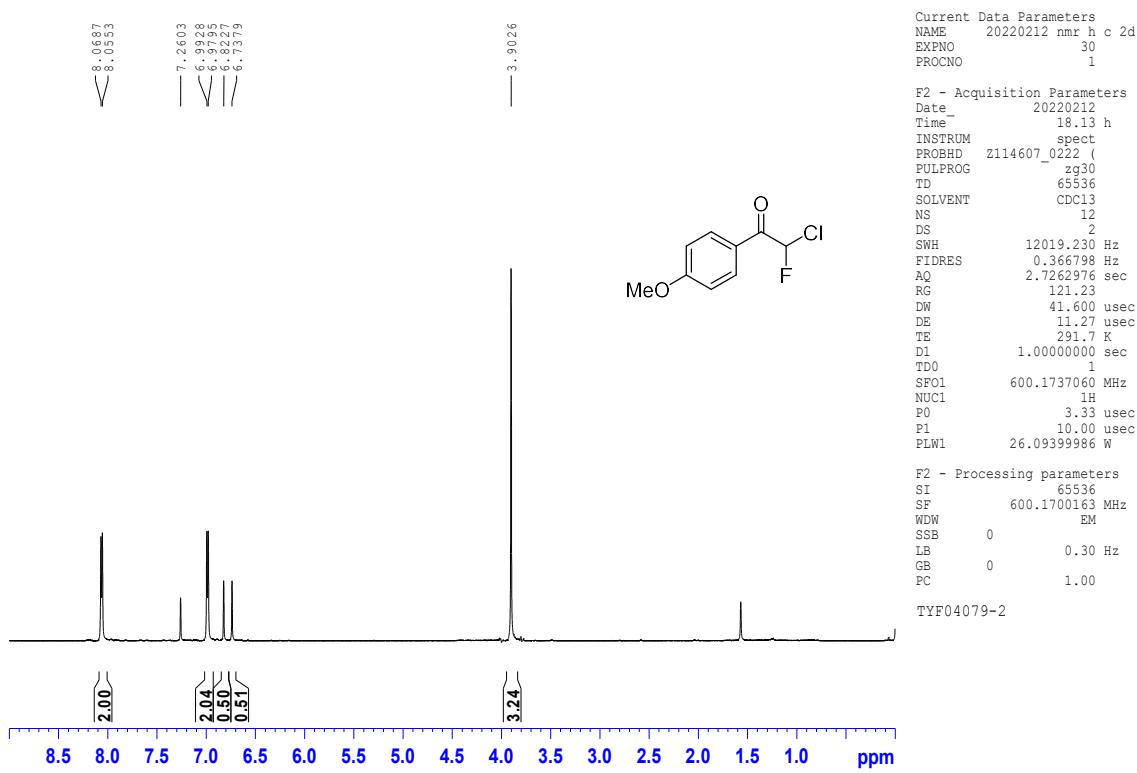
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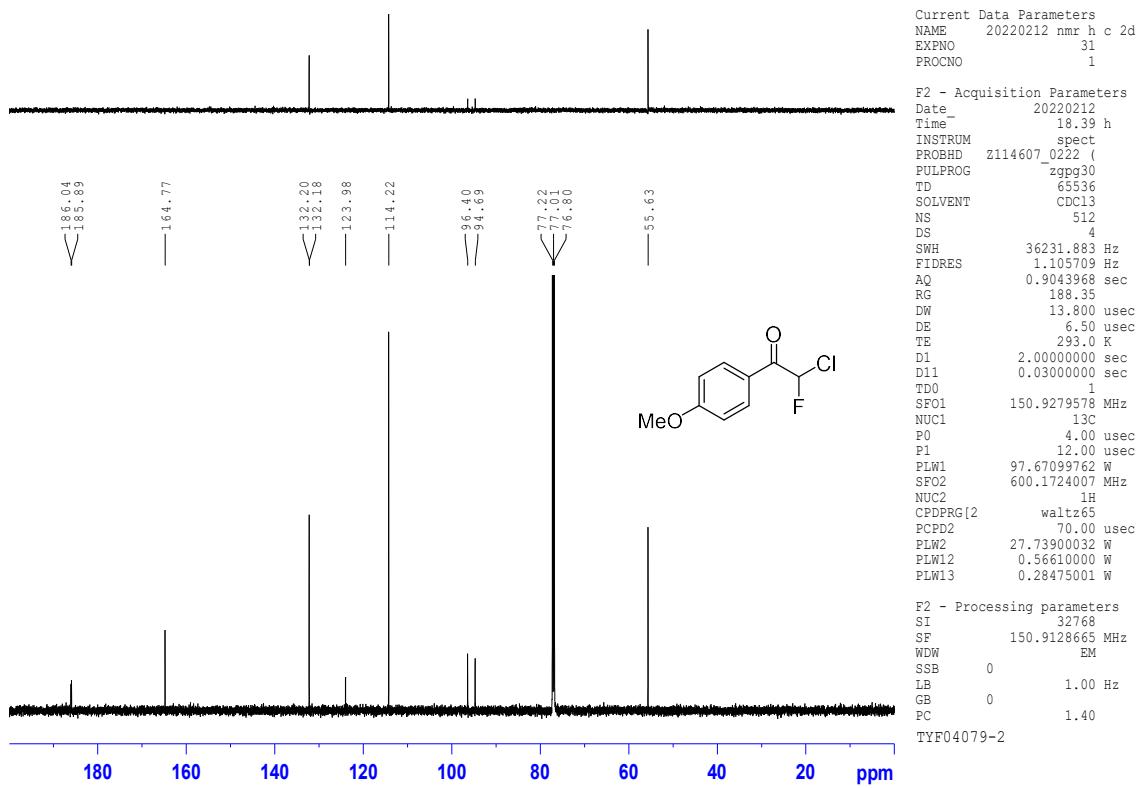
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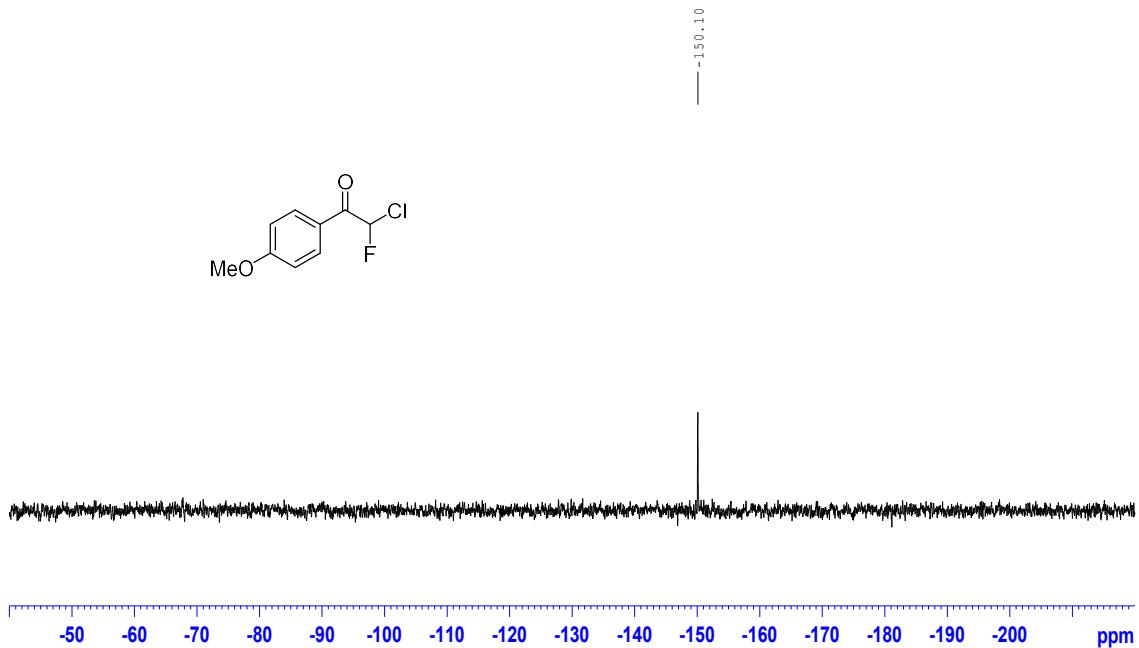
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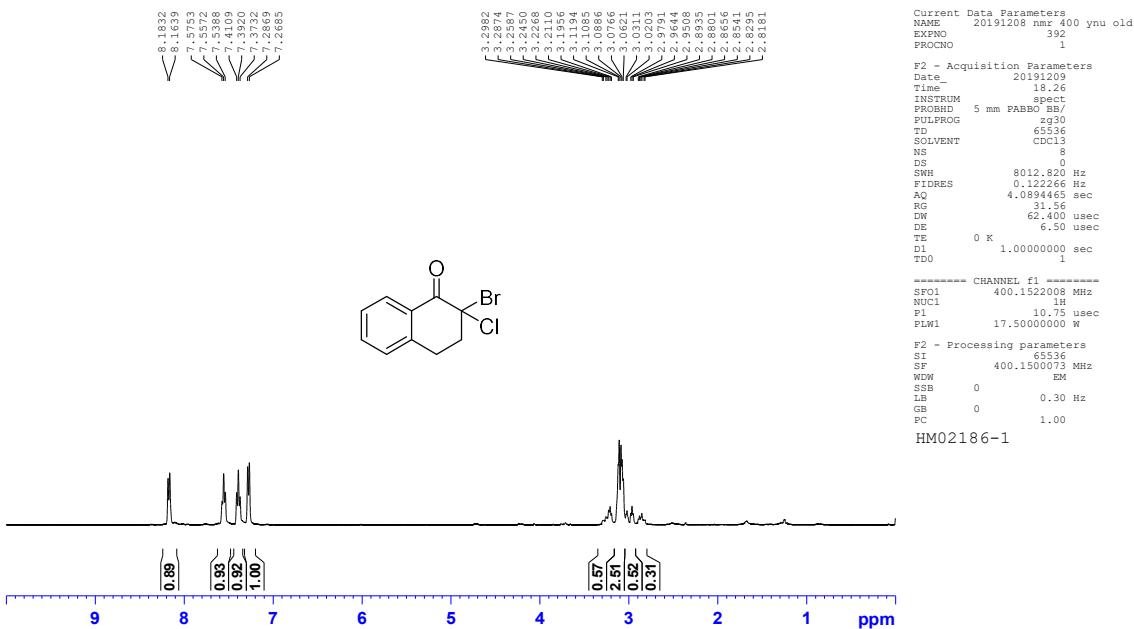
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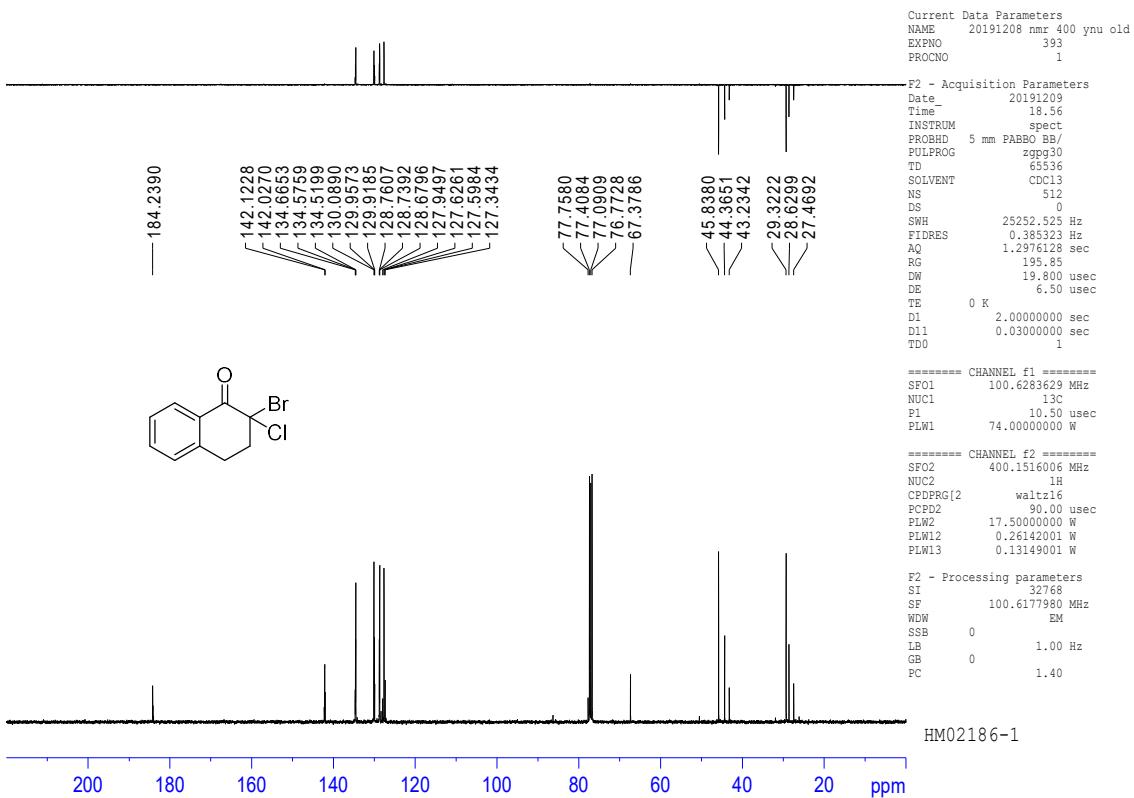
¹⁹F NMR (376 MHz, CDCl₃)



¹H-NMR (400 MHz, CDCl₃)

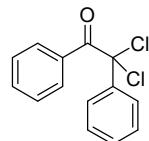


¹³C-NMR (100 MHz, CDCl₃)



¹H-NMR (400 MHz, CDCl₃)

7.7986
7.7789
7.6601
7.6423
7.4807
7.4622
7.4429
7.4055
7.4052
7.3186
7.3000
7.2820

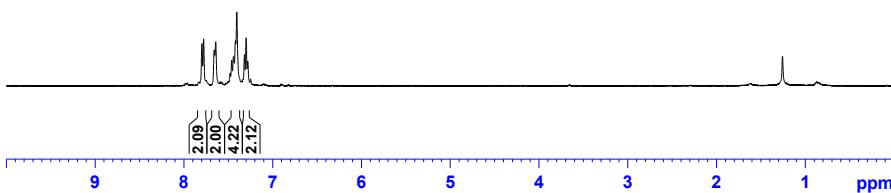


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PROCNO 1

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PULPROG zg30
TD 65536
SOLVENT CDC13
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DS 0
SWH 8012.820 Hz
FIDRES 0.122266 Hz
AQ 4.0894465 sec
RG 31.56
DW 62.000 usec
DE 6.50 usec
TE 0 K
D1 1.0000000 sec
TDO 1

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SF01 400.1522008 MHz
NUC1 1H
PL1 10.50 usec
PLW1 17.5000000 W

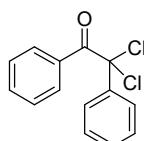
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WDW EM
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LB 0.30 Hz
GB 0
PC 1.00
HM02174-1



¹³C-NMR (100 MHz, CDCl₃)

139.54
133.36
131.19
129.92
129.04
128.17
126.02

77.40
77.39
77.37



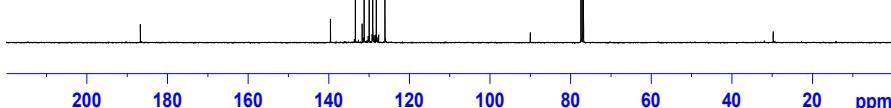
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D11 0.0300000 sec
TDO 1

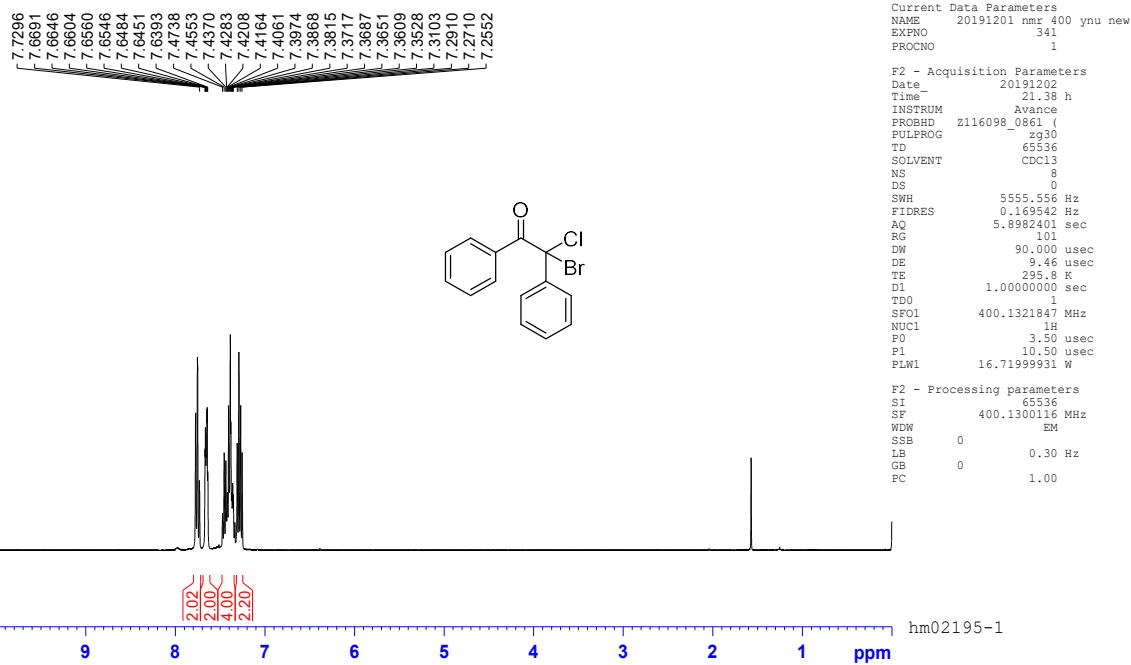
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SF01 100.6283629 MHz
NUC1 13C
PL1 10.50 usec
PLW1 74.0000000 W

===== CHANNEL f2 ======
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NUC2 1H
CPDPFG[2] waltz16
PCPD2 90.00 usec
PLW2 17.5000000 W
PLW12 0.26142001 W
PLW13 0.13149001 W

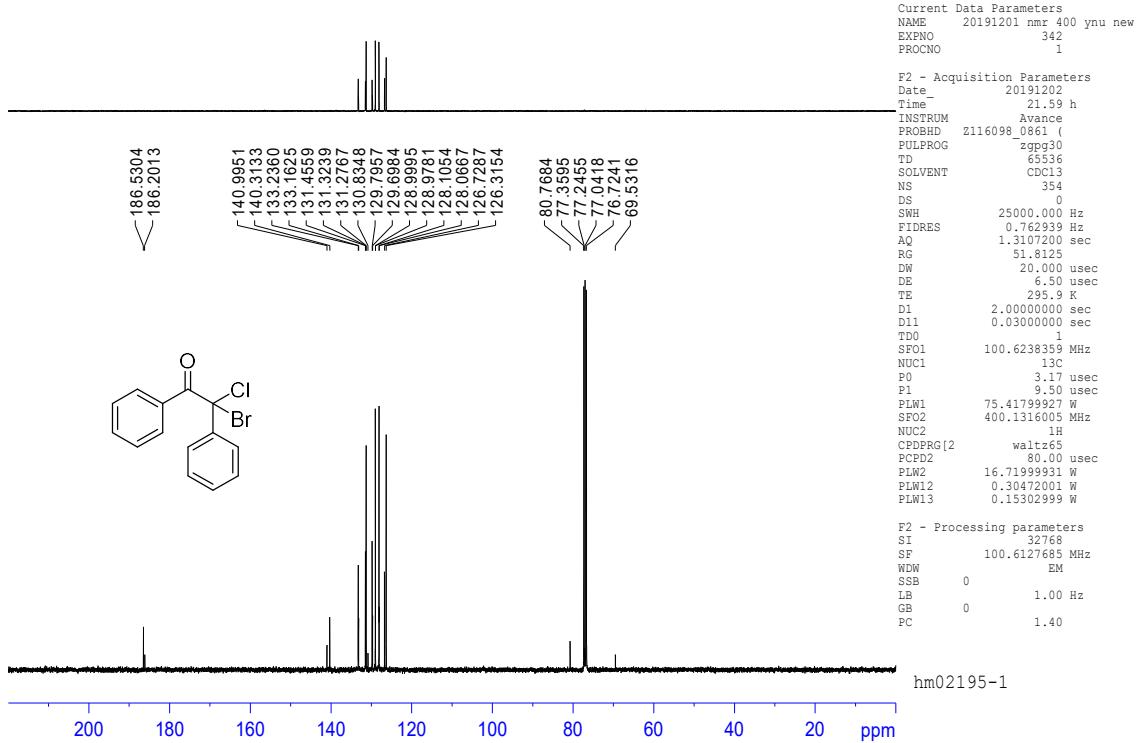
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HM02174-1



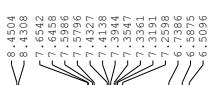
¹H-NMR (400 MHz, CDCl₃)



¹³C-NMR (100 MHz, CDCl₃)



¹H-NMR (400 MHz, CDCl₃)



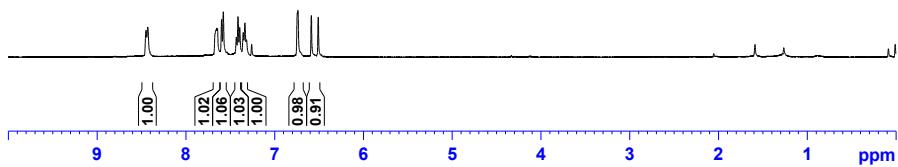
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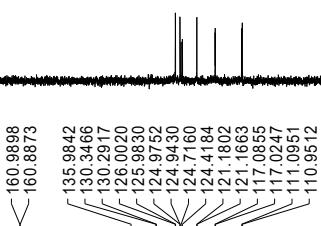
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 FIDRES 0.169542 Hz
 AQ 5.8982401 sec
 RG 90.000 usec
 DE 9.46 usec
 TE 293.8 K
 D1 1.0000000 sec
 TDO SF01 400.1321847 MHz
 NUC1 1H
 P0 3.50 usec
 P1 10.50 usec
 PLW1 16.71999931 W

F2 - Processing parameters
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 MDW EM
 SSB 0
 LB 0.30 Hz
 GB 0
 PC 1.00

HM02203-1



¹³C-NMR (100 MHz, CDCl₃)



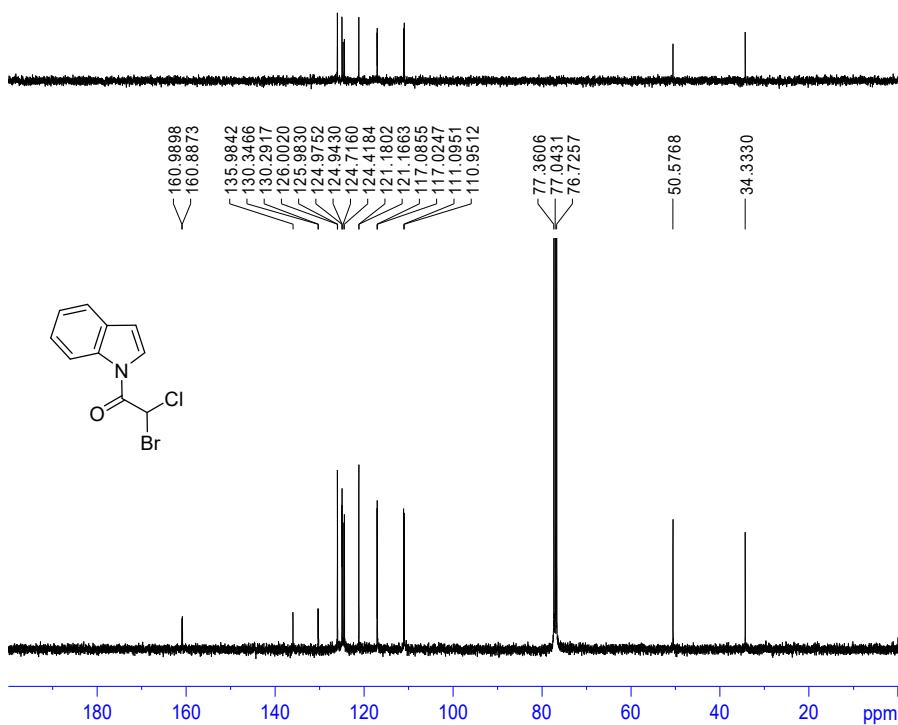
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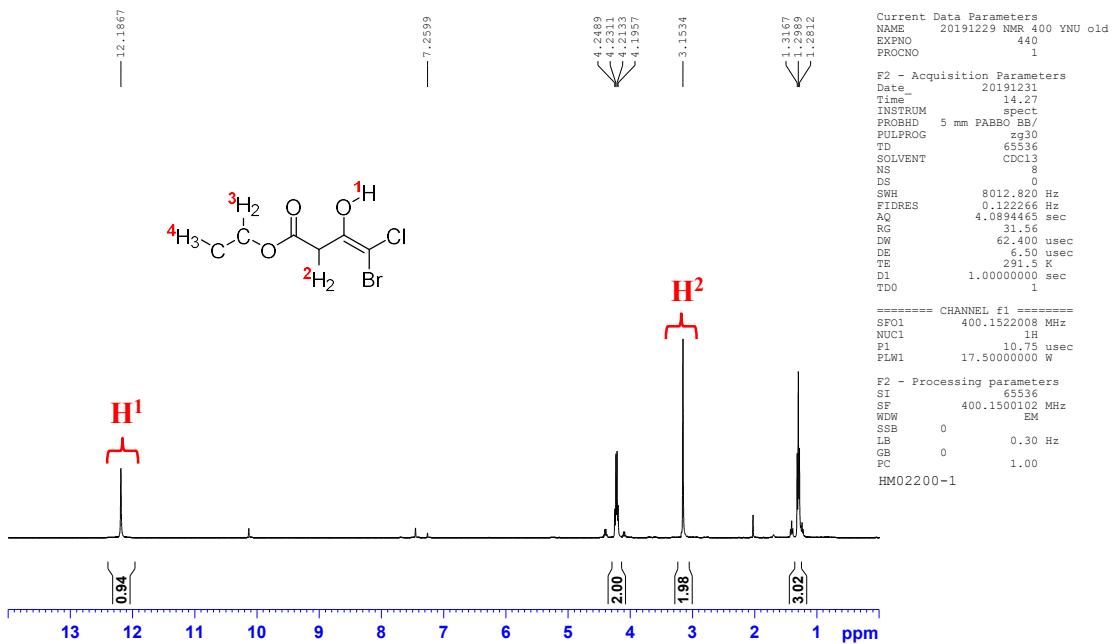
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 FIDRES 0.762939 Hz
 AQ 1.3107200 sec
 RG 54.2535
 DW 20.000 usec
 DE 6.50 usec
 TE 295.4 K
 D1 2.0000000 sec
 D11 0.03000000 sec
 TDO 1
 SF01 100.6238359 MHz
 NUC1 13C
 P0 3.17 usec
 P1 9.50 usec
 PLW1 75.41799927 W
 SF02 400.1316005 MHz
 NUC2 1H
 CPDPRG[2] waltz65
 PCPD2 80.00 usec
 PLW2 16.71999931 W
 PLW12 0.34549999 W
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F2 - Processing parameters
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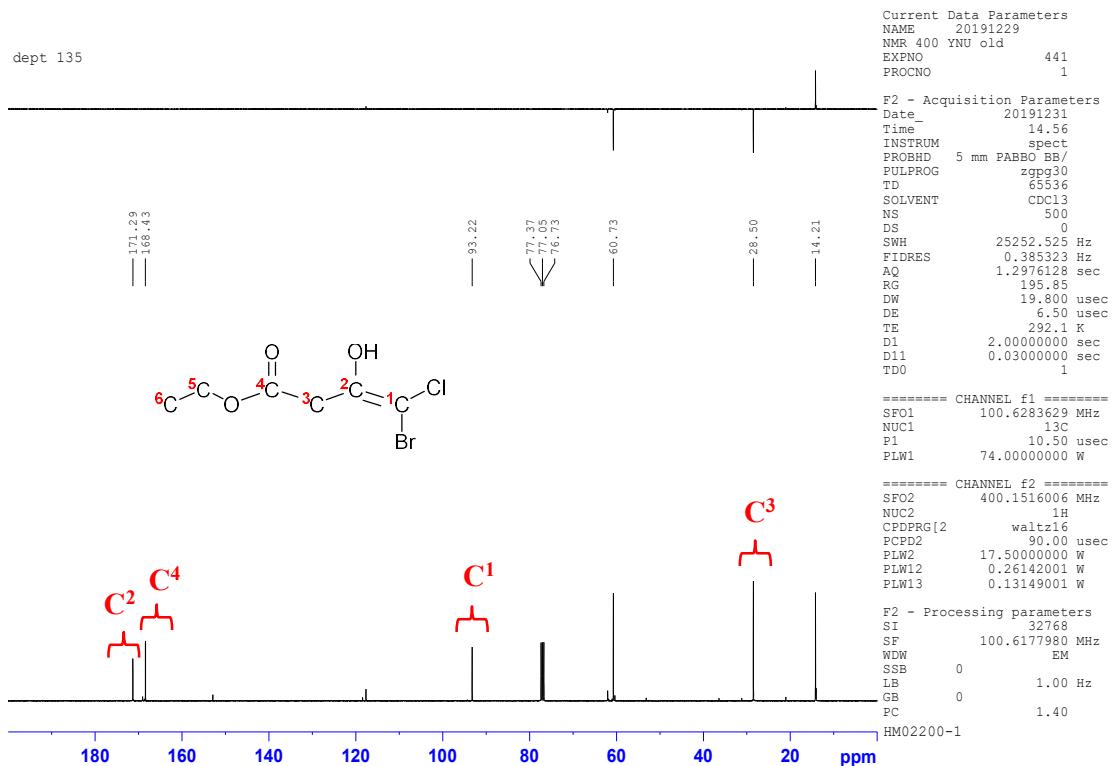
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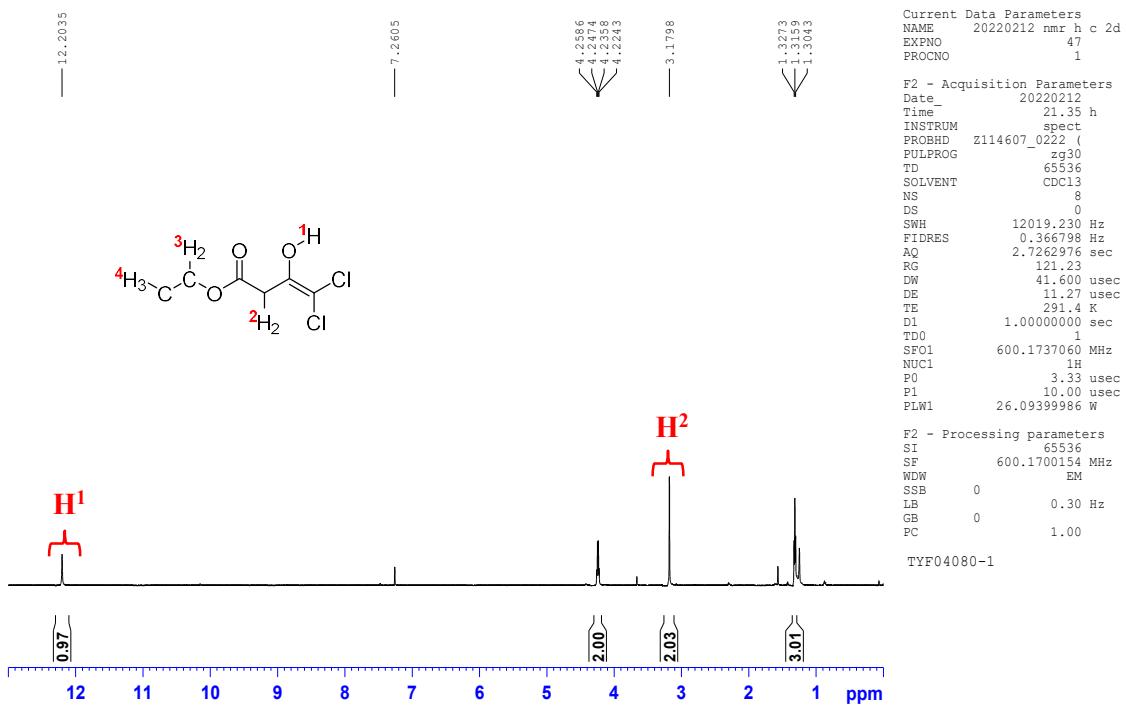
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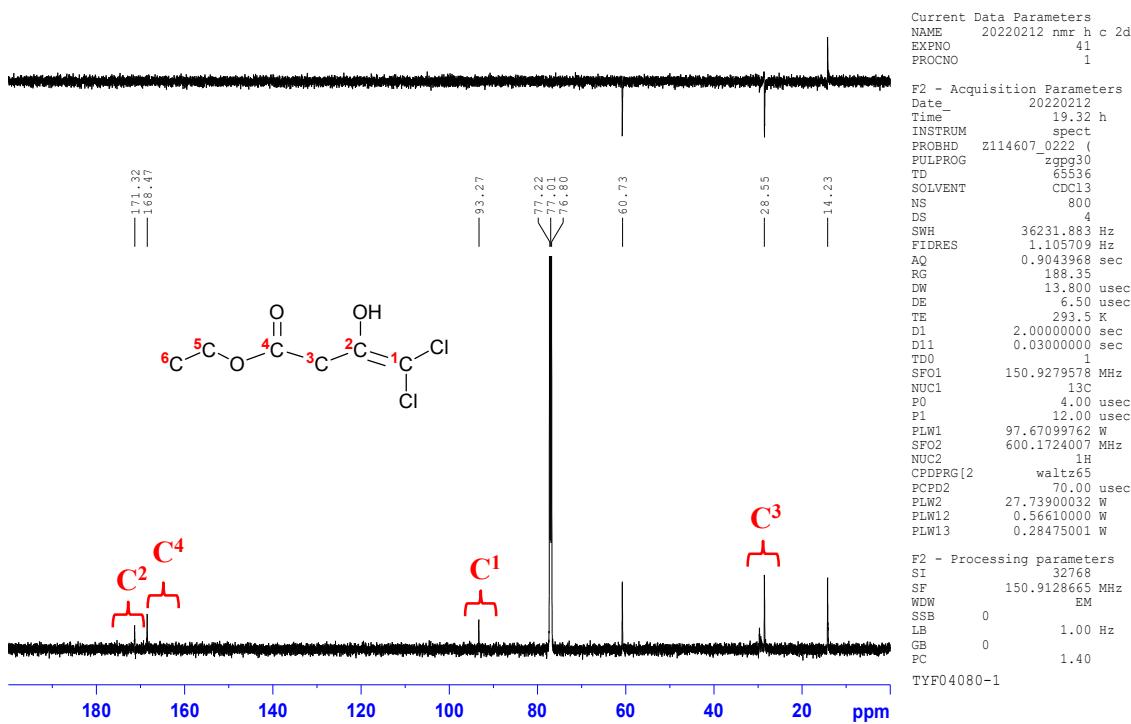
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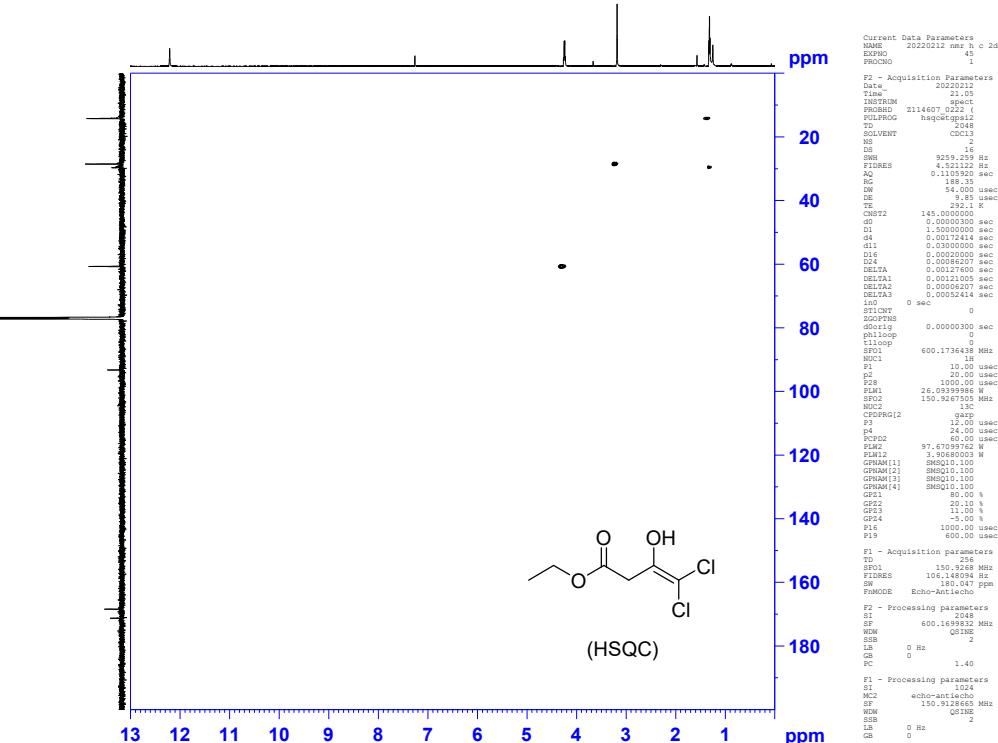
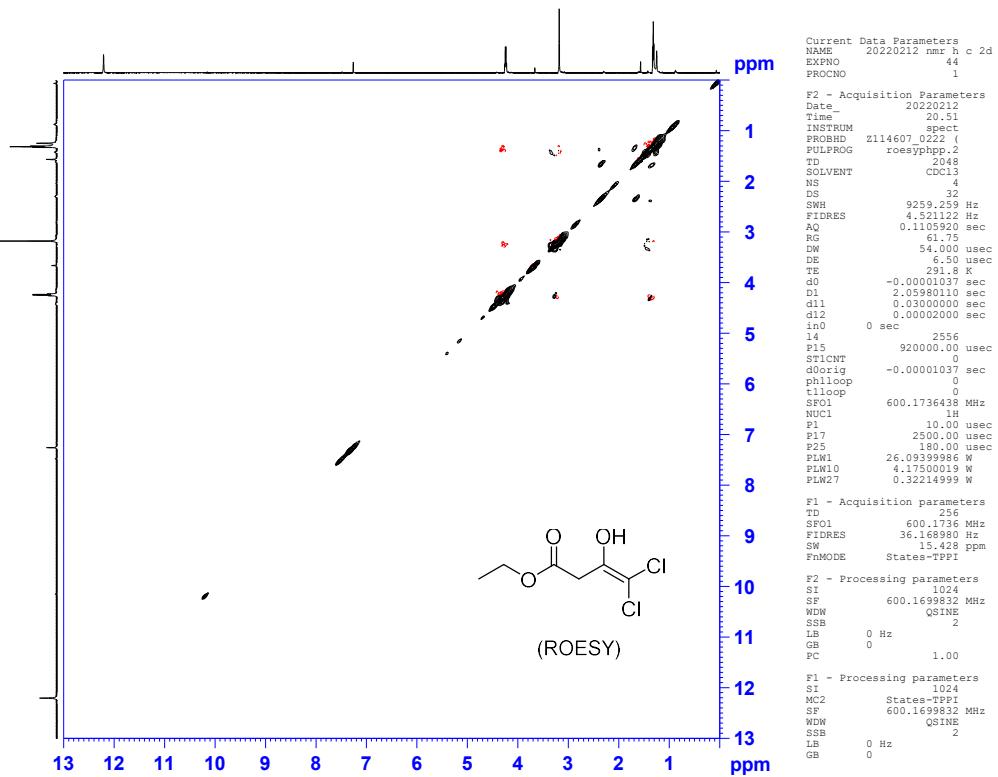


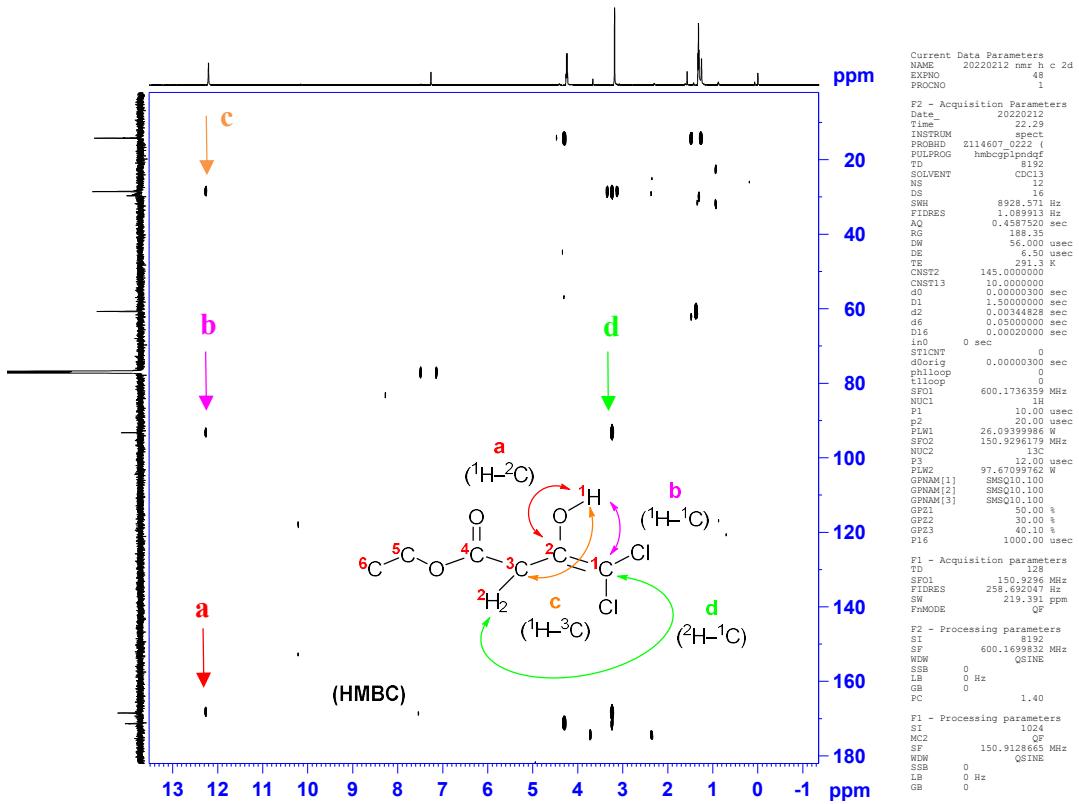
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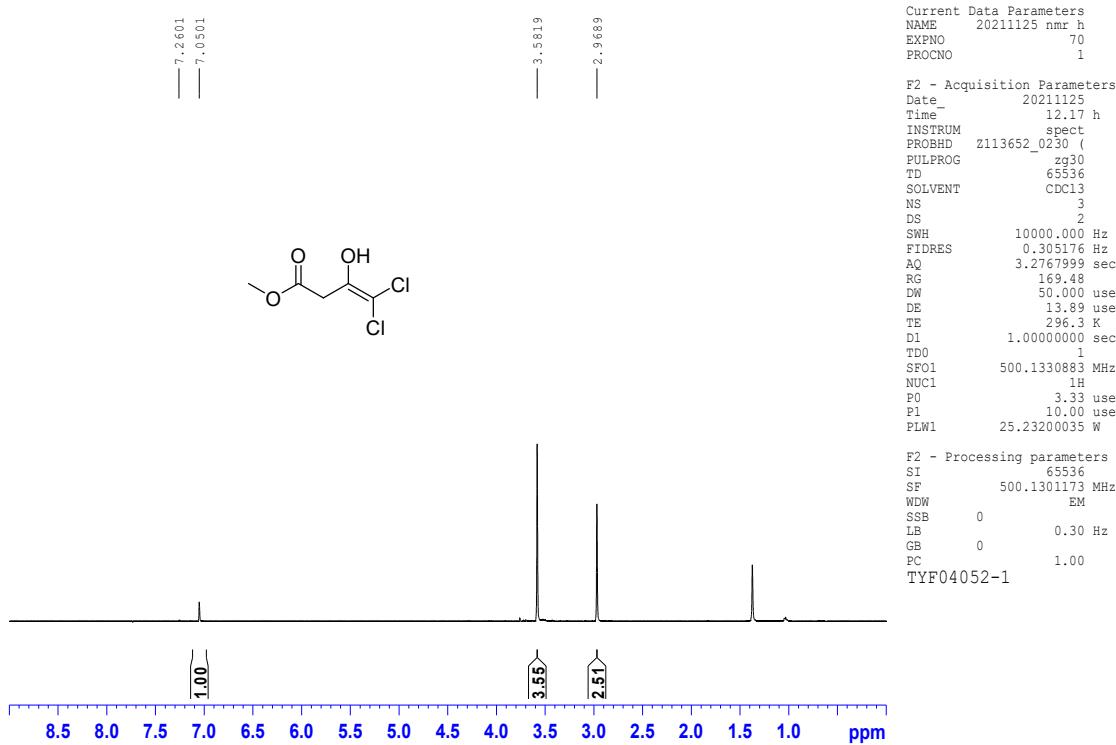
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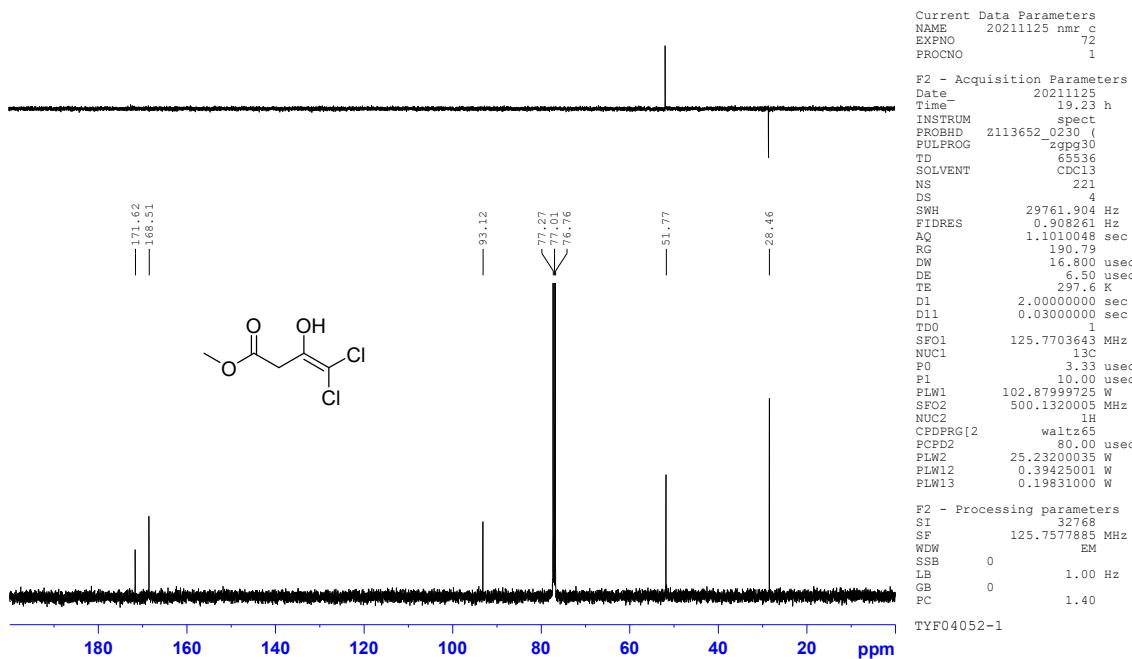




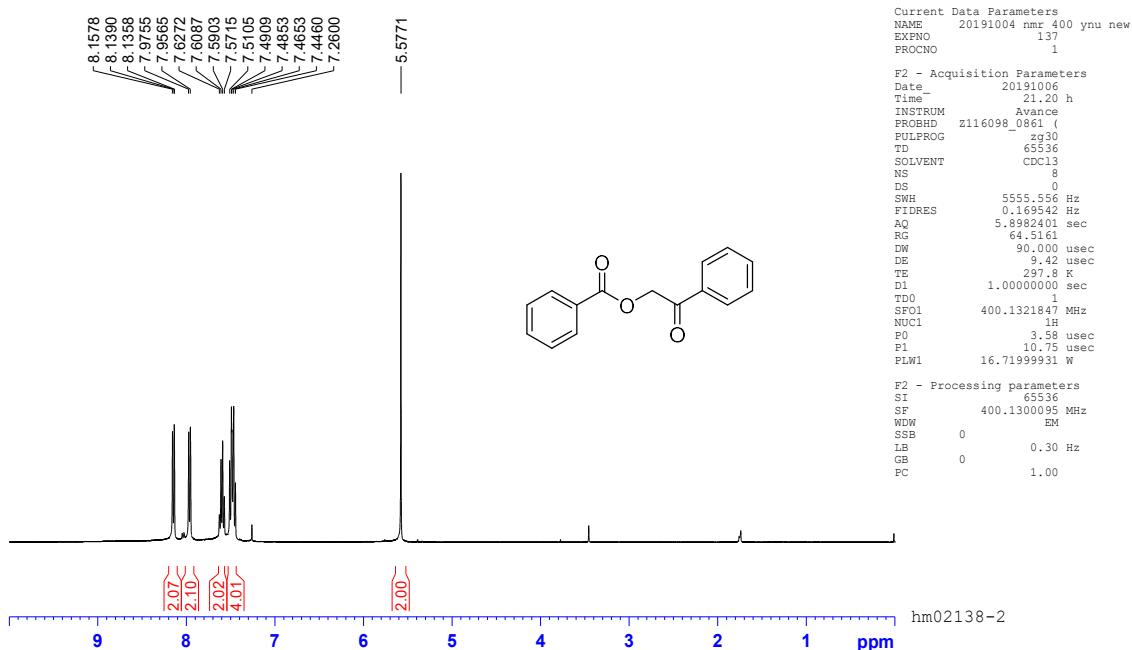
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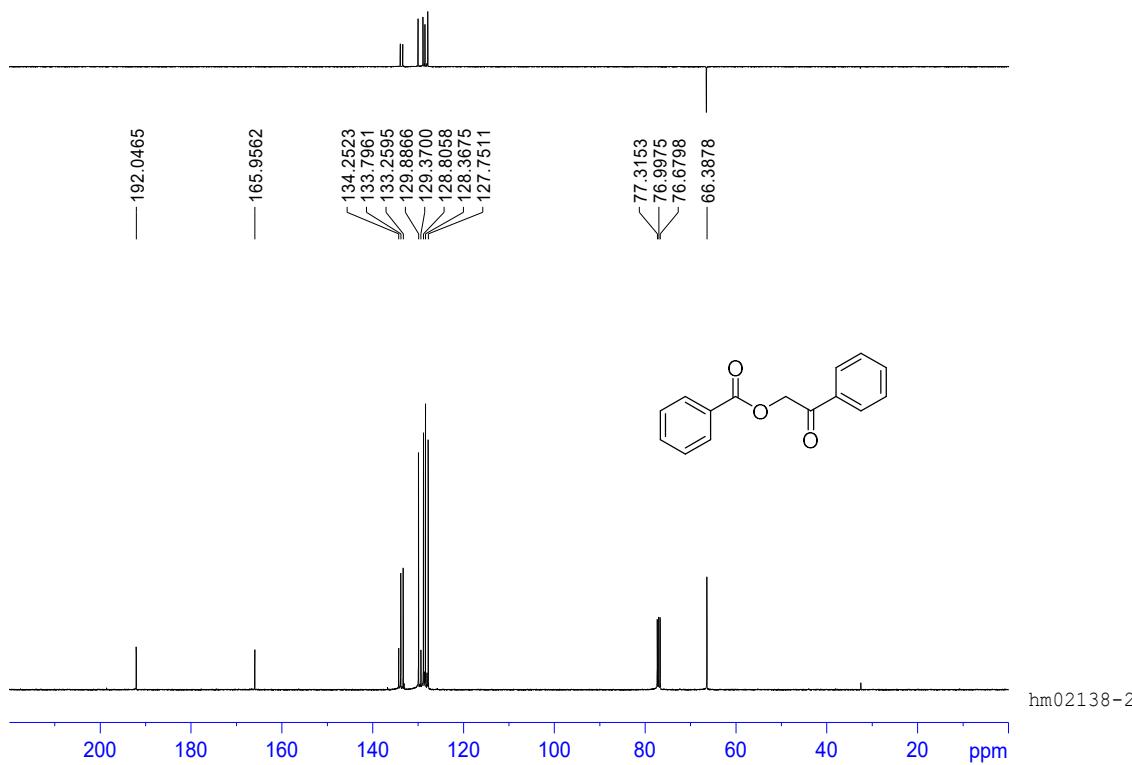
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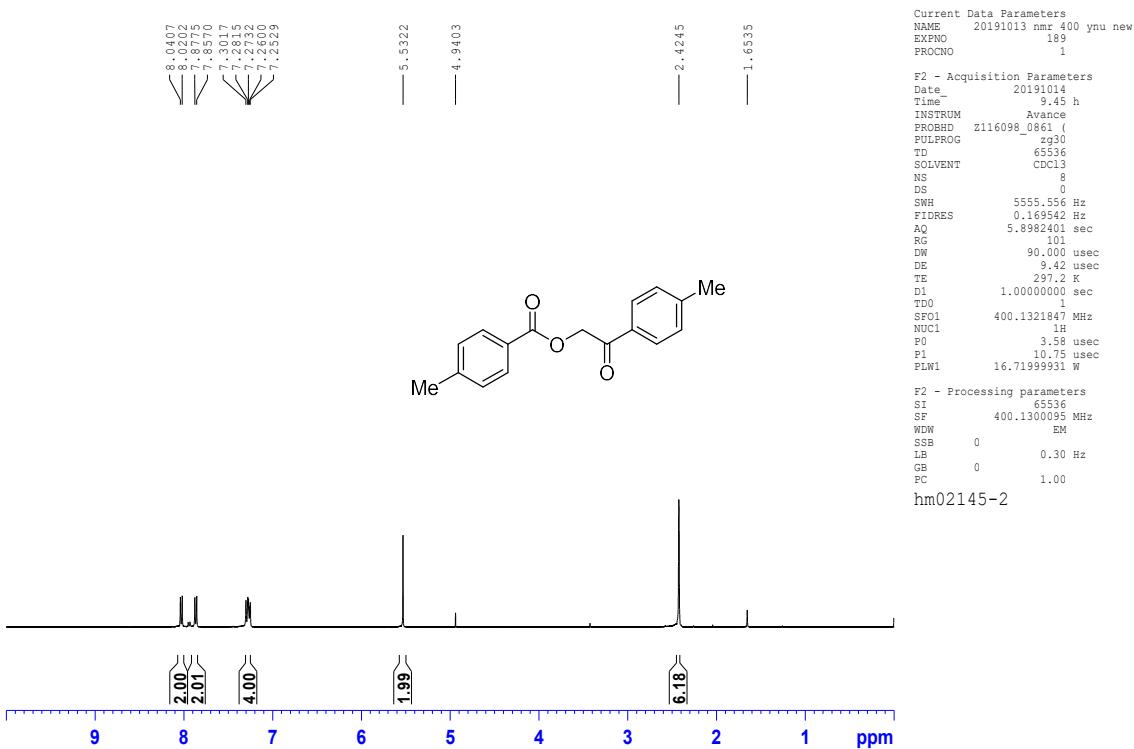
¹H-NMR (400 MHz, CDCl₃)



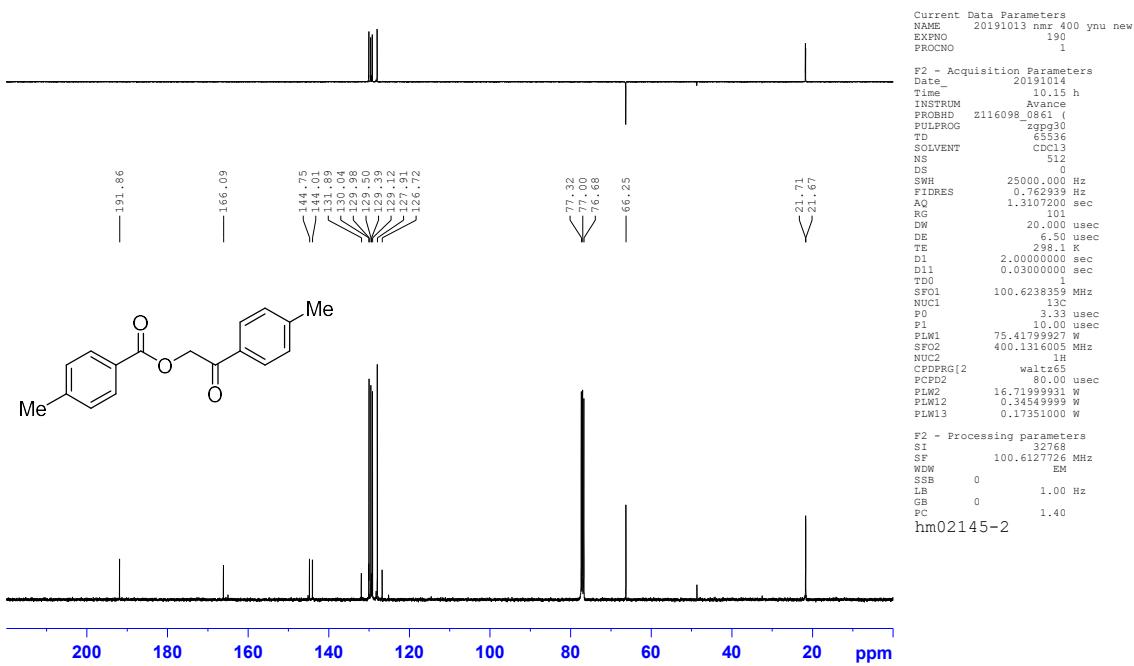
¹³C-NMR (100 MHz, CDCl₃)



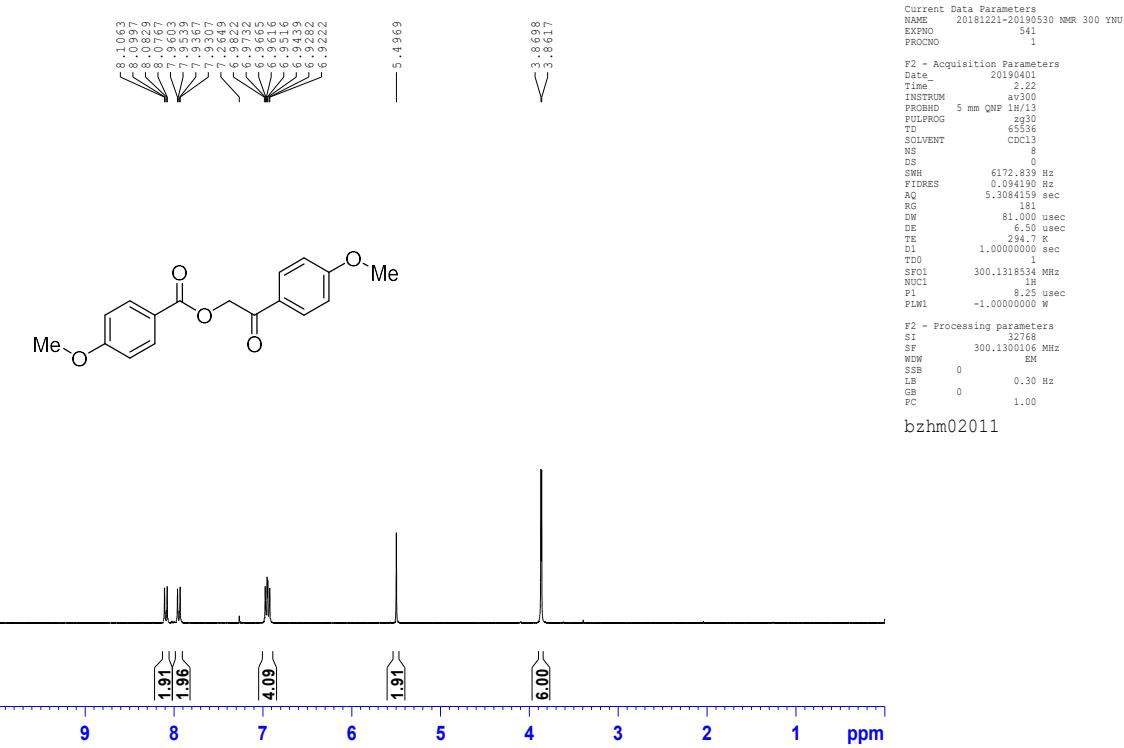
¹H-NMR (400 MHz, CDCl₃)



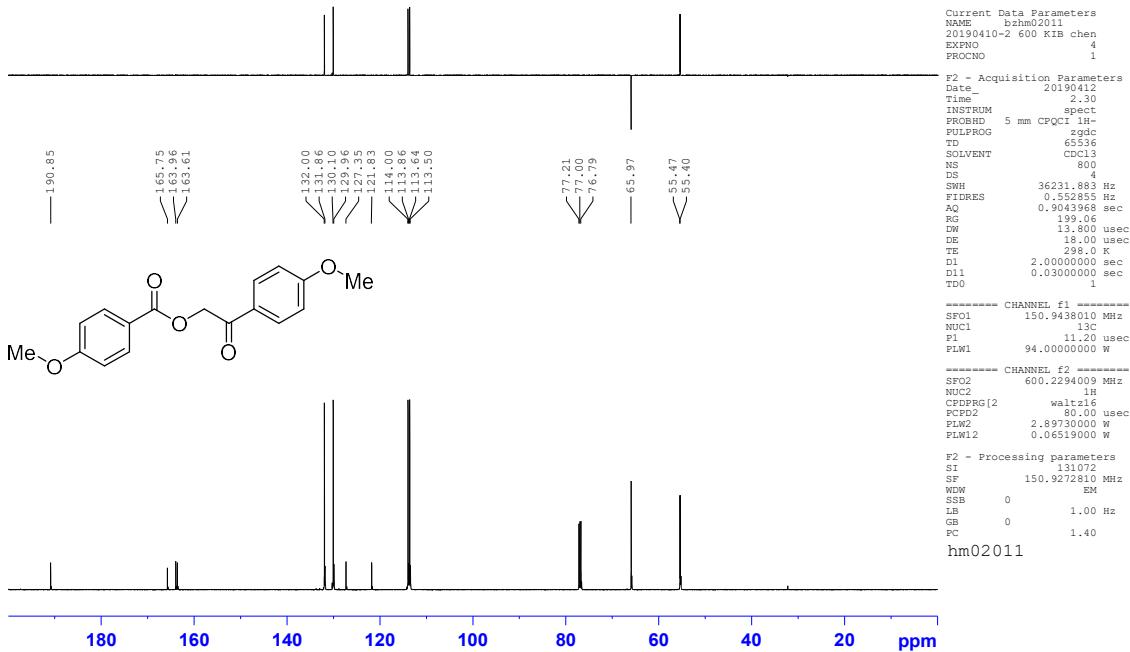
¹³C-NMR (100 MHz, CDCl₃)



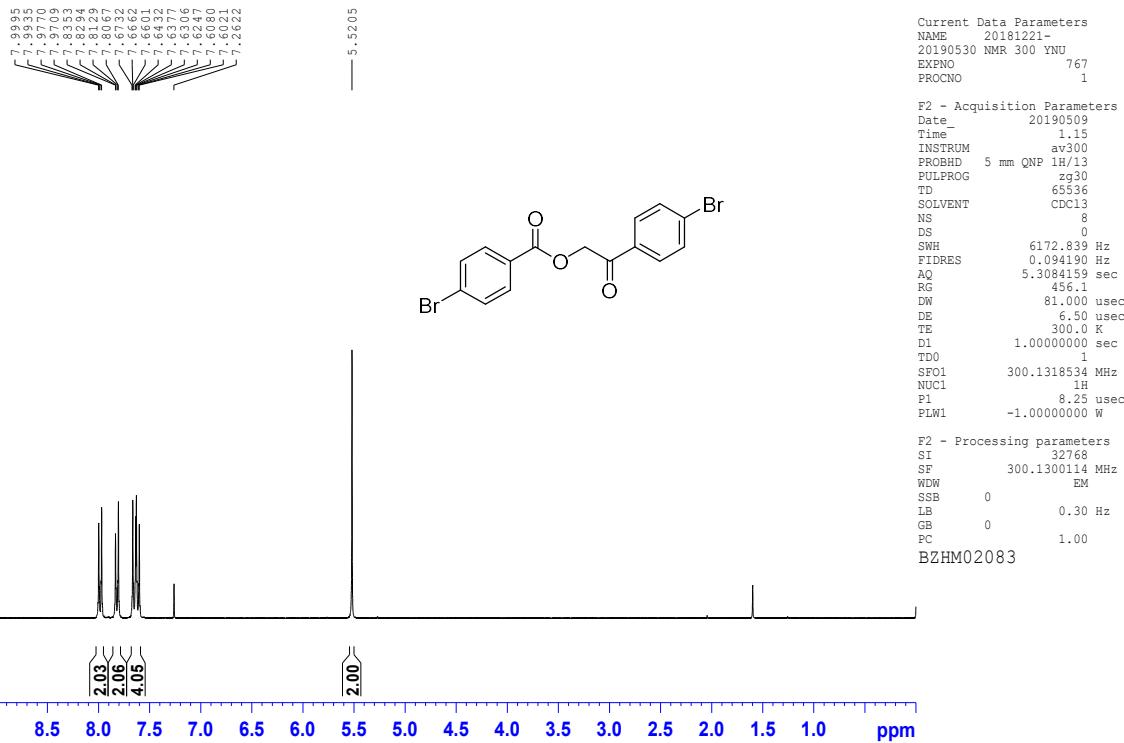
¹H-NMR (300 MHz, CDCl₃)



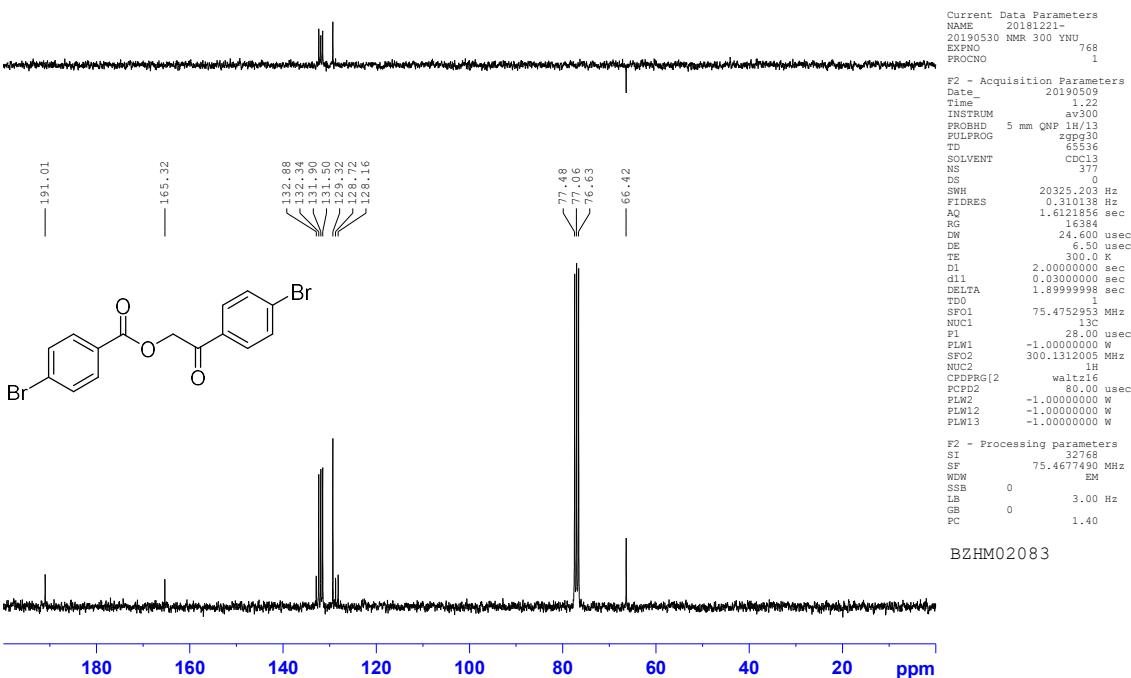
¹³C-NMR (150 MHz, CDCl₃)



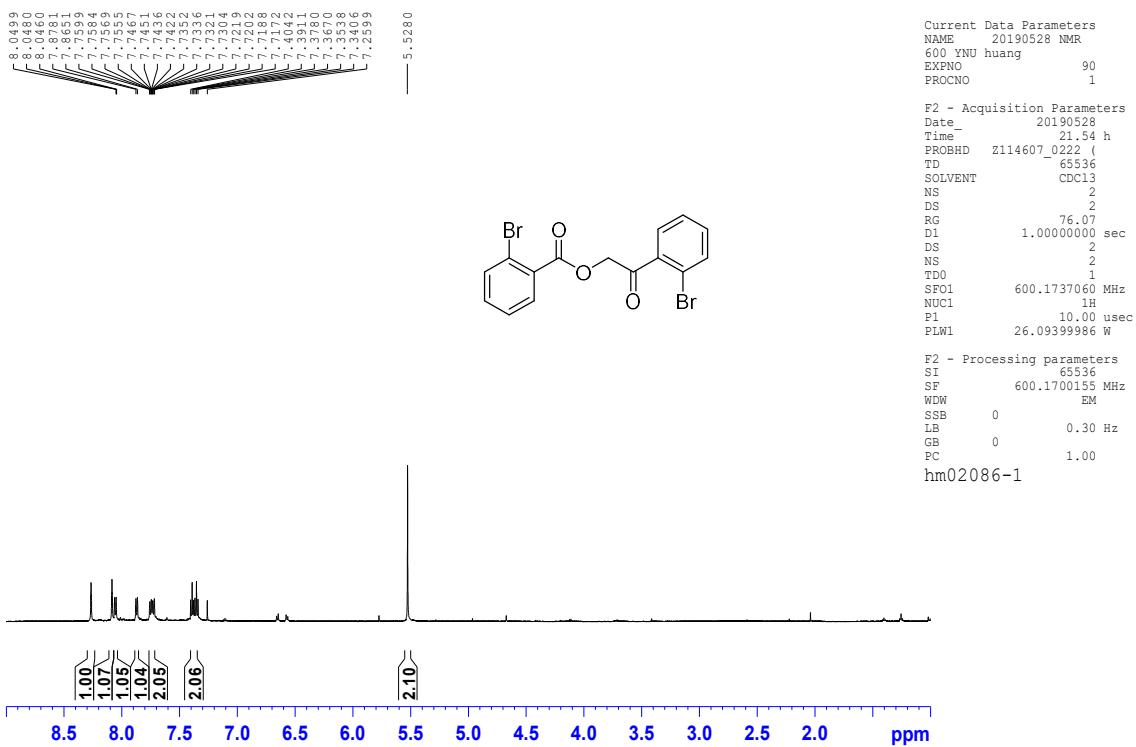
¹H-NMR (300 MHz, CDCl₃)



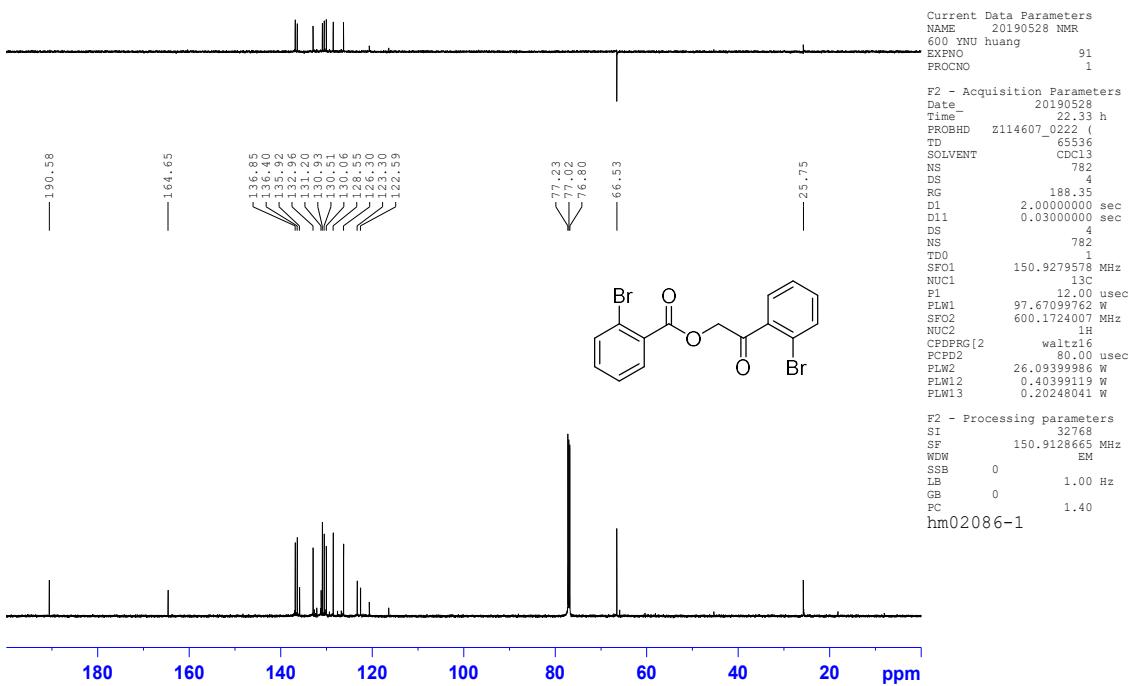
¹³C-NMR (75 MHz, CDCl₃)



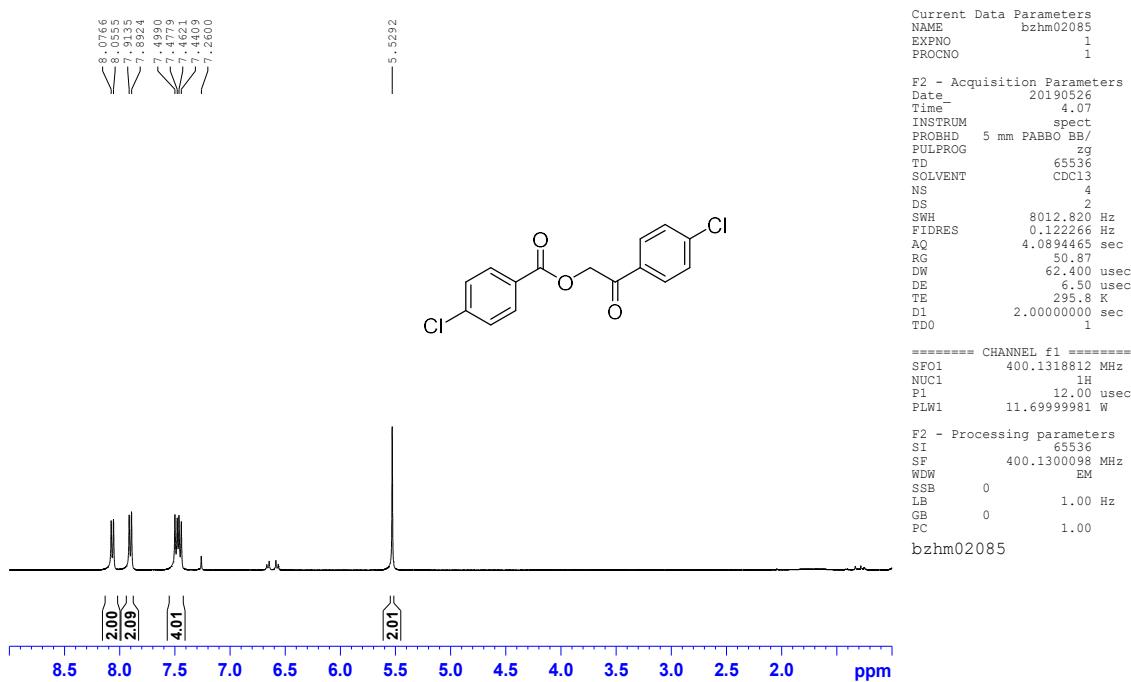
¹H-NMR (600 MHz, CDCl₃)



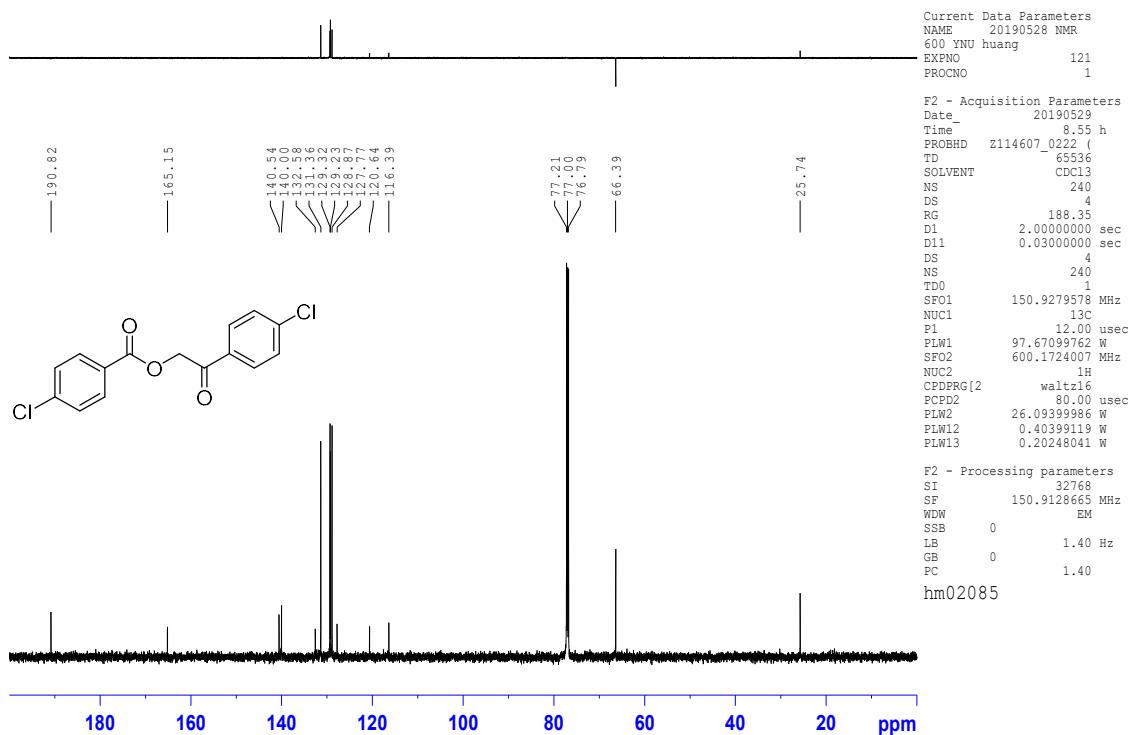
¹³C-NMR (150 MHz, CDCl₃)



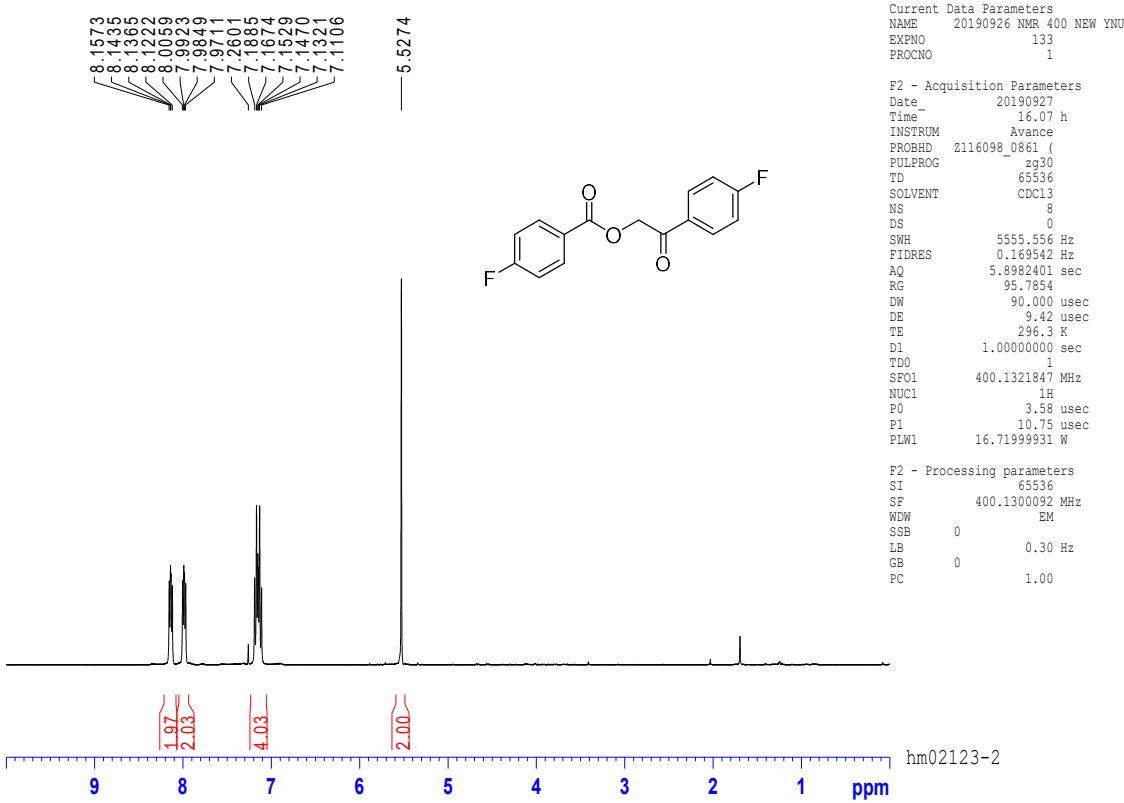
¹H-NMR (400 MHz, CDCl₃)



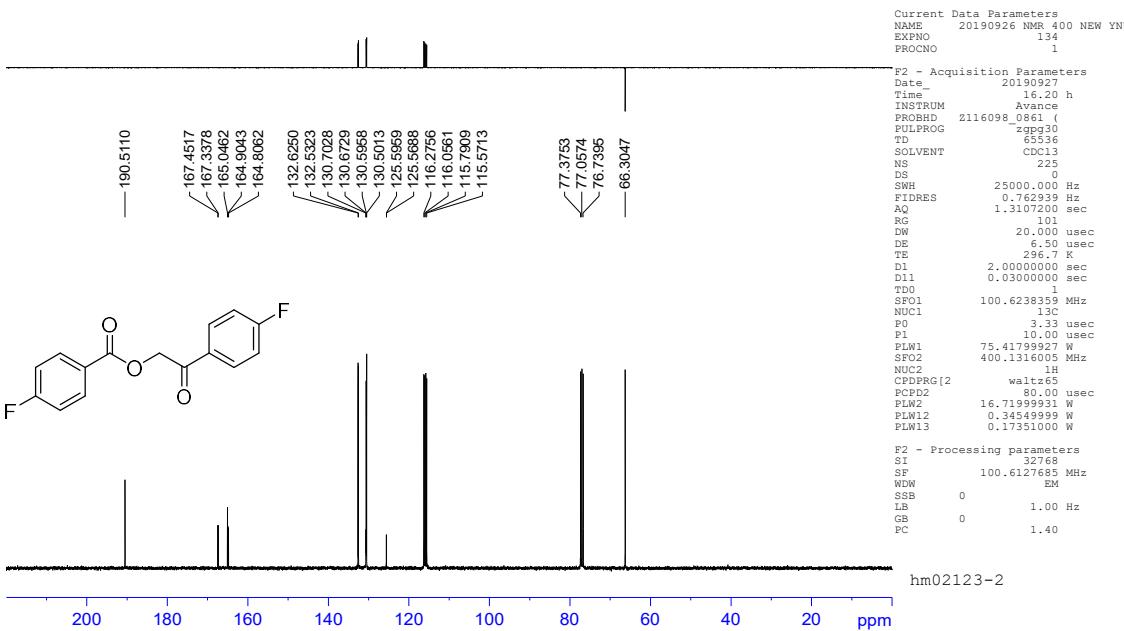
¹³C-NMR (150 MHz, CDCl₃)



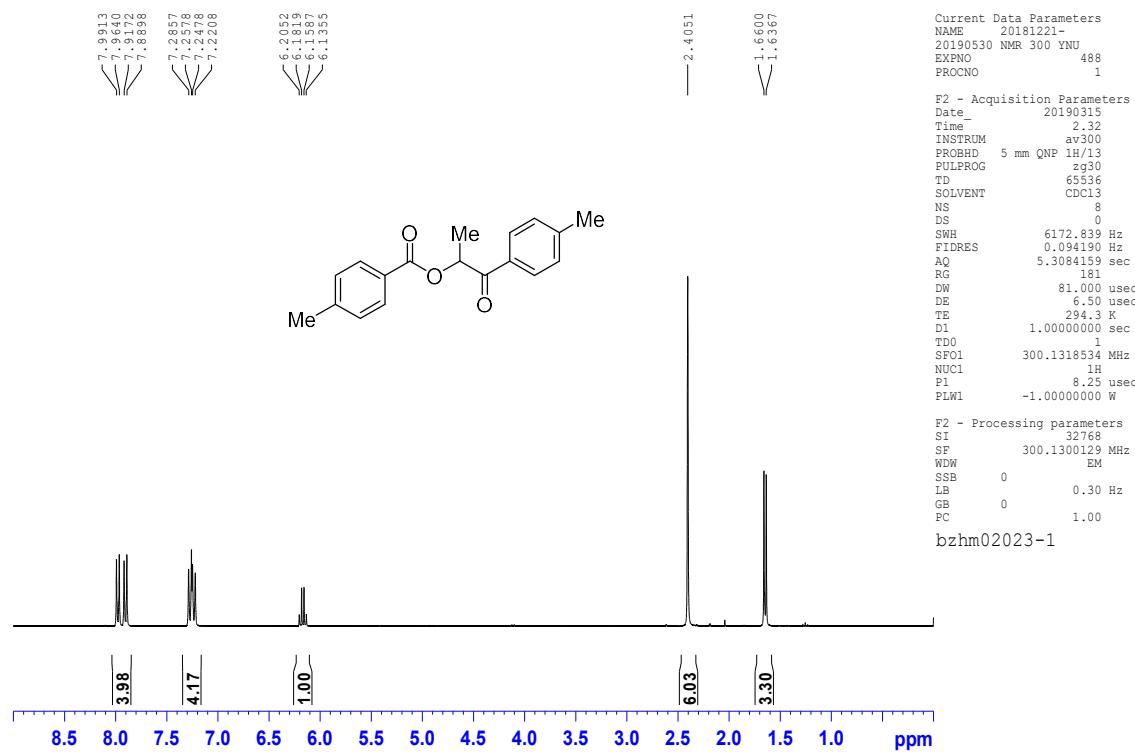
¹H-NMR (400 MHz, CDCl₃)



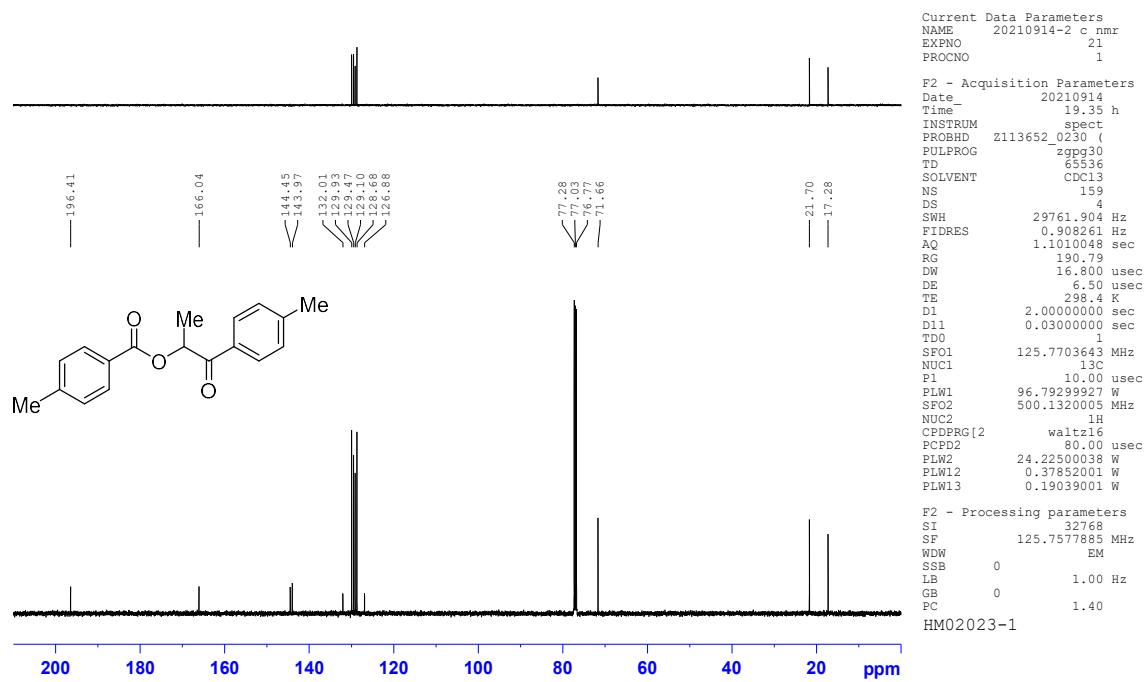
¹³C-NMR (100 MHz, CDCl₃)



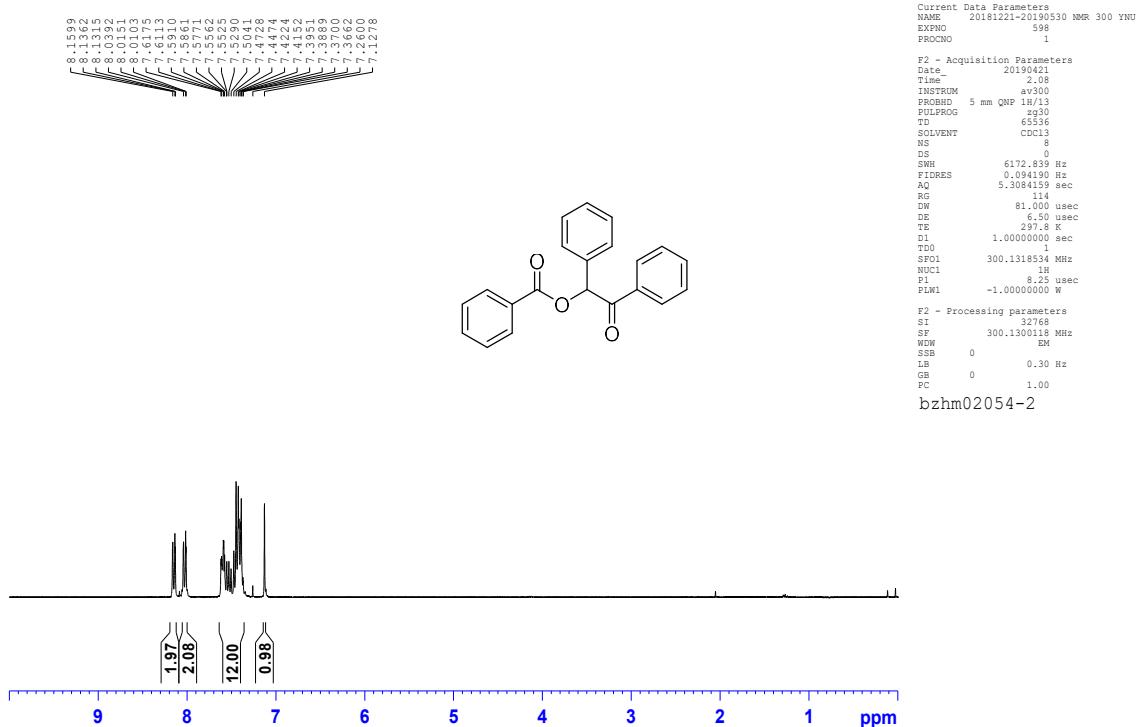
¹H-NMR (300 MHz, CDCl₃)



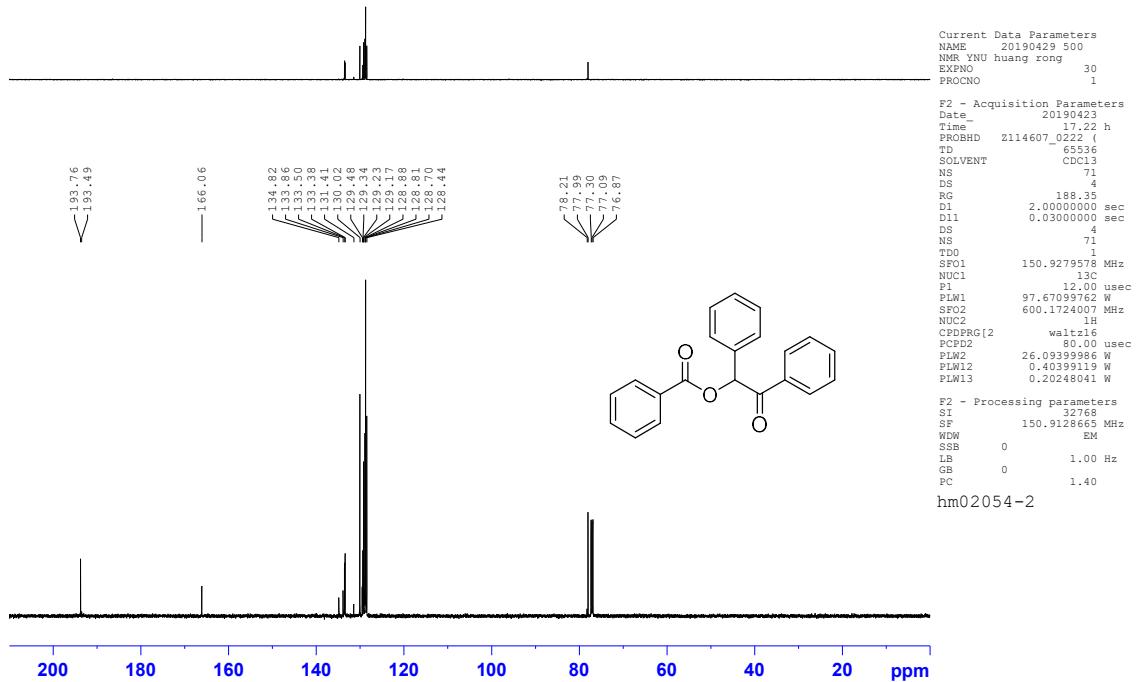
¹³C-NMR (125 MHz, CDCl₃)



¹H-NMR (300 MHz, CDCl₃)



¹³C-NMR (150 MHz, CDCl₃)



5. DFT calculations

DFT calculations of geometries and energies for all optimized starting materials, transition states, intermediates, and target materials

Single Point Energy ■, Δ (EM06, in Hartree)

Zero-point correction (E0, in Hartree)

Thermal correction to Enthalpy (H, in Hartree)

Thermal correction to Gibbs Free Energy (G, in Hartree)

Sum of electronic and zero-point Energies (EM06+E0, in Hartree)

Sum of electronic and thermal Enthalpies (EM06+H, in Hartree)

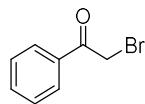
Sum of electronic and thermal Free Energies (EM06+G, in Hartree)

■ The single-point energies and solvent effects were computed at the M06-2X/def2-TZVP level

Δ The single-point energies and solvent effects were computed at the M06-2X/def2-TZVP level

Methods: Gaussian program¹ was employed for DFT calculations, and /M06-2X/def2-SVP level was used for geometry optimization and frequency calculations. Frequency calculations were performed to ensure minimums were found. The single point calculations were at M06-2X/def2-TZVP level. The SMD implicit solvent model was used to take account of the solvation effect of water.

Standard orientation, imaginary frequencies, thermodynamic energies and single-point energies of all stationary points.



A Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	3.477831	0.473945	-0.603946
2	6	0	2.216356	1.059673	-0.596556
3	6	0	1.161980	0.462069	0.107441
4	6	0	1.386682	-0.731240	0.807521
5	6	0	2.652080	-1.313692	0.801778
6	6	0	3.696379	-0.714386	0.096494
7	1	0	4.293591	0.943344	-1.155489
8	1	0	2.029960	1.988462	-1.137252
9	1	0	0.582517	-1.220107	1.358698
10	1	0	2.822820	-2.241578	1.349077
11	1	0	4.685361	-1.175566	0.093374
12	6	0	-0.164554	1.150706	0.085211
13	8	0	-0.304659	2.230759	-0.448910
14	6	0	-1.365758	0.482951	0.747584
15	1	0	-1.072326	-0.004179	1.683600
16	1	0	-2.837531	1.874794	-0.005544
17	35	0	-1.876653	-0.997663	-0.456167

Zero-point correction= 0.158338 (Hartree/Particle)

Thermal correction to Energy= 0.168492

Thermal correction to Enthalpy= 0.169436

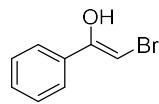
Thermal correction to Gibbs Free Energy= 0.121014

Sum of electronic and zero-point Energies= -2996.844625

Sum of electronic and thermal Energies= -2996.834471

Sum of electronic and thermal Enthalpies= -2996.833526

Sum of electronic and thermal Free Energies= -2996.881948



M1 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-3.314048	0.402462	0.770108
2	6	0	-2.076094	1.036908	0.682278
3	6	0	-1.051173	0.516807	-0.120272
4	6	0	-1.315172	-0.635516	-0.871832
5	6	0	-2.560177	-1.262260	-0.803063
6	6	0	-3.560971	-0.752347	0.025408

7	1	0	-4.092316	0.814535	1.415403
8	1	0	-1.890818	1.957329	1.239385
9	1	0	-0.540712	-1.039218	-1.525413
10	1	0	-2.752339	-2.152587	-1.405106
11	1	0	-4.531839	-1.247712	0.083523
12	6	0	0.234775	1.305531	-0.214340
13	8	0	0.081705	2.579779	-0.355365
14	6	0	1.470024	0.703673	-0.138001
15	1	0	3.327102	1.420310	0.730437
16	35	0	1.724777	-1.161371	0.258375

Zero-point correction= 0.144446 (Hartree/Particle)

Thermal correction to Energy= 0.154671

Thermal correction to Enthalpy= 0.155615

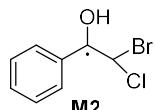
Thermal correction to Gibbs Free Energy= 0.107180

Sum of electronic and zero-point Energies= -2996.354856

Sum of electronic and thermal Energies= -2996.344632

Sum of electronic and thermal Enthalpies= -2996.343687

Sum of electronic and thermal Free Energies= -2996.392122



Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-3.770588	0.507551	0.394721
2	6	0	-2.524209	1.118025	0.449323
3	6	0	-1.356451	0.421308	0.059615
4	6	0	-1.494624	-0.918263	-0.372121
5	6	0	-2.748090	-1.517742	-0.421302
6	6	0	-3.893524	-0.813850	-0.042647
7	1	0	-4.656736	1.069669	0.694514
8	1	0	-2.442383	2.149598	0.789414
9	1	0	-0.620346	-1.499231	-0.659527
10	1	0	-2.829821	-2.552792	-0.757192
11	1	0	-4.873214	-1.291519	-0.088032
12	6	0	-0.083291	1.099442	0.140344
13	8	0	-0.093839	2.262659	0.819996
14	6	0	1.213461	0.580414	-0.327558
15	1	0	2.531673	1.880732	0.786639
16	35	0	1.781935	-1.017847	0.853783
17	17	0	1.113806	-0.047508	-2.010481
18	1	0	0.727293	2.764506	0.705623

Zero-point correction= 0.159406 (Hartree/Particle)

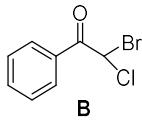
Thermal correction to Energy= 0.171064

Thermal correction to Enthalpy= 0.172008

Thermal correction to Gibbs Free Energy= 0.119970

Sum of electronic and zero-point Energies= -3456.848267

Sum of electronic and thermal Energies= -3456.836610
 Sum of electronic and thermal Enthalpies= -3456.835665
 Sum of electronic and thermal Free Energies= -3456.887703



B Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	3.734925	0.541684	-0.446020
2	6	0	2.471421	1.110777	-0.561949
3	6	0	1.334047	0.423522	-0.111416
4	6	0	1.484321	-0.850873	0.457653
5	6	0	2.750966	-1.418871	0.567902
6	6	0	3.876073	-0.726707	0.118929
7	1	0	4.611602	1.088277	-0.795938
8	1	0	2.350296	2.099982	-1.004014
9	1	0	0.625273	-1.413457	0.817489
10	1	0	2.858072	-2.410755	1.008729
11	1	0	4.865834	-1.176853	0.212251
12	6	0	0.027958	1.131216	-0.295224
13	8	0	-0.032914	2.190976	-0.873373
14	6	0	-1.307517	0.558299	0.240822
15	1	0	-2.578922	1.752603	-1.016692
16	35	0	-1.684477	-1.085288	-0.775285
17	17	0	-1.140351	0.191141	1.985596

Zero-point correction= 0.148703 (Hartree/Particle)
 Thermal correction to Energy= 0.159929
 Thermal correction to Enthalpy= 0.160873
 Thermal correction to Gibbs Free Energy= 0.109728
 Sum of electronic and zero-point Energies= -3456.292158
 Sum of electronic and thermal Energies= -3456.280932
 Sum of electronic and thermal Enthalpies= -3456.279988
 Sum of electronic and thermal Free Energies= -3456.331133



M3 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	1.268769	1.335045	0.000000
2	6	0	-0.099907	1.090144	0.000000
3	6	0	-0.559554	-0.236145	0.000000

4	6	0	0.339201	-1.307527	0.000000
5	6	0	1.708961	-1.053284	0.000000
6	6	0	2.169174	0.264600	0.000000
7	1	0	1.639817	2.360909	0.000000
8	1	0	-0.820719	1.909923	0.000000
9	1	0	-0.045636	-2.329075	0.000000
10	1	0	2.418163	-1.881793	0.000000
11	1	0	3.242583	0.461764	0.000000
12	6	0	-2.009494	-0.512125	0.000000
13	8	0	-2.917138	0.249253	0.000000

Zero-point correction= 0.098664 (Hartree/Particle)

Thermal correction to Energy= 0.104923

Thermal correction to Enthalpy= 0.105867

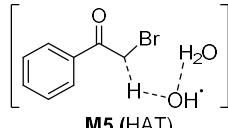
Thermal correction to Gibbs Free Energy= 0.067472

Sum of electronic and zero-point Energies= -344.419147

Sum of electronic and thermal Energies= -344.412888

Sum of electronic and thermal Enthalpies= -344.411944

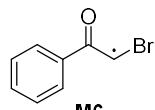
Sum of electronic and thermal Free Energies= -344.450339



Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	4.022685	0.239000	-0.081796
2	6	0	2.868538	0.987589	0.122419
3	6	0	1.630803	0.347160	0.275742
4	6	0	1.561660	-1.052141	0.229858
5	6	0	2.721113	-1.798319	0.032063
6	6	0	3.949031	-1.155053	-0.127056
7	1	0	4.982990	0.741014	-0.206911
8	1	0	2.910017	2.076866	0.162353
9	1	0	0.612827	-1.572582	0.361765
10	1	0	2.664479	-2.886973	0.001896
11	1	0	4.854576	-1.742526	-0.287431
12	6	0	0.434118	1.204535	0.510924
13	8	0	0.532490	2.390475	0.738407
14	6	0	-0.966038	0.594347	0.464091
15	1	0	-0.968770	-0.407169	1.058319
17	1	0	-2.119161	2.415108	0.407857
18	35	0	-1.325426	0.019271	-1.366053
19	1	0	-3.120705	-1.447253	1.065315
20	8	0	-3.953998	-1.280848	0.590234
21	1	0	-3.665425	-1.164141	-0.324179
22	8	0	-1.384511	-1.539974	1.842401
23	1	0	-1.479612	-1.072615	2.698958

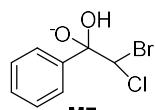
Zero-point correction= 0.189710 (Hartree/Particle)
 Thermal correction to Energy= 0.205051
 Thermal correction to Enthalpy= 0.205996
 Thermal correction to Gibbs Free Energy= 0.145352
 Sum of electronic and zero-point Energies= -3148.800360
 Sum of electronic and thermal Energies= -3148.785019
 Sum of electronic and thermal Enthalpies= -3148.784075
 Sum of electronic and thermal Free Energies= -3148.844718



Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	3.747070	0.328701	-0.458234
2	6	0	2.497608	0.923816	-0.602798
3	6	0	1.332552	0.322735	-0.100439
4	6	0	1.463420	-0.907264	0.551520
5	6	0	2.716133	-1.516189	0.689937
6	6	0	3.862309	-0.903578	0.191035
7	1	0	4.635696	0.824639	-0.853891
8	1	0	2.405393	1.880098	-1.121578
9	1	0	0.593298	-1.413662	0.963582
10	1	0	2.789095	-2.480186	1.197572
11	1	0	4.838031	-1.379370	0.305589
12	6	0	0.029506	1.120545	-0.387957
13	8	0	-0.116414	1.420920	-1.652831
14	6	0	-1.300674	0.455608	0.188597
15	1	0	-2.576315	1.495622	-1.169925
16	35	0	-1.642292	-1.290724	-0.675001
17	1	0	-0.237109	3.003167	-0.119938

Zero-point correction= 0.162515 (Hartree/Particle)
 Thermal correction to Energy= 0.174932
 Thermal correction to Enthalpy= 0.175876
 Thermal correction to Gibbs Free Energy= 0.123156
 Sum of electronic and zero-point Energies= -3532.140751
 Sum of electronic and thermal Energies= -3532.128334
 Sum of electronic and thermal Enthalpies= -3532.127390
 Sum of electronic and thermal Free Energies= -3532.180110



Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z

1	6	0	3.817764	0.161236	-0.506764
2	6	0	2.620738	0.809102	-0.807358
3	6	0	1.406613	0.327389	-0.303283
4	6	0	1.397532	-0.833932	0.481414
5	6	0	2.594823	-1.488948	0.766154
6	6	0	3.805959	-0.990326	0.281414
7	1	0	4.760869	0.551978	-0.892727
8	1	0	2.619333	1.701898	-1.435454
9	1	0	0.440641	-1.231384	0.825504
10	1	0	2.583479	-2.399638	1.367818
11	1	0	4.740407	-1.504407	0.513466
12	6	0	0.150926	1.047223	-0.702989
13	8	0	0.065962	1.513644	-1.851842
14	8	0	-0.935540	1.234271	0.219487
15	6	0	-2.160656	1.978250	-0.169521
16	1	0	-3.024937	1.595600	0.387067
17	35	0	-2.085973	-1.354491	-0.372140
18	17	0	-0.672811	1.071525	1.924610

Zero-point correction= 0.146112 (Hartree/Particle)

Thermal correction to Energy= 0.158464

Thermal correction to Enthalpy= 0.159409

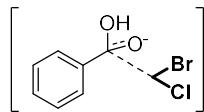
Thermal correction to Gibbs Free Energy= 0.104463

Sum of electronic and zero-point Energies= -3456.412566

Sum of electronic and thermal Energies= -3456.400214

Sum of electronic and thermal Enthalpies= -3456.399270

Sum of electronic and thermal Free Energies= -3456.454215



M8

Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	2.722365	-0.484265	1.438456
2	6	0	1.979743	0.618060	1.014177
3	6	0	1.507173	0.678216	-0.300966
4	6	0	1.796338	-0.363640	-1.188749
5	6	0	2.533710	-1.462564	-0.762599
6	6	0	2.998493	-1.525948	0.553396
7	1	0	3.094466	-0.523231	2.463646
8	1	0	1.790634	1.436705	1.710272
9	1	0	1.421576	-0.306791	-2.211593
10	1	0	2.746885	-2.276849	-1.457264
11	1	0	3.580690	-2.387043	0.886677
12	6	0	0.760494	1.863931	-0.817802
13	8	0	0.731963	2.156957	-1.996969
14	8	0	-1.448309	0.407560	-0.131736

15	1	0	-3.670082	0.326740	0.067041
16	1	0	-2.949629	1.907285	-0.359084
17	35	0	-1.382322	-1.602437	-0.299041
18	17	0	-1.343102	0.734201	1.736445
19	6	0	0.301989	2.750860	0.087865

Zero-point correction= 0.159869 (Hartree/Particle)

Thermal correction to Energy= 0.172875

Thermal correction to Enthalpy= 0.173819

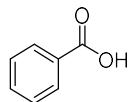
Thermal correction to Gibbs Free Energy= 0.117823

Sum of electronic and zero-point Energies= -3532.108983

Sum of electronic and thermal Energies= -3532.095977

Sum of electronic and thermal Enthalpies= -3532.095032

Sum of electronic and thermal Free Energies= -3532.151028



M9 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-1.836728	1.235735	-0.000052
2	6	0	-0.444006	1.202843	-0.000095
3	6	0	0.219012	-0.029527	-0.000028
4	6	0	-0.510420	-1.223463	0.000080
5	6	0	-1.901906	-1.185179	0.000053
6	6	0	-2.564079	0.044276	0.000023
7	1	0	-2.357900	2.193929	-0.000099
8	1	0	0.129471	2.129877	-0.000189
9	1	0	0.024348	-2.174279	0.000082
10	1	0	-2.473239	-2.114387	0.000078
11	1	0	-3.654928	0.073785	0.000035
12	6	0	1.706541	-0.114259	-0.000007
13	8	0	2.325716	-1.154471	-0.000152
14	8	0	2.305835	1.075773	0.000179
15	1	0	3.269360	0.938091	0.000038

Zero-point correction= 0.116781 (Hartree/Particle)

Thermal correction to Energy= 0.123817

Thermal correction to Enthalpy= 0.124761

Thermal correction to Gibbs Free Energy= 0.084814

Sum of electronic and zero-point Energies= -420.236407

Sum of electronic and thermal Energies= -420.229371

Sum of electronic and thermal Enthalpies= -420.228427

Sum of electronic and thermal Free Energies= -420.268375



M10 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-0.586512	0.416229	0.443946
2	1	0	-0.166776	2.502954	0.211111
3	35	0	1.219050	-0.174563	-0.031444
4	17	0	-1.781906	-0.805118	-0.074563

Zero-point correction= 0.057215 (Hartree/Particle)

Thermal correction to Energy= 0.062156

Thermal correction to Enthalpy= 0.063100

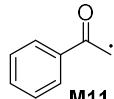
Thermal correction to Gibbs Free Energy= 0.027507

Sum of electronic and zero-point Energies= -3112.390793

Sum of electronic and thermal Energies= -3112.385852

Sum of electronic and thermal Enthalpies= -3112.384908

Sum of electronic and thermal Free Energies= -3112.420501



M11 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	2.482951	1.010446	0.080407
2	6	0	1.103742	1.195457	0.062222
3	6	0	0.237213	0.096121	-0.010039
4	6	0	0.778390	-1.195310	-0.073254
5	6	0	2.160527	-1.379069	-0.059825
6	6	0	3.014823	-0.279235	0.019661
7	1	0	3.147192	1.874020	0.139999
8	1	0	0.679347	2.199030	0.106722
9	1	0	0.131879	-2.070312	-0.143690
10	1	0	2.570371	-2.388746	-0.113939
11	1	0	4.095922	-0.427208	0.032258
12	6	0	-1.239451	0.371430	-0.032934
13	8	0	-1.641993	1.537839	-0.149381
14	6	0	-2.173775	-0.716608	0.094845
15	1	0	-4.088685	-1.114305	-0.742138
16	1	0	-4.099094	-0.831619	0.993861

Zero-point correction= 0.154471 (Hartree/Particle)

Thermal correction to Energy= 0.163555

Thermal correction to Enthalpy= 0.164499

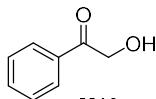
Thermal correction to Gibbs Free Energy= 0.118696

Sum of electronic and zero-point Energies= -422.906952

Sum of electronic and thermal Energies= -422.897868

Sum of electronic and thermal Enthalpies= -422.896924

Sum of electronic and thermal Free Energies= -422.942727



M12 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	2.768014	-0.935258	0.301830
2	6	0	1.407795	-1.221217	0.253424
3	6	0	0.488006	-0.221999	-0.094073
4	6	0	0.946084	1.066721	-0.400256
5	6	0	2.310238	1.348553	-0.355633
6	6	0	3.219891	0.350680	-0.004618
7	1	0	3.479525	-1.715069	0.576709
8	1	0	1.039210	-2.221004	0.487083
9	1	0	0.248408	1.856233	-0.682401
10	1	0	2.663962	2.351171	-0.599100
11	1	0	4.287367	0.574846	0.029572
12	6	0	-0.956574	-0.584801	-0.144550
13	8	0	-1.321710	-1.743130	-0.061382
14	6	0	-2.020346	0.503887	-0.241876
15	1	0	-1.180758	1.756890	1.348601
17	1	0	-1.741921	1.211279	-1.037504
18	8	0	-3.252258	-0.081400	-0.590790
19	1	0	-3.208898	-1.003515	-0.287366

Zero-point correction= 0.172953 (Hartree/Particle)

Thermal correction to Energy= 0.183003

Thermal correction to Enthalpy= 0.183948

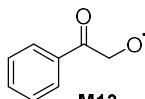
Thermal correction to Gibbs Free Energy= 0.137052

Sum of electronic and zero-point Energies= -498.678195

Sum of electronic and thermal Energies= -498.668145

Sum of electronic and thermal Enthalpies= -498.667201

Sum of electronic and thermal Free Energies= -498.714097



M13 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	2.727771	-0.915925	0.295505
2	6	0	1.371466	-1.216720	0.235887
3	6	0	0.441799	-0.224236	-0.106365
4	6	0	0.886520	1.073394	-0.392975
5	6	0	2.247115	1.369591	-0.335272
6	6	0	3.166707	0.378580	0.008283
7	1	0	3.445747	-1.691525	0.565329

8	1	0	1.013149	-2.222856	0.456946
9	1	0	0.185990	1.862686	-0.667392
10	1	0	2.589586	2.379761	-0.562785
11	1	0	4.231169	0.614838	0.052099
12	6	0	-0.996347	-0.627213	-0.152141
13	8	0	-1.339301	-1.783370	-0.062348
14	6	0	-2.056126	0.497056	-0.249917
15	1	0	-1.262850	1.647288	1.425094
16	1	0	-1.713144	1.210912	-1.019111
17	8	0	-3.231446	-0.103869	-0.654252

Zero-point correction= 0.156890 (Hartree/Particle)

Thermal correction to Energy= 0.166511

Thermal correction to Enthalpy= 0.167455

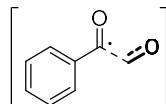
Thermal correction to Gibbs Free Energy= 0.120704

Sum of electronic and zero-point Energies= -498.009657

Sum of electronic and thermal Energies= -498.000036

Sum of electronic and thermal Enthalpies= -497.999092

Sum of electronic and thermal Free Energies= -498.045843



M14 (Bz-release) Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-2.860358	0.823331	0.229428
2	6	0	-1.531293	1.227069	0.191495
3	6	0	-0.528181	0.284247	-0.089122
4	6	0	-0.856454	-1.054266	-0.334627
5	6	0	-2.191523	-1.446918	-0.296085
6	6	0	-3.189088	-0.512718	-0.013418
7	1	0	-3.643444	1.550557	0.447014
8	1	0	-1.254137	2.265419	0.379495
9	1	0	-0.084310	-1.791259	-0.553960
10	1	0	-2.452573	-2.488110	-0.487496
11	1	0	-4.233170	-0.828400	0.016557
12	6	0	0.860896	0.788410	-0.102602
13	8	0	1.257838	1.887884	0.025122
14	6	0	2.347758	-0.578878	-0.307445
15	8	0	3.322641	0.096718	-0.698529

Zero-point correction= 0.156743 (Hartree/Particle)

Thermal correction to Energy= 0.166729

Thermal correction to Enthalpy= 0.167673

Thermal correction to Gibbs Free Energy= 0.119904

Sum of electronic and zero-point Energies= -498.000007

Sum of electronic and thermal Energies= -497.990021

Sum of electronic and thermal Enthalpies= -497.989077

Sum of electronic and thermal Free Energies= -498.036846



M15 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	1.828668	1.208071	0.000045
2	6	0	0.433854	1.205681	0.000043
3	6	0	-0.275007	0.000017	-0.000001
4	6	0	0.434033	-1.205725	-0.000047
5	6	0	1.828676	-1.208057	-0.000051
6	6	0	2.528595	0.000084	-0.000002
7	1	0	2.373521	2.153947	0.000082
8	1	0	-0.122960	2.143679	0.000072
9	1	0	-0.122914	-2.143659	-0.000083
10	1	0	2.373675	-2.153850	-0.000091
11	1	0	3.620084	0.000172	-0.000002
12	6	0	-1.805690	-0.000122	0.000001
13	8	0	-2.372419	-1.116446	-0.000010

Zero-point correction= 0.104108 (Hartree/Particle)

Thermal correction to Energy= 0.110869

Thermal correction to Enthalpy= 0.111813

Thermal correction to Gibbs Free Energy= 0.072283

Sum of electronic and zero-point Energies= -419.781501

Sum of electronic and thermal Energies= -419.774739

Sum of electronic and thermal Enthalpies= -419.773795

Sum of electronic and thermal Free Energies= -419.813326

C1 (HCHO)

M16 Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	1.160175	-0.151889	-0.000023
2	1	0	1.695389	0.227765	0.883543
3	1	0	1.144703	-1.248338	-0.000142
4	8	0	-1.230589	-0.274916	-0.000004

Zero-point correction= 0.055942 (Hartree/Particle)

Thermal correction to Energy= 0.059782

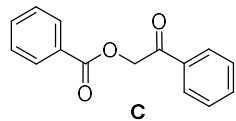
Thermal correction to Enthalpy= 0.060726

Thermal correction to Gibbs Free Energy= 0.031099

Sum of electronic and zero-point Energies= -153.586213

Sum of electronic and thermal Energies= -153.582372

Sum of electronic and thermal Enthalpies= -153.581428
 Sum of electronic and thermal Free Energies= -153.611055



Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	6	0	-2.795722	2.482452	0.871441
2	6	0	-2.758426	1.092882	0.925799
3	6	0	-2.138604	0.364396	-0.097781
4	6	0	-1.566926	1.038470	-1.185154
5	6	0	-1.618619	2.430396	-1.242180
6	6	0	-2.225690	3.152065	-0.213951
7	1	0	-3.271368	3.046158	1.675301
8	1	0	-3.203222	0.555531	1.764968
9	1	0	-1.090401	0.484495	-1.995138
10	1	0	-1.179976	2.953236	-2.093156
11	1	0	-2.255609	4.241983	-0.258255
12	6	0	-2.134125	-1.126235	0.014462
13	8	0	-2.944068	-1.709001	0.702632
14	6	0	-1.118057	-1.927138	-0.806506
15	1	0	-1.994880	-3.859602	-0.494536
16	8	0	0.167696	-1.290977	-0.826536
17	6	0	0.704246	-0.924656	0.339271
18	8	0	0.144638	-1.098368	1.399325
19	6	0	2.032899	-0.272742	0.191049
20	6	0	2.694366	0.136945	1.354233
21	6	0	2.615035	-0.062212	-1.064732
22	6	0	3.937884	0.755070	1.262434
23	1	0	2.225139	-0.033967	2.324163
24	6	0	3.859243	0.557888	-1.150292
25	1	0	2.094782	-0.382216	-1.967497
26	6	0	4.519288	0.965585	0.010257
27	1	0	4.456362	1.073669	2.167662
28	1	0	4.316346	0.723210	-2.126880
29	1	0	5.494395	1.450202	-0.062535
30	1	0	-1.452407	-1.870301	-1.853529

Zero-point correction= 0.265010 (Hartree/Particle)

Thermal correction to Energy= 0.281091

Thermal correction to Enthalpy= 0.282035

Thermal correction to Gibbs Free Energy= 0.218970

Sum of electronic and zero-point Energies= -842.601794

Sum of electronic and thermal Energies= -842.585713

Sum of electronic and thermal Enthalpies= -842.584769

Sum of electronic and thermal Free Energies= -842.647834

OH_radical_H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	-0.456857	-0.170539	0.007052
2	8	0	-1.424601	-0.095257	-0.018258
3	1	0	-1.567166	0.847908	0.132849
4	8	0	1.486171	-0.098205	0.016871
5	1	0	1.531460	0.870324	-0.128803

Zero-point correction= 0.032580 (Hartree/Particle)

Thermal correction to Energy= 0.037921

Thermal correction to Enthalpy= 0.038865

Thermal correction to Gibbs Free Energy= 0.006091

Sum of electronic and zero-point Energies= -151.951522

Sum of electronic and thermal Energies= -151.946181

Sum of electronic and thermal Enthalpies= -151.945237

Sum of electronic and thermal Free Energies= -151.978012

OH_anion_H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	-0.115645	-0.118459	0.012077
2	8	0	-1.214742	-0.078865	0.068666
3	1	0	-1.402224	0.684168	-0.490767
4	8	0	1.232769	-0.079565	-0.068811
5	1	0	1.373658	0.701728	0.479846

Zero-point correction= 0.031760 (Hartree/Particle)

Thermal correction to Energy= 0.035780

Thermal correction to Enthalpy= 0.036724

Thermal correction to Gibbs Free Energy= 0.008021

Sum of electronic and zero-point Energies= -152.136076

Sum of electronic and thermal Energies= -152.132056

Sum of electronic and thermal Enthalpies= -152.131112

Sum of electronic and thermal Free Energies= -152.159815

H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	0.753364	-0.483677	0.000000
2	8	0	0.000000	0.120916	0.000000
3	1	0	-0.753364	-0.483653	0.000000

Zero-point correction= 0.021290 (Hartree/Particle)
 Thermal correction to Energy= 0.024126
 Thermal correction to Enthalpy= 0.025070
 Thermal correction to Gibbs Free Energy= 0.002972
 Sum of electronic and zero-point Energies= -76.314093
 Sum of electronic and thermal Energies= -76.311257
 Sum of electronic and thermal Enthalpies= -76.310313
 Sum of electronic and thermal Free Energies= -76.332411

H, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	0.000000	0.000000	0.000000

Zero-point correction= 0.000000 (Hartree/Particle)
 Thermal correction to Energy= 0.001416
 Thermal correction to Enthalpy= 0.002360
 Thermal correction to Gibbs Free Energy= -0.010654
 Sum of electronic and zero-point Energies= -0.496666
 Sum of electronic and thermal Energies= -0.495249
 Sum of electronic and thermal Enthalpies= -0.494305
 Sum of electronic and thermal Free Energies= -0.507320

H₂, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	0.000000	0.000000	0.376835
2	1	0	0.000000	0.000000	-0.376835

Zero-point correction= 0.010244 (Hartree/Particle)
 Thermal correction to Energy= 0.012604
 Thermal correction to Enthalpy= 0.013548
 Thermal correction to Gibbs Free Energy= -0.001271
 Sum of electronic and zero-point Energies= -1.153268
 Sum of electronic and thermal Energies= -1.150907
 Sum of electronic and thermal Enthalpies= -1.149963
 Sum of electronic and thermal Free Energies= -1.164782

Cl radical H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	17	0	-1.275059	0.004703	-0.000004
2	1	0	1.296455	-0.214622	0.000224

3	8	0	2.254618	-0.091964	-0.000025
4	1	0	2.342604	0.870378	0.000034

Zero-point correction= 0.023422 (Hartree/Particle)
 Thermal correction to Energy= 0.027696
 Thermal correction to Enthalpy= 0.028640
 Thermal correction to Gibbs Free Energy= -0.003139
 Sum of electronic and zero-point Energies= -536.290007
 Sum of electronic and thermal Energies= -536.285733
 Sum of electronic and thermal Enthalpies= -536.284789
 Sum of electronic and thermal Free Energies= -536.316568

Cl H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	17	0	-1.085588	0.005320	0.000000
2	1	0	0.935006	-0.136464	0.000000
3	8	0	1.928785	-0.100678	0.000000
4	1	0	2.089715	0.851440	0.000000

Zero-point correction= 0.022866 (Hartree/Particle)
 Thermal correction to Energy= 0.026690
 Thermal correction to Enthalpy= 0.027634
 Thermal correction to Gibbs Free Energy= -0.001793
 Sum of electronic and zero-point Energies= -536.510141
 Sum of electronic and thermal Energies= -536.506317
 Sum of electronic and thermal Enthalpies= -536.505373
 Sum of electronic and thermal Free Energies= -536.534800

Br radical H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	-0.457141	2.099482	0.000000
2	8	0	0.026562	2.934687	0.000000
3	1	0	-0.685037	3.588904	0.000000
4	35	0	0.026562	-0.833311	0.000000

Zero-point correction= 0.022545 (Hartree/Particle)
 Thermal correction to Energy= 0.027287
 Thermal correction to Enthalpy= 0.028231
 Thermal correction to Gibbs Free Energy= -0.006056
 Sum of electronic and zero-point Energies= -2650.156221
 Sum of electronic and thermal Energies= -2650.151478
 Sum of electronic and thermal Enthalpies= -2650.150534
 Sum of electronic and thermal Free Energies= -2650.184822

Br H₂O, Standard orientation:

Center Number	Atomic Number	Atomic Type	Coordinates (Angstroms)		
			X	Y	Z
1	1	0	1.523403	-0.178407	0.000036
2	8	0	2.509309	-0.096174	-0.000004
3	1	0	2.618981	0.863277	0.000004
4	35	0	-0.691910	0.002415	0.000000

Zero-point correction= 0.023125 (Hartree/Particle)

Thermal correction to Energy= 0.026993

Thermal correction to Enthalpy= 0.027937

Thermal correction to Gibbs Free Energy= -0.002794

Sum of electronic and zero-point Energies= -2650.361806

Sum of electronic and thermal Energies= -2650.357938

Sum of electronic and thermal Enthalpies= -2650.356994

Sum of electronic and thermal Free Energies= -2650.387725