

Supporting Information

Prominence of Cu in Plasmonic Cu-Ag alloy Decorated SiO₂@S-doped C₃N₄ Core Shell Nanostructured Photocatalyst towards Enhanced Visible Light Activity

Pradeepta Babu¹, Soumya Ranjan Dash², Arjun Behera¹

T. Vijayaraghavan³, Anuradha Ashok³ and Kulamani Parida^{1*}

¹*Centre for Nanoscience and Nanotechnology (CNSNT), Siksha 'O' Anusandhan, Bhubaneswar 751030, India*

²*Physical and Material Chemistry Division, CSIR-NCL Pune, Dr. Homi Bhaba Road, Pune 411008, India*

³*Functional Materials Laboratory, PSG Institute of Advanced Studies, Coimbatore, 641004, India*

*Corresponding Author: kulamaniparida@soa.ac.in&paridakulamani@yahoo.com.

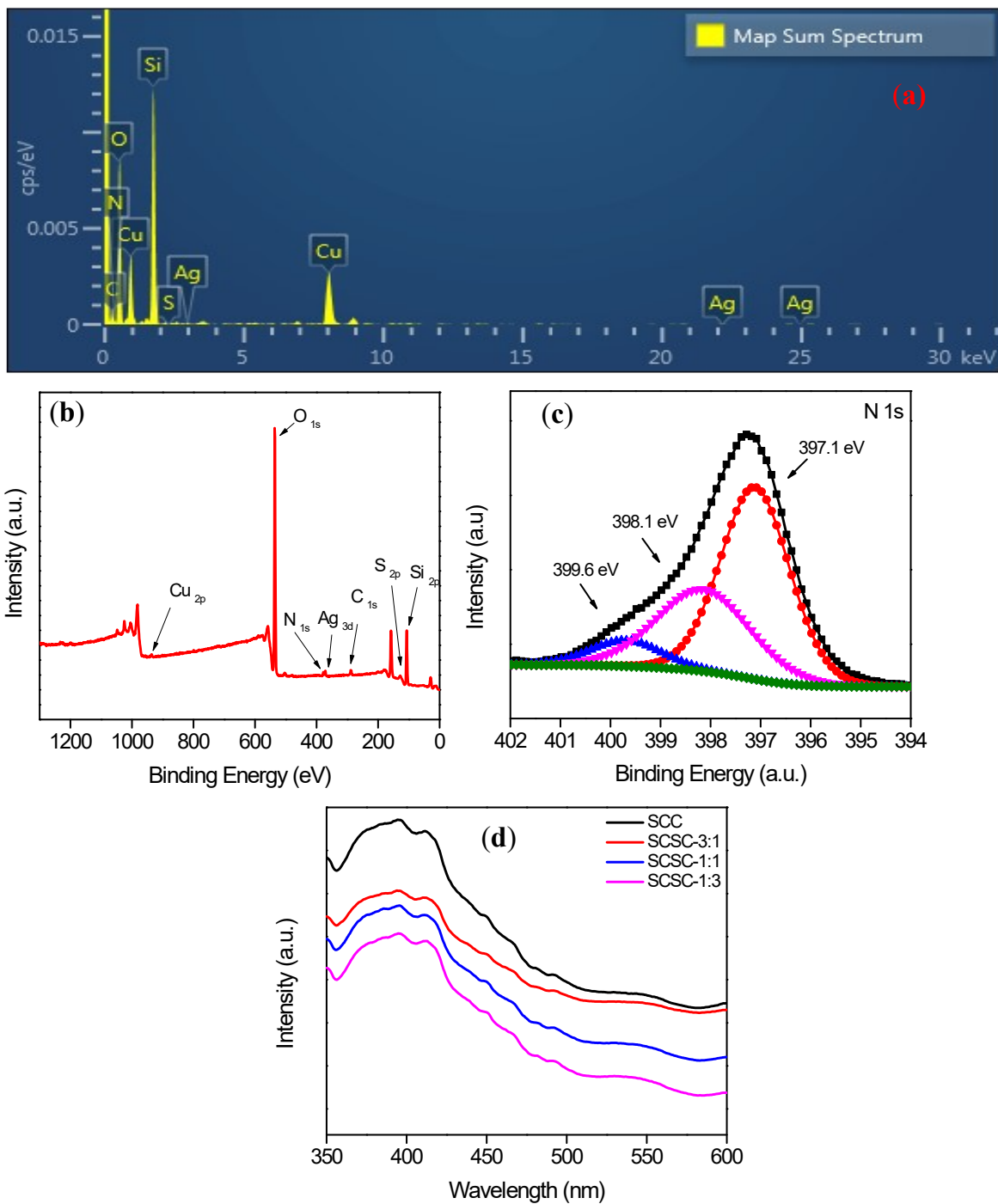
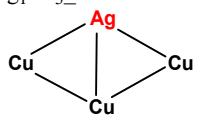
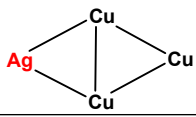


Figure S1. (a) EDX spectra of SCSC-1:3, (b) XPS survey spectrum (c) deconvoluted N 1s spectra (d) Photoluminescence spectra of SCC, SCSC-3:1, SCSC-1:1, SCSC-1:3.

Photocatalyst	Synthesis method	Activity	Reference
Ag-Cu@TiO ₂ microrod	Surfactant assisted sol-gel method	Water disinfection 8 h	1
Ag-Cu@ZrO ₂	Impregnation reduction method	Nitro benzene to azoxy benzene 75%, 16 h	2
Ag-Cu/TiO ₂	Wet impregnation method	H ₂ generation 16,650 μmol h ⁻¹ g ⁻¹	3
Ag-Cu alloy	Impregnation reduction method	4-Nitro phenol reduction 4 min	4
Ag-Cu alloy	Galvanic replacement reaction method	Rhodamine-6g degradation 20 min 88%	5
Montmorillonite decorated MWCNTs/TiO ₂	sol-gel and wet impregnation method	1888 ppm h ⁻¹ of H ₂	6
Ni-MMT/TiO ₂	sol-gel method	3470 μmole g-cat. ⁻¹ h ⁻¹ of H ₂	7
SiO ₂ @Ag-Cu@CS	chemical reduction method	Hydrogen generation 1730 μmol h ⁻¹ g ⁻¹ Phenol oxidation 81.9%, 90 min	This work

Table S1. Literature survey on various Ag-Cu alloy based photocatalyst

Models	ΔG_f (S=1)
Ag ₁ Cu ₃ _Alter 	-92.7 kcal/mol
Ag ₁ Cu ₃ 	-95.1 kcal/mol

Ag ₁ Cu ₃		
ΔG_f (S=1)	ΔG_f (S=3)	ΔG_f (S=5)
-95.1 kcal/mol	-78.8 kcal/mol	-10.8 kcal/mol

Table S2. (Left) Two probable models of Ag₁Cu₃ with their free energy of formation (ΔG_f) and (Right) Ground state spin multiplicity (S) check for the most stable model (highlighted).

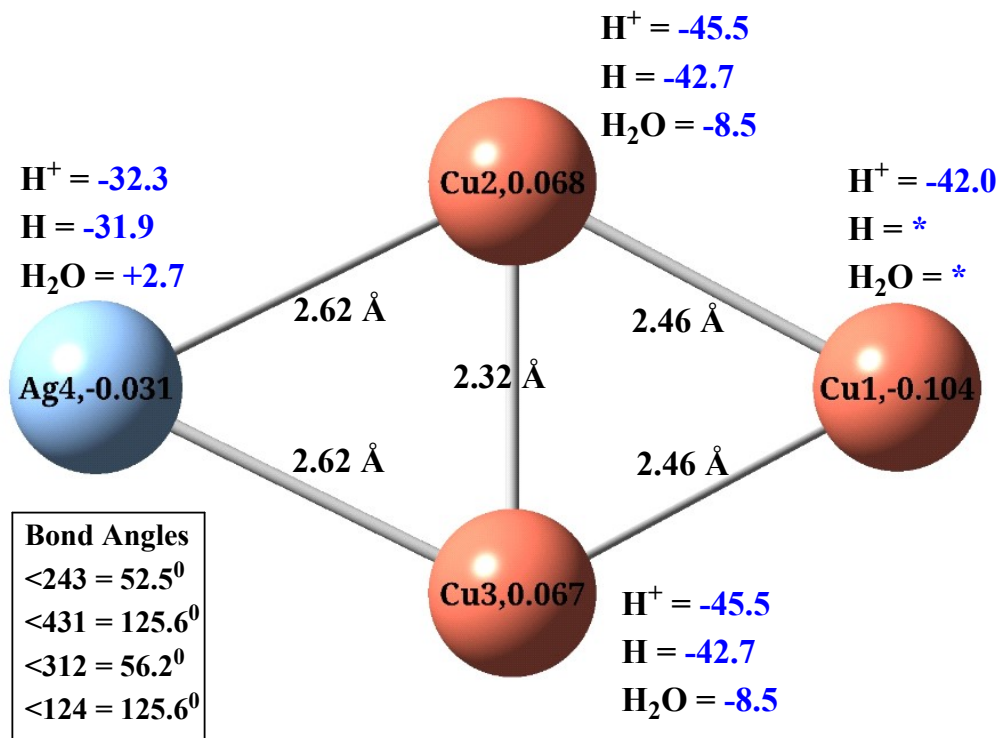


Figure S2. The most stable Ag_1Cu_3 optimized model with label, NBO values, and bond parameters. The binding energy (ΔG , in kcal/mol) for water and H^+/H at different sites are shown. *Optimized model for H_2O and H bound to Cu1 could not be obtained.

Optimized Cartesian Coordinates

Models				
Ag_1Cu_3 _Alter	Cu	2.21996230	-0.29751054	0.00019307
	Cu	0.00200260	-1.31999956	-0.00022874
	Cu	-2.21845134	-0.30077935	0.00019315
	Ag	-0.00216794	1.18362541	-0.00009717
Ag_1Cu_3	Cu	2.52858158	-0.00022633	-0.00040731
	Cu	0.35542941	1.16078177	0.00039167
	Cu	0.35485973	-1.16062738	0.00039178
	Ag	-1.99845214	0.00004439	-0.00023208
Ag_1Cu_3 _H ⁺ (Cu1)	Cu	1.21245057	1.24370500	-0.01733171
	Cu	0.51078044	-0.93569508	-1.19726966
	Cu	0.52871753	-0.90690360	1.21102745
	Ag	-1.43563358	0.31657660	0.00265116
	H	2.16827060	2.48881670	-0.02096067
Ag_1Cu_3 _H ⁺ (Cu2)	Cu	2.24172711	-1.07335156	-0.18136026
	Cu	0.05369752	-0.17305177	0.44816120
	Cu	1.98655502	1.23703255	-0.07816545
	Ag	-2.61124767	0.01507598	-0.14038396
	H	-1.44876927	-0.43681840	1.12761696
Ag_1Cu_3 _H ⁺ (Cu3)	Cu	-2.24225097	-1.07277692	-0.18142497
	Cu	-1.98551329	1.23702400	-0.07802346
	Cu	-0.05369825	-0.17337591	0.44807835
	Ag	2.61092190	0.01495188	-0.14037084
	H	1.44908354	-0.43800216	1.12716197
Ag_1Cu_3 _H ⁺ (Ag4)	Cu	0.87387761	-0.32218027	1.35497036
	Cu	0.83711928	1.34875310	-0.40911323
	Cu	0.83884787	-1.02077993	-0.96830352
	Ag	-1.50485908	-0.00355975	0.01544855
	H	-3.21712103	-0.00068586	-0.07513656
Ag_1Cu_3 _H (Cu2)	Cu	-2.45583028	-0.10981486	0.00003463
	Cu	-0.39232456	1.29461056	-0.00003757
	Cu	-0.34746202	-1.19712836	-0.00003251
	Ag	1.97356055	-0.05297529	0.00001859
	H	-0.08445717	2.84748587	0.00015399

Ag ₁ Cu ₃ _H (Cu3)	Cu	2.45583023	-0.10981190	0.00003776
	Cu	0.34746426	-1.19712851	-0.00003594
	Cu	0.39232160	1.29460937	-0.00003985
	Ag	-1.97355944	-0.05297616	0.00002033
	H	0.08442722	2.84747950	0.00014732
Ag ₁ Cu ₃ _H (Ag4)	Cu	-2.58188530	-0.00627239	0.00017448
	Cu	-0.39107946	1.17805095	-0.00016125
	Cu	-0.42219706	-1.15621124	-0.00016671
	Ag	2.01521023	-0.00523149	0.00009193
	H	3.74481171	-0.20557225	0.00013024
Ag ₁ Cu ₃ _H ₂ O (Cu2)	Cu	2.48224306	-0.43797130	-0.00391374
	Cu	0.41587804	0.97067946	-0.03361236
	Cu	0.23618434	-1.36356015	0.02336514
	Ag	-2.03974042	-0.14462262	-0.00189448
	O	0.50391326	3.01689194	-0.04295119
	H	1.23962285	3.34728485	0.50071384
	H	-0.29798690	3.40955012	0.34260411
Ag ₁ Cu ₃ _H ₂ O (Cu3)	Cu	2.48200769	-0.43893296	-0.00214574
	Cu	0.23489512	-1.36357902	0.02325259
	Cu	0.41767361	0.97043811	-0.03543715
	Ag	-2.03992503	-0.14371077	-0.00221443
	O	0.48940050	3.01681080	-0.04133851
	H	1.30097687	3.35249669	0.37642162
	H	-0.24242060	3.39756521	0.47394345
Ag ₁ Cu ₃ _H ₂ O (Ag4)	Cu	2.87033351	-0.00515582	0.00178524
	Cu	0.66291997	1.15583803	0.00058301
	Cu	0.67125713	-1.14936818	-0.00121563
	Ag	-1.69479492	0.00558987	-0.00222839
	O	-4.16872693	-0.13345876	-0.00140021
	H	-4.46795956	0.44370156	-0.72464322
	H	-4.45767151	0.32313767	0.80715322
Ag ₁ Cu ₃ _H ₂ O (TS_Cu)	Cu	2.47865031	-0.53937700	0.01018726
	Cu	-0.02269613	-1.40023577	0.00051899
	Cu	0.39050424	0.92140602	-0.03840866
	Ag	-2.08903576	0.10434912	0.00939661
	O	1.48504688	2.58170143	-0.06251963
	H	2.16014998	1.15232463	-0.01405635
	H	1.59686154	2.81765077	0.87594274
Ag ₁ Cu ₃ _H ₂ O (TS_Ag)	Cu	2.57169720	-0.08100949	0.01267018
	Cu	0.53527584	-1.34026472	0.00417181
	Cu	0.38328165	1.00925157	-0.03605578
	Ag	-2.03058033	-0.31494350	0.00373512
	O	-0.48287147	2.78646681	-0.05980607
	H	-0.57805186	3.03085911	0.87809516
	H	-1.33908636	1.42840738	-0.01799735

Ag ₁ Cu ₃ _H ₂ O (Unfavorable Product)	Cu	-0.02782248	1.17267576	1.19545425
	Cu	-0.02786914	1.17288853	-1.19531741
	Cu	1.41549583	-0.41258752	-0.00007966
	Ag	-1.34222795	-0.85212917	-0.00002930
	O	2.82023800	-1.60577910	-0.00004582
	H	3.64428776	-1.09235647	0.00018828
	H	-2.55580028	-2.06766620	-0.00010267
Ag ₁ Cu ₃ _H ₂ O (Final Product)	Cu	0.17818690	1.52362615	-0.83970663
	Cu	0.27235314	0.59020926	1.51139345
	Cu	-1.38104414	-0.38948750	-0.00337532
	Ag	1.17832789	-0.95932420	-0.30863423
	O	-3.14182579	-0.77245098	-0.40530724
	H	-3.41383093	-1.51122076	0.16279795
	H	0.15164549	2.77297649	-1.79556449
Ag ₁ Cu ₃ _2H (Int1)	Cu	0.01118499	1.45009595	-0.00005403
	Cu	-2.11642242	0.25164947	0.00004565
	Cu	2.13166815	0.30744331	0.00004966
	Ag	0.06112552	-1.25020941	-0.00002736
	H	-1.67914500	1.87884767	0.00000335
	H	-1.96024556	-1.38547864	0.00008496
Ag ₁ Cu ₃ _2H (TS)	Cu	2.51551430	-0.01527870	0.00001708
	Cu	0.28959920	1.10120901	-0.00003319
	Cu	0.32888630	-1.26088389	-0.00001325
	Ag	-1.99088644	0.01599234	0.00001259
	H	1.80746034	1.71422847	0.00011673
	H	0.87820778	2.60778547	0.00014324
Ag ₁ Cu ₃ _H ₂ (Int3)	Cu	2.51629965	-0.08261558	-0.00095255
	Cu	0.37335256	1.12815708	-0.00002578
	Cu	0.32979722	-1.19379637	0.00069779
	Ag	-2.00058240	-0.03464094	-0.00003735
	H	0.71795060	2.97451694	0.00499625
	H	-0.05461125	2.95299854	0.00489507

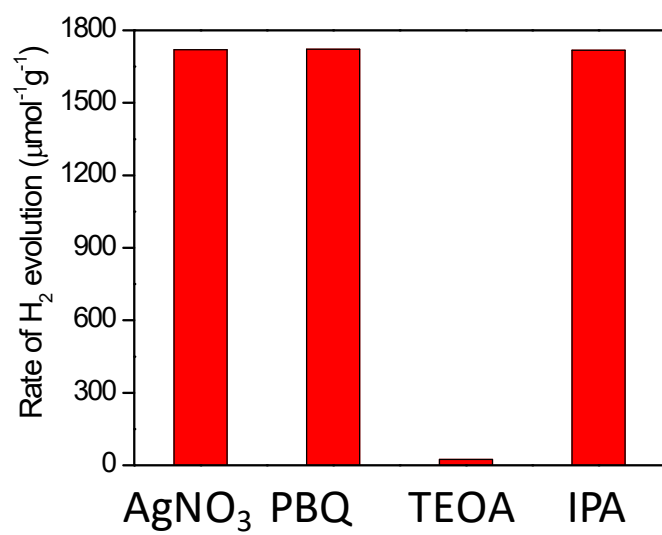


Figure S3. Effect of radical scavenger on hydrogen evolution reaction.

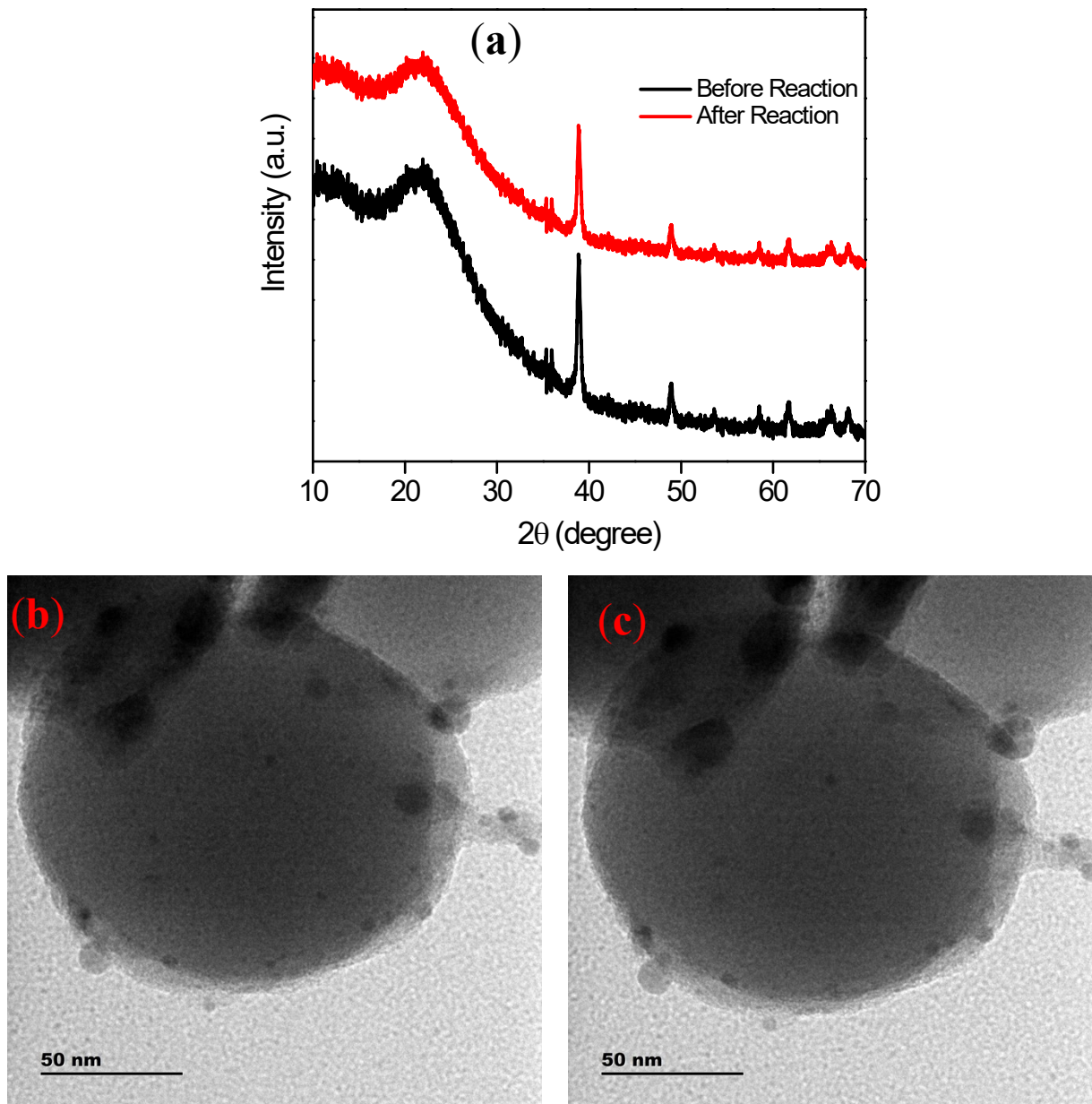
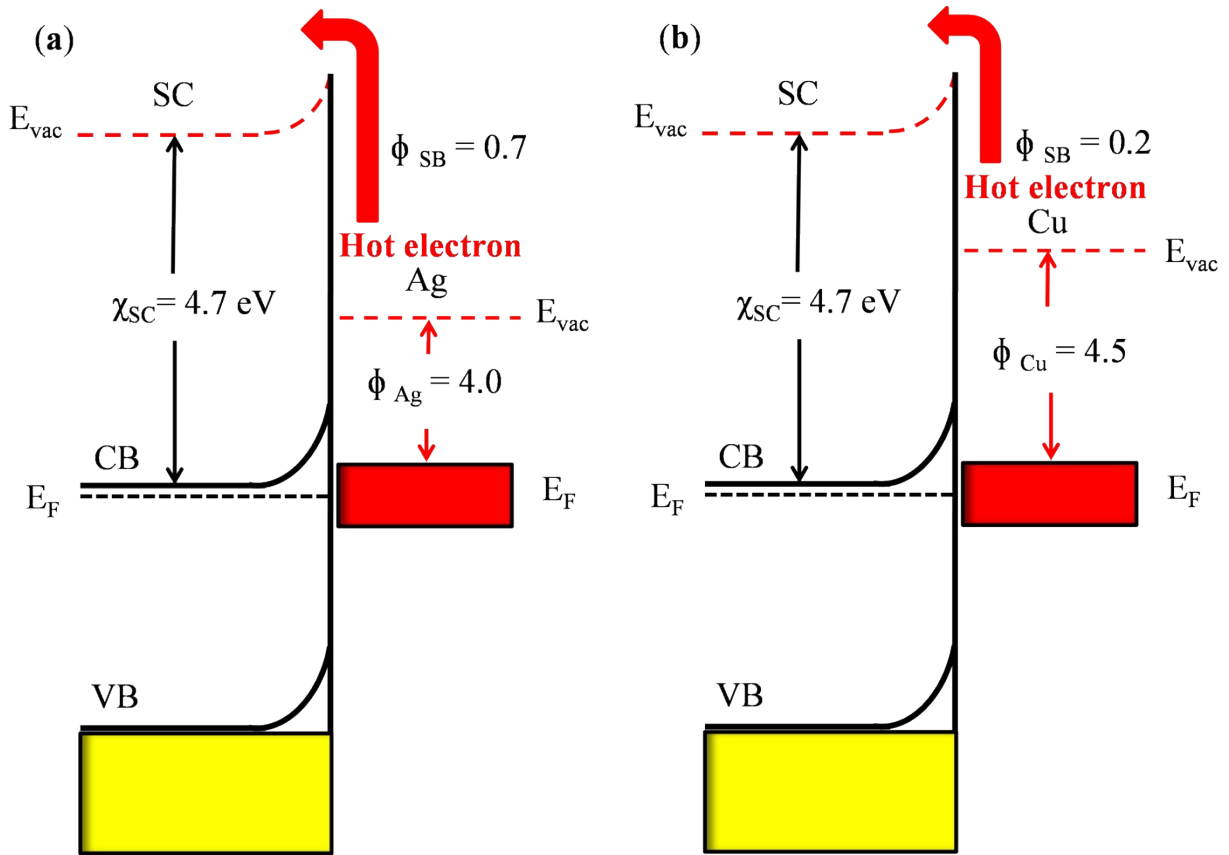


Figure S4. (a) PXRD of SCSC-1:3 before and after photocatalytic reaction, (b,c) HRTEM image of SCSC-1:3 before and after photocatalytic reaction suggesting the crystal phase and morphology remain intact after the reusability test.



Scheme S1. (a) Energy band diagram of SCS and SCC, where E_F , E_{vac} , χ_{SC} , CB, VB, ϕ_{SB} , ϕ_{Ag} , ϕ_{Cu} represents the Fermi energy level, vacuum energy level, electronegativity of conduction band of SC, conduction band and valance band of SC, Schottky barrier height, and work function of Ag and Cu on monometallic form.

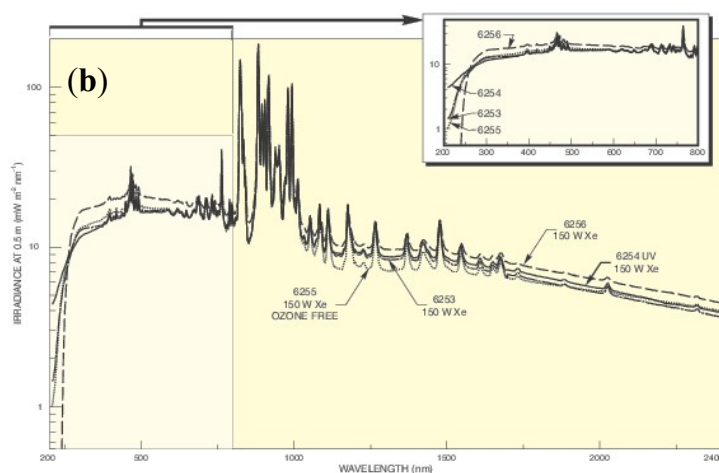
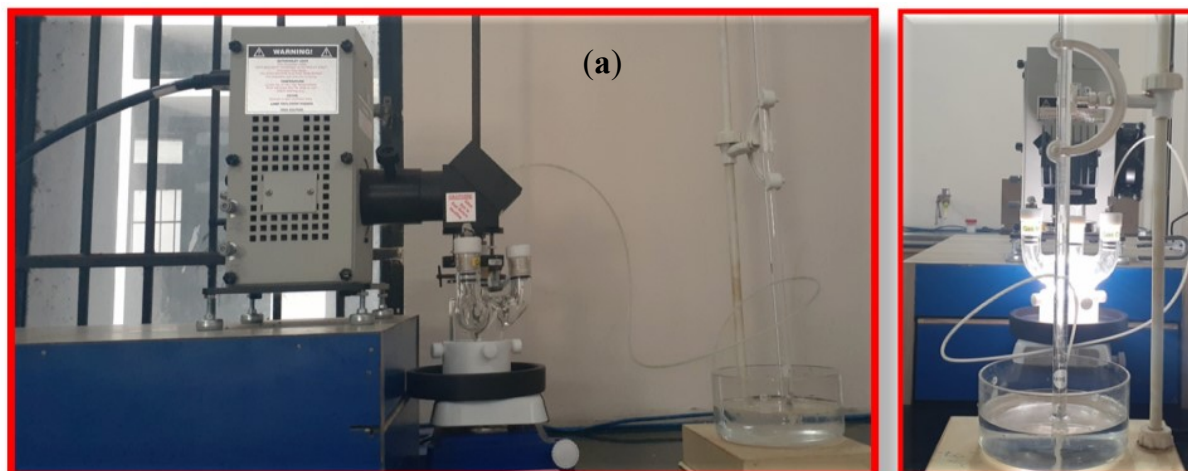


Figure S5. (a) photocatalytic reactor for water splitting reaction, (b) Spectral pattern of 150W Xenon lamp.

References

1. Zhang, Y.; Wang, L.; Kong, X.; Jiang, H.; Zhang, F.; Shi, J. Novel Ag-Cu bimetallic alloy decorated near-infrared responsive three-dimensional rod-like architectures for efficient photocatalytic water purification. *J. Colloid Interface Sci.* **2018**, *522*, 29-39 10.1016/j.jcis.2018.02.005.
2. Liu, Z.; Huang, Y.; Xiao, Q.; Zhu, H. Selective reduction of nitroaromatics to azoxy compounds on supported Ag-Cu alloy nanoparticles through visible light irradiation. *Green Chem.* **2016**, *18* (3), 817-825 10.1039/C5GC01726B.

3. Bhavani, K.; Naresh, G.; Srinivas, B.; Venugopal, A. Plasmonic resonance nature of Ag-Cu/TiO₂ photocatalyst under solar and artificial light: Synthesis, characterization and evaluation of H₂O splitting activity. *Appl. Catal., B* **2016**, *199*, 282-291
10.1016/j.apcatb.2016.06.050.
4. Wu, W.; Lei, M.; Yang, S.; Zhou, L.; Liu, L.; Xiao, X.; Jiang, C.; Roy, V.A. A one-pot route to the synthesis of alloyed Cu/Ag bimetallic nanoparticles with different mass ratios for catalytic reduction of 4-nitrophenol. *J. Mater. Chem. A* **2015**, *3* (7), 3450-3455
10.1039/C4TA06567K.
5. He, L.; Liu, C.; Tang, J.; Zhou, Y.; Yang, H.; Liu, R.; Hu, J. Self-catalytic stabilized Ag-Cu nanoparticles with tailored SERS response for plasmonic photocatalysis. *Appl. Surf. Sci.* **2018**, *434*, 265-272
10.1016/j.apsusc.2017.10.155.
6. Umer, M., Tahir, M., Azam, M.U. and Jaffar, M.M., Metals free MWCNTs@ TiO₂@ MMT heterojunction composite with MMT as a mediator for fast charges separation towards visible light driven photocatalytic hydrogen evolution. *Applied Surface Science*, **2019**, *463*, 747-757. 10.1016/j.ijhydene.2017.09.116
7. Tahir, M. Ni/MMT-promoted TiO₂ nanocatalyst for dynamic photocatalytic H₂ and hydrocarbons production from ethanol-water mixture under UV-light. *International Journal of Hydrogen Energy*, **2017**, *42*, 47, 28309-28326.
10.1016/j.ijhydene.2017.09.116