Facile Synthesis of Core-Shell PtCu Nanoparticles with Superior Electrocatalytic Activity and Stability in Hydrogen Evolution Reaction

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Figure S1 XPS wide-scan spectra of PtCu NPs.



Figure S2 (a, d, g) TEM images, (b, e, h) HRTEM images and (c, f, i) particle size distribution histograms of PtCu NPs with different Pt/Cu ratio. The insertions in (b, e, h) display the FFT images of PtCu NPs.



Figure S3 SEM images of Pt_1Cu_3/CB catalyst.



Figure S4 Nitrogen ad/desorption isotherms (a) and pore-size distribution curves (b) of carbon black and PtCu/CB samples.

Samples	ICP-OES	XPS
Pt ₂ Cu ₁ /CB	82.9:17.1	62.6:37.4
Pt_1Cu_1/CB	50.4:49.6	83.4:16.6
Pt ₁ Cu ₃ /CB	45.1:54.9	75.5:24.5
Pt ₁ Cu ₅ /CB	44.3:55.7	70.2:29.8

Table S1 Pt/Cu mass ratio determined by ICP-OES and XPS.

Samples	S _{BET}	External surface	BJH average pore size	Total pore volume of pores
	$m^{2} g^{-1}$	² -1 m g	nm	$\operatorname{cm}^{3} \operatorname{g}^{-1}$
Carbon black	815.19	706.81	9.60	0.84
Pt ₂ Cu ₁ /CB	277.15	299.71	13.98	0.37
Pt_1Cu_1/CB	285.81	318.57	11.41	0.37
Pt ₁ Cu ₃ /CB	267.71	309.03	12.34	0.37
Pt ₁ Cu ₅ /CB	231.26	282.02	18.66	0.27

Table S2 Textural properties of as-synthesized PtCu/CB catalysts.

	Pt 4f _{7/2}				Pt 4f _{5/2}			
Sample	Pt ⁰		Pt ^{II}		Pt ⁰		Pt ^{II}	
	B.E., eV	Area, %	B.E., eV	Area, %	B.E., eV	Area, %	B.E., eV	Area, %
Pt ₂ Cu ₁ /CB	71.044	19.5	71.777	9.5	74.542	34.1	75.513	36.9
Pt ₁ Cu ₁ /CB	71.129	34.9	72.009	11.9	74.499	27.7	75.359	25.5
Pt ₁ Cu ₃ /CB	71.198	31.1	72.145	9.7	74.565	27.0	75.495	32.2
Pt ₁ Cu ₅ /CB	71.112	31.2	72.094	12.1	74.502	27.7	75.444	29.0

Table S3 XPS parameters of Pt 4f.

Table S4 XPS parameters of Cu 2p.

	Cu 2p _{3/2}				Cu 2p _{1/2}			
Sample	Cu ⁰		Cu ^{II}		Cu ⁰		Cu ^{II}	
	B.E., eV	Area, %	B.E., eV	Area, %	B.E., eV	Area, %	B.E., eV	Area, %
Pt ₂ Cu ₁ /CB	932.536	35.4	934.030	32.7	951.477	16.0	953.073	15.9
Pt ₁ Cu ₁ /CB	932.216	54.5	933.650	10.9	951.966	27.7	954.069	6.9
Pt ₁ Cu ₃ /CB	932.122	44.4	933.012	20.5	951.872	20.4	953.078	14.7
Pt ₁ Cu ₅ /CB	932.151	45.6	932.969	19.4	951.901	20.4	953.075	14.6

Samples	E _{onset} ,mV	Tafel slope,	$\eta@j=10 \text{ mA} \cdot \text{cm}^{-2},$	η @j=50 mA cm ⁻² ,	
		mV dec ⁻¹	mV	mV	
20% Pt/C	0	21.89	21.2	58.6	
10% Pt/C	-5	36.80	60.1	117.1	
Pt ₂ Cu ₁ /CB	-15	55.01	50.6	111.6	
Pt_1Cu_1/CB	-12	51.56	40.5	104.1	
Pt ₁ Cu ₃ /CB	0	26.28	10.5	29.7	
Pt ₁ Cu ₅ /CB	-10	34.59	20.9	49.5	

Table S5 Comparison of onset potential, overpotential, current densities, and Tafel slope value of various samples in 0.5M H₂SO₄.

Table S6 Comparison of onset potential, overpotential, current densities, and Tafel slope value of various samples in 1.0M KOH.

Samples	E _{onset} ,mV	Tafel slope,	Tafel slope, $\eta@j=10 \text{ mA} \cdot \text{cm}^{-2}$,	
		mV dec ⁻¹	mV	mV
10% Pt/C	-2	45.76	39	130
Pt ₂ Cu ₁ /CB	-16	42.54	16	101
Pt ₁ Cu ₁ /CB	-27	42.73	56	131
Pt ₁ Cu ₃ /CB	-11	76.87	17	115
Pt ₁ Cu ₅ /CB	-10	139.49	30	168



Figure S5 The EIS Nyquist plots of PtCu/CB and 10% Pt/C samples (a) in 0.5M H_2SO_4 at -0.223 V (vs. RHE) and (b) in 1.0M KOH at -0.010 V (vs. RHE).

Electrolyte	Catalyst	Mass activity @-50 mV, A mg _{Pt} ⁻¹	E _{onset} , mV	Tafel slope, mV dec ⁻¹	η@j=10 mA·cm ⁻² , mV	Reference	
	Pt ₁ Cu ₃ /CB	1.1154	0	26.28	10	This work	
	20% Pt/C	0.2067	0	21.89	21.2	This work	
	10% Pt/C	0.0543	-5	36.80	60.1	This work	
	PtCu DNFs	0.2523	0	34	27	[1]	
	PtCu NPHs	0.0332	21	49	66	[1]	
	Pt CNFs	-	-	40	-	[2]	
	PdCo@CN	-	-	31	80	[3]	
			0.01		j= -2.75 mA cm ⁻	F 4 1	
H_2SO_4	Cu-Pt		-0.01		² @ -0.5 V	[4]	
	Cu@Pt/4- Sulfophenyl/GC	28	-	-	-	[5]	
	$Pt_1Cu_{1.03}$ -D	19.34@-0.2V	-	28.5	20	[6]	
	Pd@PtCu/C	-	-	26.2	19	[7]	
	PtNi	-	-380	119	-	[8]	
	Pt ₂₄ Cu ₇₆ NFs	0.74@-0.9V	-	52	18	[9]	
	Pt-WC- 1,10/BMZ	6.659@-0.033V	-	40.1	23	[10]	
	Pt ₁ Cu ₃ /CB	0.2141	0	122.34	17	This work	
КОН	10 % Pt/C	0.1375	-5	87.70	39	This work	
	PtCu NPHs	0.0470	-3	76	61	[1]	
	PtCu DNFs	0.1503	0	55	36	[1]	
	PdCo@CN	-	-	-	250	[3]	
	Cu@Ni	-	-	79	140	[11]	

Table S7 Comparison of HER performance of Pt-based catalysts in acid and alkaline electrolytes.



Figure S6 HER polarization curves of PtCu/CB catalysts before and after 1000 cycles in (a) 0.5M H₂SO₄ and in (c) 1.0M KOH.



Figure S7 SEM images of Pt_1Cu_3/CB catalyst after (a, b) 9 h alkaline HER and (c, d)

24 h acidic HER stability test.



Figure S8 Comparison of Pt 4f XPS spectra between the as-prepared Pt_1Cu_3/CB catalyst and the corresponding one after 9 h alkaline HER stability test.

Reference

- Zhang, X. F., et al., Solvothermal synthesis of monodisperse PtCu dodecahedral nanoframes with enhanced catalytic activity and durability for hydrogen evolution reaction. ACS Applied Energy Materials, 2018. 1(9): p. 5054-5061.
- [2] Ghasemi, S., S.R. Hosseini, and S. Nabipour, Preparation of nanohybrid electrocatalyst based on reduced graphene oxide sheets decorated with Pt nanoparticles for hydrogen evolution reaction. Journal of the Iranian Chemical Society, 2019. 16(1): p. 101-109.
- [3] Chen, J., et al., *Active and durable hydrogen evolution reaction catalyst derived from Pd-doped metal–organic frameworks*. ACS applied materials & interfaces, 2016. **8**(21): p. 13378-13383.
- [4] Mandegarzad, S., et al., *Cu-Pt bimetallic nanoparticles supported metal organic framework-derived nanoporous carbon as a catalyst for hydrogen evolution reaction.* Electrochimica Acta, 2016. **190**: p. 729-736.
- [5] Mirzaei, P., et al., *Bimetallic Cu–Rh Nanoparticles on Diazonium-Modified Carbon Powders for the Electrocatalytic Reduction of Nitrates*. Langmuir, 2019. 35(45): p. 14428-14436.
- [6] Jia, Y., et al., Sandwich-Type Electrochemical Immunosensor Based on Signal Amplification System of Porous PtCu Decorated FCuS Nanospheres for CEA Detection. Journal of The Electrochemical Society, 2019. 166(8): p. B713-B719.
- Bao, M., et al., Surface evolution of PtCu alloy shell over Pd nanocrystals leads to superior hydrogen evolution and oxygen reduction reactions. ACS Energy Letters, 2018. 3(4): p. 940-945.
- [8] Zheng, J., *Seawater splitting for high-efficiency hydrogen evolution by alloyed PtNix electrocatalysts.* Applied Surface Science, 2017. **413**: p. 360-365.
- [9] Huang, X.-Y., et al., L-proline assisted solvothermal preparation of Cu-rich rhombic dodecahedral PtCu nanoframes as advanced electrocatalysts for oxygen reduction and hydrogen evolution reactions. Electrochimica Acta, 2019. 299: p. 89-97.
- [10] Chen, X., et al., Coupling low platinum and tungsten carbide supported on ZIFs-Derived porous carbon for efficient hydrogen evolution. Electrochimica Acta, 2019. 328: p. 135077.
- [11] Li, Z., et al., Mesoporous Hollow Cu–Ni Alloy Nanocage from Core–Shell Cu@ Ni Nanocube for Efficient Hydrogen Evolution Reaction. ACS Catalysis, 2019. 9(6): p. 5084-5095.