

## Electronic Supporting Information for

### **Co-precipitation polymerization of dual functional monomers and polystyrene- co-divinylbenzene for ciprofloxacin imprinted polymer preparation**

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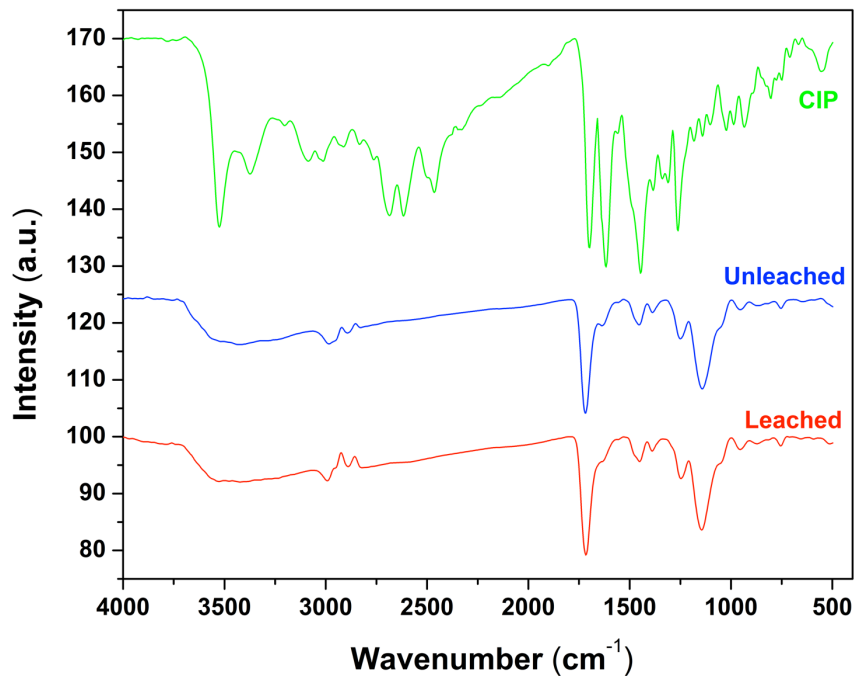


Figure S 1. FT-IR of ciprofloxacin (CIP), unleached and leached imprinted polymer

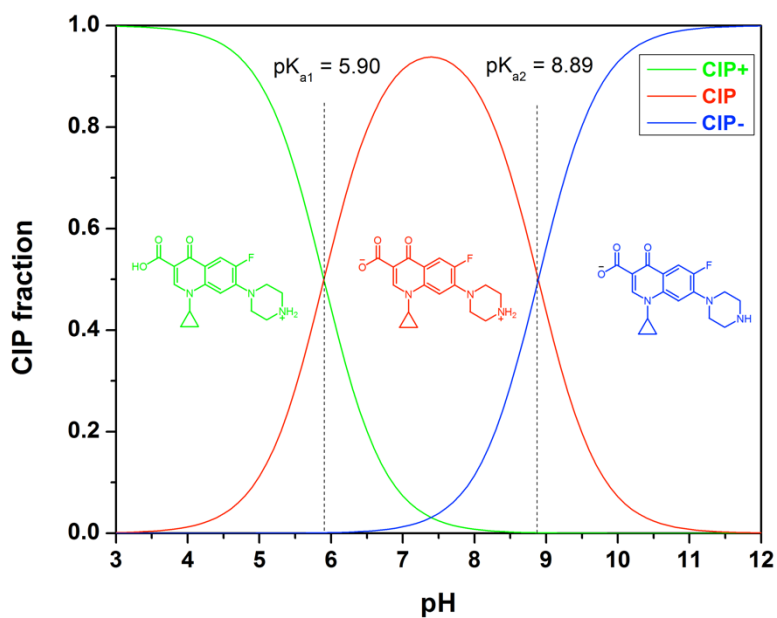


Figure S 2. Molecular structure of CIP and its ionic forms as a function of pH, calculated with  $pK_{a1} = 5.90$  and  $pK_{a2} = 8.89$

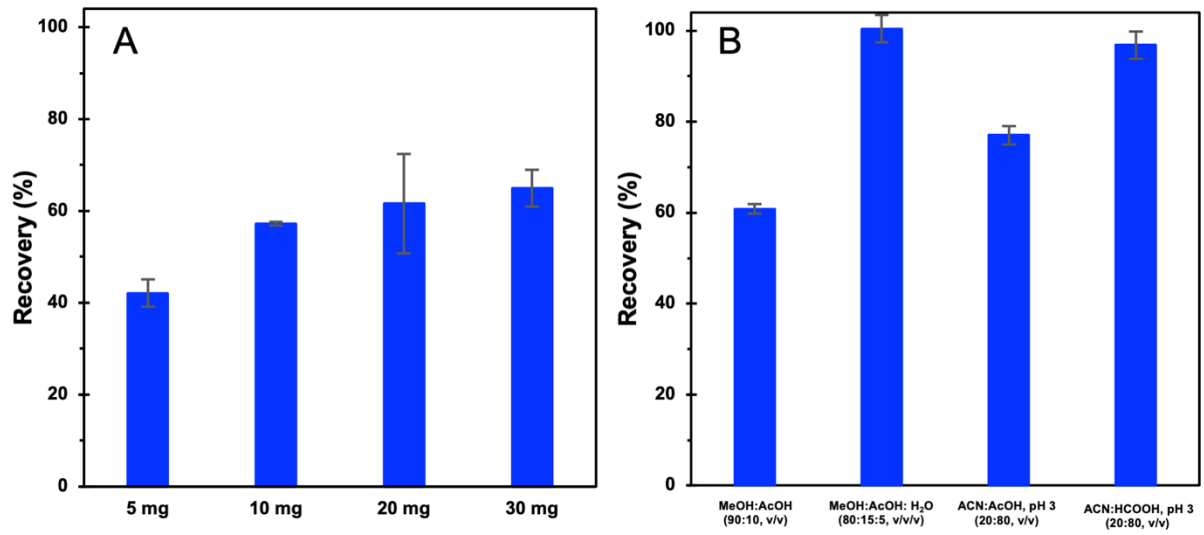


Figure S 3. Influence of mass of polymer (A) and elution solvent (B) on CIP extraction recovery

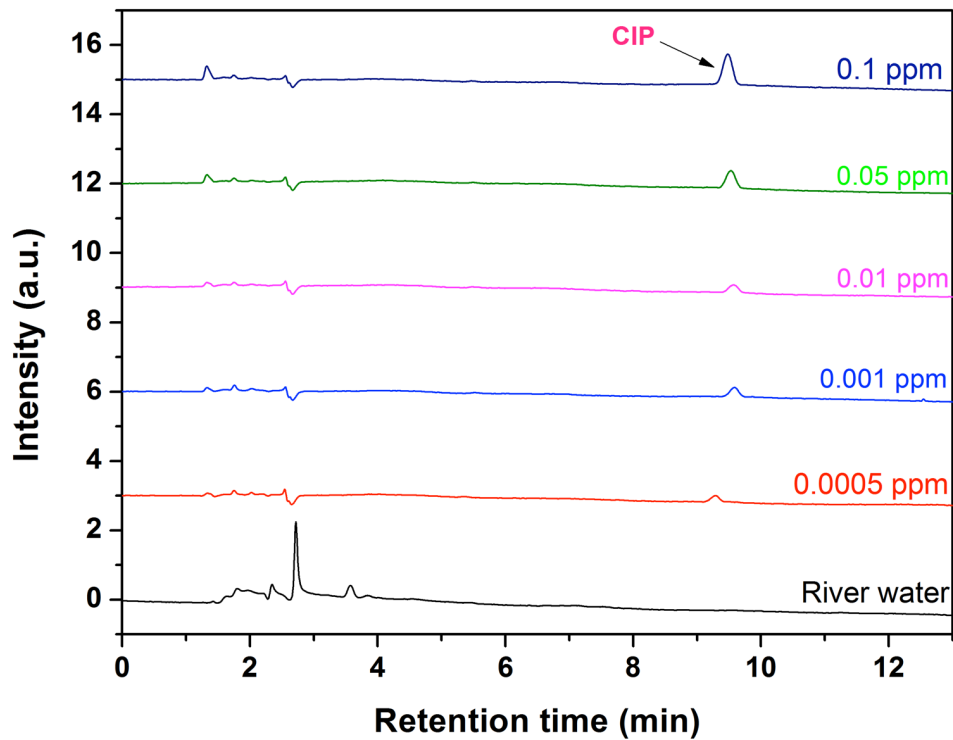
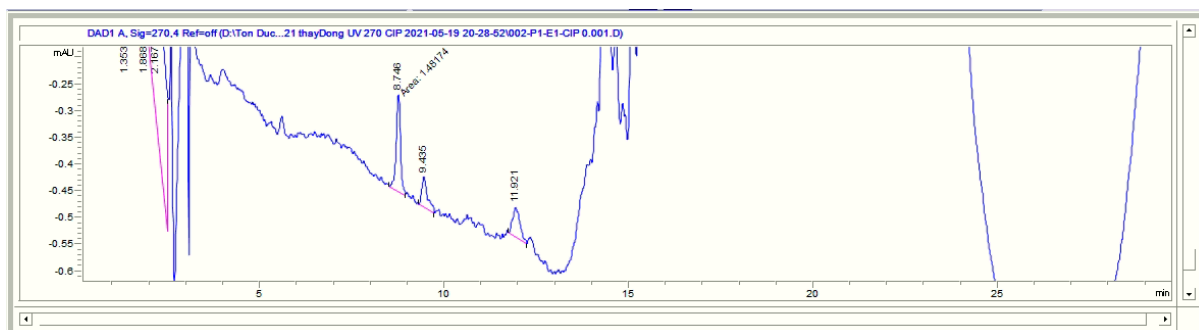


Figure S 4. Chromatograms of real river water samples and extraction of spiked samples at different concentrations

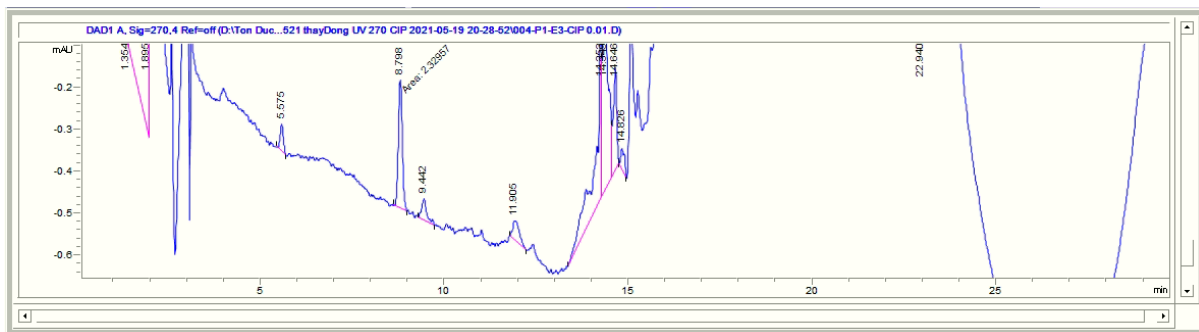


Noise determination:

Time range	Noise	Noise	Noise	Wander	Drift
from [min]   to [min]	(6*SD) [mAU]	(PtoP) [mAU]	(ASTM) [mAU]	[mAU]	[mAU/h]
6.000   7.000	2.395e-2	1.823e-2	-	-	-5.853e-1
10.000   11.000	3.011e-2	2.143e-2	-	-	-8.803e-1

RetTime [min]	k'	Area [mAU*s]	Height [mAU]	Symm.	Width [min]	Plates	Resol	Signal /Noise
1.353	-	5.03349e-1	1.26174e-1	1.06	0.0650	2393	-	5.3
1.868	-	2.66074	3.53756e-1	1.23	0.1316	1120	3.08	14.8
2.167	-	14.13431	7.43272e-1	0.76	0.3396	225	0.74	31.0
<u>8.746</u>	-	1.48174	1.81036e-1	1.07	0.1167	31121	16.95	<u>6.0</u>

Figure S 5. Reported chromatogram of 0.001 ppm CIP standrad solution



Noise determination:

Time range	Noise	Noise	Noise	Wander	Drift
from [min]   to [min]	(6*SD) [mAU]	(PtoP) [mAU]	(ASTM) [mAU]	[mAU]	[mAU/h]
6.000   7.000	2.205e-2	1.554e-2	-	-	-1.121
10.000   11.000	3.208e-2	2.229e-2	-	-	-1.056

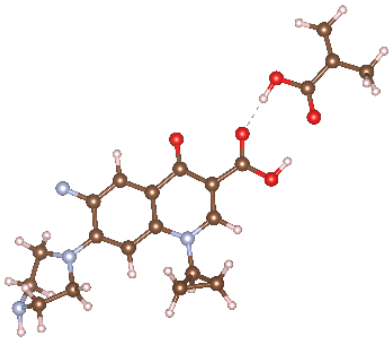
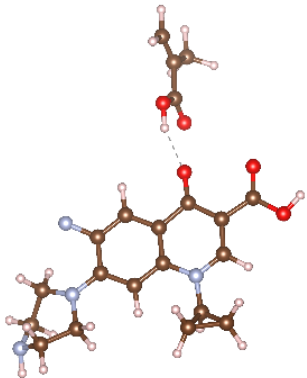
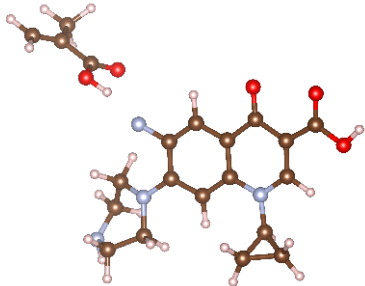
RetTime [min]	k'	Area [mAU*s]	Height [mAU]	Symm.	Width [min]	Plates	Resol	Signal /Noise
1.354	-	2.22928	2.77190e-1	1.41	0.1033	950	-	12.6
1.895	-	10.65650	6.67580e-1	3.41	0.5418	68	0.99	30.3
5.575	-	3.97873e-1	6.62914e-2	0.91	0.0944	19290	6.79	3.0
<u>8.798</u>	-	2.32957	3.07315e-1	0.92	0.1160	31880	18.00	<u>9.6</u>

Figure S 6. Reported chromatogram of 0.01 ppm CIP standrad solution

Table S 1. Pseudo-first-order and pseudo-second-order parameters for CIP adsorption on polymers MIP4 and NIP4

Polymer	Lagergren first-order			Pseudo second-order			Interparticle mass transfer diffusion	
	$q_e$ ( $\text{mg g}^{-1}$ )	$k_1$ ( $\text{min}^{-1}$ )	$R^2$	$q_e$ ( $\text{mg g}^{-1}$ )	$k_2$ ( $\text{mg g}^{-1} \text{min}^{-1}$ )	$R^2$	$k_{id}$ ( $\text{mg g}^{-1} \text{min}^{-1}$ )	$R^2$
MIP4	5.1103	0.0988	0.9740	5.4371	0.0361	0.9926	0.1425	0.9710
NIP4	4.3379	0.0629	0.9312	4.8370	0.0183	0.9391	0.2610	0.9344

Table S 2 : Interaction configurations of methacrylic acid (MAA) with CIP molecule

Interaction energies $E_{int}$ (kJ/mol)	-72.12	-36.24	-19.30
Configurations	<p>I- Double hydrogen bonds between carboxylic groups</p> 	<p>II- Hydrogen bond between H atom of MAA carboxylic group and carbonyl group of CIP</p> 	<p>III- Hydrogen bond between H atom of MAA carboxylic group and fluorine atom of CIP</p> 

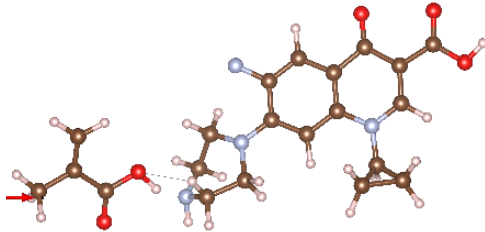
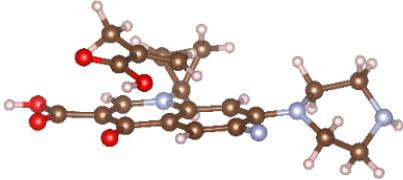
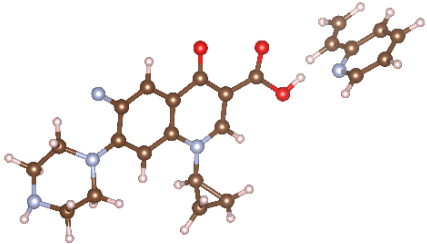
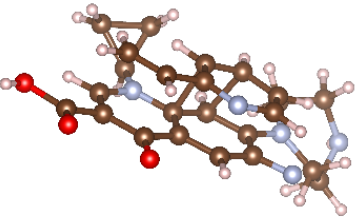
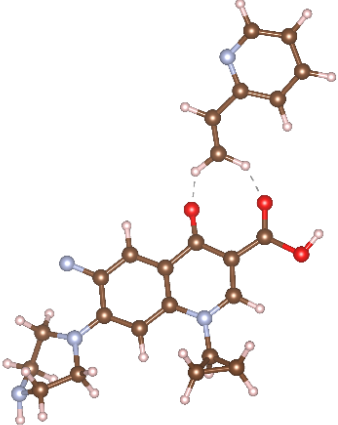
-55.55	-25.98
<p data-bbox="539 451 1012 635">IV- Hydrogen bond between H of MAA carboxylic group and one nitrogen atom of CIP</p>  <p>The image shows a ball-and-stick model of a hydrogen bond. On the left, a carboxylic acid group (MAA) is shown with a red arrow pointing to a hydrogen atom. On the right, a nitrogen atom (CIP) is shown. A dashed line represents the hydrogen bond between the H and the N. The atoms are color-coded: carbon is brown, oxygen is red, and nitrogen is blue.</p>	<p data-bbox="1061 451 1525 560">V- flat interaction between MAA and CIP</p>  <p>The image shows a ball-and-stick model of a flat interaction. A carboxylic acid group (MAA) is shown on the left, and a nitrogen-containing ring (CIP) is shown on the right. The two molecules are oriented such that their planes are parallel to each other, indicating a pi-pi or similar flat interaction. The atoms are color-coded: carbon is brown, oxygen is red, and nitrogen is blue.</p>



Table S 3 : Interaction configurations of 2-vinylpyridine (2-VP) with CIP molecule

Interaction energies $E_{int}$ (kJ/mol)	-48.96	-53.43	-19.58
Configurations	<p>I- Hydrogen bond between H atom of carboxylic group and nitrogen atom of 2-VP</p> 	<p>II- Pi-pi stacking between benzyl groups of CIP and 2-VP</p> 	<p>III- Hydrogen bond between H atom of ethylene group of 2-VP and oxygen atoms of CIP</p> 

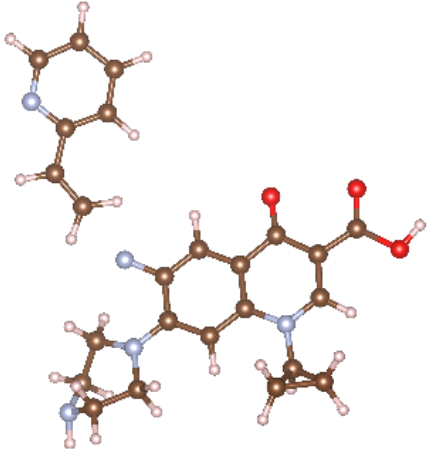
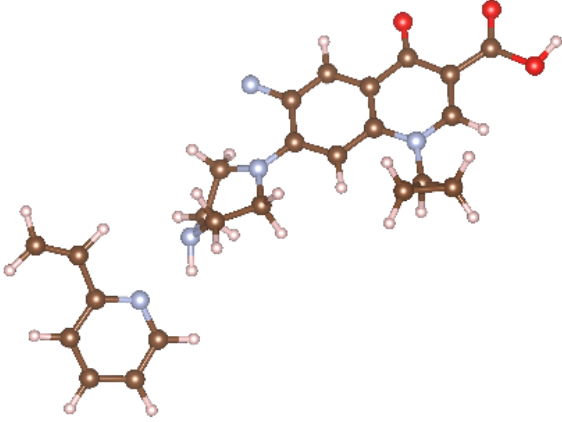
-8.46	-33.30
<p data-bbox="443 416 1048 600">IV- Hydrogen bond between H atom of ethylene group of 2-VP and nitrogen atom of CIP</p> 	<p data-bbox="1137 416 1697 600">V- Hydrogen bond between N atom of 2-VP and h atom bonded to nitrogen atom in CIP</p> 

Table S 4 : Monomer-monomer interactions

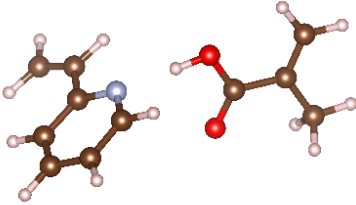
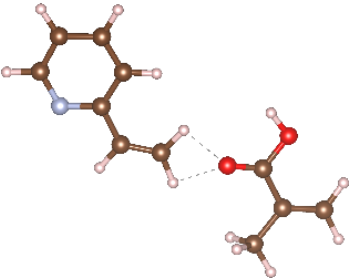
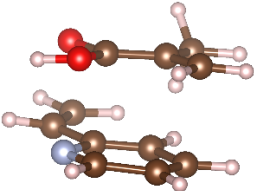
Interaction energies $E_{int}$ (kJ/mol)	<b>-50.88</b>	-6.00	-17.30
Configurations	<p>I- H bond between 2VP nitrogen atom and H of MAA carboxylic group</p> 	<p>II- H bond between H atom of ethylene of 2VP atom and O atom of MAA carboxylic group</p> 	<p>III- Flat interaction between MAA and 2VP</p> 

Table S 5 : Monomers arrangement in MIP and NIP polymers

