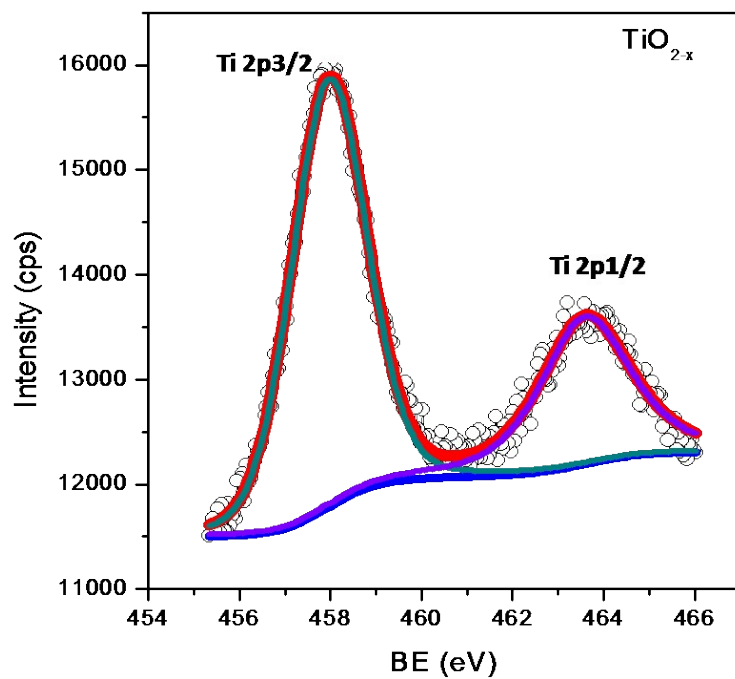
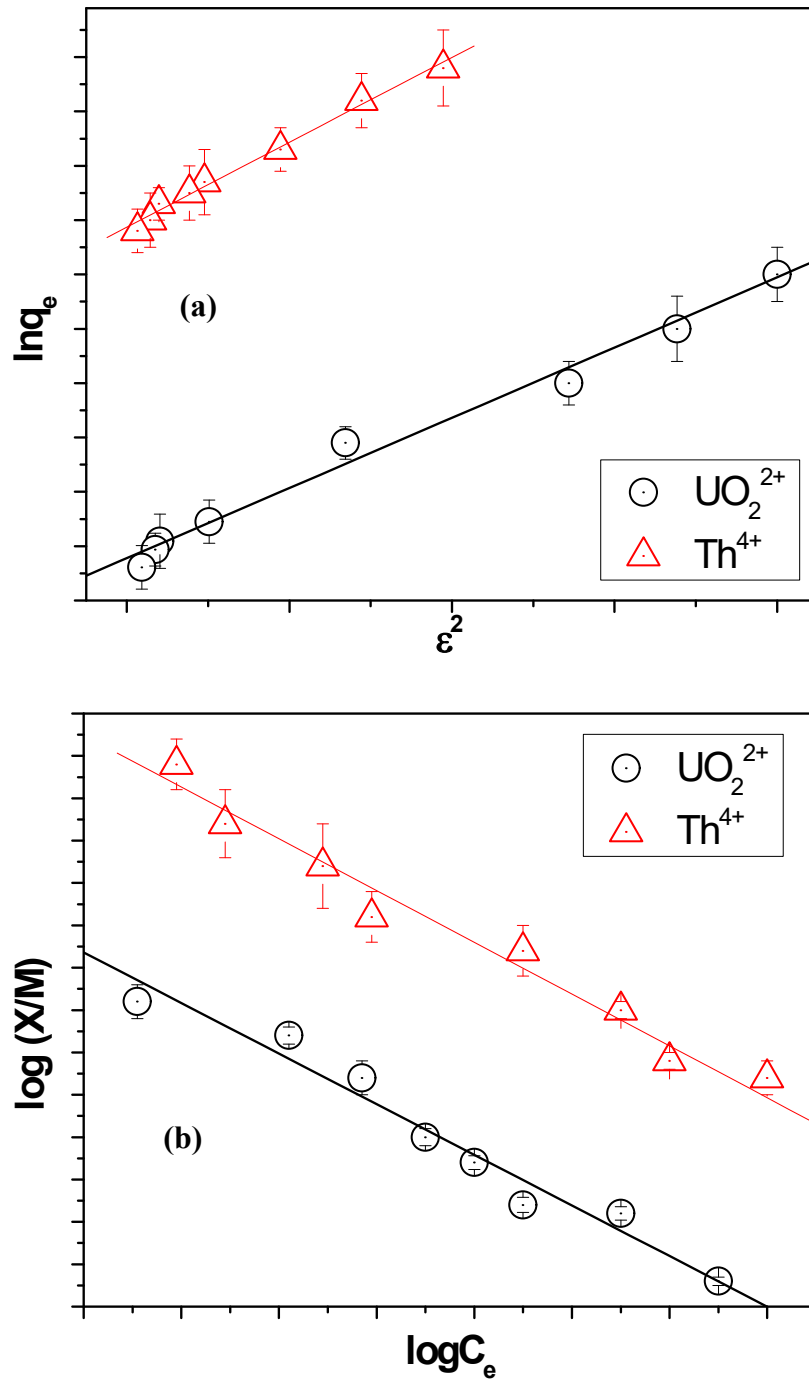


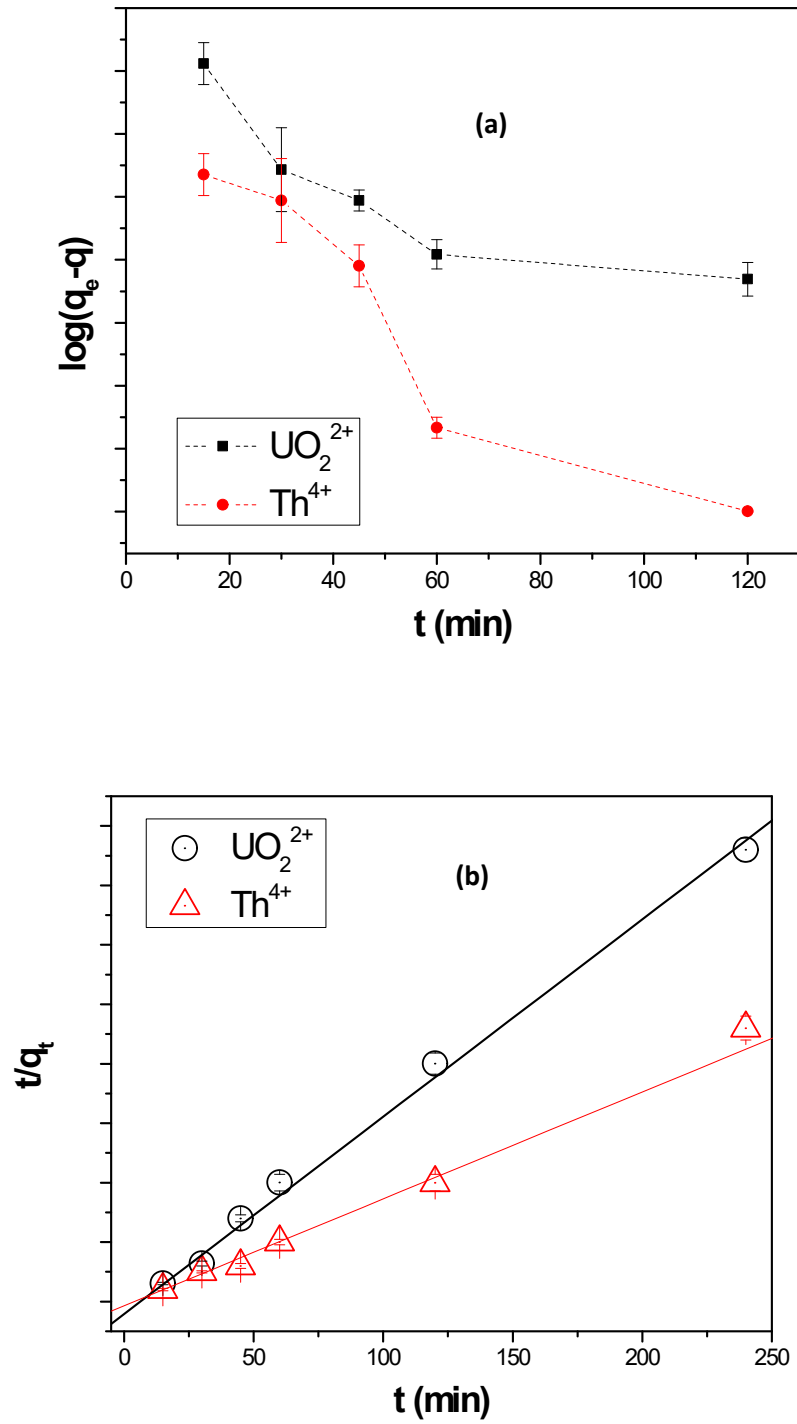
Supporting Information



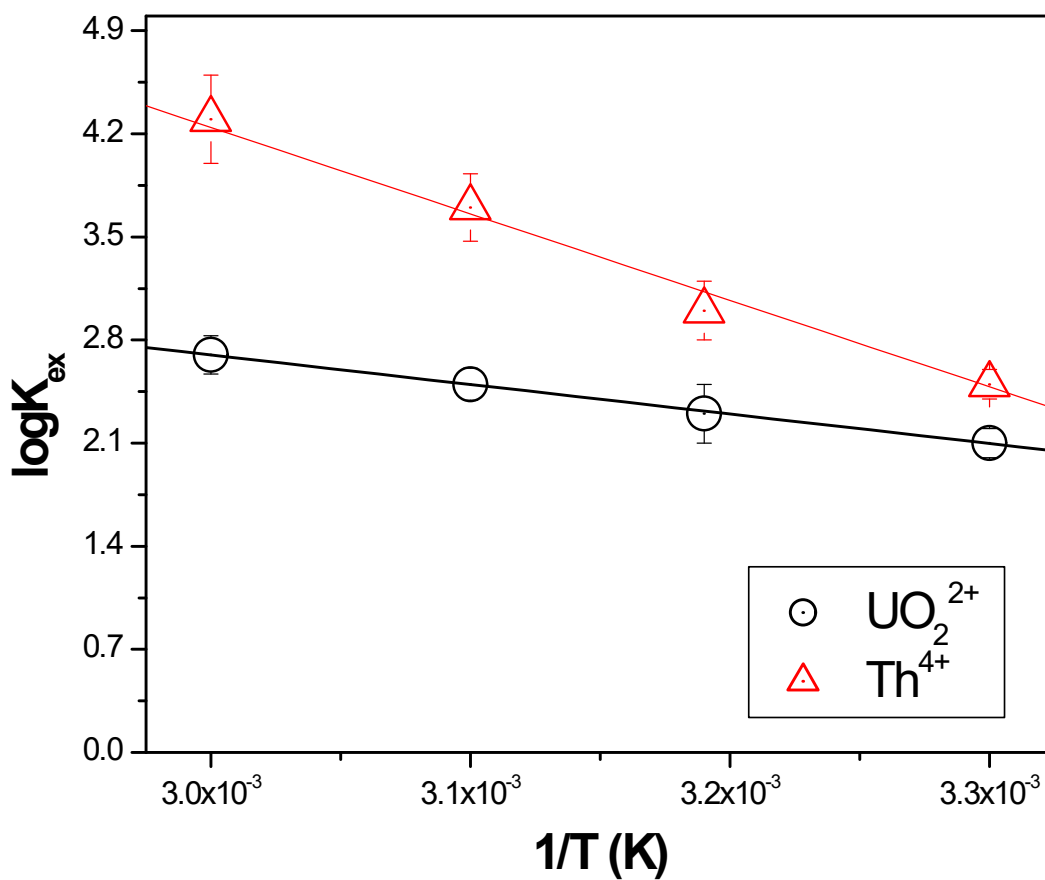
SI Fig. 1: XPS spectra for TiO₂ precursor for Ti peak



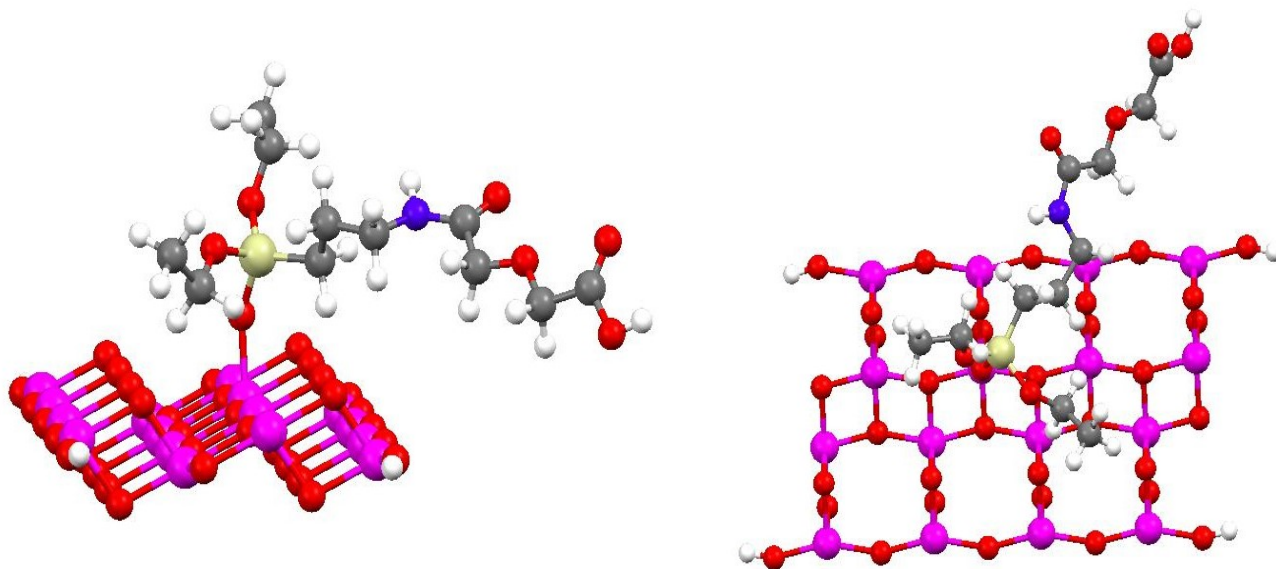
SI Fig. 3: The linear regression analysis for (a) Dubbin-Radushkevich (D-R) isotherm; and (b) Freundlich isotherm



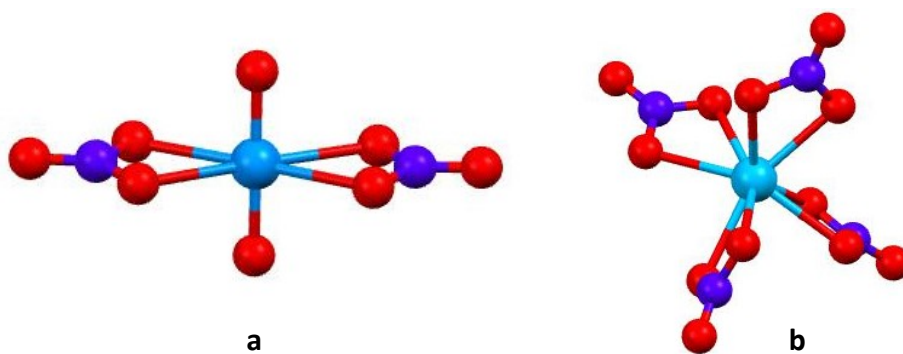
SI Fig. 4: Linear regression analysis for (a) Lagergren pseudo 1st order kinetics;
 (b) Pseudo 2nd order kinetics model



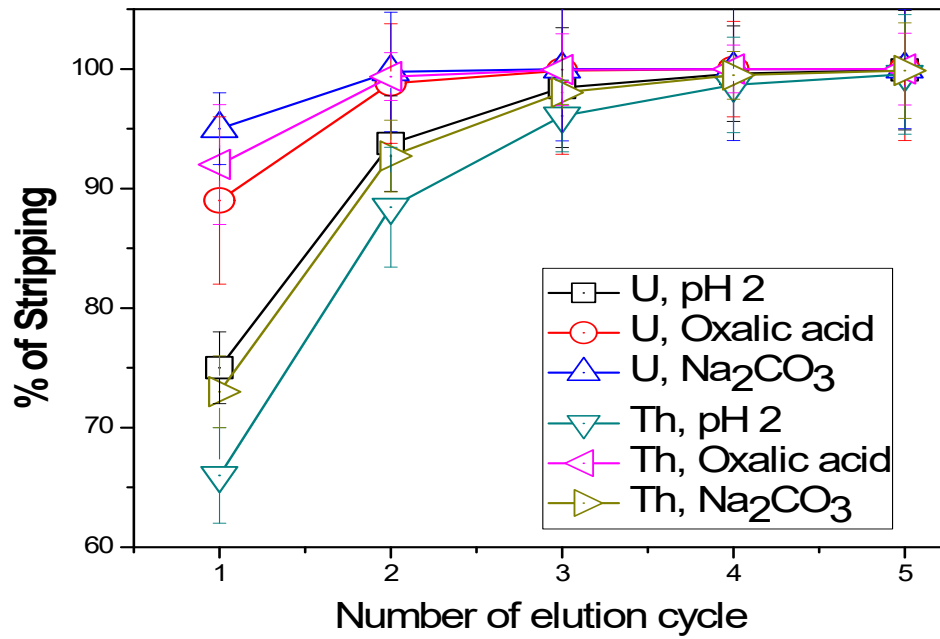
SI Fig. 5: The variation in conditional extraction constants for UO₂²⁺ and Th⁴⁺ on SAPOGA-TiO₂ as a function of reciprocal of absolute temperature



SI Fig. 6: The optimized structure for the sorbent SAPOGA-TiO₂



SI Fig. 7: The optimized structure of (a) UO₂(NO₃)₂ and (b) Th(NO₃)₄



SI Fig. 8: Cumulative % stripping of metal ions using multiple contacts of aqueous phase complexing agents

SI Table.1: Analytical results obtained by linear regression analysis of the experimentally obtained data using Langmuir, D-R, and Freundlich isotherm models

	Langmuir			Dubinin-Radushkevich			Freundlich		
	q_e (mg/g)	b (l/mol)	χ^2	x_m (mg/g)	E (kJ/mol)	χ^2	K_f (mmol/g)	n	χ^2
UO_2^{2+}	231	0.05	0.9989	200	15.7	0.9875	115	9	0.989 0
Th^{4+}	458	0.08	0.9972	381	19.0	0.9842	154	13	0.986 7

SI Table.2: The linear regression analysis of the experimentally obtained data using Lagergren 1st order, Intra particle diffusion and Pseudo 2nd order kinetics models

	Lagergren first order kinetics model			Intra particle diffusion model			Pseudo 2 nd order model		
	q_e	k_{ads}	χ^2	k_p (mg g ⁻¹ min ⁻¹)	C	χ^2	q_e (mg g ⁻¹)	k_2 (mg g ⁻¹ min ⁻¹)	χ^2
UO_2^{2+}	66	0.08	0.7817	51	23	0.9997 0.	88	5.8E-04	0.9973
Th^{4+}	89	0.04	0.6866	48	29	9982	99	3.6E-04	0.9956

SI Table.3: Calculated structural parameters (bond length in Å) of complexes of UO_2^{2+}/Th^{4+} ion with SAPOGA-TiO₂ in presence of nitrate ion.

System	M-O _{C=O} (Å)	M-O _{ethereal} (Å)	M-O _{NO3} (Å)
$UO_2(NO_3)_2-L$	2.423, 2.593	2.835	2.442, 2.427, 2.504, 2.517
$Th(NO_3)_4-L$	2.441, 2.600	2.783	2.515, 2.519, 2.453, 2.480, 2.597, 2.578, 2.518

SI Table.4: Calculated value of Gibbs free energy (kcal/mol) in the gas and aqueous phase

Equilibrium reaction	Free energy of adsorption (kcal/mol)	
	Gas phase	Aqueous phase
$[\text{UO}_2(\text{H}_2\text{O})_5]^{2+}_{(\text{aq})} + 2\text{NO}_3^-_{(\text{aq})} + \text{L}_{(\text{aq})} \longrightarrow \text{UO}_2(\text{NO}_3)_2\text{-L}_{(\text{aq})} + 5\text{H}_2\text{O}$	-312.5	-30.2
$[\text{Th}(\text{H}_2\text{O})_{10}]^{4+}_{(\text{aq})} + 4\text{NO}_3^-_{(\text{aq})} + \text{L}_{(\text{aq})} \longrightarrow \text{Th}(\text{NO}_3)_4\text{-L}_{(\text{aq})} + 10\text{H}_2\text{O}$	-915.1	-83.7
$\text{UO}_2(\text{NO}_3)_2_{(\text{aq})} + \text{L}_{(\text{aq})} \longrightarrow \text{UO}_2(\text{NO}_3)_2\text{-L}_{(\text{aq})}$	-24.9	-6.6
$\text{Th}(\text{NO}_3)_4_{(\text{aq})} + \text{L}_{(\text{aq})} \longrightarrow \text{Th}(\text{NO}_3)_4\text{-L}_{(\text{aq})}$	-87.8	-36.8

SI Table.5: Calculated charge and orbital population using NBO analysis in aqueous phase at the B3LYP/TZVP level of theory.

System	charge	s	p	d	f
$\text{UO}_2(\text{NO}_3)_2\text{-L}$	1.803	4.17	11.77	11.46	2.78
$\text{Th}(\text{NO}_3)_4\text{-L}$	1.694	4.20	11.99	11.12	0.981
$[\text{UO}_2(\text{H}_2\text{O})_5]^{2+}$	2.088	4.14	11.75	11.38	2.61
$[\text{Th}(\text{H}_2\text{O})_{10}]^{4+}$	2.175	4.18	11.99	10.91	0.73
$\text{UO}_2(\text{NO}_3)_2$	2.195	4.11	11.74	11.30	2.62
$\text{Th}(\text{NO}_3)_4$	2.382	4.15	11.99	10.86	0.59

SI Table 6: Average value of Wiberg bond Indices for metal-ligand and metal nitrate bonding.

System	M-O(-C=O _{DGA})	M-O(-ether O _{DGA})	M-O (NO ₃)
$\text{UO}_2(\text{NO}_3)_2\text{-L}$	0.334	0	0.461
$\text{Th}(\text{NO}_3)_4\text{-L}$	0.369	0.073	0.431