

Electronic Supplementary Information

Evaluating the efficacy of nanosized CuZnAl and CuZnZr mixed oxides for electrocatalytic CO₂ reduction

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Table

S1:

Catalyst	Theoretical (mol%)				Actual (mol%)			
	Cu	Zn	Al	Zr	Cu	Zn	Al	Zr
CuZnAl	0.675	0.275	0.05	--	0.674	0.274	0.06	--
CuZnZr	0.675	0.275	--	0.05	0.673	0.273	--	0.07

Theoretical and actual composition of the synthesized catalysts.

Table S2: The size of different crystallites in CuZnAl and CuZnZr samples calcined at different temperature.

Phase	*Crystallite size (nm)					
	CuZnAl-500	CuZnAl-600	CuZnAl-700	CuZnZr-500	CuZnZr-600	CuZnZr-700
CuO	111.8	44.7	40.6	28.6	47.3	55.3
ZnO	112.7	71.4	46.2	51.7	57.2	57.4
Al ₂ O ₃	-	-	-	-	-	-
CuAl ₂ O ₄	-	-	34.5	-	-	-
ZrO ₂	-	-	-	35.4	42.8	50.6

*Calculated using Scherrer equation and XRD data

Table S3: Textural properties of catalysts from N₂-physisorption and chemical composition of catalysts

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Catalyst	Calcination temperature (°C)	S_{BET} (m ² /g)	V_t (cc/g)	Av. pore diameter (nm)
CuZnAl	500	17	0.130	18.4
	600	12	0.002	3.0
	700	3	0.005	2.8
CuZnZr	500	26	0.075	7.8
	600	8	0.003	5.7
	700	2	0.001	3.5

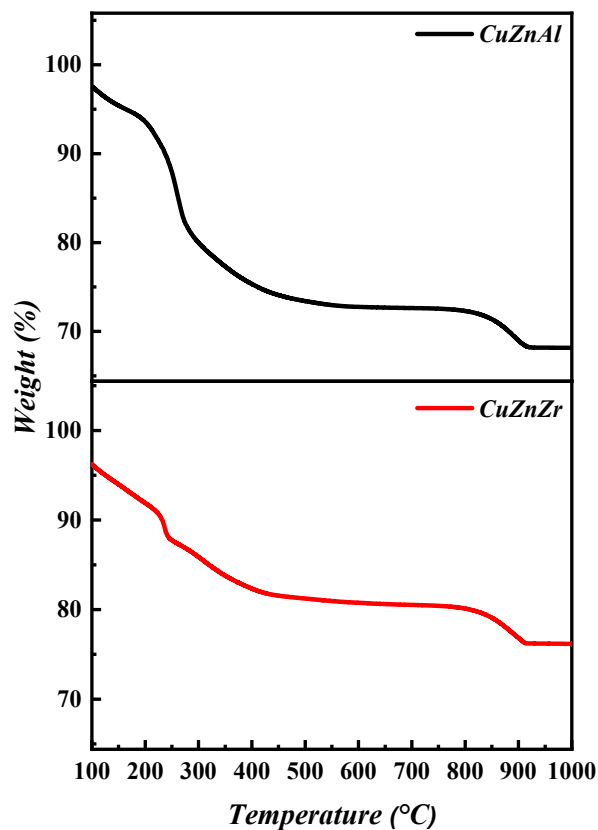


Fig S1: Thermograms of as-synthesized CuZnAl, and CuZnZr catalysts

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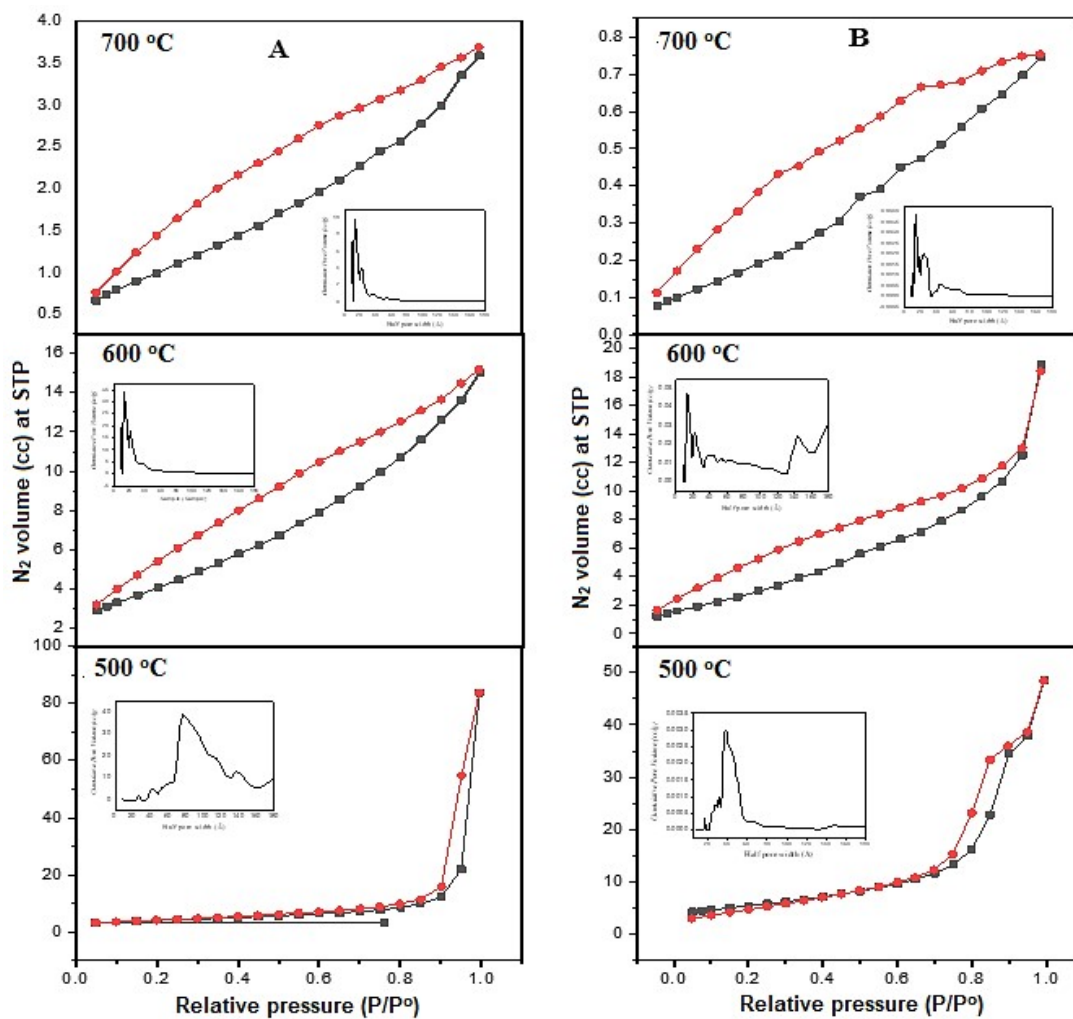


Fig. S2: N_2 adsorption-desorption isotherms for (A) CuZnAl and (B) CuZnZr catalysts calcined in air at 500, 600, and 700 °C.

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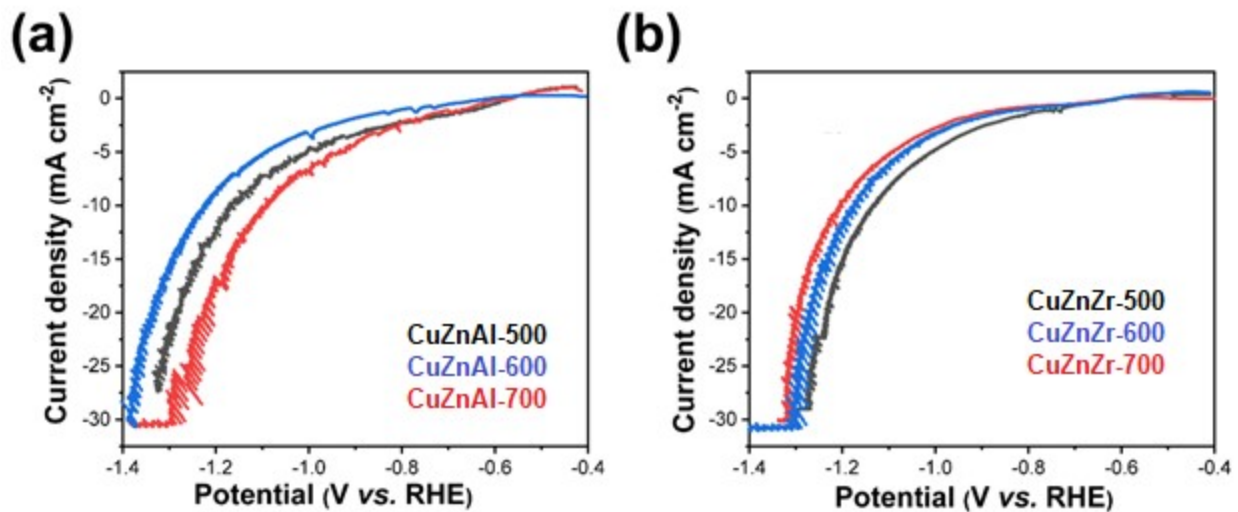


Fig. S3: (a) LSV curves of CuZnAl-x (temperature x=500/600/700 °C) in CO₂-saturated 0.1 KHCO₃ solution, (b) LSV curves of CuZnZr-x (temperature x=500/600/700 °C) in CO₂-saturated 0.1 KHCO₃ solution.

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Table S4: Comparison between the performance of nanosized and non-nanosized CuZn-based catalysts

Catalyst	Nanosized copper catalysts						Catalyst	Non-nanosized copper catalysts					
	CO ₂ R products							CO ₂ R products					
	CO, FE(%)	CH ₃ OH, FE(%)	CH ₄ , FE(%)	C ₂ H ₅ OH, FE(%)	C ₂ H ₄ , FE(%)	Ref.		CO, FE(%)	CH ₃ OH, FE(%)	CH ₄ , FE(%)	C ₂ H ₅ OH, FE(%)	C ₂ H ₄ , FE(%)	Ref.
7CuO/3ZnO/C	50	-	3	-	47	[1]	PcCu-O ₈ -Zn/CNT	89	-	-	-	-	[5]
8CuO/2ZnO/C	38	at - 0.46 V (vs RHE)	5	at - 0.76 V (vs RHE)	50								
9CuO/1ZnO/C	35	-	29	46 at - 0.9 V (vs RHE)									
Cu/ZnO x	-	-	36% at 1.1 V (vsRHE)	-	-	[2]	Cu/ZnO	5.4	2.8	1	10.2	10	[6]
CuZn	38	at - 1.46 V (vs RHE)	-	-	-	[3]	Cu/ZnO	5	2	-	12	6	[7]
CuZn_240	36		-	-	-								
CuZn_300	37		-	-	-								
CuO-ZnO ₁₀	-	15 at - 0.6 V (vs RHE)	-	22 at -0.8 V (vs RHE)	-	[4]							

[1] Z. Li *et al.*, “CuO/ZnO/C electrocatalysts for CO₂-to-C₂₊ products conversion with high yield: On the effect of geometric structure and composition,” *Appl. Catal. A: Gen.*, 606 (2020) 117829, doi: 10.1016/j.apcata.2020.117829.

[2] R. Wang *et al.*, “Engineering a Cu/ZnO_x Interface for High Methane Selectivity in CO₂ Electrochemical Reduction,” *Ind. Eng. Chem. Res.*, 60 (2021) 273–280, doi: 10.1021/acs.iecr.0c04718.

[3] C. Azenha *et al.*, “Tuning the selectivity of Cu₂O/ZnO catalyst for CO₂ electrochemical reduction,” *J. CO₂ Utilization* 68 (2023) doi: 10.1016/j.jcou.2022.102368.

[4] S. Dongare, *et al.*, “Electrochemical reduction of CO₂ using oxide based Cu and Zn bimetallic catalyst,” *Electrochim. Acta*, 392 (2021) doi: 10.1016/j.electacta.2021.138988.

[5] H. Zhong *et al.*, “Synergistic electroreduction of carbon dioxide to carbon monoxide on bimetallic layered conjugated metal-organic frameworks,” *Nat. Commun.*, 11 (2020) 1409, doi: 10.1038/s41467-020-15141-y.

[6] E. Andrews *et al.*, “Electrochemical Reduction of CO₂ at Cu Nanocluster/(10 $\bar{1}$ 0) ZnO Electrodes,” *J Electrochem Soc.*, 160 (2013) H841–H846, doi: 10.1149/2.105311jes.

[7] T. Liang *et al.*, “Molecular dynamics simulations of CO₂ reduction on Cu(111) and Cu/ZnO(10 $\bar{1}$ 0) using charge optimized many body potentials,” *Catal. Commun.*, 52 (2014) 84–87, doi: 10.1016/j.catcom.2013.11.033.