1	New Journal of Chemistry			
2	Electronic Supplementary Information			
3	A stable and easily regenerable solid amine adsorbent derived from			
4	polyethylenimine-impregnated dialdehyde-cellulose/graphene-oxide			
5	composite			
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Titration method to quantify the aldehyde content

The aldehyde content on the surface of DAC was estimated by the titration method as reported in the literature.¹ After freeze drying, 0.1 g of pieced DAC aerogel was weighed and placed into 0.25 M Hydroxylamine hydrochloride (NH₂OH·HCl) solution (25 mL), which was adjusted to pH 4.5 with HCl. All the mixture was placed in dark, shaken in the thermostatic water bath at ambient temperature for 48 h. The conversion of aldehyde to oxime was determined by the consumption of 0.1 M NaOH solution, each sample was quantified twice.

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47 of GO (0.01 wt%, 0.03 wt%, 0.05 wt%, 0.07 wt%, 0.1 wt%) before and after

48 periodate oxidation.



51 **Fig. S2** Photo images of the CNF aerogel and DAC aerogel when hydrated again after

52 freeze-drying: (a) hydrated samples at the beginning; (b) hydrated samples after shaking.



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- **Fig. S3** Photo images of water solution with different concentration of GO before (a)
- and after (b) periodate oxidation: A: 0.05 wt% GO; B: 0.6 wt% GO.
- 57
- 58 Mechanical property of DAC-0.05GO gel in humid conditions



60 Fig. S4 Stress-strain curves during 100 cycles of compression-releasing at a constant

strain of 30%, showing the stress–strain curves of selected cycles (0th, 30th, 60th,

62 100th).

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66 Fig. S5 XPS spectra of DAC-0.05GO after amination: (a) N1s spectrum of DAC-

67 0.05GO-cPEI; (b) N1s spectrum of DAC-0.05GO@PEI.

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69 Structural characterization

70 Nitrogen adsorption–desorption isotherms, characterized at 77 K in the range of

relative pressure from 10⁻⁶ to 1 by using adsorption instrument (JW-BK200C, JEGB 71 SCI&TECH., China), were used to determine the surface area and pore size distribution 72 of the samples. V _{Total} was calculated based on the nitrogen amount adsorbed at $P/P_0 =$ 73 0.95. Before measuring, non-aminated aerogels were degassed under a vacuum at 120 °C 74 for 6 h, aminated aerogels were degassed under a vacuum at 100 °C for 6 h. Specific 75 surface area was calculated by using the Brunauer-Emmett-Teller (BET) method from 76 the isotherm in the relative pressure range between 0.05 and 0.15. Pore parameters were 77 calculated by using density functional theory (DFT) method. 78



Fig. S6. The N₂ adsorption–desorption isotherms (a) and pore size distributions (b) of

81 DAC, DAC-0.05GO, DAC-0.05GO-cPEI and DAC-0.05GO@PEI.

Samples	S _{BET} (m^2/g)	V _{Total} (cm ³ /g)	D Pore (nm)
DAC	11.487	0.039	10.625
DAC-0.05GO	10.271	0.039	13.894
DAC-0.05GO-cPEI	11.291	0.045	15.641
DAC-0.05GO@PEI	10.783	0.032	10.687

82 **Table S1** Summary of pore parameters of the gels.

As it was shown in Fig. S6, all N₂ adsorption-desorption isotherms belonged to 83 Type IV with hysteresis loops (Fig. S6a) and all the samples had similar pore size 84 distributions (Fig. S6b), which showed that these materials possessed mesoporous 85 structure. As shown in Table S1, the specific surface area and total pore volume of all 86 87 the gels were close, indicating that composite or amination showed little influence on pore characteristic. Therefore, the slight-increased CO₂ adsorption amount on 88 composite gel might be attributed to slight-increased space inside the support, while the 89 increased CO₂ adsorption amount on DAC-0.05GO-cPEI might be attributed to 90 increased effective amino sites. 91





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Fig. S7 CO₂ amount adsorbed on DAC-0.05GO-cPEI at 25 °C before conducting
desorption test at 70 °C (a), 80 °C (b) and 90 °C (c).



- **Fig. S8** Photo images of DAC-cPEI: (a) fresh sample; (b) regenerated sample after
- 99 CO₂ desorption at 70 $^{\circ}$ C.

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Fig. S9 SEM images of DAC-0.05GO-cPEI: (a) fresh sample; (b) regenerated sample.

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104 **References:**

105 1 U. Kim, M. Wada and S. Kuga, 2004, **56**, 7–10.