

Electronic Supplementary Information (ESI)

Covalent Triazine Frameworks for the Dynamic Adsorption/Separation of Benzene/Cyclohexane Mixtures

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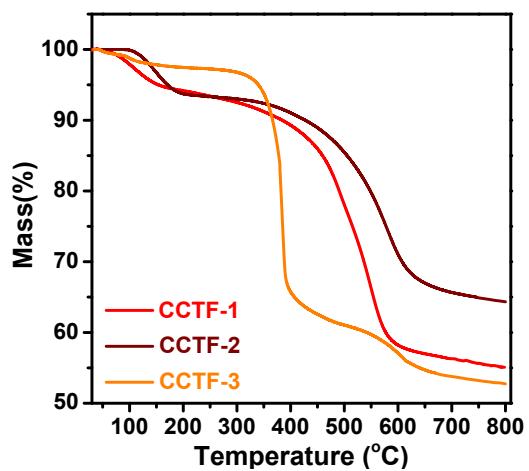


Fig. S1 The TGA curves of CCTF-1, CCTF-2, and CCTF-3.

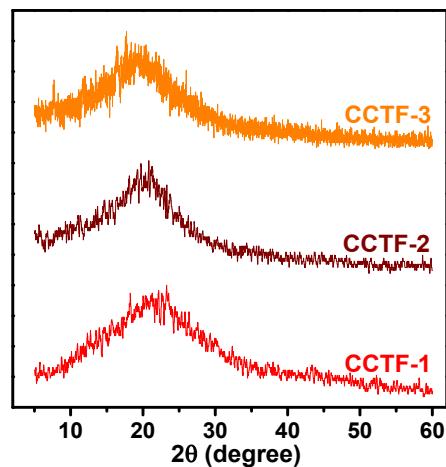


Fig. S2 X-ray diffractions of CCTFs.

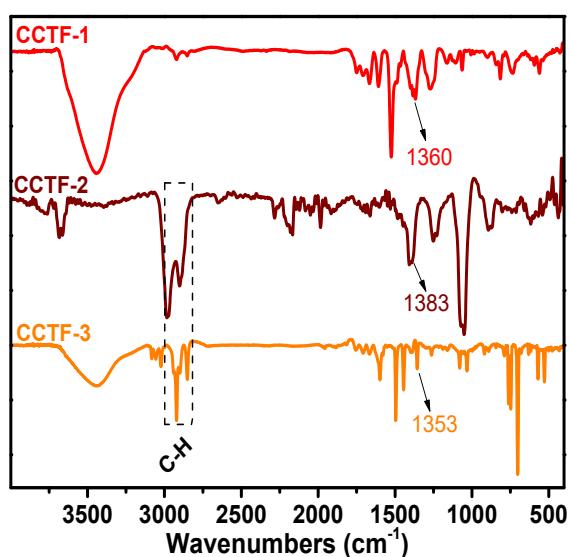


Fig. S3 The FT-IR spectrum of CCTF-1, CCTF-2, and CCTF-3.

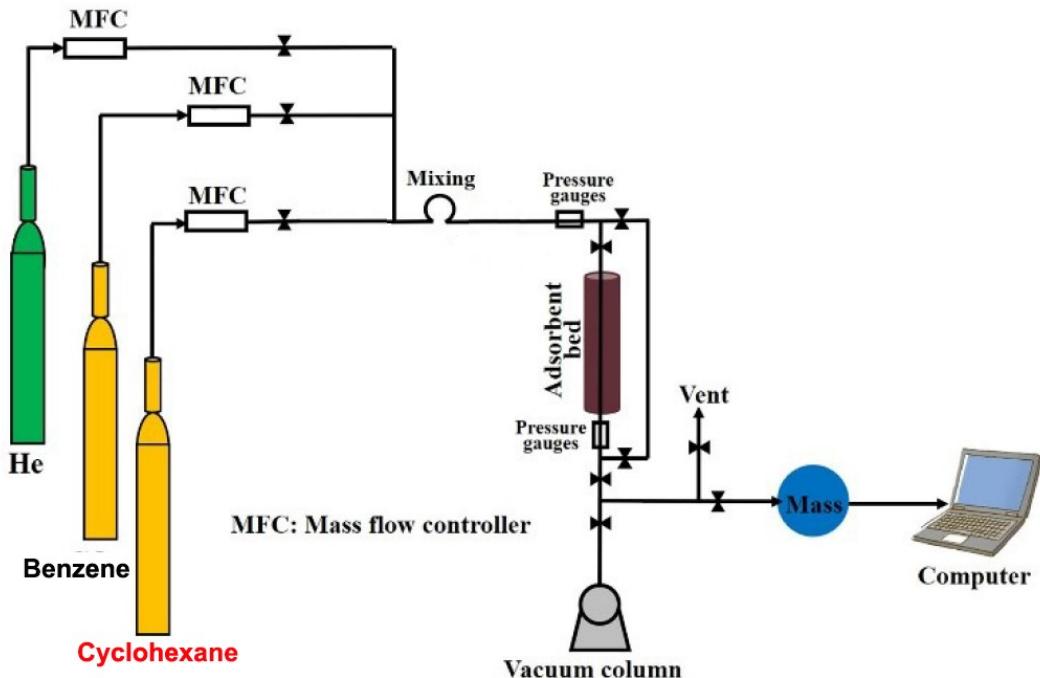


Fig S4. Scheme of the dynamic breakthrough apparatus.

The adsorption amount of adsorbent n is calculated according to the equation S1 as:

$$Q_{nad} = Q_{nin} - Q_{nout} = q_{in} \times C_{n0} \times \Delta T - \int_0^t [(q / (1 - \sum_1^N C_{nt}))] * C_{nt} dt \quad S1$$

Q_{nad} is the adsorption amount of adsorbent n; Q_{nin} is total flow rate of adsorbent n flowing in of the penetrating column at ΔT ; Q_{nout} is total flow rate of adsorbent n flowing out of the penetrating column at ΔT ; q_{in} is total flow rate of gas through the column inlet; C_{n0} is the total flow rate of gas through the outlet of the column; ΔT is the total adsorption time (min); q is the flow rate of carrier gas; C_{nt} is the percentage concentration of adsorbent N at the entrance of the penetrating column;

The selectivity of gas (1) over gas (2) is calculated according to the equation S2 as:

$$S = \left(\frac{X_1}{Y_1} \right) \div \left(\frac{X_2}{Y_2} \right) \quad S2$$

X_1/Y_1 : Molar fraction of component 1 in the adsorbed phase/molar fraction of component 1 in the gas phase

X_2/Y_2 : Molar fraction of component 2 in the adsorbed phase/molar fraction of component 2 in the gas phase

The diffusion coefficient of vapor

$$D = \frac{L^2}{6t_{eq}} \quad S3$$

D is the diffusion coefficient of vapor; L is the loading length; t_{eq} is the delay time.

Table S1. Comparison of uptakes of benzene and cyclohexane vapors, Bz/Cy selectivity between CCTFs, and other porous materials reported in the literature

Samples	S_{BET} (m ² /g)	Bz (mg/g)	Cy (mg/g)	Ideal Bz/Cy Selectivity	Ref
CCTF-1	512	878	441	1.99	This work
CCTF-2	538	650	311	2.09	This work
CCTF-3	207	136	61	2.23	This work
PAF-2	891	138	7	19.71	1
Hybrid[3]arene	0.9	80.4	7.5	10.72	2
MALP-1	1179	585	492	1.19	3
MALP-2	1126	545	472	1.15	3
MALP-3	1141	571	488	1.17	3
MALP-4	1093	558	465	1.20	3
CE-P1	630	478	-	-	4
CE-P2	195	151	-	-	4
PAN-1	925	726	527	1.38	5
PAN-2	1242	692	383	1.81	5
PAN-F	702	544	433	1.26	6
PAN-T	795	570	518	1.10	6
CMP-S-1	873	647.4	378	1.71	7
MP1	1020	765	440	1.74	8
PCN-AD	843	980	574	1.71	9
PBI-Ad-1	1023	980	536	1.83	10
PBI-Ad-2	926	765	463	1.65	10
PCN-TPC	686	176	74	2.38	11
PCN-TPPC	662	778	107	7.27	11
UMC-600	1980	663	527	1.26	12
UMC-700	2212	735	605	1.21	12
UMC-800	2406	868	734	1.18	12
FJU-P6	1066	630.4	313.6	2.01	13
FJU-P7	1425	768.6	640.4	1.20	13
PCN-TA	721	1009	206	4.90	14
PCN-TC	674	867	244	3.55	14
PCN-DC	393	635	129	4.92	14
POP-1	486	1020	658	1.55	15
MPI-1	1454	1198	501	2.39	16
MPI-2	814	766	448	1.71	16
MPI-3	586	549	415	1.32	16
sPI-1	1108	1597	851	1.88	17
sPI-2	900	1760	778	2.26	17

Reference

1. H. Ren, T. Ben, E. Wang, X. Jing, M. Xue, B. Liu, Y. Cui, S. Qiu and G. Zhu, *Chem. Commun.*, 2010, **46**, 291-293.
2. J. Zhou, G. Yu, Q. Li, M. Wang and F. Huang, *J Am Chem Soc*, 2020, **142**, 2228-2232.
3. M. Rong, L. Yang, L. Wang, H. Xing, J. Yu, H. Qu and H. Liu, *Ind. Eng. Chem. Res.*, 2019, **58**, 17369-17379.
4. H. Yu, C. Shen and Z. Wang, *ChemPlusChem*, 2013, **78**, 498-505.
5. G. Li, B. Zhang, J. Yan and Z. Wang, *Macromolecules*, 2014, **47**, 6664-6670.
6. G. Li, B. Zhang, J. Yan and Z. Wang, *Chem. Commun.*, 2016, **52**, 1143-1146.
7. T. Chen, W. Zhang, B. Li, W. Huang, C. Lin, Y. Wu, S. Chen and H. Ma, *ACS Appl. Mater. Interfaces*, 2020, **12**, 56385-56392.
8. X. Ma, Y. Wang, K. Yao, Z. Ali, Y. Han and I. Pinnau, *ACS Omega*, 2018, **3**, 15966-15974.
9. C. Shen, H. Yu and Z. Wang, *Chem. Commun.*, 2014, **50**, 11238-11241.
10. B. Zhang, G. Li, J. Yan and Z. Wang, *J. Phys. Chem. C*, 2015, **119**, 13080-13087.
11. G. Deng and Z. Wang, *ACS Appl. Mater. Interfaces*, 2017, **9**, 41618-41627.
12. J. Yan, B. Zhang and Z. Wang, *J. Phys. Chem. C*, 2017, **121**, 22753-22761.
13. L. Chen, H. Zhang, Y. Ye, Z. Yuan, J. Wang, Y. Yang, S. Lin, F. Xiang, S. Xiang and Z. Zhang, *New J. Chem.*, 2021, **45**, 22437-22443.
14. C. Shen, J. Yan, G. Deng, B. Zhang and Z. Wang, *Polym. Chem.*, 2017, **8**, 1074-1083.
15. H. Tan, Q. Chen, T. Chen and H. Liu, *ACS Appl. Mater. Interfaces*, 2018, **10**, 32717-32725.
16. G. Li and Z. Wang, *Macromolecules*, 2013, **46**, 3058-3066.
17. J. Yan, B. Zhang and Z. Wang, *Polym. Chem.*, 2016, **7**, 7295-7303.