

Support Materials

Sulfur Vacancy on MoS₂ Enhanced the Activation of Peroxymonosulfate through the Co-existence of Radical and Non-radical Pathways to Degrade Organic Pollutants in Wastewater

Cai-Wu Luo ^{a, b, c, *}, Lei Cai ^c, Chao Xie ^c, Gang Li ^a, Tian-Jiao Jiang ^c

^a State Key Laboratory of Safety and Health for Metal Mines, Sinosteel Maanshan General Institute of Mining Research Co., Ltd, 243000, China

^b Fujian Provincial Key Lab of Coastal Basin Environment, Fujian Polytechnic Normal University, 350300, China

^c School of Resource Environmental and Safety Engineering, University of South China, 421000, China

* Corresponding author: Cai-Wu Luo;

Tel: +86-734-8282345;

E-mail addresses: luocaiwu00@126.com.

Figure S1

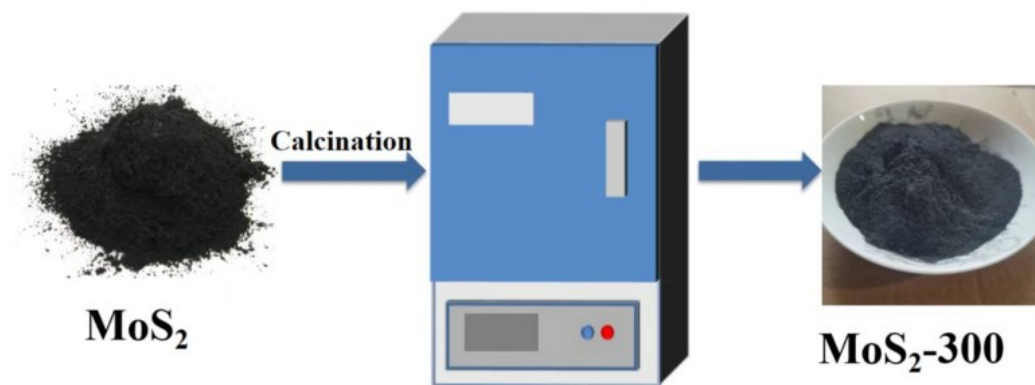


Figure S1 Preparation step of MoS₂-300 catalyst.

Figure S2

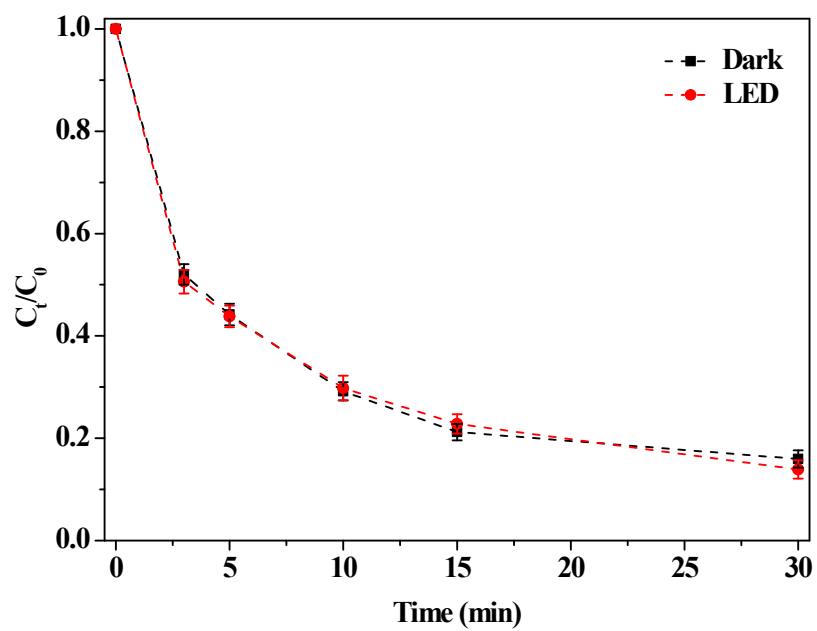


Figure S2 Comparison of darkness and LED illumination on degradation of RhB.

Reaction conditions: [RhB] = 10 mg/L, [PMS] = 1.0 mM, [MoS₂-300] = 1.0 g/L,
initial pH = 3.0.

Figure S3

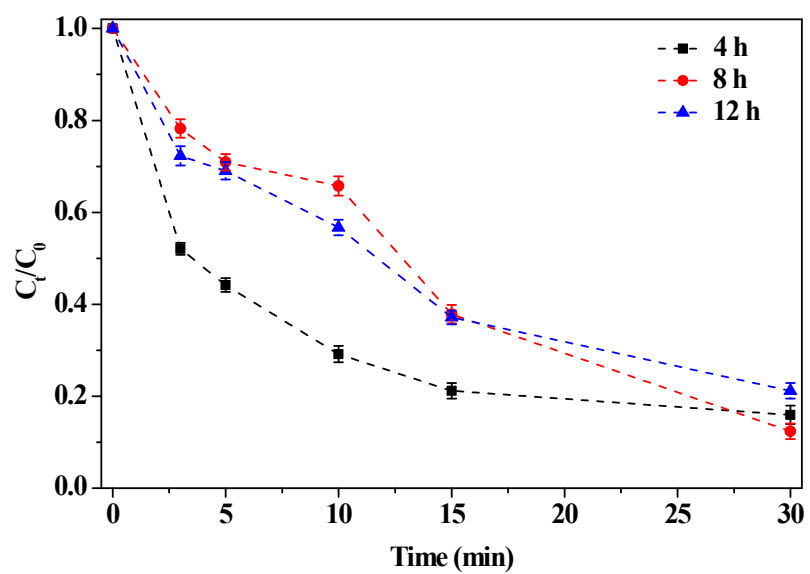


Figure S3 Effect of the treated time on the degradation of RhB. Reaction conditions: [RhB] = 10 mg/L, [PMS] = 1.0 mM, [MoS₂-300] = 1.0 g/L, initial pH = 3.0, in the darkness.

Figure S4

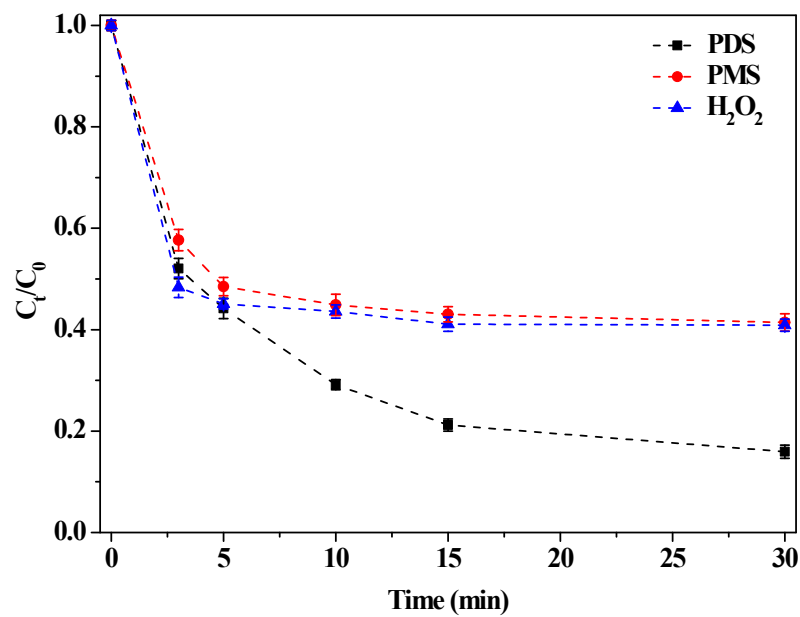


Figure S4 Removal of RhB with different oxidation system. Reaction conditions: [RhB] = 10 mg/L, [Oxidants] = 1.0 mM, [MoS₂-300] = 1.0 g/L, initial pH = 3.0, in the darkness.

Figure S5

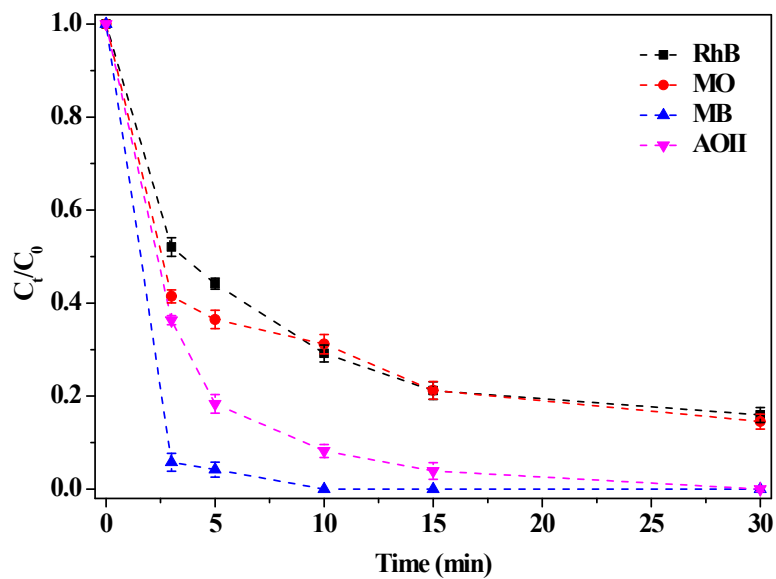


Figure S5 Removal of organic pollutants. Reaction conditions: [Pollutants] = 10 mg/L, [MoS₂-300] = 1.0 g/L, [PMS] = 1.0 mM, initial pH = 3.0, in the darkness.

Figure S6

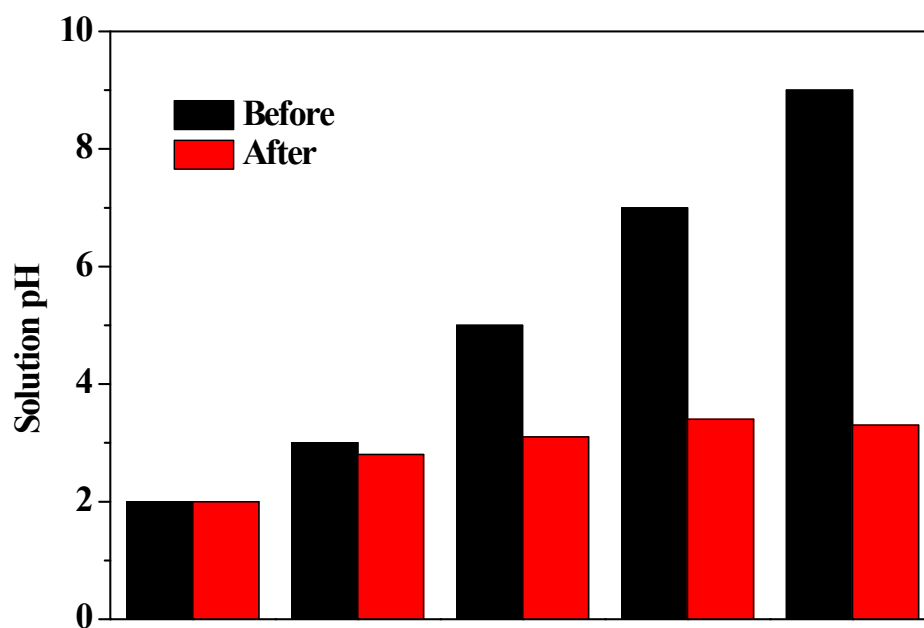


Figure S6 Changes of solution pH before and after the reaction.

Figure S7

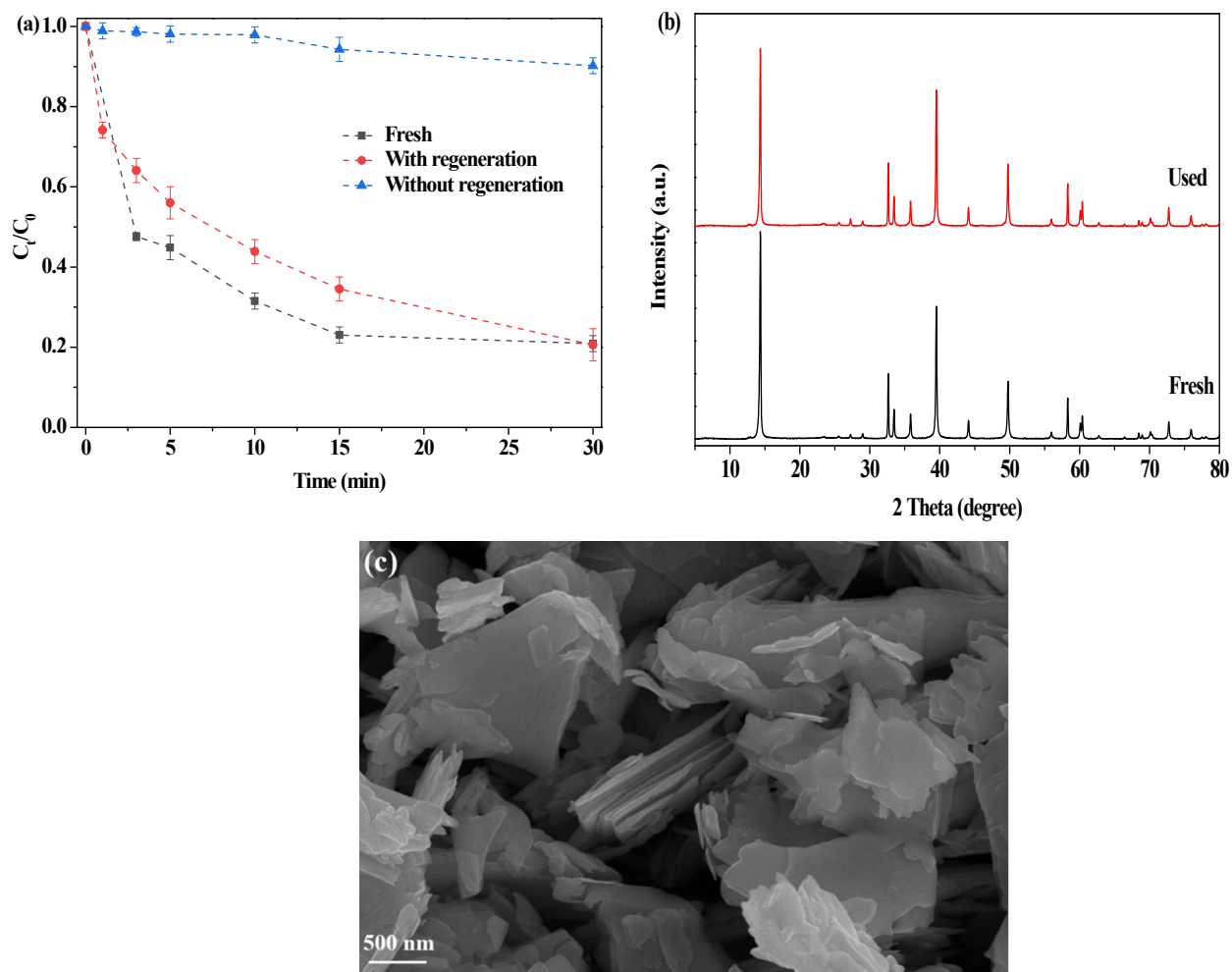


Figure S7 (a) Removal of RhB without or with regeneration; (b) XRD patterns of fresh and used MoS₂ catalysts; (c) SEM image of used MoS₂-300 catalyst. Reaction conditions: [RhB] = 10 mg/L, [PMS] = 0.5 mM, [MoS₂-300] = 1.0 g/L, initial pH = 3.0, in the darkness.

Figure S8

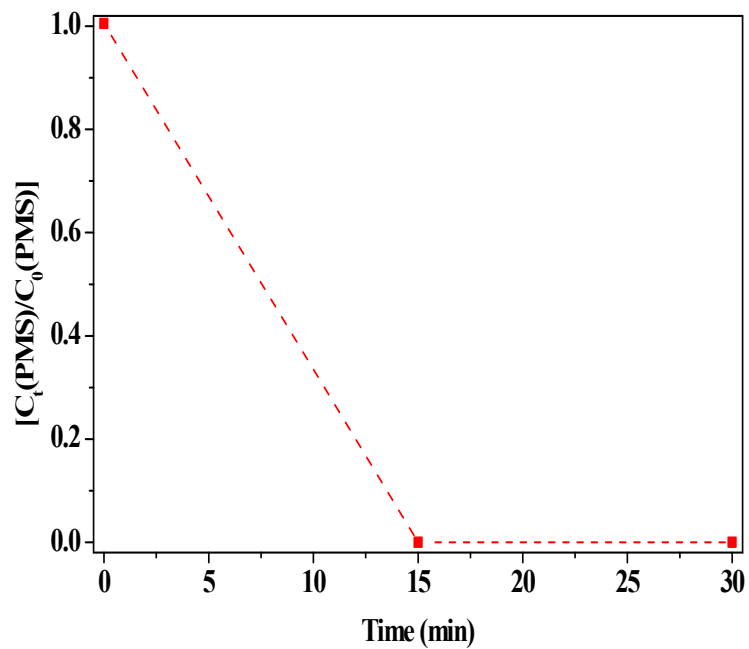


Figure S8 Decomposition of PMS. Reaction conditions: $[\text{RhB}] = 10 \text{ mg/L}$, $[\text{MoS}_2\text{-}300] = 1.0 \text{ g/L}$, $[\text{PMS}] = 1.0 \text{ mM}$, initial $\text{pH} = 3.0$, in the darkness.

Table S1 Changes of surface Mo(IV) concentration, ratios of Mo(IV) and Mo(VI), and 2H-MoS₂ and 1T-MoS₂ phases in catalysts with different temperatures by XPS characterization ^a.

Catalysts	Mo(IV)	Mo(IV)/Mo(VI)	2H-MoS ₂ /1T-MoS ₂ ^b	2H-MoS ₂ /1T-MoS ₂ ^c
MoS ₂	98.3%	57.8	0.48	0.50
MoS ₂ -300	85.1%	5.7	0.62	0.66
MoS ₂ -500	58.8%	1.4	0.28	no

^a: Based on the peak area of each component; ^b: Mo 3d; ^c: S 2p.

Table S2 The concentrations of detected elements in reaction solution after the fifth reaction by ICP-MS.

Catalysts	Mo (mg/L)	S (mg/L)
MoS ₂ -300	36.2	91.6