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Supplementary Information

for

Gravimetric and electrochemical statistical optimizations for improving copper corrosion resistance in hydrochloric acid using thiosemicarbazone-linked 3acetylpyridine

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Sample	Chemical species	Photoemission angle (°)	Assignment	BE (eV)	Relative intensities
					(%)
Copper (Blank)	Cl 2p	0	H-Cl	198.91	77.05
			CuCl ₂	200.56	22.95
	Cu 2p	0	Cu ⁺ /CuCl	932.20	83.49
			Cu^+/Cu_2O	932.53	16.51
Copper with T3AP	C 1s	0	C=C	284.74	63.51
			C=S	287.89	17.18
			C-C	282.48	11.77
			C-N/C=N	286.97	7.54
	N 1s	0	N-H	399.17	61.41
			$N-N/C=N^+$	399.11	20.13
			N-Cu	398.55	12.10
			N-O	397.22	6.35
	S 2p	0	Cu-S	162.40	71.40
			C-S	163.25	15.06
			S-O	167.52	13.76
	Cu 2p	0	Cu-S	932.02	39.00
			Cu-N	933.23	21.40
			Cu ⁺ /CuCl	932.26	15.62

Table S1. Binding energies (eV), relative intensity, and their assignment for Cl 2p, C1s, N 1s, S 2p, and Cu 2p peak observed for copper strip in HCl at an optimum concentrationof T3AP



Thiosemicarbazide

3-acetylpyridine

3-acetylpyridine thiosemicarbazone

Figure S1. The general reaction synthesis of T3AP







Figure S4. The graph of predicted vs actual data for both responses (a) WL IE, (b) EIS IE



Figure S5. 3D graph of EIS inhibition efficiency response of T3AP



Figure S6. The XPS spectrum for Cl 2p of untreated samples



Figure S7. The XPS spectrum for Cu 2p of untreated samples



Figure S8. The XPS spectrum for Cu 2p of samples with inhibitor T3AP



Figure S9. The XPS spectrum for C 1s of samples with inhibitor T3AP



Figure S10. The XPS spectrum for N 1s of samples with inhibitor T3AP



Figure S11. The XPS spectrum for S 2p of samples with inhibitor T3AP



Figure S12. The experimental vs theoretical bond length of T3AP compound