

Supplementary Information for

**Revealing the crystal facet effect on N<sub>2</sub>O formation during the NH<sub>3</sub>-SCR over  $\alpha$ -MnO<sub>2</sub> catalysts**

*Jundong Guo,<sup>a</sup> Fengli Gan,<sup>a</sup> Yifan Zhao,<sup>a</sup> Jinglin He,<sup>a</sup> Tao Gao,<sup>d</sup>, Xia Jiang,<sup>a, b, c</sup> Shenggui Ma,<sup>a, b, c\*</sup>*

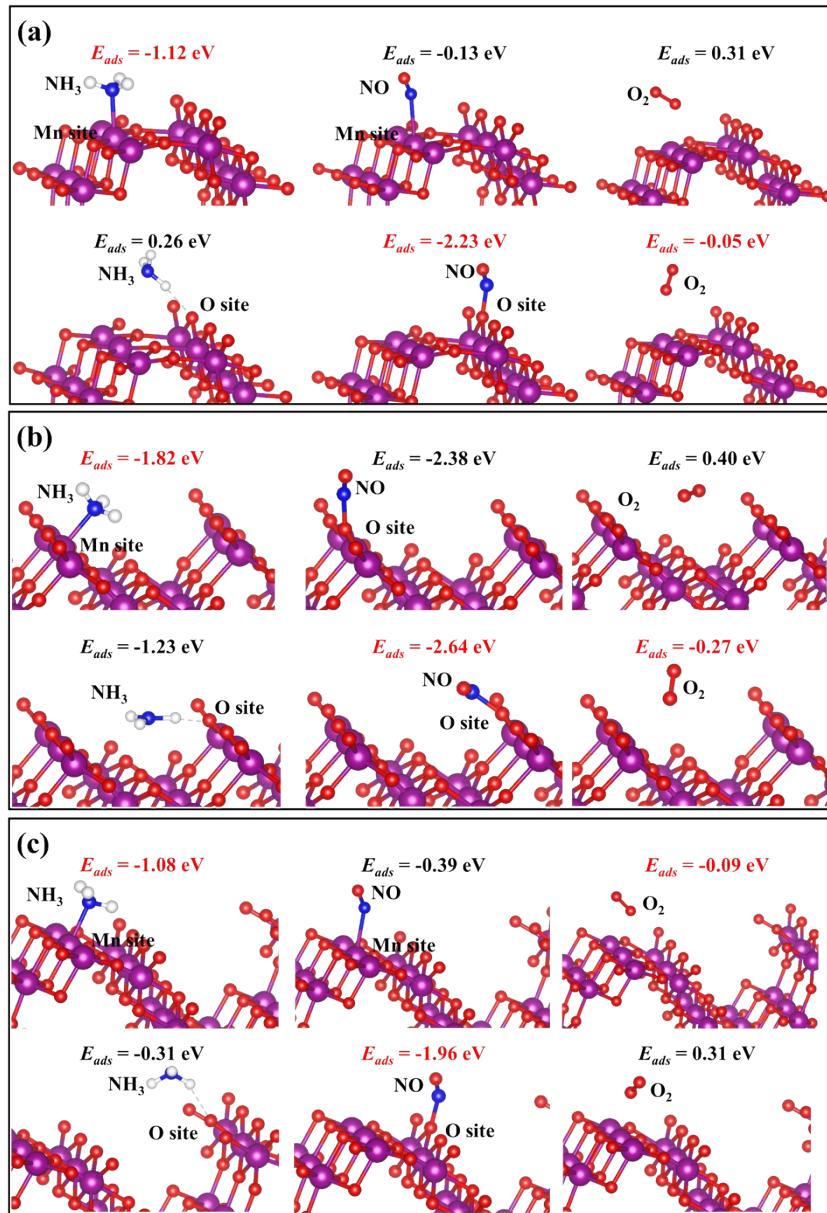
<sup>a</sup> College of Architecture and Environment, Sichuan University, Chengdu 610065, China.

<sup>b</sup> College of Carbon Neutrality Future Technology, Sichuan University, Chengdu 610065, China.

<sup>c</sup> National Engineering Research Center for Flue Gas Desulfurization, Sichuan University, Chengdu 610065, China.

<sup>d</sup> Institute of Atomic and Molecular Physics, Sichuan University, Chengdu 610065, China.

\*Corresponding author: Shenggui Ma. E-mail: masgui@scu.edu.cn.



**Fig. S1** Adsorption configurations of  $\text{NH}_3$ ,  $\text{NO}$  and  $\text{O}_2$  molecules on (a)  $\alpha\text{-MnO}_2$  (100); (b)  $\alpha\text{-MnO}_2$  (110); (c)  $\alpha\text{-MnO}_2$  (310). Purple, red, dark blue and white circles denote Mn, O, N and H atoms respectively.

**Table S1** Reaction rate constant ( $k$ ) of each elementary reaction on  $\alpha\text{-MnO}_2$  under different temperatures.

Facets	Reaction rate constants ( $k$ , $\text{s}^{-1}$ )	Temperature (K)				
		350	375	400	425	450
(100)	$k$ (R1)	$2.79 \times 10^{-7}$	$7.37 \times 10^{-6}$	$1.33 \times 10^{-4}$	$1.75 \times 10^{-3}$	$1.76 \times 10^{-2}$
	$k$ (R2)	$4.56 \times 10^4$	$1.73 \times 10^5$	$5.59 \times 10^5$	$1.57 \times 10^6$	$3.95 \times 10^6$
	$k$ (R3)	$2.09 \times 10^{11}$	$2.47 \times 10^{11}$	$2.87 \times 10^{11}$	$3.28 \times 10^{11}$	$3.69 \times 10^{11}$
(110)	$k$ (R1)	$3.59 \times 10^1$	$1.75 \times 10^2$	$7.05 \times 10^2$	$2.41 \times 10^3$	$7.19 \times 10^3$
	$k$ (R2)	$1.18 \times 10^{-2}$	$9.09 \times 10^{-2}$	$5.36 \times 10^{-1}$	2.54	10.03
	$k$ (R3)	$3.59 \times 10^1$	$1.82 \times 10^2$	$7.57 \times 10^2$	$2.66 \times 10^3$	$8.13 \times 10^3$
(310)	$k$ (R1)	$2.47 \times 10^9$	$3.72 \times 10^9$	$5.33 \times 10^9$	$7.32 \times 10^9$	$9.71 \times 10^9$
	$k$ (R2)	$2.15 \times 10^{-1}$	1.65	9.89	48.05	$1.96 \times 10^2$
	$k$ (R3)	$1.55 \times 10^{-4}$	$2.09 \times 10^{-3}$	$2.06 \times 10^{-2}$	$2.16 \times 10^{-1}$	$9.41 \times 10^{-1}$