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Supporting Information

Content:

1. Experimental Section

2. NMR and IR Characterization

Figure S1-S4. ¹H, ¹³C, ³¹P NMR and IR spectra of 2a. Figure S5-S8. ¹H, ¹³C, ³¹P NMR and IR spectra of 2b. Figure S9-S12. ¹H, ¹³C, ³¹P NMR and IR spectra of 3a. Figure S13-S16. ¹H, ¹³C, ³¹P NMR and IR spectra of 3b. Figure S17-S18. IR spectra of 4a, and 4b. Figure S19-S23. ¹H, ¹³C, ³¹P NMR and IR spectra of 5a. Figure S24-S28. ¹H, ¹³C, ³¹P NMR and IR spectra of 5b.

3. Cyclic Voltammetry

Figure S29-S32. CV spectra of 2a, 2b, 3a, and 3b.

4. X-Ray Crystallographic Analysis

Table S1–S3. Crystal data and structure refinement parameters of 2–5.Figure S33-S40. Molecular structures of 2–5.

5. Computational Details

Table S4–S6. Selected calculated bond lengths, bond orders, and NPA charges of 2–4.

Table S7–S8. Calculated g-factors, coupling constants and spin densities of 2–4.

Figure S41-S46. Selected molecular orbitals of 2–4.

6. EPR Spectroscopy

Figure S47. Stacked X-, Q, and W-band EPR spectra of 4a.

7. References

1. Experimental Section

General procedure. All experiments and manipulations were carried out under a dry argon atmosphere using either standard Schlenk techniques or an MBraun LABmaster Pro glovebox. Toluene, benzene, flurobenzene, and n-hexane were dried using Braun solvent drying system, degassed, and stored over 4Å molecular sieve. Toluene- d_8 , benzene- d_6 , and THF- d_8 were dried by refluxing over NaK, distilled prior to use, and stored over 4Å molecular sieve. Commercial reagents were purchased from Aldrich, Acros, or Alfa-Aesar Chemical Co. and used as received. $L(Cl)GaPCO (L = HC[C(Me)N(Ar)]_2, Ar = 2,6-iPr_2C_6H_3)^{1} IMe_{4,2}^{2}$ and $Fc[B(C_6F_5)_4]^3$ were prepared according to literature procedures. NMR spectra (¹H, ¹³C, ³¹P) were recorded on Bruker Avance II 500 MHz spectrometer and were referenced to internal C₆D₅H (¹H δ = 7.16; ¹³C δ = 128.06), C₇D₇H (¹H δ = 2.08, 6.97, 7.01, 7.09; ¹³C δ = 137.19, 129.26, 127.96, 125.96), and THF- d_8 $({}^{1}\text{H}\ \delta = 1.72, 3.58; {}^{13}\text{C}\ \delta = 67.21, 25.31)$. ${}^{31}\text{P}\ \text{NMR}$ spectra are referenced to internal C₆D₅H (${}^{1}\text{H}\ \delta = 7.16$), C_7D_7H (¹H $\delta = 2.08, 6.97, 7.01, 7.09$), and THF- d_8 (¹H $\delta = 1.72, 3.58$) using chi-values (γ).⁴ Elemental analyses were performed at the *Elementaranalyse Labor* of the University of Duisburg-Essen. IR spectra were recorded by a Bruker ALPHA-T FT-IR spectrometer equipped with a single-reflection ATR sampling module. Melting points were measured in sealed glass capillaries. Cyclic voltammetry studies were performed in a glovebox using a Metrohm Autolab PGSTAT 204 potentiostat with a three electrodes setup consisting of a Pt disc (d 1/4 1 mm) working electrode, Pt wire counter electrode, and Ag wire pseudoreference electrode, and ferrocene as internal standard. Positive feedback compensation was utilized to reduce solvent resistance effects.

Synthesis of [L(Cl)GaP]₂ (2a)

A toluene (10 mL) solution of L(Cl)GaPCO 1 (2.0 g, 3.44 mmol) was irradiated for 5 hours and the resulting green solid was collected by filtration, washed with toluene (2 mL) and dried. Yield: 65% (1.23 g). M.p. 120 °C (dec.). Single crystals suitable for X-ray diffraction were grown by cooling a hot toluene solution of **2a** to ambient temperature. Anal. calcd. (%) for C₅₈H₈₂Cl₂Ga₂N₄P₂ (1104.39): C, 62.89; H, 7.46; N, 5.06. Found: C, 62.83; H, 7.52; N, 5.11. ¹H NMR (500 MHz, C₇D₈, 298 K) δ = 0.99 (d, ³*J*_{HH} = 6.7 Hz, 6H, CH(*CH*₃)₂), 1.13 (d, ³*J*_{HH} = 6.8 Hz, 6H, CH(*CH*₃)₂), 1.19 (t, ³*J*_{HH} = 7.2 Hz, 12H, CH(*CH*₃)₂), 1.64 (s, 6H, CC*H*₃), 3.04 (sept, ³*J*_{HH} = 6.8 Hz, 2H, C*H*(CH₃)₂), 3.74 (sept, ³*J*_{HH} = 6.7 Hz, 2H, C*H*(CH₃)₂), 4.93 (s, 1H, C*H*), 6.90-7.10 (m, 6H, C₆*H*₃). ¹³C{¹H} NMR (125 MHz, C₇D₈, 298 K) δ = 23.1, 23.7, 24.1, 24.5 (CH(*CH*₃)₂), 27.1, 27.6, 28.6, 29.3 (CH(CH₃)₂), 97.0 (CH), 140.3, 142.7, 145.6, 168.9 (C₆H₃). ³¹P{¹H} (202 MHz, C₇D₈, 298 K) δ = 761.6 ppm. ATR-IR: *v* 2957, 2908, 2851, 1552, 1454, 1432, 1379, 1309, 1256, 1170, 1013, 931, 863, 788, 774, 749, 632, 608, 525 cm⁻¹.

Synthesis of [L(Br)GaP]₂ (2b)

TMSBr (2 mL, 13.54 mmol) was added to a solid sample of 2a (1.0 g, 0.902 mmol) at ambient temperature and stirred overnight. All volatiles were then removed in *vacuo* and the green residue was washed with toluene (1 mL). Yield: 99% (1.06 g). Single crystals suitable for X-ray diffraction were grown by cooling a hot toluene solution of **2b** to ambient temperature. M.p. 123 °C (dec.). Anal. calcd. (%) for $C_{58}H_{82}Br_2Ga_2N_4P_2$ (1192.29): C, 58.22; H, 6.91; N, 4.68. Found: C, 58.33; H, 7.05; N, 4.79. ¹H NMR (500 MHz, C₇D₈, 298 K) $\delta = 0.99$ (d, ³*J*_{HH} = 6.5 Hz, 6H, CH(C*H*₃)₂), 1.13 (d, ³*J*_{HH} = 6.6 Hz, 6H, CH(C*H*₃)₂), 1.22 (d, ³*J*_{HH} = 6.8 Hz, 6H, CH(C*H*₃)₂), 1.26 (d, ³*J*_{HH} = 6.6 Hz, 6H, CH(C*H*₃)₂), 1.63 (s, 6H, CC*H*₃), 3.00 (sept, ³*J*_{HH} = 6.8 Hz, 2H, C*H*(CH₃)₂), 3.68 (sept, ³*J*_{HH} = 6.8 Hz, 2H, C*H*(CH₃)₂), 4.86 (s, 1H, C*H*), 6.90-7.08 (m, 6H, C₆*H*₃). ¹³C{¹H} NMR (125 MHz, C₇D₈, 298 K) $\delta = 23.6$, 24.2, 24.7, 25.0 (CH(CH₃)₂), 27.7, 28.2, 29.8 (CH(CH₃)₂), 97.6 (CH), 140.8, 143.3, 146.2, 169.5 (C₆H₃). ³¹P{¹H} (202 MHz, C₇D₈, 298 K) $\delta = 766.8$. ATR-IR: ν 2967, 2919, 2859, 1558, 1519, 1459, 1436, 1384, 1311, 1258, 1173, 1101, 1015, 935, 867, 791, 774, 754, 710, 633, 527 cm⁻¹.

Synthesis of [L(Cl)GaP]₂[IMe₄] (3a)

To a toluene (10 mL) suspension of 2a (100 mg, 0.09 mmol), IMe₄ (11 mg, 0.09 mmol) was added at ambient temperature. The initial green suspension immediately dissolved and a dark red solution formed. It was then filtered over celite and dried in vacuo. It was washed with n-hexane (2 mL), and dried to yield a red crystalline powder. Single crystals suitable for X-ray diffraction were grown by diffusing n-hexane to a saturated toluene solution of 3a at ambient temperature. Yield: 95% (105 mg). M.p. 141 °C (dec.). Anal. calcd. (%) for C₆₅H₉₄Cl₂Ga₂N₆P₂ (1228.49): C, 63.38; H, 7.69; N, 6.82. Found: C, 63.42; H, 7.73; N, 6.87. ¹H NMR (500 MHz, C_7D_8 , 298 K) $\delta = 0.45$ (d, ${}^{3}J_{HH} = 6.4$ Hz, 3H, $CH(CH_3)_2$), 0.69 (d, ${}^{3}J_{HH} = 6.4$ Hz, 3H, $CH(CH_3)_2$, 0.92 (t, ${}^{3}J_{HH} = 7.1 Hz$, 3H, $CH(CH_3)_2$), 1.05 (d, ${}^{3}J_{HH} = 6.8 Hz$, 3H, $CH(CH_3)_2$), 1.09 (d, ${}^{3}J_{HH} = 6.7 Hz$) Hz, 3H, CH(CH₃)₂), 1.12 (d, ${}^{3}J_{HH} = 6.7$ Hz, 3H, CH(CH₃)₂), 1.16-1.28 (m, 12H, CH(CH₃)₂), 1.29 (t, ${}^{3}J_{HH} =$ 6.5 Hz, 6H, CH(CH₃)₂), 1.38-1.40 (m, 12H, CH(CH₃)₂), 1.40 (s, 3H, CCH₃), 1.42 (s, 3H, CCH₃), 1.51 (s, 3H, CCH₃), 1.53 (s, 3H, CCH₃), 1.58 (s, 3H, CCH₃), 1.61 (s, 3H, CCH₃), 2.77 (s, 3H, NCH₃), 3.06 (sept, ${}^{3}J_{HH} =$ 6.7 Hz, 1H, $CH(CH_3)_2$), 3.16 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 3.23 (s, 3H, NCH_3), 3.23 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 3.29 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 3.38 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 3.58 (sept, ${}^{3}J_{\text{HH}} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.80 (sept, ${}^{3}J_{\text{HH}} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.95 (sept, ${}^{3}J_{\text{HH}} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 4.87 (s, 1H, CH), 4.88 (s, 1H, CH), 6.86 (d, ${}^{3}J_{HH} = 7.8$ Hz, 1H, C_6H_3), 7.05-7.08 (m, 2H, C_6H_3 , 7.10-7.15 (m, 5H, C_6H_3), 7.18-7.23 (m, 2H, C_6H_3), 7.26-7.32 (m, 2H, C_6H_3). ¹³C{¹H} NMR (125) MHz, C_7D_8 , 298 K) $\delta = 7.8$, 14.0 (NCH₃), 22.7, 23.2, 23.5, 23.8, 23.9, 24.0, 24.2, 24.6, 24.7, 25.0, 25.3, 25.4, 25.6, 26.2, 26.3, 26.8, 27.2, 27.5, 27.7, 27.9 (CH(CH₃)₂), 28.6, 29.1, 29.5, 29.9, 30.0, 31.7, 32.4, 32.6, 33.3, 33.5 (CH(CH₃)₂), 96.5, 97.9 (CH), 122.9, 123.3, 123.7, 123.9, 124.0, 124.1, 124.3, 125.0, 125.2, 125.3, 126.0, 127.0, 142.0, 143.2, 143.3, 143.5, 143.7, 144.0, 144.4, 145.0, 145.1, 145.3, 151.3, 152.5, 166.4, 166.8, 168.3, 168.8 (C₆H₃). ³¹P{¹H} (202 MHz, C₇D₈, 298 K) $\delta = -85.8$ (d, ¹J_{P-P} = 353.0 Hz), -258.5 (d, ¹J_{P-P} = 353.0 Hz). ATR-IR: v 2951, 2909, 2852, 1541, 1515, 1451, 1431, 1378, 1312, 1253, 1173, 1094, 1014, 934, 854, 792, 755, 724, 689, 635, 526 cm⁻¹.

Synthesis of [L(Br)GaP]₂[IMe₄] (3b)

Compound **3b** was synthesized following the similar protocols used for **3a** using **2b** (100 mg, 0.084 mmol) and IMe₄ (10 mg, 0.084 mmol). Single crystals suitable for X-ray diffraction were grown by diffusing n-

hexane to a saturated toluene solution of 3a at ambient temperature. Yield: 92% (100 mg). M.p. 147 °C (dec.). Anal. calcd. (%) for C₆₅H₉₄Br₂Ga₂N₆P₂ (1316.39): C, 59.11; H, 7.17; N, 6.36. Found: C, 59.15; H, 7.21; N, 6.39. ¹H NMR (500 MHz, C₇D₈, 298 K) $\delta = 0.42$ (d, ³J_{HH} = 6.4 Hz, 3H, CH(CH₃)₂), 0.80 (d, ³J_{HH} = 6.4 Hz, 3H, CH(CH₃)₂), 0.89 (t, ${}^{3}J_{HH} = 7.1$ Hz, 3H, CH(CH₃)₂), 1.01 (d, ${}^{3}J_{HH} = 6.8$ Hz, 3H, CH(CH₃)₂), 1.06 (m, 6H, CH(CH₃)₂), 1.13 (d, ${}^{3}J_{HH} = 6.8$ Hz, 3H, CH(CH₃)₂), 1.15 (d, ${}^{3}J_{HH} = 6.8$ Hz, 6H, CH(CH₃)₂), 1.20 (d, ${}^{3}J_{\text{HH}} = 6.6 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}, 1.21-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, \text{CH}(\text{C}H_{3})_{2}), 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31 \text{ (d, }{}^{3}J_{\text{HH}} = 6.8 \text{ Hz}, 3\text{H}, 1.34-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1.31-1.24 \text{ (m, 6H, CH}(\text{C}H_{3})_{2}), 1$ 1.36 (m, 6H, CH(CH₃)₂),1.40 (s, 3H, CCH₃), 1.42 (s, 3H, CCH₃), 1.47 (s, 3H, CCH₃), 1.50 (s, 3H, CCH₃), 1.54 (s, 3H, CCH₃), 1.59 (s, 3H, CCH₃), 2.86 (s, 3H, NCH₃), 3.04 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.12 (s, 3H, NCH₃), 3.17 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.24 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 2H, CH(CH₃)₂), 3.43 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.67 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, CH(CH₃)₂), 3.83 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$, 3.96 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 4.89 (s, 1H, CH), 4.91 (s, 1H, CH), 6.83 (d, ${}^{3}J_{HH} = 7.5$ Hz, 1H, C₆H₃), 7.02-7.05 (m, 2H, C₆H₃), 7.06-7.12 (m, 5H, C₆H₃), 7.14-7.18 (m, 2H, C₆H₃), 7.22-7.27 (m, 2H, C₆H₃). ¹³C{¹H} NMR (125 MHz, C₇D₈, 298 K) δ = 8.3, 14.3 (NCH₃), 23.1, 23.6, 24.0, 24.3, 24.5, 24.6, 25.1, 25.2, 25.5, 25.6, 26.8, 27.1, 27.2, 27.5, 27.8, 28.0 (CH(CH₃)₂), 28.8, 29.0, 29.5, 29.9, 30.3, 32.0, 33.4, 33.9 (CH(CH₃)₂), 97.2, 98.7 (CH), 123.3, 123.6, 123.9, 124.2, 124.4, 124.5, 124.7, 125.4, 125.7, 126.4, 127.6, 142.3, 142.6, 143.4, 143.6, 143.9, 144.1, 144.4, 144.5, 145.0, 145.6, 145.7, 145.8, 151.6, 152.6, 167.0, 167.4, 169.0, 169.4 (C_6H_3). ³¹P{¹H} (202 MHz, C_7D_8 , 298 K) $\delta = -77.7$ (d, ¹ $J_{P-P} = 358.9$ Hz), -247.5 (d, ¹ J_{P-P}) = 358.9 Hz). ATR-IR: v 2945, 2908, 2855, 1544, 1518, 1454, 1431, 1378, 1312, 1252, 1172, 1096, 1014, 934, 854, 791, 755, 636, 531 cm⁻¹.

Synthesis of {[L(Cl)GaP]₂[IMe₄]}[BAr^F₄] (4a)

To a red toluene solution (5 mL) of **3a** (50 mg, 0.04 mmol), $[FeCp_2][BAr^{F_4}]$ (35 mg, 0.04 mmol) was added at room temperature. Immediately the red solution turned light orange and **4a** was obtained as orange crystals in 83% isolated yield. Yield: 64 mg. M.p.: 100 °C (dec.). Anal. calcd. (%) for C₈₉H₉₄BCl₂F₂₀Ga₂N₆P₂ (1907.47): C, 55.94; H, 4.96; N, 4.40. Found: C, 56.05; H, 5.01; N, 4.47. ATR-IR: *v* 2955, 2911, 2855, 1634, 1521, 1452, 1405, 1365, 1312, 1252, 1172, 1096, 1019, 976, 867, 797, 754, 661, 656 cm⁻¹.

Synthesis of {[L(Br)GaP]₂[IMe₄]}[BAr^F₄] (4b)

Compound **4b** was synthesized following the similar protocols used for **4a** using **3b** (50 mg, 0.038 mmol), $[FeCp_2][BAr^F_4]$ (33 mg, 0.038 mmol). Single crystals suitable for X-ray diffraction were grown by storing a saturated toluene solution of **4b** at ambient temperature. Yield: 79% (60 mg, orange crystals). M.p.: 107 °C (dec.). Anal. calcd. (%) for C₈₉H₉₄BBr₂F₂₀Ga₂N₆P₂ (1999.72): C, 53.46; H, 4.74; N, 4.20. Found: C, 53.54; H, 4.83; N, 4.31. ATR-IR: *v* 2951, 2911, 2855, 1634, 1521, 1504, 1455, 1365, 1311, 1252, 1172, 1079, 1019, 974, 864, 797, 754, 681, 656 cm⁻¹.

Synthesis of [L(Cl)GaP(IMe₄)HPGa(Cl)L][BAr^F₄] (5a)

 $HSn(nBu)_3$ (30 µL, 0.030 mmol; 1M solution in fluorobenzene) was added to a solution of **4a** (50 mg, 0.026 mmol) in flurobenzene (5 mL) at ambient temperature. Upon addition the solution turned colourless. Then

the mixture was stirred for 5a min and all the volatiles were removed in vacuo. The colourless residue was then washed with *n*-hexane (1 mL) and dried in *vacuo* to afford **5a** as a colourless crystalline powder. Yield 73% (36 mg). Single crystals suitable for an X-ray diffraction analysis were obtained by slow diffusion of *n*hexane to a saturated flurobenzene solution of 5a at ambient temperature. M.p.: 153 °C (dec.). Anal. calcd. (%) for C₈₉H₉₅BCl₂F₂₀Ga₂N₆P₂ (1908.48): C, 55.91; H, 5.01; N, 4.40. Found: C, 56.01; H, 5.12; N, 4.47. ¹H NMR (300 MHz, THF- d_8 , 298 K) $\delta = 0.25$ (d, ${}^{3}J_{HH} = 6.3$ Hz, 3H, CH(CH₃)₂), 0.45 (d, ${}^{3}J_{HH} = 6.1$ Hz, 3H, CH(CH₃)₂), 0.80 (d, ${}^{3}J_{HH} = 6.6$ Hz, 3H, CH(CH₃)₂), 0.92 (d, ${}^{3}J_{HH} = 6.5$ Hz, 3H, CH(CH₃)₂), 1.00–1.06 (m, 6H, CH(CH₃)₂), 1.08–1.12 (m, 9H, CH(CH₃)₂), 1.18–1.23 (m, 9H, CH(CH₃)₂), 1.28 (t, ${}^{3}J_{HH} = 6.7$ Hz, 6H, $CH(CH_3)_2$, 1.39 (t, ${}^{3}J_{HH} = 6.6$ Hz, 6H, $CH(CH_3)_2$), 1.69 (s, 3H, CCH_3), 1.77 (s, 6H, CCH_3), 1.81 (s, 3H, CCH₃), 1.93 (s, 3H, CCH₃), 1.97 (s, 3H, CCH₃), 2.56 (s, 3H, NCH₃), 2.72 (s, 3H, NCH₃), 2.78 (sept, 2H, $CH(CH_3)_2$, 3.03 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 2H, $CH(CH_3)_2$), 3.37 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 1H, $CH(CH_3)_2$), 3.49 (sept, ${}^{3}J_{\text{HH}} = 6.7$ Hz, 1H, CH(CH₃)₂), 5.42 (s, 1H, CH), 5.45 (s, 1H, CH), 6.91 (d, ${}^{3}J_{\text{HH}} = 6.8$ Hz, 1H, C₆H₃), 7.10–7.17 (m, 2H, C₆H₃), 7.19–7.27 (m, 4H, C₆H₃), 7.31–7.40 (m, 4H, C₆H₃), 7.53 (t, ${}^{3}J_{\text{HH}} = 7.6$ Hz, 1H, C_6H_3). ¹³C{¹H} NMR (125 MHz, THF- d_8 , 298 K) $\delta = 7.7$, 8.6, 9.1, 23.9, 24.4 (CH(CH_3)_2), 28.3, 29.0, 29.3, 29.4, 30.3, 33.1 (CH(CH₃)₂), 98.8, 99.4 (CH), 115.9, 116.0 (NC), 125.0, 125.4, 126.0, 128.2, 128.8, 129.2, 131.0, 136.4, 138.0, 140.1, 145.4, 148.4, 150.0, 163.2, 164.7, 171.7, 171.9, 172.7 (*C*₆H₃). ³¹P{¹H} (121 MHz, THF- d_8 , 298 K) $\delta = -148.8$ (d, ${}^{1}J_{PP} = 84.9$ Hz), -229.1 (d, ${}^{1}J_{PP} = 82.0$ Hz). ${}^{31}P$ (121 MHz, THF- d_8 , 298 K) $\delta =$ -147.5 (dd, ${}^{1}J_{PP} = 25.3$, 84.9 Hz), -228.0 (dd, ${}^{1}J_{PP} = 85.6$, 180.9 Hz). ATR-IR: v 2964, 2926, 2867, 2156, 1640, 1510, 1456, 1371, 1258, 1178, 1082, 1021, 977, 941, 871, 800, 758, 684, 659, 604, 571, 533 cm⁻¹.

Synthesis of [L(Br)GaP(IMe₄)HPGa(Br)L][BAr^F₄] (5b)

Compound **5b** was synthesized following the similar protocols used for **5a** using **4b** (50 mg, 0.025 mmol), HSn(nBu)₃ (30 µL, 0.030 mmol; 1M solution in fluorobenzene). Yield: 70% (34 mg, colorless crystals). M.p.: 151 °C (dec.). Anal. calcd. (%) for C₈₉H₉₅BBr₂F₂₀Ga₂N₆P₂ (1996.37): C, 53.43; H, 4.79; N, 4.20. Found: C, 53.61; H, 4.96; N, 4.27. ¹H NMR (300 MHz, THF- d_8 , 298 K) $\delta = 0.28$ (d, ³ $J_{\text{HH}} = 6.0$ Hz, 3H, CH(CH₃)₂), 0.48 (d, ${}^{3}J_{HH} = 5.9$ Hz, 3H, CH(CH₃)₂), 0.82 (d, ${}^{3}J_{HH} = 6.2$ Hz, 3H, CH(CH₃)₂), 0.93 (d, {}^{3}J_{HH} = 6.2 Hz, 3H 6.5 Hz, 3H, CH(CH₃)₂), 1.03 (t, ${}^{3}J_{HH} = 6.1$ Hz, 6H, CH(CH₃)₂), 1.11–1.19 (m, 9H, CH(CH₃)₂), 1.18–1.23 (m, 9H, CH(CH₃)₂), 1.28 (br, 6H, CH(CH₃)₂), 1.39 (t, ${}^{3}J_{HH} = 6.5$ Hz, 6H, CH(CH₃)₂), 1.69 (s, 3H, CCH₃), 1.77 (s, 3H, CCH₃), 1.78 (s, 3H, CCH₃), 1.82 (s, 3H, CCH₃), 1.94 (s, 3H, CCH₃), 1.97 (s, 3H, CCH₃), 2.51 (s, 3H, NCH₃), 2.72 (s, 3H, NCH₃), 2.86 (sept, 3H, CH(CH₃)₂), 3.09 (sept, ${}^{3}J_{HH} = 6.7$ Hz, 2H, CH(CH₃)₂), 3.41 (sept, ${}^{3}J_{\text{HH}} = 6.7$ Hz, 1H, CH(CH₃)₂), 5.41 (s, 1H, CH), 5.46 (s, 1H, CH), 6.89 (d, ${}^{3}J_{\text{HH}} = 7.4$ Hz, 1H, C₆H₃), 7.14–7.18 (m, 2H, C₆H₃), 7.27–7.33 (m, 4H, C₆H₃), 7.35–7.40 (m, 4H, C₆H₃), 7.54 (t, ${}^{3}J_{\text{HH}} = 7.6$ Hz, 1H, C_6H_3). ¹³C{¹H} NMR (125 MHz, THF- d_8 , 298 K) $\delta = 8.8, 9.1, 24.1, 24.4$ (CH(CH₃)₂), 27.4, 28.4, 28.7, 30.1, 30.4 (CH(CH₃)₂), 99.1, 99.7 (CH), 115.9, 116.0 (CN), 125.4, 125.6, 125.6, 125.9, 126.2, 126.4, 129.0, 129.2, 129.4, 129.7, 130.9, 136.4, 138.0, 138.4, 140.0, 140.4, 140.7, 141.1, 142.5, 144.1, 146.4, 148.5, 150.1, 163.1, 164.8, 171.2, 171.7, 172.0 (C_6H_3). ³¹P{¹H} (121 MHz, THF- d_8 , 298 K) $\delta = -142.5$ (d, ¹ $J_{PP} = 87.9$ Hz), -225.5 (d, ${}^{1}J_{PP} = 87.2$ Hz). ${}^{31}P$ (121 MHz, THF- d_8 , 298 K) $\delta = -142.4$ (dd, ${}^{1}J_{PP} = 24.7, 87.9$ Hz), -225.3 (dd, ${}^{1}J_{PP} = -142.4$ (

88.5, 181.1 Hz). ATR-IR: *v* 2967, 2926, 2872, 2155, 1641, 1512, 1456, 1367, 1315, 1257, 1177, 1085, 1021, 979, 938, 870, 800, 757, 685, 658, 608, 571, 531 cm⁻¹.

2. Spectroscopic Characterization



Figure S1. ¹H NMR (500 MHz, C₇D₈, 298 K) spectrum of compound 2a.



Figure S2. $^{13}\mathrm{C}\{^{1}\mathrm{H}\}$ NMR (125 MHz, $\mathrm{C}_{7}\mathrm{D}_{8},$ 298 K) spectrum of compound 2a.



Figure S3. $^{31}P\{^{1}H\}$ NMR (202 MHz, $C_{7}D_{8},$ 298 K) spectrum of compound 2a.



Figure S4. ATR-IR spectrum of 2a.



Figure S5. ¹H NMR (500 MHz, C₇D₈, 298 K) spectrum of compound **2b**.



Figure S6. ${}^{13}C{}^{1}H$ NMR (125 MHz, C_7D_8 , 298 K) spectrum of compound 2b.



Figure S7. $^{31}P\{^{1}H\}$ NMR (202 MHz, $C_7D_8,$ 298 K) spectrum of compound 2b.



Figure S8. ATR-IR spectrum of 2b.





Figure S9. ¹H NMR (300 MHz, C₇D₈, 298 K) spectrum of compound 3a.



Figure S10. ¹³C $\{^{1}H\}$ NMR (150 MHz, C₇D₈, 298 K) spectrum of compound **3a**.



Figure S11. $^{31}P\{^{1}H\}$ NMR (161 MHz, $C_7D_8,$ 298 K) spectrum of compound 3a.



Figure S12. ATR-IR spectrum of 3a.



Figure S13. ¹H NMR (500 MHz, C₇D₈, 298 K) spectrum of compound 3b.



Figure S14. ${}^{13}C{}^{1}H$ NMR (125 MHz, C₇D₈, 298 K) spectrum of compound **3b**.



Figure S16. ATR-IR spectrum of 3b.







Figure S18. ATR-IR spectrum of 4b.



Figure S20. ¹³C{¹H} NMR (125 MHz, THF- d_8 , 298 K) spectrum of compound 5a.



Figure S22. ³¹P NMR (121 MHz, THF-*d*₈, 298 K) spectrum of compound 5a.









Figure S24. ¹H NMR (300 MHz, THF-*d*₈, 298 K) spectrum of compound 5b.



Figure S26. ³¹P{¹H} NMR (121 MHz, THF- d_8 , 298 K) spectrum of compound 5b.



Figure S27. ³¹P NMR (121 MHz, THF-*d*₈, 298 K) spectrum of compound **5b**.



Figure S28. ATR-IR spectrum of 5b.

3. Cyclic Voltammetry



Figure S29a. Cyclic voltammogram of **2a** in THF (0.01 M *n*-Bu₄N[PF₆]) as a supporting electrolyte, $0.1Vs^{-1}$, *vs* Fc/Fc⁺).



Figure S29b. Cyclic voltammogram of 2a in THF (0.01 M n-Bu₄N[B(C₆F₅)₄]) as a supporting electrolyte, 0.1Vs⁻¹, vs Fc/Fc⁺).



Figure S30a. Cyclic voltammogram of **2b** in THF (0.01 M *n*-Bu₄N[PF₆]) as a supporting electrolyte, 0.1Vs⁻¹, *vs* Fc/Fc⁺).



Figure S30b. Cyclic voltammogram of **2b** in THF (0.01 M *n*-Bu₄N[B(C₆F₅)₄]) as a supporting electrolyte, 0.1Vs⁻¹, vs Fc/Fc⁺).



Figure S31. Cyclic voltammogram of **3a** in THF (0.01 M *n*-Bu₄N[B(C₆F₅)₄]) as a supporting electrolyte, 0.05Vs⁻¹, *vs* Fc/Fc⁺). The cycle for Fc/Fc⁺ couple has been removed for clarity.



Figure S32. Cyclic voltammogram of **3b** in THF (0.01 M *n*-Bu₄N[B(C₆F₅)₄]) as a supporting electrolyte, $0.1Vs^{-1}$, vs Fc/Fc⁺). The cycle for Fc/Fc⁺ couple has been removed for clarity.

4. UV-vis Spectroscopy



Figure S33. UV-visible spectrum of 2a (10⁻⁴ M) recorded in flurobenzene.



Figure S34. UV-visible spectrum of 2b (10⁻⁴ M) recorded in flurobenzene.



Figure S35. UV-visible spectrum of 3a (10⁻⁴ M) recorded in flurobenzene.



Figure S36. UV-visible spectrum of 3b (10⁻⁴ M) recorded in flurobenzene.



Figure S37. UV-visible spectrum of **4a** (10⁻⁴ M) recorded in flurobenzene.



Figure S38. UV-visible spectrum of 4b (10⁻⁴ M) recorded in flurobenzene.

5. X-Ray Crystallographic Analysis

The crystals were mounted on nylon loops in inert oil. Data of **2a**, **3a**, and **3b** were collected on a Bruker AXS D8 Kappa diffractometer with APEX2 detector (monochromated Mo_{*Ka*} radiation, $\lambda = 0.71073$ Å) at 100(2) K while those of **2b**, **4a**, **4b**, **5a**, and **5b** were collected on a Bruker AXS D8 Venture diffractometer with Photon II detector (monochromated Cu_{*Ka*} radiation, $\lambda = 1.54178$ Å, microfocus source) at 100(2) K. The structures were solved by Direct Methods (SHELXS-2013)^[5] and refined anisotropically by full-matrix least-squares on F^2 (SHELXL-2017).^[6-8] Absorption corrections were performed semi-empirically from equivalent reflections on basis of multi-scans (Bruker AXS APEX3). Hydrogen atoms were refined using a riding model or rigid methyl groups.

In 2a an isopropyl group is disordered over two positions. The structure of 3a contains highly disordered solvent - possibly *n*-hexane. The final refinement was done with a solvent free dataset from a PLATON/SQUEEZE run. (For details see: A. L. Spek, Acta Cryst. A46 (1990), 194-201). Since the nature and amount of the solvent is not clear it was not included in the sum formula. In 3b an isopropyl group is disordered over two positions. All its corresponding bond lengths and angles were restrained to be equal (SADI). RIGU and SIMU restraints were applied to its atoms' anisotropic displacement parameters. The solvent molecule is disordered over a centre of inversion. Its bond lengths and angles were restrained to be equal (SADI) and RIGU restraints were used for the displacement parameters. Due to their close proximity C3 1 and C3 2 were refined with common displacement parameters (EADP). The central core of the cation of 4a is disordered over two positions. The corresponding bond lengths and angles were restrained to be equal (SADI) and RIGU restraints were applied to the displacement parameters. Due to their close proximity C60 and C60' as well as C61 and C61' were refined with additional SIMU restraints. The toluene molecule is disordered over two positions. RIGU restraints were used for the refinement of its atoms' displacement parameters. In 4b a second orientation of the central Br-Ga-P-P-Ga-Br chain could be identified and refined. Hints on the missing carbene of this component can be found in the residual electron density but due to the low occupancy (approx. 5%) the second orientation of the carbene could not be identified completely nor refined. RIGU and SIMU restraints were used to refine the displacement parameters of the minor component. For Ga2' and Br2' an additional ISOR ($\sigma = 0.003$) restraint was necessary. In 5a the PH hydrogen atoms were refined freely with their P-H bond lengths restrained to be equal (SADI). The central Ga(Cl)-PH-P(carbene)-Ga(Cl) moiety is disordered over two positions. For Ga1 and Cl1 a third orientation can be found but a refinement failed due to the low occupancy. The corresponding bond lengths of the carbene were restrained to be equal (SADI) and RIGU restraints were applied to its displacement parameters. For atoms in close proximity additional SIMU restraints were used. The structure contains highly disordered solvent - fluoro benzene or toluene. The final refinement was done with a solvent free dataset from a PLATON/SQUEEZE run. (For details see: A. L. Spek, Acta Cryst. A46 (1990), 194-201). Since the nature and amount of the solvent is not clear it was not included in the sum formula. In 5b the PH hydrogen atoms could not be identified in the residual electron density. Their positions were taken from the isomorphous chloro compound. The H atom of the larger disorder component was refined freely, the co-ordinates of the one of the smaller were refined by a riding model (AFIX 3) and its displacement parameter was refined freely. The P–H bond lengths were restrained to be equal to 1.42 Å (DFIX). The proper assignment of these hydrogen atoms should be confirmed by other means. The central Ga(Br)–PH–P(carbene)–Ga(Br) moiety is disordered over two positions. The corresponding bond lengths of the carbene were restrained to be equal (SADI) and RIGU restraints were applied to its displacement parameters. For atoms in close proximity additional SIMU restraints were used. The alternate positions of N5 were refined with common displacement parameters (EADP). The structure also contains highly disordered solvent – possibly fluoro benzene or toluene. The final refinement was done with a solvent free dataset from a PLATON/SQUEEZE run. (For details see: A. L. Spek, *Acta Cryst. A46* (**1990**), 194–201). Since the nature and amount of the solvent is not clear, it was not included in the sum formula. The crystal was a very thin plate and diffracted rather poorly, esp. at higher angles. In addition, the highly anisotropic shape of the crystal hampered the absorption correction. This leads to a rather high R_{int} . Due to the weak data quantitative results should be carefully scrutinized and may be unreliable.

CCDC-2184009 (**2a**; mks_001m), -2184010 (**2b**; mks_005m), -2184011 (**3a**; mks_037am_sq), -2184012 (**3b**; mks_049m), -2184320 (**4a**; mks_057m), -2184142 (**4b**; mks_056m), -2184019 (**5a**; mks_073), and -2184020 (**5b**; mks_077) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via <u>www.ccdc.cam.ac.uk/data_request/cif</u>.

Compound	2a	2b	3a
Emp. formula	$C_{58}H_{82}Cl_2Ga_2N_4P_2$	$C_{58}H_{82}Br_2Ga_2N_4P_2$	$C_{65}H_{94}Cl_2Ga_2N_6P_2$
Formula weight	1107.55	1196.47	1231.74
Temperature [K]	100(2)	100(2)	$0.152 \times 0.098 \times 0.094$
Crystal system	monoclinic	monoclinic	triclinic
Space group	$P2_{1}/n$	$P2_1/n$	<i>P</i> -1
<i>a</i> [Å]	13.7010(18)	13.6417(6)	14.863(2)
<i>b</i> [Å]	13.8715(16)	13.9614(6)	20.330(5)
<i>c</i> [Å]	15.7041(18)	15.7579(7)	25.712(6)
α [°]	90	90	83.603(4)
β[°]	105.336(6)	105.1021(15)	75.258(6)
γ [°]	90	90	72.565(4)
V[Å ³]	2878.3(6)	2897.6(2)	7163(3)
Ζ	2	2	4
ρ [Mgm ⁻³]	1.278	1.371	1.142
μ [mm ⁻¹]	1.124	3.569	0.910
F (000)	1168	1240	2608
Crystal size [mm]	0.452 x 0.402 x 0.368	0.151 × 0.136 × 0.096	$0.152 \times 0.098 \times 0.094$
$\theta \max [^{\circ}]$	40.215	80.790	30.528
Index ranges	$-16 \le h \le 24$	$-17 \le h \le 17$	$-21 \le h \le 21$
	$-25 \le k \le 25$	$-17 \le k \le 17$	$-29 \le k \le 29$
	$-28 \le l \le 27$	$-20 \le l \le 19$	$-36 \le l \le 36$
No. of reflect. collected	72672	113993	264320
Unique reflect.	18029	6312	43665
R _{int}	0.0346	0.0431	0.0782
Data / restraints / params.	18029 / 0 / 347	6312 / 0 / 317	43665 / 0 / 1435
Goodness-of-fit on F ²	1.043	1.032	1.032
$R1 [I > 2\sigma(I)]$	0.0330	0.0228	0.0436
$wR2 [I > 2\sigma(I)]$	0.0822	0.0593	0.1016
R1 [all data]	0.0495	0.0243	0.0756
wR2 [all data]	0.0910	0.0593	0.1119
Largest diff. peak and hole max./min.[e·Å ⁻³]	1.096 and -0.610	0.502 and -0.377	1.550 and -0.921

Table S1. Crystal data and structure refinement of 2a, 2b, and 3a.

Compound	3b	4a	4b
Emp. formula	$C_{68}H_{101}Br_2Ga_2N_6P_2$	$C_{96}H_{102}BCl_2F_{20}Ga_2N_6P_2$	$C_{89}H_{94}BBr_2F_{20}Ga_2N_6P_2$
Formula weight	1363.74	2002.92	1999.71
Temperature [K]	100(2)	100(2)	100(2)
Crystal system	monoclinic	triclinic	monoclinic
Space group	$P2_{1}/n$	<i>P</i> -1	$P2_{1}/c$
<i>a</i> [Å]	14.7131(10)	12.6151(10)	21.557(2)
<i>b</i> [Å]	25.5509(18)	16.6393(13)	20.432(2)
<i>c</i> [Å]	18.5397(13)	23.1348(18)	20.684(2)
α [°]	90	77.844(3)	90
β[°]	99.6843(16)	77.731(3)	20.684(2)
γ [°]	90	86.897(3)	90
V[Å ³]	6870.4(8)	4638.6(6)	8979.6(16)
Ζ	4	2	4
ρ [Mgm ⁻³]	1.318	1.434	1.479
μ [mm ⁻¹]	2.037	2.344	2.888
F (000)	2852	2062	4068
Crystal size [mm]	$0.477 \times 0.204 \times 0.122$	$0.326 \times 0.149 \times 0.080$	$0.130 \times 0.112 \times 0.096$
$\theta \max [^{\circ}]$	33.212	80.659	81.277
Index ranges	$-22 \le h \le 22$	$-16 \le h \le 16$	$-27 \le h \le 27$
	$-39 \le k \le 39$	$-21 \le k \le 21$	$-26 \le k \le 25$
	$-28 \le l \le 28$	$-29 \le l \le 29$	$-26 \le l \le 26$
No. of reflect. collected	228807	283933	405733
Unique reflect.	26297	20142	19684
R _{int}	0.0514	0.0434	0.0949
Data / restraints / params.	20984 / 117 / 799	18999 / 266 / 1392	16685 / 39 / 1178
Goodness-of-fit on F ²	1.043	1.171	1.011
$R1 [I > 2\sigma(I)]$	0.0314	0.0382	0.0281
wR2 [I>2 σ (I)]	0.0641	0.0850	0.0665
R1 [all data]	0.0501	0.0407	0.0369
wR2 [all data]	0.0703	0.0862	0.0706
Largest diff. peak and hole max./min.[e·Å ⁻³]	0.774 and -0.463	0.580 and -0.506	0.430 and -0.708

Table S2. Crystal data and structure refinement of 3b, 4a, and 4b.

Compound	5a	5b
Emp. formula	$C_{89}H_{95}BCl_2F_{20}Ga_2N_6P_2$	$C_{89}H_{95}BBr_2F_{20}Ga_2N_6P_2$
Formula weight	1911.79	2000.71
Temperature [K]	100(2)	100(2)
Crystal system	triclinic	triclinic
Space group	<i>P</i> -1	<i>P</i> -1
a [Å]	12.5203(3)	12.5498(4)
<i>b</i> [Å]	16.6728(5)	16.6890(7)
c [Å]	23.1865(7)	23.2672(9)
α [°]	78.0634(15)	77.921(2)
β[°]	77.1747(12)	77.403(2)
γ [°]	87.3124(14)	87.595(2)
V[Å ³]	4617.3(2)	4650.5(3)
Ζ	2	2
ρ [Mgm ⁻³]	1.375	1.429
μ [mm ⁻¹]	2.327	2.789
F (000)	1964	2036
Crystal size [mm]	0.261 x 0.171 x 0.142	0.144 x 0.124 x 0.015
$\theta \max [^{\circ}]$	2.71 to 79.55	1.988 to 75.216
Index ranges	$-15 \le h \le 15$	$-15 \le h \le 15$
	$-21 \le k \le 21$	$-20 \le k \le 20$
	$-29 \le l \le 29$	$-29 \le l \le 29$
No. of reflect. collected	267036	234649
Unique reflect.	18937	14631
R _{int}	0.0426	0.1279
Data / restraints / params.	20020 / 200 / 1274	19037 / 195 / 1263
Goodness-of-fit on F ²	1.230	1.192
$R1 \left[I > 2\sigma(I)\right]$	0.0584	0.0793
$wR2 \left[I > 2\sigma(I)\right]$	0.1201	0.1684
R1 [all data]	0.0610	0.1015
wR2 [all data]	0.1212	0.1771
Largest diff. peak and hole max./min.[Å ⁻³]	1.084 and -0.641	0.577 and -0.795

Table S3. Crystal data and structure refinement of 5a, and 5b.



Figure S39. Molecular structure of **2a** with thermal ellipsoids at 50% probability level. The hydrogen atoms are omitted for clarity. Selected bond length (Å) and angels (°): Ga(1)–N(2) 1.9363(7), Ga(1)–N(1) 1.9436(7), Ga(1)–Cl(1) 2.2207(3), Ga(1)–P(1) 2.3131(3), P(1)–P(1)^{#1} 2.0381(5); N(2)–Ga(1)–N(1) 97.40(3), N(2)–Ga(1)–Cl(1) 105.39(2), N(1)–Ga(1)–Cl(1) 103.88(2), N(2)–Ga(1)–P(1) 111.98(2), N(1)–Ga(1)–P(1) 118.22(2), Cl(1)–Ga(1)–P(1) 117.464(13), P(1)^{#1}–P(1)–Ga(1) 95.182(18) ^{#1}coordinates used to generate equivalent atoms: -x,-y,-z+1.



Figure S40. Molecular structure of **2b** with thermal ellipsoids at 50% probability level. The hydrogen atoms are omitted for clarity. Selected bond length (Å) and angels (°): Br(1)-Ga(1) 2.3680(2), Ga(1)-N(2) 1.9331(12), Ga(1)-N(1) 1.9439(12), Ga(1)-P(1) 2.3120(4), $P(1)-P(1)^{\#1} 2.0282(8)$; N(2)-Ga(1)-N(1) 97.58(5), N(2)-Ga(1)-P(1) 112.51(3), N(1)-Ga(1)-P(1) 117.80(3), N(2)-Ga(1)-Br(1) 105.41(3), N(1)-Ga(1)-Br(1) 104.03(3), P(1)-Ga(1)-Br(1) 117.114(13), $P(1)^{\#1}-P(1)-Ga(1) 94.92(2)$.



Figure S41. Molecular structure of 3a with thermal ellipsoids at 50% probability level. The hydrogen atoms are omitted for clarity. Selected bond length (Å) and angels (°): (molecule 1) Ga(1)-N(1) 1.9506(16), Ga(1)-N(2) 1.9722(17), Ga(1)–Cl(1) 2.2249(6), Ga(1)–P(1) 2.3328(7), Ga(2)–N(3) 1.9851(17), Ga(2)–N(4) 2.0167(17), Ga(2)–Cl(2) 2.2590(7), Ga(2)–P(2) 2.2747(6), P(1)–C(59) 1.862(2), P(1)–P(2) 2.1871(8); N(1)– Ga(1)-N(2) 96.53(7), N(1)-Ga(1)-Cl(1) 101.43(5), N(2)-Ga(1)-Cl(1) 97.97(5), N(1)-Ga(1)-P(1) 120.93(5), N(2)-Ga(1)-P(1) 116.64(5), Cl(1)-Ga(1)-P(1) 118.81(2), N(3)-Ga(2)-N(4) 92.72(7), N(3)-Ga(2)-Cl(2)101.85(5), N(4)-Ga(2)-Cl(2) 98.33(5), N(3)-Ga(2)-P(2) 121.02(5), N(4)-Ga(2)-P(2) 118.33(5), Cl(2)-Ga(2)-P(2) 119.33(2), C(59)-P(1)-P(2) 110.67(7), C(59)-P(1)-Ga(1) 95.21(6), P(2)-P(1)-Ga(1) 103.70(3), P(1)–P(2)–Ga(2) 95.93(3); (molecule 2) Ga(1)–N(2) 1.9551(16), Ga(1)–N(1) 1.9714(15), Ga(1)–Cl(1) 2.2113(6), Ga(1)-P(1) 2.3433(6), Ga(2)-N(4) 2.0023(17), Ga(2)-N(3) 2.0113(16), Ga(2)-Cl(2) 2.2579(6), Ga(2)-P(2) 2.2695(7), P(1)-C(59) 1.864(2), P(1)-P(2) 2.1835(8); N(1)-Ga(1)-N(2) 96.53(7), N(1)-Ga(1)-Cl(1) 101.43(5), N(2)-Ga(1)-Cl(1) 97.97(5), N(1)-Ga(1)-P(1) 120.93(5), N(2)-Ga(1)-P(1) 116.64(5), Cl(1)-Ga(1)-P(1) 118.81(2), N(3)-Ga(2)-N(4) 92.72(7), N(3)-Ga(2)-Cl(2) 101.85(5), N(4)-Ga(2)-Cl(2) 98.33(5), N(3)-Ga(2)-P(2) 121.02(5), N(4)-Ga(2)-P(2) 118.33(5), Cl(2)-Ga(2)-P(2) 119.33(2), C(59)–P(1)–P(2) 110.67(7), C(59)–P(1)–Ga(1) 95.21(6), P(2)–P(1)–Ga(1) 103.70(3),P(1)-P(2)-Ga(2) 95.93(3).



Figure S42. Molecular structure of **3b** with thermal ellipsoids at 50% probability level. The hydrogen atoms, and the alternate position of the disordered part are omitted for clarity. Selected bond length (Å) and angels (°): Br(1)–Ga(1) 2.3981(2), Br(2)–Ga(2) 2.4083(2), Ga(1)–N(1) 1.9513(11), Ga(1)–N(2) 1.9732(11),Ga(1)–P(1) 2.3205(4), Ga(2)–N(4) 1.9751(11), Ga(2)–N(3) 2.0028(11), Ga(2)–P(2) 2.2664(4), P(1)–C(59) 1.8570(13), P(1)–P(2) 2.1623(5); N(1)–Ga(1)–N(2) 96.36(5), N(1)–Ga(1)–P(1) 114.82(3), N(2)–Ga(1)–P(1) 134.28(4), N(1)–Ga(1)–Br(1) 102.46(4), N(2)–Ga(1)–Br(1) 100.16(4), P(1)–Ga(1)–Br(1) 104.349(11), N(4)–Ga(2)–N(3) 93.36(5), N(4)–Ga(2)–P(2) 121.61(3), N(3)–Ga(2)–P(2) 117.42(3), N(4)–Ga(2)–Br(2) 101.94(3), N(3)–Ga(2)–Br(2) 99.17(3), P(2)–Ga(2)–Br(2) 118.447(12), C(59)–P(1)–P(2) 110.55(5), C(59)–P(1)–Ga(1) 96.46(4), P(2)–P(1)–Ga(1) 113.099(18), P(1)–P(2)–Ga(2) 94.977(17).



Figure S43. Molecular structure of **4a** with thermal ellipsoids at 50% probability level. The hydrogen atoms, a disordered solvent molecule (toluene), and the alternate position of the disordered part are omitted for clarity. Selected bond length (Å) and angels (°): Ga(1)-N(2) 1.9261(19), Ga(1)-N(1) 1.983(2), Ga(1)-Cl(1) 2.2021(10), Ga(1)-P(1) 2.3479(13), Ga(2)-N(3) 1.8686(18), Ga(2)-N(4) 1.9669(18), Ga(2)-Cl(2) 2.2074(10), Ga(2)-P(2) 2.3456(10), P(1)-C(59) 1.823(3), P(1)-P(2) 2.1423(8); N(2)-Ga(1)-N(1) 96.57(8), N(2)-Ga(1)-Cl(1) 106.28(6), N(1)-Ga(1)-Cl(1) 105.95(7), N(2)-Ga(1)-P(1) 112.02(7), N(1)-Ga(1)-P(1) 118.73(6), Cl(1)-Ga(1)-P(1) 115.17(6), N(3)-Ga(2)-N(4) 99.03(8), N(3)-Ga(2)-Cl(2) 105.72(6), N(4)-Ga(2)-Cl(2) 101.77(6), N(3)-Ga(2)-P(2) 108.10(6), N(4)-Ga(2)-P(2) 124.00(6), Cl(2)-Ga(2)-P(2) 115.82(5), C(59)-P(1)-P(2) 110.32(11), C(59)-P(1)-Ga(1) 102.17(9), P(2)-P(1)-Ga(1) 110.25(4), P(1)-P(2)-Ga(2) 100.12(3).



Figure S44. Molecular structure of **4b** with thermal ellipsoids at 50% probability level. The hydrogen atoms and the alternate position of the disordered part are omitted for clarity. Selected bond length (Å) and angels (°):Br(1)-Ga(1) 2.3498(5), Br(2)-Ga(2) 2.3403(6), Ga(1)-N(2) 1.9308(15), Ga(1)-N(1) 1.9422(15), Ga(1)-P(1) 2.3591(8), Ga(2)-N(4) 1.9336(15), Ga(2)-N(3) 1.9553(15), Ga(2)-P(2) 2.3522(7), P(1)-C(59) 1.8262(19), P(1)-P(2) 2.1336(7); N(2)-Ga(1)-N(1) 97.57(6), N(2)-Ga(1)-Br(1) 108.74(5), N(1)-Ga(1)-Br(1) 104.00(5), N(2)-Ga(1)-P(1) 119.49(5), N(1)-Ga(1)-P(1) 111.16(5), Br(1)-Ga(1)-P(1) 113.71(3), N(4)-Ga(2)-N(3) 96.91(6), N(4)-Ga(2)-Br(2) 105.42(5), N(3)-Ga(2)-Br(2) 104.47(5), N(4)-Ga(2)-P(2) 111.79(5), N(3)-Ga(2)-P(2) 120.45(5), Br(2)-Ga(2)-P(2) 115.40(2), C(59)-P(1)-P(2) 111.85(6), C(59)-P(1)-Ga(1) 100.18(6), P(2)-P(1)-Ga(1) 113.07(3), P(1)-P(2)-Ga(2) 101.59(3).



Figure S45. Molecular structure of **5a** with thermal ellipsoids at 50% probability level. The hydrogen atoms and the alternate position of the disordered part are omitted for clarity. Selected bond length (Å) and angels (°): Ga(1)-N(2) 1.913(3), Ga(1)-N(1) 1.991(3), Ga(1)-Cl(1) 2.2000(14), Ga(1)-P(1) 2.3655(16), Ga(2)-N(3) 1.860(3), Ga(2)-N(4) 1.988(3), Ga(2)-Cl(2) 2.2032(15), Ga(2)-P(2) 2.3466(16), P(1)-C(59) 1.838(4), P(1)-P(2) 2.2227(14), N(5)-C(59) 1.351(5), N(5)-C(60) 1.378(8), N(5)-C(62) 1.468(6), N(6)-C(59) 1.348(6); N(2)-Ga(1)-N(1) 96.48(14), N(2)-Ga(1)-Cl(1) 105.43(10), N(1)-Ga(1)-Cl(1) 104.99(10), N(2)-Ga(1)-P(1) 110.68(10), N(1)-Ga(1)-P(1) 117.68(7), N(3)-Ga(2)-N(4) 98.82(12), N(3)-Ga(2)-Cl(2) 106.26(10), N(4)-Ga(2)-Cl(2) 102.64(9), N(3)-Ga(2)-P(2) 114.16(9), N(4)-Ga(2)-P(2) 118.06(10), Cl(2)-Ga(2)-P(2) 114.93(7), C(59)-P(1)-P(2) 103.18(16), C(59)-P(1)-Ga(1) 99.66(14), P(2)-P(1)-Ga(1) 107.23(6), P(1)-P(2)-Ga(2) 99.12(5), N(6)-C(59)-P(1) 132.6(4), N(5)-C(59)-P(1) 121.3(4).

Figure S46. Molecular structure of **5b** with thermal ellipsoids at 50% probability level. The hydrogen atoms and the alternate position of the disordered part are omitted for clarity. Selected bond length (Å) and angels (°): Ga(1)-N(2) 1.914(5), Ga(1)-N(1) 2.007(6), Ga(1)-Br(1) 2.356(2), Ga(1)-P(1) 2.367(3), Ga(2)-N(3) 1.859(6), Ga(2)-N(4) 2.004(6), Ga(2)-P(2) 2.339(3), Ga(2)-Br(2) 2.354(2), P(1)-C(59) 1.829(8), P(1)-P(2) 2.223(3), N(5)-C(59) 1.340(10), N(5)-C(60) 1.387(10), N(5)-C(62) 1.452(9), N(6)-C(59) 1.54(9); N(2)-Ga(1)-N(1) 96.3(2), N(2)-Ga(1)-Br(1) 105.73(17), N(1)-Ga(1)-Br(1) 104.58(17), N(2)-Ga(1)-P(1) 110.26(19), N(1)-Ga(1)-P(1) 118.78(17), Br(1)-Ga(1)-P(1) 118.34(11), N(3)-Ga(2)-N(4) 98.9(2), N(3)-Ga(2)-P(2) 113.78(18), N(4)-Ga(2)-P(2) 117.50(19), N(3)-Ga(2)-Br(2) 106.04(19), N(4)-Ga(2)-Br(2) 102.13(18), P(2)-Ga(2)-Br(2) 116.34(13), C(59)-P(1)-P(2) 104.1(3), C(59)-P(1)-Ga(1) 98.9(3), P(2)-P(1)-Ga(1) 107.51(11), P(1)-P(2)-Ga(2) 99.43(11), N(5)-C(59)-P(1) 121.2(6), N(6)-C(59)-P(1) 131.9(6).

6. Computational Details

All quantum chemical calculations were performed using the program package ORCA 5.03. The structural optimizations were done using the B3LYP (12-15) functional and a def2-TZVP basis set⁹⁻¹⁶ while the subsequent spin and EPR analyses included relativistic effects using the Zero Order Regular Approximation (ZORA)¹⁷. The calculations include atom pairwise dispersion correction^{18,19} Natural Population Analysis was performed using the NBO7.0 Program (NBO7.0) in combination with ORCA4.2.1.²⁰

Bond/Angle	2a	2b	3 a	3b	4a	4b
P1=P2	2.040	2.040	2.158	2.158	2.130	2.133
Ga1–P1	2.332	2.333	2.325*	2.327*	2.358*	2.359*
Ga2–P2	2.334	2.333	2.280	2.280	2.354	2.357
P1-C59			1.852	1.852	1.816	1.817
Ga1–P1–P2	95.8	95.3	111.6*	111.0*	119.5*	118.1*
P1-P2-Ga2	94.9	95.3	93.7	94.0	99.6	100.3
Ga1-P1-C59			96.0	96.5	99.5	99.9
*carbene side		•	•	•	•	•

Table S4. Calculated (selected) bond distances (Å) and angles (°) of compounds 2a-b, 3a-b, and 4a-b.

Table S5. Mayer and Löwdin bond orders (in parenthesis) of compounds 2a-b, 3a-b, and 4a-b.

Bond	2a	2b	3a	3b	4 a	4b
P1=P2	1.8810	1.8830	1.2075	1.2193	1.2938	1.2534
	(2.0810)	(2.0849)	(1.4480)	(1.4461)	(1.4820)	(1.4746)
Ga1–P1	0.9573	0.9632	0.9295*	0.9346*	0.8369*	0.8718*
	(1.2032)	(1.2067)	(1.1291*)	(1.1275*)	(1.0189*)	(1.0194*)
Ga2–P2	0.9506	0.9628	1.1797	1.1647	0.9554	0.9781
	(1.2093)	(1.2066)	(1.4726)	(1.4728)	(1.1927)	(1.1920)
P1-C59			1.0712	1.0670	1.0013	1.0268
			(1.2002)	(1.1960)	(1.1979)	(1.2004)
*carbene side	;					

atom	2a	2b	3 a	3 b	4a	4b
P1	-0.31175	-0.30720	-0.09235*	-0.09188*	-0.04921*	-0.04625*
P2	-0.31816	-0.30720	-0.84128	-0.83266	-0.33600	-0.32474
Gal	1.41333	1.33854	1.45620*	1.39219*	1.48556*	1.41174*
Ga2	1.41224	1.33854	1.37116	1.30478	1.42213	1.34646
C (carbene)			0.16318	0.16437	0.16668	0.16319
N1 (carbene)			-0.31110	-0.31191	-0.30599	-0.30681
N2 (carbene)			-0.30965	-0.30810	-0.28370	-0.28485
*carbene side	2					

Table S6. Natural population analysis (NPA) atomic charges of compounds 2a-b, 3a-b, and 4a-b.

Table S7. FMO energies as well as the HOMO-LUMO gap of compounds 2a, 2b, 3a, and 3b calculated at M06-2X/def2-TZVPP//def2-SVP level of theory.

Orbital	Energy / eV				
	2a	2b	3 a	3b	
НОМО	-5.22	-5.26	-3.91	-3.95	
LUMO	-2.37	-2.44	-1.34	-1.37	
HOMO-LUMO gap	-7.60	-7.70	-5.25	-5.32	

Table S8. Calculated *g*-tensor and hyperfine coupling constants (A in MHz) for the radical cations **4a** and **4b** calculated at the ZORA B3LYP D3BJ def2-tzvp level of theory.

	4a				4b			
	x	У	Z	iso	x	У	Z	iso
g-tensor	1.997	2.004	2.019	2.006	2.000	2.002	2.016	2.006
<i>A</i> (P1)	263.1	274.7	426.9	321.6	263.2	275.1	421.3	319.9
<i>A</i> (P2)	-126.9	-145.8	594.7	107.4	-129.0	-147.8	598.8	107.3
A(Ga1)	83.2	84.2	97.4	88.3	84.1	84.8	97.5	88.8
<i>A</i> (Ga2)	-26.8	-43.0	-46.1	-38.6	-29.7	-45.7	-48.9	-41.4

Table S9. Mulliken spin populations of the radical cations 4a and 4b.

	4a	4b
P1	0.122	0.117
P2	0.797	0.804
Gal	0.016	0.015
Ga2	-0.006	-0.007
C59 _{NHC}	0.019	0.019
N1 _{NHC}	0.004	0.004
N2 _{NHC}	0.010	0.010

Figure S47. Selected molecular orbitals (from HOMO-3 to LUMO+3) of compound **2a** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

Figure S48. Selected molecular orbitals (from HOMO-3 to LUMO+3) of compound **2b** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

Figure S49. Selected molecular orbitals (from HOMO–3 to LUMO+3) of compound **3a** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

Figure S50. Selected molecular orbitals (from HOMO-3 to LUMO+3) of compound **3b** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

LUMO+1 (-0.1290 au)

Spin density

Figure S51. Selected molecular orbitals (from SOMO-3 to LUMO+1) and spin density plot of compound **4a** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

Figure S52. Selected molecular orbitals (from SOMO-3 to LUMO+1) and spin density plot of compound **4b** calculated at M06-2X/def2-TZVPP//def2-SVP. The isovalue was arbitrarily chosen to be 0.0432. Hydrogen atoms were omitted for clarity reasons.

6. EPR Spectroscopy

X-band CW EPR spectra were obtained using an X-band Bruker Elexsys E500 EPR spectrometer equipped with a ER4116DM dual mode resonator and an ESR 900 He cryostat. EPR spectra at 290 K were obtained using a microwave frequency of ~9.64 GHz, microwave power of 20 mW, 100 mT field sweep centered at 340 mT, a Lock-In modulation amplitude of 0.1 mT, a time constant of 40.96 ms, a sweep time of 83.89 s with 10 averages and a modulation frequency of 100 KHz. The temperature was stabilized at 80 K for low temperature measurements using a He flow cryostat (Oxford Instruments). The EPR spectra at 80 K were obtained under identical conditions using same parameters with an exception of 120 mT sweep width, 0.3 mT of modulation amplitude and 5 averages.

Q-band pulse EPR measurements were performed using a Bruker Elexsys E580 EPR X-band Spectrometer with homebuilt Q-band extension capable to deliver 34 GHz microwave pulses up to 10W. The experiments were conducted in a homebuilt Q-band Pulse-ENDOR resonator.²¹ Electron spin echo (ESE)-detected field-swept spectra were measured at 40 K using the pulse sequence: $t_p-\tau - 2t_p-\tau$ –echo with typical values $t_p=20$ ns, $\tau=240$ ns, repetition time 5ms.

W-band pulse EPR measurements were performed using a Bruker ELEXSYS E680 spectrometer equipped with a homebuilt W-band extension and a cryogen free Cryogenic 6 T magnet with a variable temperature insert. ESE-detected field-swept spectra were measured at 32 K using the pulse sequence: $t_{p}-\tau - 2t_{p}-\tau$ –echo with typical values tp=20 ns, τ =260 ns, repetition time 5ms. Small hyperfine couplings were detected using the Mims ENDOR sequence $t_{p}-\tau - t_{p}-T_{RF}-t_{p}-\tau$ –echo. With $T_{RF} = 20 \ \mu$ s. ELDOR detected NMR experiments aimed at uncovering large hyperfine couplings were performed at 32 K using the sequence T_{HTA} - $t_{p}-\tau$ –echo, with T_{HTA} (variable ELDOR frequency pulse) of 800 ns. However, the phosphorus couplings were larger than the bandwidth of the resonator, and no clear resonance signals could be observed.

All the simulations were obtained using Easyspin²² 5.2.30 with the function "pepper" and "garlic" for low temperature and room temperature spectra, respectively. Both complexes **4a** and **4b** showed identical X-band CW EPR spectra with a small variation in their line widths and therefore were simulated using identical spin Hamiltonian parameters. The negative signs were introduced to the hyperfine coupling values in spin-Hamiltonian parameters from the complementary computational calculations (Table S8) and to match the isotropic values obtained from liquid solution simulations. The discrepancy in the simulation parameters between RT and 80 K results from fact that the low temperature simulations require a very large number of parameters, which are strongly correlated. Therefore, it may be difficult to obtain a unique solution. Nevertheless, the spin-Hamiltonian parameters obtained from the simulations are in good agreement with the parameters calculated using computational calculations (Table S9).

		4a			4b				
		x	У	Z	iso	x	У	Ζ	iso
	DFT	1.997	2.004	2.019	2.006	2.000	2.002	2.016	2.006
g-tensor	Exp (80 K)				2.017				2.017
	Exp (290 K)				2.015				2.015
A(P 1)	DFT	263.1	274.7	426.9	321.6	263.2	275.1	421.3	319.9
/MHz	Exp (80 K)	229.9	255.9	352.0	279.3	229.9	255.9	352.0	279.3
/ 1011 12	Exp (290 K)				298.7				298.7
4(P2)	DFT	-126.9	-145.8	594.7	107.4	-129.0	-147.8	598.8	107.3
/MHz	Exp (80 K)	-55.3	-114.4	595.0	141.8	-55.3	-114.4	595.0	141.8
/ 101112	Exp (290 K)				192.7				192.7
4(Ga1)	DFT	83.2	84.2	97.4	88.3	84.1	84.8	97.5	88.8
/MHz	Exp (80 K)	72.2	72.6	104.3	83.1	72.2	72.6	104.3	83.1
/ 1011 12	Exp (290 K)				92.6				92.6
4(Ga2)	DFT	-26.8	-43.0	-46.1	-38.6	-29.7	-45.7	-48.9	-41.4
/MHz	Exp (80 K)	-21.2	-95.3	-98.1	-71.5	-21.2	-95.3	-98.1	-71.5
	Exp (290 K)				-59.7				-59.7

 Table S10. Spin Hamiltonian parameters obtained from the simulation of X-band EPR spectra and calculated from the computational calculations for the radicals 4a and 4b.

Table S11. P1 and P2 Mulliken Spin populations (p-orbitals) of the SOMO of compound 4a.

P1	pz	0.123207	
	px	-0.008936	0.085762
	ру	-0.028509	
P2	pz	0.327342	
	px	0.029450	0.767553
	ру	0.410761	

Figure S53. X-band CW (~9.64 GHz, black) EPR, Pseudo modulated Q-band (~34 GHz, green) and Pseudo modulated W-band (~94 GHz, blue) Hahn-Echo detected field sweep spectra of **4a** (1 mM 1:9 mixture of fluorobenzene and toluene) at 80 K, 40 K and 32 K, respectively.

Coordinates of all the optimized geometries:

2a			
Cl	-3.635127000	7.823076000	4.184875000
Ga	-2.276410000	8.750283000	5.690977000
Ν	-3.349031000	10.198185000	6.495120000
С	-3.427384000	11.369828000	5.881452000
С	-2.607843000	11.747786000	4.803085000
Н	-2.825186000	12.716003000	4.379407000
С	-1.462074000	11.118657000	4.304114000
С	-4.442660000	12.381861000	6.344145000
Η	-4.613163000	12.316870000	7.416298000
Н	-5.395635000	12.183137000	5.850001000
Η	-4.130576000	13.391108000	6.085278000
Η	-0.842238000	11.445955000	2.291314000
С	-4.080600000	9.902177000	7.698249000
С	-5.400554000	9.429132000	7.636893000
С	-6.025734000	9.061722000	8.829307000

Η	-7.037680000	8.680162000	8.798200000
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