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Supporting Information

Building Up "Genome" of Bi-atom Catalysts toward the Efficient HER/OER/ORR

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DFT-D3 method

Since GGA correlation functionals do not account for long-range correlations, the DFT-D method proposed by Grimme¹ is widely used to solve the dispersion problem of DFT. Compared with the DFT-D1 and DFT-D2, the DFT-D3 takes the the number of neighbors in the atomic environment into account, and further enhancing the accuracy of the results. Patel et al ² applied the DFT-D3 to all hybrid and GGA functional calculations, and found that the PBE-D3 can better describe the vdW interactions. Moreover, some literatures^{3,4} have performed comparisons between DFT-D3 correction and experiments, and validate the reliability of DFT-D3.

Gibbs free energy computations

With regard to HER, the overall reaction scheme can be described as follows⁵:

$$H^{+} + e^{-} + * \rightarrow H *$$

(S1)
 $H * + H^{+} + e^{-} \rightarrow H_{2}(g) + *$
(S2)

where * denotes the active site, and (g) denotes the gas phases. Under standard conditions (pH = 0, U = 0 V vs standard hydrogen electrode (SHE)), the Gibbs free energy change for hydrogen adsorption (ΔG_{H*}) can be calculated by: $\Delta G_{H*} = \Delta E_{H*} + \Delta E_{ZPE} - T\Delta S_H$ (S3)

where ΔE_{H*} stands for the hydrogen adsorption energy. ΔE_{ZPE} and ΔS_{H} are the zeropoint energy and entropy differences between gas phase and adsorbed state for hydrogen, respectively. *T* represents a temperature of 298.15 K herein.

The OER proceeds via four elementary reaction steps as follows⁶:

$$H_2 O(l) + * \to OH * + e^- + H^+$$
 (S4)

$$OH * \to O *+ e^- + H^+ \tag{S5}$$

$$H_2 O(l) + 0 * \xrightarrow{\Delta G_3} 00H * + e^- + H^+$$
 (S6)

$$00H * \to O_2(g) + e^- + H^+$$
(S7)

where (1) represents the liquid phases. Moreover, OH*, O*, and OOH* stand for the oxygenated intermediates.

The ORR can be treated as the inverse of ORR, and proceeds according following four elementary reaction steps⁷:

$$*+ O_2(g) + e^- + H^+ \xrightarrow{\Delta G_a} OOH$$
 (S8)

$$00H *+ e^{-} + H^{+} \xrightarrow{\Delta G_{b}} H_{2}O(l) + 0$$
(S9)

$$0 *+ e^{-} + H^{+} \xrightarrow{\Delta G_{c}} OH *$$
(S10)

$$OH *+ e^{-} + H^{+} \xrightarrow{\Delta G_d} H_2 O(l) +*$$
 (S11)

The Gibbs free energy change (ΔG) for the elementary steps can be calculated as follows:

$$\Delta G = \Delta E + \Delta Z P E - T \Delta S + \Delta G_U + \Delta G_{pH} + G_{solv}$$
(S12)

where ΔE , ΔZPE , and ΔS stand for the energy difference in adsorption energy, zeropoint energy, and entropy between the product and reactant of each elemental step, respectively. $\Delta G_U = -eU$, of which U is the applied electrode potential, and e stands for the number of transferred electrons. $\Delta G_{pH} = -k_B T ln 10 pH$, of which k_B means the Boltzmann constant. The ZPE and TS can be defined as follows:

$$ZPE = \frac{1}{2} \sum_{i} hv_{i}$$
(S13)
$$-TS = k_{B}T \sum_{i} ln(1 - e^{-hv_{i}/k_{B}T}) - \sum_{i} hv_{i}(\frac{1}{\frac{hv_{i}}{k_{B}T}})$$

where v_i is the vibration frequency of the ith vibrational mode, and h is Planck's constant. The ZPE and TS of the reaction intermediates can be obtained from the vibration analysis using the VASPKIT code⁸. The ZPE and TS are calculated for the adsorbates only, as the contribution of the surface to the ZPE and entropy remains

(S14)

almost constant under different adsorbed intermediates^{9,10}. Moreover, G_{solv} means a solvation correction to the reaction intermediates, which was estimated to be 0.30 eV.¹¹ The correction accounts for the solvation of the OOH* and OH* by the solution surrounding the metal site. The reason for this correction is the following: the enthalpy difference between the gas and liquid phases of H₂O is ~0.45 eV at 298.15 K and 0.035 bar, but their free energies are identical. This principle behind this phenomenon can be ascribed to the presence of hydrogen bonds in the liquid phase, in which each H₂O creates three hydrogen bonds (include two donors and one acceptor) with adjacent molecules. Since OH* is able to create two hydrogen bonds, it will be stabilized by ~0.3 eV when solvated, and we expected OOH* to be stabilized in the same order of magnitude due to solvation effects.¹²

Generally, pH effects on electrochemical systems can be divided into three aspects¹³: i) change in the proton donor or acceptor with the electrolyte pH, ii) affect adsorbate dipolefiled interactions, and iii) induce solution-phase reactions. Since the solvation-induced adsorption free energy change is primarily associated with the formation of H-bonds with H₂O under the aqueous environment, and contribution of H-bonds might not be changed significantly with the variations of pH values. To this end, we did not consider more details about the pH effect just as many literatures¹⁴⁻¹⁶ in terms of electrochemical catalysis in acidic conditions.

The overpotential (η) is adopted to evaluate the OER and ORR performance according to following equations:

$$\eta^{OER} = max\{\Delta G_1, \Delta G_2, \Delta G_3, \Delta G_4\}/e^{-1.23}$$
(S15)

$$\eta^{OER} = max\{\Delta G_a, \Delta G_b, \Delta G_c, \Delta G_d\}/e + 1.23$$
(S16)

Kinetic computations

The transition states (TS) are calculated by employing the climbing-image nudged elastic band (CI-NEB) method.¹⁷ There are eight images between the initial state (IS) and final state (FS).

The Pt calculations are carried on a 3×3 two layers FCC(111) slab separated by

three equivalent layers of vacuum. The bottom three layers are fixed, and the top two layers are allowed to relax.

Table S1. The energies (*E*) and bond lengths for M-M and M-N in M₂/g-CN (M=Sc, Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Y, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Hf, Ta, W, Re, Os, Ir, Pt, Au).

М		E (eV)		The bond	lengths (Å	Å)
	M-2N		M-3N	M-M	M-N1	M-N2	M-N3
Sc		-	-421.0724599	2.224	2.179	2.009	2.168
Ti		-	-424.5841186	1.975	2.187	2.031	2.148
V		-	-426.4234747	1.830	2.174	2.061	2.122
Cr		-	-425.8382477	2.078	2.098	2.007	2.071
Mn		-	-425.405069	2.395	2.056	1.906	2.005
Fe		-	-423.5734277	2.191	2.064	1.982	2.022
Co		-	-420.979397	2.253	2.032	1.935	1.998
Ni		-	-418.865846	2.403	2.011	1.888	1.978
Cu		-	-414.6877747	2.379	2.070	1.968	2.051
Zn	-410.6453	237	-	2.243	2.069	2.031	-
Y		-	-423.5839285	3.095	2.435	2.147	2.365
Zr		-	-424.1222809	2.225	2.200	2.047	2.174
Nb		-	-427.8507333	2.060	2.183	2.063	2.146
Mo		-	-428.2043986	2.129	2.165	2.127	2.127
Ru	-424.1746	241	-	2.193	-	1.988	1.950
Rh		-	-421.1317353	2.316	1.959	1.941	2.240
Pd		-	-417.7710435	2.407	2.077	1.979	2.055
Ag		-	-411.546058	2.633	2.406	2.234	2.420
Hf		-	-426.05604263	2.313	2.155	2.005	2.135

Au	-	-411.7909616	2.513	2.089	2.072	2.075
Pt	-	-420.0054452	2.482	2.044	1.996	2.024
Ir	-	-423.458621	2.345	1.952	1.970	2.155
Os	-	-427.335763	2.250	2.057	2.040	2.053
Re	-	-430.5398592	2.201	2.090	2.092	2.061
W	-	-431.2287911	2.271	2.128	2.108	2.103
Та	-	-429.3836541	2.157	2.156	2.023	2.123

Table S2. The computed formation energy $(E_{\rm f})$, dissolution potential $(U_{\rm diss})$ of metals, the number of transferred electrons $(N_{\rm e})$ during the dissolution, and the standard dissolution potential $(U_{\rm diss}^0)$ of metal atoms.

М	$U_{\rm diss}^0$ (V)	N_{e}	$E_{\rm f}({\rm eV})$	$U_{ m diss}({ m eV})$
Sc	-2.08	3	-4	-0.48
Ti	-1.63	2	-6	5.27 1.50
V	-1.18	2	-6	2.03
Cr	-0.91	2	-6	5.92 2.55
Mn	-1.19	2	-3	0.76
Fe	-0.45	2	-5	5.20 2.15
Co	-0.28	2	-5	2.58
Ni	-0.26	2	-5	2.51
Cu	0.34	2	-3	2.06
Zn	-0.76	2	-1	.65 0.06
Y	-2.37	3	-6	-0.37
Zr	-1.45	4	-6	0.11
Nb	-1.10	3	-8	1.69
Мо	-0.20	3	-10	0.78 3.39
Ru	0.46	2	-6	5.26 3.59
Rh	0.60	2	-5	5.79 3.49

Pd	0.95	2	-3.76	2.83
Ag	0.80	1	-1.91	2.71
Hf	-1.55	4	-5.89	-0.08
Та	-0.6	3	-7.58	1.93
W	0.1	3	-7.38	2.56
Re	0.3	3	-7.00	2.63
Os	0.84	8	-7.14	1.73
Ir	1.16	3	-6.95	3.48
Pt	1.18	2	-5.86	4.11
Au	1.50	3	-2.05	2.18

Table S3. The adsorption energies (E_{ad}) for H on possible sites in M₂/g-CN, the corresponding M-H bond lengths, the Gibbs free energy change for H adsorption (ΔG_{H^*}) , and the charge transfer between H* and the M₂ atoms.

М	$E_{\rm ad}$ -T _I (eV)	$E_{\rm ad}$ -B _{II} (eV)	The bond $\Delta \mathbf{G}$	G _{H*} (eV)	$Q_e\left(e ight)$
			length for		
			M-H (Å)		
Ti	-	-1.01	1.859	-0.83	0.42
V	-	-0.72	1.808	-0.53	0.46
Cr	-	-1.56	1.742	-1.37	0.45
Mn	-	-1.05	1.676	-0.84	0.39
Fe	-	-0.82	1.603,1.667	-0.61	0.29
Co	-	-1.10	1.582	-0.87	0.27
Ni	-	-0.26	1.584	-0.05	0.22
Cu	-	-0.36	1.599	-0.16	0.29
Zn	-	-0.23	1.647	-0.03	0.30
Zr	-	-2.00	2.072	-1.85	0.63
Nb	-	-0.78	1.970	-0.62	0.48

Mo	-0.49	-	1.738	-0.35 0.32
Ru	-	-0.68	1.760	-0.47 0.21
Rh	-	-1.10	1.710	-0.88 0.18
Pd	-	-0.20	1.678	0.03 0.08
Ag	-	-0.63	1.747	-0.45 0.19
Та	-	-1.63	1.937	-1.44 0.54
W	-0.52	-	1.739	-0.37 0.34
Re	-0.62	-	1.669	-0.44 0.20
Os	-	-0.63	1.838,1.741	-0.44 0.12
Ir	-	-1.22	1.733	-1.00 0.14
Pt	-	-0.38	1.695	-0.15 0.04
Au	-	-1.15	1.727	-0.94 0.07

Table S4. The adsorption energies (E_{ad}) for I^* (I=O, OH, OOH) on possible sites in M_2 /g-CN, and the corresponding M-I bond lengths.

Μ	E_{ad} -T	_I - <i>I</i> * (e	V)	E_{ad} -B	11- <i>I</i> * (e'	V)	The bond	The bond length for M - I (Å)			
	I=O	<i>I</i> =	<i>I</i> =00	<i>I</i> =0	I=O	<i>I</i> =00	<i>I</i> =OH	I=O	<i>I</i> =00H		
	Н	0	Н	Н		Н					
Ti	-	-	-	-	-	-4.97	1.992	1.828	1.817		
				2.46	3.08						
V	-	-	-	-	-	-2.57	2.003	1.796	1.711,1.8		
				1.63	1.52				57		
Cr	-	-	-	-	-	-2.10	1.953	1.780	1.784		
				2.27	1.92						
М	-	-	-	-	-	-1.23	1.930	1.778	1.776,1.7		
n				1.52	1.39				92		
Fe	-	-	-	-	-	1.99	1.962	1.752	1.757		
				0.88	0.45						

Co	-	-	-	-	-	1.84	1.855	1.720	1.792
				1.28	0.40				
Ni	-	-	-	0.20	1.12	3.38	1.903	1.751	1.875
Cu	-	-	-	-	1.26	3.02	1.917,1.8	1.787	1.876
				0.21			60		
Zn	-	-	3.56	-	0.35	-	1.942	1.804	1.955
				0.43					
Zr	-	-	-	-	-	-5.78	2.165	1.928,1.9	1.955
				3.72	4.30			41	
N	-	-		-	-	-3.01	2.144	1.953	2.423,1.7
b				1.73	1.78				69
Μ	-	-	1.83	-	-	-	1.922	1.854,2.0	1.899
0	1.30				1.10			67	
Ru	-	-	-	-	0.12	2.40	2.044,2.0	1.874,1.9	1.938,1.8
				0.53			76	23	84
Rh	-	-	-	-	0.33	2.34	2.010	1.881	1.970
				0.77					
Pd	-	-	-	1.07	1.94	4.21	2.040	1.914	2.145,2.2
									82
A	-	-	3.12	-	2.08		2.437,2.0	1.995	2.043
g				0.12			53		
Та	-	-	-	-	-	-4.77	2.129	1.957	2.125,1.8
				2.71	3.21				30
W	-	-	-2.64	-	-	-	1.910	1.864,2.0	1.745
	1.49				1.35			86	
Re	-	-	-	-	-	-0.32	1.929	1.952,1.9	2.029,1.8
	0.74				0.26			72	80
Os	-	-	-	-	0.37	0.51	2.124,2.0	1.966,1.9	1.977,1.8
				0.01			51	15	91

Ir	-	-	-	-	0.18	2.47	2.022	1.908	1.973
				0.63					
Pt	-	-	2.49	1.29	1.78		2.057	1.945	2.507
А	-	-	3.34	-	-	-	1.953	1.961	1.964
u	0.12				1.08				

Table S5. The Gibbs free energy change for oxygenated intermediate adsorption $(\Delta G_{O^*}, \Delta G_{OH^*}, \Delta G_{OOH^*})$ on M₂/g-CN, the overpotentials for OER (η^{OER}), and the overpotentials for ORR (η^{ORR}).

М	ΔG_{OH^*}	ΔG_{OH^*}	$\Delta G_{O^{\ast}}$		ΔG_{OOH*}	η^{OER}	η^{ORR}
	(eV)	under	(eV)	ΔG_{OOH^*}	(eV)	(V)	(V)
		solvation		(eV)	under		
		correction			solvation		
		(eV)			correction		
Ti	-2.16	-2.46	-3.01	-4.64	-4.94	8.33	3.39
V	-1.29	-1.59	-1.47	-2.22	-2.52	5.92	2.52
Cr	-1.92	-2.22	-1.87	-1.75	-2.05	5.44	3.15
Mn	-1.18	-1.48	1.32	-0.87	-1.17	4.56	2.41
Fe	-0.54	-0.84	-0.38	2.28	1.98	1.43	1.77
Co	-0.90	-1.20	-0.33	2.20	1.90	1.49	2.13
Ni	0.56	0.26	1.20	3.74	3.44	1.31	0.67
Cu	0.14	-0.16	1.31	3.37	3.07	0.82	1.09
Zn	-0.07	-0.307	0.41	3.92	3.62	2.28	1.30
Zr	-3.42	-3.72	-4.23	-5.48	-5.78	9.17	4.65
Nb	-1.42	-1.72	-1.73	-2.68	-2.98	6.37	2.65
Mo	-0.99	-1.29	-1.05	2.16	1.86	1.97	2.22
Ru	-0.17	-0.47	0.18	2.72	2.42	1.31	1.40
Rh	-0.39	-0.69	0.40	2.70	2.40	1.07	1.62
Pd	1.44	1.14	2.00	4.54	4.24	1.31	0.85

Ag	0.17	-0.13	2.11	3.43	3.13	0.71	1.06
Та	-2.38	-2.68	-3.16	-4.40	-4.70	8.09	3.61
W	-1.18	-1.48	-1.30	-2.26	-2.56	5.95	2.41
Re	-0.42	-0.72	-0.19	0.09	-0.21	3.60	1.65
Os	0.34	0.04	0.43	0.91	0.61	2.78	1.14
Ir	-0.25	-0.55	0.25	2.83	2.53	1.34	1.48
Pt	1.67	1.37	1.85	5.03	4.73	1.95	1.34
Au	0.18	-0.12	1.14	3.65	3.35	1.29	1.05

Table S6. The bond lengths for M_I - M_{II} , M_I -N, and M_{II} -N in M_IM_{II}/g -CN (M_I = Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; M_{II} = V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; $M_I \neq M_{II}$).

M _I M	II/g-CN			The bond lengths (Å)						
M_{I}	M_{II}	M_{I} - M_{II}	M _I -N1	M _I -N2	M _I -N3	M _{II} -N1	M _{II} -N2	M _{II} -N3		
Ti	Fe	1.958	2.207	2.177	2.164	2.046	2.020	2.092		
Ti	Co	2.036	2.175	2.159	2.131	2.028	1.995	2.069		
Ti	Ni	2.233	2.118	2.076	2.085	2.023	1.943	2.049		
Ti	Cu	2.435	2.101	2.008	2.090	2.046	1.931	2.043		
Ti	Pd	2.281	2.129	2.00	2.100	2.149	2.137	2.174		
Ti	Pt	2.323	2.131	2.048	2.086	2.077	2.103	2.098		
V	Ti	1.915	2.145	2.002	2.100	2.155	2.080	2.198		
V	Fe	1.931	2.151	2.151	2.100	2.044	2.058	2.097		
V	Co	2.110	2.130	2.084	2.099	2.027	1.981	2.063		
V	Ni	2.220	2.052	2.042	2.120	2.062	1.928	1.979		
V	Cu	2.455	2.085	1.960	2.083	2.023	1.930	2.046		
V	Pd	2.281	2.093	1.967	2.100	2.125	2.100	2.123		
V	Pt	2.320	2.086	2.011	2.108	2.076	2.121	2.063		

Cr	Ti	1.835	2.122	1.982	2.071	2.181	2.142	2.228
Cr	V	1.877	2.148	2.063	2.098	2.125	2.091	2.168
Cr	Mn	2.001	2.120	2.078	2.082	2.055	2.051	2.101
Cr	Fe	2.424	2.068	1.945	2.048	1.995	1.853	2.008
Cr	Co	2.217	2.029	2.015	2.159	2.112	1.926	1.945
Cr	Ni	2.343	2.084	1.986	2.061	1.987	1.899	2.026
Cr	Cu	2.435	2.059	1.936	2.049	2.007	1.916	2.031
Cr	Zn	2.401	2.076	1.931	2.052	2.126	1.918	2.120
Cr	Zr	1.975	2.089	1.923	2.053	2.222	2.183	2.258
Cr	Nb	1.915	2.116	1.958	2.072	2.167	2.125	2.216
Cr	Mo	2.017	2.131	2.039	2.088	2.119	2.137	2.163
Cr	Ru	2.089	2.122	2.023	2.086	2.102	2.106	2.129
Cr	Rh	2.302	2.106	1.961	2.108	2.097	2.028	2.096
Cr	Pd	2.387	2.113	1.955	2.082	2.098	2.098	2.125
Cr	Ag	2.461	2.075	1.909	2.062	2.136	2.131	2.156
Cr	Та	2.032	2.103	1.964	2.062	2.139	2.092	2.173
Cr	W	2.119	2.111	2.037	2.067	2.099	2.133	2.136
Cr	Re	2.107	2.121	2.046	2.079	2.064	2.125	2.097
Cr	Os	2.137	2.108	2.038	2.073	2.055	2.089	2.080
Cr	Ir	2.289	2.088	1.983	2.103	2.059	2.036	2.056
Cr	Pt	2.413	2.109	1.969	2.074	2.045	2.098	2.064
Cr	Au	2.466	2.036	1.871	2.004	2.071	2.150	2.093
Mn	Ti	1.911	2.111	2.010	2.071	2.172	2.159	2.216
Mn	V	2.000	2.124	2.055	2.082	2.114	2.102	2.157
Mn	Fe	2.291	2.071	1.980	2.039	2.002	1.927	2.038
Mn	Co	2.251	2.090	2.027	2.056	1.982	1.902	2.022
Mn	Ni	2.335	2.052	1.975	2.022	1.987	1.900	2.020
Mn	Cu	2.407	2.044	1.949	2.020	2.035	1.939	2.053
Mn	Pd	2.348	2.058	1.943	2.026	2.084	2.073	2.108

Mn	Pt	2.366	2.057	1.958	2.023	2.042	2.060	2.059
Fe	Co	2.206	2.064	2.003	2.021	1.991	1.927	2.032
Fe	Ni	2.327	2.033	1.946	2.005	1.989	1.898	2.015
Fe	Cu	2.361	2.052	1.952	2.028	2.050	1.950	2.062
Fe	Pd	2.315	2.026	1.911	2.006	2.072	2.033	2.089
Fe	Pt	2.343	2.023	1.927	1.997	2.040	2.034	2.055
Co	Ni	2.397	2.007	1.872	1.973	1.984	1.907	2.017
Cu	Co	2.456	2.038	1.947	2.003	1.964	1.808	1.973
Cu	Ni	2.389	2.076	1.963	2.043	2.003	1.874	2.024
Cu	Pd	2.409	2.069	1.923	2.043	2.074	2.016	2.088
Cu	Pt	2.425	2.051	1.917	2.020	2.041	1.986	2.049
Zn	Ti	2.556	2.344	2.137	-	2.245	2.026	2.108
Zn	V	2.403	2.103	1.906	2.153	2.089	1.945	2.069
Zn	Mn	2.369	2.209	1.987	2.245	2.038	1.944	2.060
Zn	Fe	2.309	2.093	1.984	2.246	2.105	1.898	1.966
Zn	Co	2.315	2.168	2.023	2.158	1.987	1.836	2.011
Zn	Ni	2.265	2.182	2.036	2.177	1.992	1.866	2.043
Zn	Cu	2.320	2.150	1.977	2.144	2.048	1.921	2.073
Zn	Pd	2.335	2.222	1.977	2.230	2.097	2.053	2.125
Zn	Pt	2.351	2.231	1.994	2.221	2.052	2.048	2.070
Zr	Ti	2.111	2.221	2.090	2.188	2.122	1.972	2.153
Zr	V	2.042	2.236	2.133	2.200	2.083	1.945	2.116
Zr	Mn	1.968	2.255	2.180	2.216	2.038	1.883	2.068
Zr	Fe	2.161	2.283	2.252	2.255	2.033	1.954	2.062
Zr	Co	2.182	2.257	2.242	2.230	2.021	1.935	2.050
Zr	Ni	2.325	2.209	2.188	2.187	2.025	1.921	2.047
Zr	Cu	2.498	2.209	2.211	2.376	-	2.034	2.074
Zr	Zn	2.722	2.266	2.252	2.399	-	-	-
Zr	Nb	2.151	2.219	2.087	2.186	2.130	2.018	2.163

Zr	Rh	2.284	2.273	2.200	2.258	2.129	2.086	2.141
Zr	Pd	2.393	2.222	2.136	2.202	2.139	2.116	2.158
Zr	Ag	2.677	2.277	2.257	2.415	-	-	-
Zr	Та	2.190	2.211	2.074	2.181	2.116	2.003	2.146
Zr	Ir	2.359	2.270	2.203	2.250	2.064	2.065	2.080
Zr	Pt	2.444	2.232	2.151	2.205	2.073	2.106	2.091
Zr	Au	2.577	2.205	2.082	2.183	2.095	2.148	2.102
Nb	Ti	2.031	2.181	2.061	2.143	2.136	2.024	2.174
Nb	V	1.941	2.204	2.110	2.162	2.102	2.005	2.145
Nb	Mn	2.054	2.209	2.188	2.170	2.065	2.006	2.103
Nb	Fe	2.100	2.190	2.194	2.155	2.044	1.987	2.078
Nb	Co	2.151	2.157	2.175	2.125	2.025	1.950	2.057
Nb	Ni	2.287	2.131	2.125	2.110	2.033	1.921	2.041
Nb	Cu	2.432	2.136	2.102	2.254	-	2.010	2.026
Nb	Zn	2.606	2.136	2.111	2.233	-	2.164	-
Nb	Rh	2.229	2.161	2.118	2.130	2.115	2.068	2.139
Nb	Pd	2.354	2.132	2.067	2.120	2.140	2.100	2.146
Nb	Pt	2.392	2.135	2.081	2.110	2.075	2.081	2.092
Mo	Ti	1.962	2.154	2.042	2.112	2.153	2.071	2.194
Mo	V	1.926	2.171	2.064	2.127	2.103	2.015	2.154
Mo	Mn	2.080	2.148	2.156	2.113	2.073	2.015	2.108
Mo	Fe	2.070	2.134	2.157	2.102	2.029	1.988	2.072
Mo	Co	2.146	2.114	2.130	2.091	2.027	1.944	2.044
Mo	Ni	2.235	2.032	2.060	2.165	2.177	1.893	1.912
Mo	Cu	2.495	2.112	1.990	2.115	2.046	1.931	2.064
Mo	Zn	2.487	2.028	2.038	-	-	2.031	2.030
Mo	Zr	2.086	2.132	1.998	2.098	2.200	2.125	2.235
Mo	Nb	2.032	2.149	2.024	2.110	2.146	2.074	2.189
Mo	Ru	2.162	2.131	2.103	2.100	2.099	2.072	2.124

Mo	Rh	2.221	2.114	2.065	2.092	2.111	2.053	2.125
Mo	Pd	2.439	2.138	1.996	2.127	2.131	2.071	2.144
Mo	Ag	2.670	2.128	1.991	2.232	-	2.392	-
Mo	Ta	2.089	2.140	2.006	2.103	2.130	2.039	2.166
Mo	W	2.223	2.154	2.127	2.131	2.105	2.120	2.131
Mo	Re	2.201	2.147	2.125	2.112	2.065	2.094	2.099
Mo	Os	2.222	2.124	2.108	2.092	2.058	2.066	2.083
Mo	Ir	2.271	2.109	2.078	2.083	2.063	2.048	2.081
Mo	Pt	2.455	2.138	2.011	2.111	2.075	2.044	2.097
Mo	Au	2.560	2.137	1.982	2.115	2.084	2.136	2.097
Ru	Ti	2.051	2.118	2.091	2.094	2.157	2.113	2.196
Ru	V	2.050	2.132	2.125	2.103	2.098	2.073	2.139
Ru	Mn	2.275	2.101	2.004	2.090	2.082	1.973	2.110
Ru	Fe	2.146	2.003	2.028	2.221	2.234	1.921	1.916
Ru	Co	2.205	1.991	2.005	2.219	2.228	1.873	1.887
Ru	Ni	2.311	2.030	1.998	2.096	2.050	1.901	1.981
Ru	Cu	2.484	2.050	1.904	2.058	2.040	1.948	2.079
Ru	Zn	2.454	2.065	1.903	2.123	-	2.020	2.189
Ru	Zr	2.264	2.103	2.023	2.114	2.273	2.234	2.285
Ru	Nb	2.170	2.123	2.080	2.103	2.167	2.159	2.197
Ru	Rh	2.333	1.994	1.937	2.214	2.222	1.959	2.009
Ru	Pd	2.496	2.103	1.910	2.082	2.116	2.048	2.145
Ru	Ag	2.590	2.080	1.875	2.104	2.537	2.285	2.518
Ru	Ta	2.230	2.112	2.065	2.083	2.138	2.143	2.169
Ru	Ir	2.432	2.059	1.944	2.120	2.080	1.966	2.060
Ru	Pt	2.502	2.096	1.920	2.068	2.068	2.003	2.090
Ru	Au	2.535	2.070	1.906	2.047	2.076	2.134	2.098
Rh	Ti	2.115	2.138	2.117	2.114	2.132	2.089	2.170
Rh	V	2.123	2.140	2.107	2.116	2.089	2.033	2.122

Rh	Mn	2.312	2.095	2.015	2.075	2.071	1.996	2.114
Rh	Fe	2.256	2.068	2.017	2.065	2.022	1.966	2.043
Rh	Co	2.295	2.023	1.991	2.121	2.082	1.903	1.964
Rh	Ni	2.391	2.055	1.959	2.041	1.982	1.885	2.031
Rh	Cu	2.446	2.041	1.917	2.042	2.006	1.921	2.042
Rh	Zn	2.378	2.046	1.925	2.055	2.142	1.944	2.127
Rh	Pd	2.475	2.086	1.937	2.060	2.093	2.046	2.126
Rh	Pt	2.494	2.087	1.952	2.052	2.049	2.029	2.072
Pd	Co	2.364	2.098	2.041	2.070	1.973	1.846	2.009
Pd	Ni	2.374	2.088	2.013	2.067	1.982	1.858	2.020
Pd	Pt	2.464	2.086	2.006	2.062	2.036	1.993	2.052
Ag	Ti	2.594	-	-	-	2.273	2.053	2.195
Ag	V	2.604	2.478	2.256	-	2.176	1.946	2.123
Ag	Mn	2.395	2.117	2.070	2.096	2.001	1.860	2.016
Ag	Fe	2.397	2.128	2.089	2.100	1.986	1.818	1.998
Ag	Co	2.389	2.137	2.106	2.107	1.967	1.782	1.979
Ag	Ni	2.360	2.157	2.110	2.133	2.007	1.825	2.024
Ag	Cu	2.386	2.160	2.174	2.150	2.056	1.909	2.059
Ag	Zn	2.367	2.193	2.122	2.170	2.157	1.916	2.140
Ag	Nb	2.638	-	-	-	2.257	2.099	2.181
Ag	Rh	2.432	2.123	2.058	2.093	2.033	1.889	2.036
Ag	Pd	2.474	2.181	2.117	2.159	2.078	1.992	2.087
Ag	Pt	2.486	2.143	2.090	2.119	2.039	1.969	2.046
Ta	Ti	2.078	2.160	2.037	2.126	2.130	2.012	2.166
Ta	V	2.003	2.179	2.072	2.139	2.092	1.986	2.136
Ta	Mn	2.180	2.183	2.160	2.151	2.054	1.995	2.090
Ta	Fe	2.169	2.157	2.168	2.125	2.025	1.977	2.064
Ta	Co	2.209	2.132	2.150	2.102	2.012	1.940	2.050
Та	Ni	2.301	2.110	2.110	2.082	2.013	1.918	2.039

Ta	Cu	2.411	2.118	2.073	2.214	-	2.031	2.029
Та	Zn	2.557	2.094	2.084	2.243	-	2.083	2.173
Та	Nb	2.111	2.163	2.035	2.130	2.139	2.044	2.175
Та	Rh	2.276	2.135	2.103	2.106	2.100	2.058	2.128
Та	Pd	2.368	2.114	2.062	2.092	2.130	2.102	2.144
Та	Ag	2.622	2.157	2.070	2.217	-	-	-
Та	Pt	2.402	2.116	2.080	2.087	2.067	2.080	2.090
W	Ti	2.025	2.129	2.011	2.094	2.144	2.050	2.184
W	V	1.994	2.139	2.021	2.104	2.094	1.994	2.144
W	Mn	2.150	2.112	2.133	2.079	2.057	2.011	2.096
W	Fe	2.162	2.099	2.125	2.068	2.022	1.981	2.063
W	Co	2.189	2.083	2.113	2.056	2.009	1.948	2.043
W	Ni	2.257	2.038	2.076	2.054	2.049	1.913	1.974
W	Cu	2.441	2.062	1.986	2.063	2.059	1.927	2.068
W	Zn	2.482	2.016	2.024	-	-	2.023	2.019
W	Zr	2.139	2.114	1.980	2.083	2.192	2.103	2.225
W	Nb	2.091	2.126	1.996	2.092	2.140	2.053	2.178
W	Ru	2.215	2.095	2.093	2.066	2.085	2.065	2.114
W	Rh	2.259	2.086	2.057	2.056	2.095	2.052	2.117
W	Pd	2.353	2.044	2.013	2.077	2.161	2.093	2.091
W	Ag	2.609	2.104	2.002	2.176	-	-	-
W	Та	2.142	2.117	1.986	2.085	2.123	2.027	2.156
W	Re	2.242	2.119	2.112	2.087	2.066	2.087	2.092
W	Os	2.257	2.093	2.097	2.063	2.055	2.063	2.082
W	Ir	2.297	2.076	2.075	2.050	2.054	2.051	2.074
W	Pt	2.378	2.051	2.040	2.056	2.079	2.077	2.065
W	Au	2.545	2.091	1.974	2.073	2.085	2.122	2.094
Re	Ti	2.032	2.102	2.008	2.074	2.141	2.066	2.182
Re	V	2.095	2.104	2.116	2.072	2.096	2.084	2.145

Re	Mn	2.141	2.089	2.094	2.063	2.049	2.016	2.091
Re	Fe	2.105	2.076	2.106	2.053	2.023	2.003	2.061
Re	Co	2.156	2.020	2.061	2.115	2.179	1.938	1.916
Re	Ni	2.308	2.048	2.028	2.044	2.003	1.909	2.015
Re	Cu	2.436	2.051	1.963	2.046	2.043	1.930	2.063
Re	Zn	2.336	1.956	1.971	-	-	1.908	1.899
Re	Zr	2.136	2.090	1.955	2.066	2.186	2.098	2.222
Re	Nb	2.253	2.093	2.082	2.076	2.189	2.174	2.213
Re	Ru	2.210	2.080	2.053	2.058	2.089	2.056	2.111
Re	Rh	2.246	1.985	1.994	2.153	2.247	2.024	1.977
Re	Pd	2.423	2.078	1.977	2.061	2.108	2.075	2.134
Re	Ag	2.587	2.074	1.934	2.108	-	2.315	-
Re	Та	2.287	2.093	2.073	2.072	2.159	2.151	2.190
Re	Os	2.227	2.076	2.073	2.053	2.053	2.063	2.077
Re	Ir	2.278	2.055	2.039	2.054	2.065	2.039	2.063
Re	Pt	2.446	2.076	1.993	2.058	2.062	2.051	2.084
Re	Au	2.505	2.054	1.963	2.039	2.069	2.100	2.085
Os	Ti	2.156	2.074	2.077	2.051	2.151	2.113	2.188
Os	V	2.122	2.083	2.108	2.057	2.080	2.080	2.125
Os	Mn	2.236	2.041	2.022	2.080	2.146	1.988	2.030
Os	Fe	2.164	1.999	2.032	2.129	2.206	1.958	1.918
Os	Co	2.211	1.990	2.005	2.138	2.211	1.904	1.891
Os	Ni	2.319	2.027	1.995	2.036	2.013	1.915	2.009
Os	Cu	2.471	2.032	1.907	2.036	2.039	1.948	2.088
Os	Zn	2.432	1.968	1.924	-	-	2.024	2.031
Os	Zr	2.349	2.071	2.026	2.072	2.276	2.230	2.290
Os	Nb	2.246	2.080	2.073	2.058	2.158	2.160	2.189
Os	Ru	2.272	2.045	2.001	2.090	2.140	2.010	2.057
Os	Rh	2.348	1.996	1.942	2.138	2.219	1.979	2.007

Os	Pd	2.482	2.061	1.924	2.057	2.126	2.057	2.125
Os	Ag	2.585	2.058	1.883	2.073	2.518	2.280	2.507
Os	Та	2.277	2.079	2.062	2.052	2.133	2.147	2.159
Os	Ir	2.368	2.000	1.963	2.125	2.129	1.980	2.013
Os	Pt	2.497	2.069	1.934	2.045	2.063	2.008	2.091
Os	Au	2.534	2.042	1.911	2.026	2.078	2.118	2.099
Ir	Ti	2.199	2.078	2.096	2.057	2.123	2.099	2.159
Ir	V	2.189	2.081	2.087	2.062	2.072	2.045	2.113
Ir	Mn	2.330	2.053	2.005	2.035	2.063	1.995	2.103
Ir	Fe	2.282	2.042	2.014	2.031	2.006	1.975	2.046
Ir	Co	2.294	2.026	2.004	2.046	2.025	1.932	2.000
Ir	Ni	2.403	2.031	1.961	2.019	1.977	1.889	2.027
Ir	Cu	2.449	2.023	1.923	2.023	2.000	1.915	2.041
Ir	Zn	2.386	2.028	1.935	2.029	2.138	1.949	2.133
Ir	Nb	2.290	2.086	2.059	2.066	2.123	2.130	2.155
Ir	Rh	2.330	1.960	1.942	2.167	2.242	1.963	1.948
Ir	Pd	2.484	2.053	1.939	2.033	2.087	2.045	2.123
Ir	Ag	2.497	2.032	1.907	2.026	2.119	2.103	2.147
Ir	Ta	2.317	2.084	2.053	2.062	2.100	2.115	2.128
Ir	Pt	2.505	2.051	1.953	2.026	2.044	2.029	2.069
Pt	Co	2.388	2.053	2.031	2.035	1.967	1.866	2.004
Pt	Ni	2.399	2.050	2.007	2.035	1.971	1.867	2.015
Au	Ti	2.483	2.095	2.150	2.082	2.085	1.978	2.107
Au	V	2.407	2.075	2.082	2.059	2.038	1.939	2.051
Au	Mn	2.410	2.074	2.074	2.056	1.991	1.885	2.016
Au	Fe	2.414	2.076	2.086	2.058	1.973	1.845	1.998
Au	Co	2.412	2.087	2.101	2.059	1.953	1.803	1.978
Au	Ni	2.403	2.082	2.073	2.062	1.979	1.837	2.008
Au	Cu	2.419	2.089	2.171	2.079	2.052	1.920	2.062

Au	Zn	2.338	2.101	2.066	2.075	2.148	1.929	2.137
Au	Nb	2.519	2.096	2.136	2.109	2.113	2.028	2.113
Au	Rh 2	2.464	2.072	2.061	2.048	2.014	1.906	2.029
Au	Pd 2	2.450	2.069	2.035	2.047	2.041	1.949	2.058
Au	Ag	2.604	2.141	2.209	2.126	2.477	2.213	2.491
Au	Ta 2	2.512	2.096	2.129	2.099	2.079	2.020	2.095
Au	Ir 2	2.483	2.066	2.048	2.042	2.001	1.919	2.013
Au	Pt 2	2.475	2.062	2.038	2.046	2.022	1.954	2.035

Table S7. The adsorption energies (E_{ad}) for H on possible sites in M_IM_{II}/g-CN (M_I = Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; M_{II} = V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; M_I \neq M_{II}), the corresponding M_I-H and M_{II}-H bond lengths, and Δ G_{H*}.

$M_I M_I$	_I /g-	The	bond	length	for	The	bond	length	for	$\Delta G_{\mathrm{H}^{\ast}}$	$E_{\rm ad}$
CN		M _I -H	I (Å)			M _{II} -I	M_{II} -H (Å)			(eV)	(eV)
M_{I}	M_{II}										
Ti	Fe	1.868	8			1.706	5			-0.35	-0.53
Ti	Co	1.869	9			1.63	5			-0.19	-0.38
Ti	Ni	1.83	1			1.606	5			-0.49	-0.69
Ti	Cu	1.908	8			1.573	3			-0.92	-1.11
Ti	Pd	1.724	4			-				-0.26	-0.38
Ti	Pt	1.930	0			1.644	1			-0.28	-0.50
V	Ti	1.759	9			1.913	3			-0.74	-0.92
V	Fe	1.765	5			1.756	5			-0.50	-0.68
V	Co	1.80	1			1.558	3			-0.49	-0.71
V	Ni	1.777	7			1.59	l			-0.41	-0.62
V	Cu	1.816	6			1.59	l			-0.55	-0.55
V	Pd	1.646	6			-				-0.15	-0.35

V	Pt	1.869	1.636	-0.42	-0.65
Cr	Ti	1.724	1.880	-0.61	-0.80
Cr	V	1.745	1.762	-0.68	-0.86
Cr	Mn	1.746	1.683	-0.92	-1.12
Cr	Fe	1.735	1.616	-0.80	-1.01
Cr	Co	1.760	1.554	-0.92	-1.14
Cr	Ni	1.745	1.567	-0.50	-0.71
Cr	Cu	1.770	1.593	-0.44	-0.64
Cr	Zn	1.788	1.609	-0.83	-1.03
Cr	Zr	1.706	-	-0.62	-0.80
Cr	Nb	1.745	1.895	-0.66	-0.85
Cr	Mo	1.812	1.785	-0.58	-0.78
Cr	Ru	1.911	1.681	-0.68	-0.88
Cr	Rh	1.841	1.649	-0.96	-1.18
Cr	Pd	1.778	1.639	-0.16	-0.38
Cr	Ag	1.874	1.697	-0.54	-0.71
Cr	Ta	1.789	1.812	-1.04	-1.25
Cr	W	1.860	1.740	-0.73	-0.93
Cr	Re	-	1.676	-0.28	-0.47
Cr	Os	-	1.667	-0.86	-1.06
Cr	Ir	1.893	1.656	-1.14	-1.36
Cr	Pt	1.825	1.627	-0.53	-0.77
Cr	Au	-	1.584	-1.74	-1.94
Mn	Ti	1.658	1.874	-0.78	-0.98
Mn	V	1.752	1.779	-0.48	-0.67
Mn	Fe	1.714	1.605	-0.82	-1.04
Mn	Co	1.740	1.546	-0.69	-0.91
Mn	Ni	1.690	1.588	-0.41	-0.61
Mn	Cu	1.691	1.613	-0.38	-0.58

Mn	Pd	1.683	1.699	-0.20	-0.40
Mn	Pt	1.746	1.642	-0.33	-0.55
Fe	Co	1.653	1.565	-0.62	-0.84
Fe	Ni	1.621	1.588	-0.39	-0.61
Fe	Cu	1.632	1.622	-0.39	-0.60
Fe	Pd	1.616	1.689	-0.24	-0.46
Fe	Pt	1.657	1.656	-0.27	-0.50
Co	Ni	1.571	1.602	-0.77	-0.99
Cu	Co	1.632	1.575	-0.37	-0.59
Cu	Ni	1.603	1.575	-0.16	-0.37
Cu	Pd	1.621	1.655	-0.02	-0.24
Cu	Pt	1.666	1.652	-0.29	-0.52
Zn	Ti	1.587	-	-0.46	-0.65
Zn	V	1.591	-	-1.26	-1.46
Zn	Mn	1.621	1.779	-0.35	-0.54
Zn	Fe	1.658	1.622	-0.68	-0.89
Zn	Co	1.672	1.559	-0.85	-1.07
Zn	Ni	1.665	1.574	-0.26	-0.47
Zn	Cu	1.657	1.602	-0.43	-0.63
Zn	Pd	1.702	1.664	-0.06	-0.26
Zn	Pt	1.791	1.621	-0.31	-0.54
Zr	Ti	1.954	1.857	-1.05	-1.24
Zr	V	2.056	1.767	-0.88	-1.06
Zr	Mn	1.977	1.665	-1.07	-1.27
Zr	Fe	2.060	1.663	-0.32	-0.48
Zr	Co	2.063	1.627	-0.05	-0.23
Zr	Ni	1.875	-	-0.40	-0.54
Zr	Cu	2.042	1.581	-1.20	-1.39
Zr	Zn	-	1.569	0.24	0.10

Zr	Nb	-	1.889	-1.21	-1.37
Zr	Rh	-	1.648	0.26	0.08
Zr	Pd	1.882	-	-0.33	-0.54
Zr	Ag	1.847	-	0.36	-0.04
Zr	Ta	-	1.886	-1.65	-1.82
Zr	Ir	-	1.643	0.14	-0.08
Zr	Pt	2.050	1.655	-0.18	-0.40
Zr	Au	1.876	-	-0.44	-0.69
Nb	Ti	1.862	-	-0.85	-1.03
Nb	V	1.958	1.816	-0.49	-0.67
Nb	Mn	1.927	1.776	-0.45	-0.61
Nb	Fe	1.913	1.781	-0.32	-0.49
Nb	Co	1.935	1.684	-0.10	-0.28
Nb	Ni	1.769	-	-0.69	-0.85
Nb	Cu	1.909	1.613	-1.06	-1.25
Nb	Zn	1.729	-	-0.33	-0.51
Nb	Rh	1.789	-	-0.24	-0.38
Nb	Pd	1.779	-	-0.60	-0.76
Nb	Pt	1.911	1.705	-0.12	-0.33
Mo	Ti	1.841	-	-0.53	-0.70
Mo	V	1.887	1.770	-0.70	-0.89
Mo	Mn	1.793	-	-0.51	-0.67
Mo	Fe	1.809	1.859	-0.48	-0.65
Mo	Co	1.829	1.667	-0.31	-0.50
Mo	Ni	1.750	1.770	-0.58	-0.79
Mo	Cu	1.799	1.658	-0.61	-0.80
Mo	Zn	1.732	-	-0.24	-0.41
Mo	Zr	1.812	-	-0.79	-0.95
Mo	Nb	1.899	1.874	-0.63	-0.83

Mo	Ru	-	1.662	0.09	-0.08
Мо	Rh	1.699	-	-0.62	-0.79
Мо	Pd	1.686	-	-0.56	-0.74
Мо	Ag	1.677	-	-0.45	-0.63
Мо	Ta	1.935	1.815	-1.19	-1.40
Mo	W	-	1.736	-0.61	-0.81
Mo	Re	-	1.671	-0.31	-0.50
Mo	Os	-	1.650	0.05	-0.13
Mo	Ir	1.701	-	-0.62	-0.80
Mo	Pt	1.878	1.685	-0.25	-0.46
Мо	Au	1.701	-	-0.44	-0.60
Ru	Ti	1.628	-	-0.08	-0.26
Ru	V	1.644	-	0	-0.18
Ru	Mn	1.690	1.854	-0.92	-1.13
Ru	Fe	1.693	1.795	-0.93	-1.12
Ru	Co	1.705	1.643	-0.76	-0.97
Ru	Ni	1.660	1.694	-0.76	-0.97
Ru	Cu	1.634	-	-0.68	-0.87
Ru	Zn	1.731	1.740	-0.81	-1.03
Ru	Zr	1.618	-	-0.18	-0.38
Ru	Nb	1.619	-	0.06	-0.13
Ru	Rh	1.695	1.814	-0.36	-0.55
Ru	Pd	1.709	1.770	-0.30	-0.50
Ru	Ag	1.588	-	-0.37	-0.56
Ru	Ta	1.650	-	0.03	-0.16
Ru	Ir	1.783	1.736	-0.47	-0.67
Ru	Pt	1.697	1.741	-0.62	-0.84
Ru	Au	1.577	-	-0.53	-0.72
Rh	Ti	1.671	1.904	-0.59	-0.80

Rh	V	1.659	-	-0.51	-0.72
Rh	Mn	1.639	1.809	-0.80	-1.02
Rh	Fe	1.658	1.683	-0.64	-0.85
Rh	Co	1.673	1.618	-0.90	-1.13
Rh	Ni	1.658	1.626	-0.85	-1.07
Rh	Cu	1.670	1.666	-0.32	-0.53
Rh	Zn	1.655	1.743	-1.12	-1.33
Rh	Pd	1.653	1.732	-0.53	-0.74
Rh	Pt	1.693	1.694	-0.75	-0.98
Pd	Co	1.698	1.567	-0.66	-0.88
Pd	Ni	1.676	1.582	0	-0.22
Pd	Pt	1.731	1.657	-0.08	-0.30
Ag	Ti	1.719	-	-0.51	-0.68
Ag	V	1.717	-	-0.70	-0.87
Ag	Mn	1.696	1.906	-0.68	-0.84
Ag	Fe	1.732	1.683	-0.35	-0.54
Ag	Co	1.780	1.582	-0.20	-0.40
Ag	Ni	1.721	1.651	-0.20	-0.39
Ag	Cu	1.722	1.630	-0.29	-0.48
Ag	Zn	1.709	1.710	-0.62	-0.81
Ag	Nb	-	1.755	-0.05	-0.21
Ag	Rh	1.759	1.711	-0.38	-0.57
Ag	Pd	2.070	1.571	0.13	-0.09
Ag	Pt	1.766	1.673	0	-0.21
Та	Ti	1.850	-	-1.18	-1.36
Та	V	1.893	1.853	-0.96	-1.15
Ta	Mn	1.880	1.850	-0.67	-0.84
Ta	Fe	1.867	1.901	-0.51	-0.67
Ta	Co	1.881	1.789	-0.31	-0.47

Та	Ni	1.767	-	-0.96	-1.13
Та	Cu	1.900	1.609	-1.05	-1.24
Ta	Zn	1.720	-	-0.36	-0.56
Та	Nb	1.893	-	-1.08	-1.26
Ta	Rh	1.785	-	-0.50	-0.66
Ta	Pd	1.777	-	-0.94	-1.10
Та	Ag	1.759	-	-0.37	-0.54
Ta	Pt	1.925	1.694	-0.10	-0.30
W	Ti	1.817	-	-1.00	-1.19
W	V	1.846	1.802	-1.36	-1.56
W	Mn	1.743	1.822	-0.38	-0.57
W	Fe	1.712	-	-0.71	-0.88
W	Co	1.790	1.665	-0.43	-0.62
W	Ni	1.689	-	-0.92	-1.10
W	Cu	1.812	1.672	-0.79	-0.98
W	Zn	1.677	-	-0.56	-0.75
W	Zr	1.801	-	-1.25	-1.44
W	Nb	-	1.898	-1.31	-1.50
W	Ru	1.712	-	-0.41	-0.58
W	Rh	1.850	1.724	-0.31	-0.51
W	Pd	1.699	-	-0.92	-1.10
W	Ag	1.702	-	-0.79	-0.98
W	Ta	-	1.838	-1.66	-1.87
W	Re	-	1.718	-0.23	-0.40
W	Os	-	1.705	-0.05	-0.23
W	Ir	1.711	-	-0.90	-1.09
W	Pt	1.707	-	-0.97	-1.16
W	Au	1.696	-	-0.99	-1.17
Re	Ti	1.674	-	-0.67	-0.86

Re	V	1.674	-	-0.44	-0.64
Re	Mn	1.645	-	-0.44	-0.60
Re	Fe	1.720	-	-0.69	-0.84
Re	Co	1.745	1.710	-0.78	-0.97
Re	Ni	1.635	-	-1.01	-1.21
Re	Cu	1.644	-	-0.90	-1.09
Re	Zn	1.641	-	-1.23	-1.43
Re	Zr	1.773	2.046	-1.74	-1.95
Re	Nb	1.670	-	-0.37	-0.57
Re	Ru	1.790	1.799	-0.30	-0.48
Re	Rh	1.655	-	-1.04	-1.23
Re	Pd	1.651	-	-0.78	-0.96
Re	Ag	1.656	-	-0.78	-0.97
Re	Та	1.679	-	-0.36	-0.56
Re	Os	1.659	-	-0.82	-1.02
Re	Ir	1.787	1.775	-0.30	-0.50
Re	Pt	1.742	1.788	-0.37	-0.56
Re	Au	1.632	-	-0.60	-0.81
Os	Ti	1.624	-	-0.22	-0.41
Os	V	1.637	-	-0.08	-0.26
Os	Mn	1.690	-	-1.02	-1.22
Os	Fe	1.692	1.864	-1.07	-1.27
Os	Co	1.702	1.709	-0.99	-1.20
Os	Ni	1.674	1.747	-0.88	-1.08
Os	Cu	1.685	1.797	-0.83	-1.03
Os	Zn	1.711	-	-0.99	-1.19
Os	Zr	1.710	2.086	-1.28	-1.48
Os	Nb	1.735	1.977	-0.83	-1.02
Os	Ru	1.620	-	-0.63	-0.82

Os	Rh	1.715	1.848	-0.59	-0.79
Os	Pd	1.604	-	-0.71	-0.90
Os	Ag	1.619	-	-0.67	-0.85
Os	Ta	1.635	-	0.05	-0.14
Os	Ir	1.745	1.773	-0.66	-0.87
Os	Pt	1.701	1.777	-0.69	-0.91
Os	Au	1.607	-	-0.88	-1.06
Ir	Ti	1.665	-	-0.85	-1.07
Ir	V	1.657	-	-0.78	-1.00
Ir	Mn	1.646	1.874	-0.97	-1.19
Ir	Fe	1.655	1.743	-0.81	-1.02
Ir	Co	1.681	1.648	-0.98	-1.20
Ir	Ni	1.665	1.663	-0.98	-1.20
Ir	Cu	1.671	1.713	-0.48	-0.70
Ir	Zn	1.655	1.823	-1.30	-1.52
Ir	Nb	1.693	-	0.26	0.07
Ir	Rh	1.707	1.752	-1.00	-1.21
Ir	Pd	1.660	1.766	-0.65	-0.87
Ir	Ag	1.632	-	-0.21	-0.41
Ir	Та	1.739	-	0.19	0.02
Ir	Pt	1.693	1.726	-0.80	-1.02
Pt	Co	1.668	1.608	-0.87	-1.10
Pt	Ni	1.656	1.626	-0.19	-0.42
Au	Ti	-	1.718	-0.13	-0.45
Au	V	1.580	-	-1.18	-1.38
Au	Mn	1.579	-	-0.51	-0.71
Au	Fe	-	1.542	0.02	-0.16
Au	Co	1.797	1.565	0.05	-0.15
Au	Ni	1.615	-	0.08	-0.12

Au	Cu 1.628	1.812	-0.43	-0.63
Au	Zn 1.599	-	-0.68	-0.88
Au	Nb -	1.772	-0.63	-0.78
Au	Rh 1.843	1.640	-0.09	-0.29
Au	Pd 1.569	-	0.19	-0.01
Au	Ag 1.568	-	-1.15	-1.36
Au	Ta -	1.770	-0.97	-1.14
Au	Ir -	1.586	-0.13	-0.30
Au	Pt 1.812	1.664	0.56	0.35

Table S8. The adsorption energies (E_{ad}) for OH on possible sites in M_IM_{II}/g-CN (M_I = Fe, Co, Ni, Cu, Zn, Mo, Ru, Rh, Pd, Ag, Re, Os, Ir, Pt, Au; M_{II} = Co, Ni, Cu, Zn, Mo, Ru, Rh, Pd, Ag, Re, Os, Ir, Pt, Au; M_I \neq M_{II}), the corresponding M_I-OH and M_{II}-OH bond lengths, and the Gibbs free energy change for OH adsorption (Δ G_{OH}*).

M _I M _I	∏∕g-	$E_{\rm ad}$	The	bond	length	The bond length for	ΔG_{OH^*}	ΔG_{OH*}
CN		(eV)	for M _I	-OH (Å)	M_{II} -OH (Å)	(eV)	under
M _I	M_{II}							solvation
								correction
								(eV)
Fe	Co	-0.85	1.921			1.930	-0.51	-0.81
Fe	Ni	-0.59	1.883			1.965	-0.24	-0.54
Fe	Cu	-0.49	1.867			1.998	-0.14	-0.44
Fe	Pd	0.00	1.816			-	0.25	-0.05
Fe	Pt	0.03	1.906			2.064	0.41	0.11
Co	Ni	-0.69	1.858			1.910	-0.31	-0.61
Cu	Co	-0.45	1.968			1.863	-0.09	-0.39
Cu	Ni	0.06	1.936			1.892	0.41	0.11
Cu	Pd	0.48	1.917			2.073	0.82	0.52

Cu	Pt	0.33	2.012	2.053	0.67	0.37
Zn	Fe	-0.80	2.059	1.897	-0.46	-0.76
Zn	Co	-1.14	1.994	1.869	-0.78	-1.08
Zn	Ni	-0.29	1.948	1.929	0.01	-0.29
Zn	Cu	-0.63	1.950	1.888	-0.16	-0.46
Zn	Pd	0.25	1.971	2.139	0.70	0.40
Zn	Pt	0.13	2.012	2.065	0.43	0.13
Mo	Fe	-1.32	1.989	2.220	-1.02	-1.32
Mo	Co	-1.43	1.898	-	-1.15	-1.45
Mo	Ni	-1.49	1.866	-	-1.22	-1.52
Mo	Cu	-1.49	1.870	-	-1.21	-1.51
Mo	Zn	-1.26	1.905	-	-0.98	-1.28
Mo	Ru	-1.02	1.918	-	-0.73	-1.03
Mo	Rh	-1.39	1.909	-	-1.11	-1.41
Mo	Pd	-1.62	1.894	-	-1.33	-1.63
Mo	Ag	-1.41	1.878	-	-1.11	-1.41
Mo	Re	-0.77	2.146	2.055	-1.02	-1.32
Mo	Os	-1.23	1.904	-	-0.92	-1.22
Mo	Ir	-1.44	1.905	-	-1.14	-1.44
Mo	Pt	-0.10	2.029	2.065	0.27	-0.03
Mo	Au	-1.55	1.865	-	-1.27	-1.57
Ru	Fe	-1.21	2.039	2.014	-0.85	-1.15
Ru	Co	-0.97	2.010	1.956	-0.61	-0.91
Ru	Ni	-0.77	1.998	2.008	-0.42	-0.72
Ru	Cu	-0.61	2.007	2.033	-0.26	-0.56
Ru	Zn	-1.23	2.034	2.111	-0.88	-1.18
Ru	Rh	-0.48	2.031	2.027	-0.11	-0.41
Ru	Pd	-0.62	1.915	-	-0.31	-0.61
Ru	Ag	-0.61	1.925	-	-0.30	-0.60

Ru	Ir	-0.45	2.041	2.035	-0.08	-0.38
Ru	Pt	-0.11	2.044	2.037	0.26	-0.04
Ru	Au	-0.12	1.931	-	0.19	-0.11
Rh	Fe	-0.74	2.039	1.910	-0.37	-0.67
Rh	Co	-1.00	2.013	1.860	-0.62	-0.92
Rh	Ni	-0.49	2.018	1.927	-0.13	-0.43
Rh	Cu	-0.10	2.030	1.971	0.25	-0.05
Rh	Zn	-1.20	2.046	2.014	-0.85	-1.15
Rh	Pd	0.24	1.991	2.227	0.59	0.29
Rh	Pt	0.09	2.023	2.039	0.46	0.16
Pd	Co	-0.26	2.046	1.840	0.12	-0.18
Pd	Ni	0.68	2.046	1.891	1.04	0.74
Pd	Pt	1.15	2.215	2.039	1.50	1.20
Ag	Fe	-0.57	-	1.816	-0.25	-0.55
Ag	Co	-0.28	-	1.797	0.05	-0.25
Ag	Ni	-0.20	-	1.787	0.08	-0.22
Ag	Cu	-0.37	2.405	1.815	-0.06	-0.36
Ag	Zn	-0.65	2.116	1.967	-0.31	-0.61
Ag	Rh	-0.29	-	1.957	0.04	-0.26
Ag	Pd	0.42	-	1.973	0.75	0.45
Ag	Pt	0.80	2.069	2.122	1.15	0.85
Re	Fe	-1.03	1.993	2.199	-0.74	-1.04
Re	Co	-1.27	1.913	-	-0.99	-1.29
Re	Ni	-1.26	1.904	-	-0.95	-1.25
Re	Cu	-1.39	1.918	-	-1.09	-1.39
Re	Zn	-1.82	1.913	-	-1.53	-1.83
Re	Ru	-0.13	2.040	2.086	0.23	-0.07
Re	Rh	-1.54	1.922	-	-1.24	-1.54
Re	Pd	-1.31	1.906	-	-1.00	-1.30

Re	Ag	-1.40	1.918	-	-1.11	-1.41
Re	Os	-1.06	1.937	-	-1.06	-1.36
Re	Ir	-0.12	2.048	2.068	0.25	-0.05
Re	Pt	-1.40	1.909	-	-1.08	-1.38
Re	Au	-1.12	1.901	-	-0.81	-1.11
Os	Fe	-1.23	2.042	2.046	-0.88	-1.18
Os	Co	-0.97	2.012	1.995	-0.62	-0.92
Os	Ni	-0.79	1.997	2.066	-0.45	-0.75
Os	Cu	-0.85	1.995	2.113	-0.51	-0.81
Os	Zn	-1.41	2.012	2.252	-1.07	-1.37
Os	Ru	-0.43	2.035	2.126	-0.08	-0.38
Os	Rh	-0.86	1.926	-	-0.54	-0.84
Os	Pd	-0.91	1.922	-	-0.59	-0.89
Os	Ag	-0.99	1.935	-	-0.66	-0.96
Os	Ir	-0.81	1.928	-	-0.49	-0.79
Os	Pt	-1.05	1.919	-	-0.72	-1.02
Os	Au	-0.73	1.933	-	-0.41	-0.71
Ir	Fe	-0.64	2.043	1.932	-0.27	-0.57
Ir	Co	-0.89	2.014	1.895	-0.51	-0.81
Ir	Ni	-0.52	2.023	1.956	-0.15	-0.45
Ir	Cu	-0.18	2.028	2.020	0.17	-0.13
Ir	Zn	-1.26	2.051	2.059	-0.90	-1.20
Ir	Rh	-0.75	2.004	2.046	-0.37	-0.67
Ir	Pd	0.16	1.999	2.351	0.50	0.20
Ir	Ag	-0.26	1.969	-	0.07	-0.23
Ir	Pt	0.17	2.018	2.074	0.55	0.25
Pt	Co	-0.33	2.040	1.859	0.05	-0.25
Pt	Ni	0.64	2.048	1.932	1.01	0.71
Au	Fe	0.08	-	1.829	0.24	-0.06

Au	Co	0.46	-	1.844	0.75	0.45
Au	Ni	1.02	-	1.912	1.33	1.03
Au	Cu	-0.14	2.038	1.963	0.21	-0.09
Au	Zn	0.20	-	1.843	0.48	0.18
Au	Rh	0.58	-	1.974	0.90	0.60
Au	Pd	1.39	-	2.009	1.71	1.41
Au	Ag	-0.73	1.961	-	-0.44	-0.74
Au	Ir	0.47	-	1.991	0.79	0.49
Au	Pt	1.34	-	2.041	1.64	1.34

Table S9. The adsorption energies (E_{ad}) for I^* (I=O, OH, OOH) on possible sites in M_IM_{II}/g -CN ($M_IM_{II} =$ FePd, FePt, CuNi, CuPd, CuPt, ZnNi, ZnPd, ZnPt, MoPt, RuPt, RuAu, RhCu, RhPd, RhPt, PdCo, PdNi, PdPt, AgCo, AgNi, AgCu, AgRh, AgPd, AgPt, ReIr, IrCu, IrPd, IrAg, IrPt, PtCo, PtNi, AuFe, AuCo, AuNi, AuCu, AuZn, AuRh, AuPd, AuIr, AuPt), and the corresponding M_I -I and M_{II} -I bond lengths.

M _I M _{II} /g-		$E_{\rm ad}$ - I^* (eV)			The bond length for			The bond length for		
CN			M_{I} - $I(Å)$			M_{II} - $I(Å)$				
M _I	M_{II}	<i>I</i> =OH	I=O	<i>I</i> =00	I=O	<i>I</i> =0	I=OO	I=O	<i>I</i> =0	<i>I</i> =00
				Н	Н		Н	Н		Н
Fe	Pd	0.70	0.00	3.01	1.816	1.701	1.705	-	1.993	-
Fe	Pt	0.03	0.70	1.46	1.906	1.723	1.743	2.064	1.985	1.986
Cu	Ni	0.06	1.35	3.13	1.936	1.804	1.940	1.892	1.729	1.875
Cu	Pd	0.48	1.81	3.83	1.917	1.783	1.832	2.073	1.918	-
Cu	Pt	0.33	1.59	3.47	2.012	1.850	2.070	2.053	1.923	2.034
Zn	Ni	-0.29	0.82	3.03	1.948	1.847	1.999	1.929	1.746	1.881
Zn	Pd	0.25	1.55	3.46	1.971	1.828	2.029	2.139	1.932	2.121
Zn	Pt	0.13	1.36	3.37	2.012	1.860	2.064	2.065	1.945	2.049
Mo	Pt	-0.10	-1.90	-1.99	2.029	1.693	1.700	2.065	-	-

Ru	Pt	-0.11	0.51	1.26	2.044	1.861	1.873	2.037	1.970	1.982
Ru	Au	-0.12	0.88	2.48	1.931	1.867	1.840	-	2.009	-
Rh	Cu	-0.10	1.21	3.03	2.030	1.857	1.981	1.971	1.873	1.999
Rh	Pd	0.23	1.28	3.38	1.991	1.877	1.948	2.227	1.930	-
Rh	Pt	0.09	1.01	1.94	2.023	1.895	1.882	2.039	1.938	1.932
Pd	Co	-0.26	0.83	2.92	2.046	1.939	2.043	1.840	1.723	1.792
Pd	Ni	0.68	1.64	3.84	2.046	1.916	2.037	1.891	1.751	1.887
Pd	Pt	1.15	1.89	4.44	2.215	1.933	-	2.039	1.927	2.038
Ag	Co	-0.28	1.04	2.73	-	2.072	-	1.797	1.717	1.757
Ag	Ni	-0.20	1.34	2.89	-	2.078	-	1.787	1.705	1.779
Ag	Cu	-0.37	1.70	2.91	2.405	2.022	-	1.815	1.773	1.811
Ag	Rh	-0.29	1.36	2.64	-	2.035	-	1.957	1.858	1.913
Ag	Pd	0.42	2.37	3.47	-	2.392	-	1.973	1.851	1.964
Ag	Pt	0.80	2.05	4.10	2.069	2.000	2.066	2.122	1.916	-
Re	Ir	-0.12	-1.12	-0.45	2.048	1.721	1.843	2.068	-	2.048
Ir	Cu	-0.18	0.85	2.94	2.028	1.863	1.976	2.020	1.920	2.055
Ir	Pd	0.16	1.10	3.31	1.999	1.881	1.936	2.351	1.960	-
Ir	Ag	-0.26	1.06	2.68	1.969	1.796	1.921	-	-	-
Ir	Pt	0.17	0.93	1.72	2.018	1.898	1.947	2.074	1.963	1.948
Pt	Co	-0.33	0.66	2.87	2.040	1.947	2.024	1.859	1.734	1.809
Pt	Ni	0.64	1.49	3.82	2.048	1.927	2.032	1.932	1.766	1.917
Au	Fe	0.01	1.18	2.92	-	-	-	1.829	1.656	1.770
Au	Co	0.46	1.94	3.44	-	2.104	-	1.844	1.722	1.838
Au	Ni	1.02	2.28	4.02	-	1.961	-	1.912	1.753	1.876
Au	Cu	-0.14	1.46	3.89	2.038	1.949	-	1.963	1.814	1.807
Au	Zn	0.20	0.83	3.52	-	1.955	-	1.843	1.846	1.884
Au	Rh	0.58	1.72	3.54	-	1.960	-	1.974	1.926	1.949
Au	Pd	1.39	2.74	3.99	-	1.959	1.974	2.009	1.918	-
Au	Ir	0.47	1.81	3.46	-	2.258	-	1.991	1.856	1.953

Table S10. The Gibbs free energy change for oxygenated intermediates adsorption $(\Delta G_{O^*}, \Delta G_{OH^*}, \Delta G_{OOH^*})$ on $M_I M_{II}/g$ -CN $(M_I M_{II} =$ FePd, FePt, CuNi, CuPd, CuPt, ZnNi, ZnPd, ZnPt, MoPt, RuPt, RuAu, RhCu, RhPd, RhPt, PdCo, PdNi, PdPt, AgCo, AgNi, AgCu, AgRh, AgPd, AgPt, ReIr, IrCu, IrPd, IrAg, IrPt, PtCo, PtNi, AuFe, AuCo, AuNi, AuCu, AuZn, AuRh, AuPd, AuIr, AuPt), the overpotentials for OER (η^{OER}) , and the overpotentials for ORR (η^{ORR}) .

$M_I M_{II}/g$ -CN		$\Delta G_{OH^*}(eV)$	$\Delta G_{O^*}(eV)$	ΔG_{OOH^*}	$\eta^{\text{OER}}(V)$	$\eta^{ORR}(V)$
M_{I}	M _{II}			(eV)		
Fe	Pd	0.25	0.77	3.32	1.33	0.98
Fe	Pt	0.41	0.77	1.92	1.77	0.87
Cu	Ni	0.41	1.41	3.47	0.83	0.82
Cu	Pd	0.82	1.87	4.16	1.06	0.47
Cu	Pt	0.67	1.64	3.85	0.97	0.56
Zn	Ni	0.06	0.88	3.38	1.28	1.17
Zn	Pd	0.70	1.61	3.83	0.99	0.53
Zn	Pt	0.49	1.42	3.76	1.11	0.74
Mo	Pt	0.27	-1.87	-1.61	5.30	3.37
Ru	Pt	0.26	0.58	1.72	1.97	0.97
Ru	Au	0.19	0.94	2.78	0.91	1.04
Rh	Cu	0.25	1.27	3.39	0.89	0.98
Rh	Pd	0.59	1.35	3.72	1.14	0.64
Rh	Pt	0.46	1.08	2.37	1.32	0.77
Pd	Co	0.12	0.90	3.30	1.17	1.11

Pd	Ni	1.04	1.71	4.22	1.28	0.56
Pd	Pt	1.50	1.96	4.80	1.61	1.11
Ag	Co	0.05	1.09	3.08	0.77	1.18
Ag	Ni	0.08	1.37	3.21	0.61	1.15
Ag	Cu	-0.06	1.73	3.25	0.56	1.29
Ag	Rh	0.04	1.41	2.98	0.71	1.19
Ag	Pd	0.75	2.41	3.83	0.43	0.48
Ag	Pt	1.15	2.10	4.40	1.06	0.71
Re	Ir	0.25	-1.07	-0.05	3.74	2.55
Ir	Cu	0.17	0.91	3.26	1.16	1.06
Ir	Pd	0.50	1.17	3.63	1.23	0.73
Ir	Ag	0.07	1.09	3.03	0.71	1.16
Ir	Pt	0.55	1.00	2.11	1.58	0.78
Pt	Co	0.05	0.73	3.24	1.28	1.18
Pt	Ni	1.01	1.57	4.20	1.40	0.67
Au	Fe	0.24	1.22	3.22	0.78	0.99
Au	Co	0.75	2.00	3.75	0.52	0.48
Au	Ni	1.33	2.35	4.35	0.76	0.66
Au	Cu	0.21	1.51	4.21	1.47	1.02
Au	Zn	0.48	0.88	3.83	1.72	0.83
Au	Rh	0.90	1.78	3.86	0.85	0.35
Au	Pd	1.71	2.81	4.32	0.48	0.63
Au	Ir	0.78	1.86	3.79	0.70	0.44
Au	Pt	1.64	2.81	4.99	0.95	1.30
Table S11. The initial data set of intrinsic descriptors including the distance of two metal atoms (L, Å), the average bond lengths between the surrounding N atoms and metal atoms $(\overline{d}, Å)$, the covalent radius $(R_{cov-ul}, R_{cov-lr}, Å)$, the number of valance electrons (N_{va-ul}, N_{va-lr}) , the number of d electrons (N_{d-ul}, N_{d-lr}) , the Pauling electronegativity (E_{p-ul}, E_{p-dr}) , the first ionization energy (E_{I-ul}, E_{Id-r}, eV) , the electron affinity (A_{e-ul}, A_{e-dr}, eV) , and the d-band center $(\varepsilon_{d-ul}, \varepsilon_{d-dr}, eV)$ for 23 homonuclear and 253 heteronuclear metal dimers adsorbed on g-CN. Herein, -ul and -lr denote the characteristis for metal atoms situated at the upper left and the down right sites of g-CN, respectively.

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R u Z n	2 4 5 4	2 1 0 9	1 4 2	1 2 2	8	1 2	7	1 0	2 2	1 6 5	7 3 6 1	9 3 9 4	1 0 5	0	- 4. 2 6 2	- 6. 7 9 0	3. 85	0. 55	2. 64	0.2
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R u N b	2 1 7	2 1 3 8	1 4 2	1 6 4	8	5	7	5	2 2	1 6	7 3 6 1	6 7 5 8	1 0 5	0 8 9 3	- 2. 9 9 6	- 3. 1 7 9	3. 8	0. 6	3. 06	0.2 2
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Table S12. The calculated coefficient of determination (R^2) and the root-mean-square-error (RMSE) values for training samples and testing samples.

DFT	calculated	Training samples	5	Testing samples	
VS	predicted	R ²	RMSE	R ²	RMSE
value	S				
ΔG_{H^*}	(eV)	0.978	0.333	0.965	0.012
η^{OER} ((V)	0.987	0.243	0.982	0.007
η^{ORR}	(V)	0.982	0.245	0.977	0.012

Table S13. The calculated ΔG_{H^*} , η^{OER} and η^{ORR} values for Ni₂/g-CN, Pd₂/g-CN, CuPd/g-CN, PdNi/g-CN, AgPt/g-CN, AuFe/g-CN, AuCo/g-CN, AuNi/g-CN, AuRh/g-CN, AgPd/g-CN, AuPd/g-CN, and AuIr/g-CN after considering solvation correction via Vaspsol.

Systems	$\Delta G_{\mathrm{H}^{\ast}}$		η^{OER}		η^{OER}	
	vaspsol	vasp	vaspsol	vasp	vaspsol	vasp
Ni ₂ /g-CN	-0.07	-0.05	-	-	-	-
Pd ₂ /g-CN	-0.01	0.03	-	-	-	-
CuPd/g-CN	-0.03	-0.02	-	-	0.56	0.47
PdNi/g-CN	-0.03	0	-	-	-	-
AgPt/g-CN	0.05	0	-	-	-	-

AuFe/g-CN	0.11	0.02	-	-	-	-
AuCo/g-CN	0.02	0.05	0.44	0.52	0.62	0.48
AuNi/g-CN	0.11	0.08	-	-	-	-
AuRh/g-CN	-0.11	-0.09	-	-	0.47	0.35
AgPd/g-CN	-	-	0.46	0.43	0.67	0.48
AuPd/g-CN	-	-	0.39	0.48	-	-
AuIr/g-CN	-	-	-	-	0.57	0.44



Fig. S1. The stable configurations for (a) Sc_2/g-CN, (b) Ti_2/g-CN, (c) V_2/g-CN, (d)

Cr₂/g-CN, (e) Mn₂/g-CN, (f) Fe₂/g-CN, (g) Co₂/g-CN, (h) Ni₂/g-CN, (i) Cu₂/g-CN, (j) Zn₂/g-CN, (k) Y₂/g-CN, (l) Zr₂/g-CN, (m) Nb₂/g-CN, (n) Mo₂/g-CN, (o) Ru₂/g-CN, (p) Rh₂/g-CN, (q) Pd₂/g-CN, (r) Ag₂/g-CN, (s) Hf₂/g-CN, (t) Ta₂/g-CN, (u) W₂/g-CN, (v) Re₂/g-CN, (w) Os₂/g-CN, (x) Ir₂/g-CN, (y) Pt₂/g-CN, and (z) Au₂/g-CN.



Fig. S2. The stable configurations for H adsorption on (a) Ti_2/g -CN, (b) V_2/g -CN, (c) Cr_2/g -CN, (d) Mn_2/g -CN, (e) Fe_2/g -CN, (f) Co_2/g -CN, (g) Ni_2/g -CN, (h) Cu_2/g -CN, (i) Zn_2/g -CN, (j) Zr_2/g -CN, (k) Nb_2/g -CN, (l) Mo_2/g -CN, (m) Ru_2/g -CN, (n) Rh_2/g -CN, (o) Pd_2/g -CN, (p) Ag_2/g -CN, (q) Ta_2/g -CN, (r) W_2/g -CN, (s) Re_2/g -CN, (t) Os_2/g -CN, (u) Ir_2/g -CN, (v) Pt_2/g -CN, and (w) Au_2/g -CN.


Fig. S3. The stable configurations for OH adsorption on (a) Ti_2/g -CN, (b) V_2/g -CN, (c) Cr_2/g -CN, (d) Mn_2/g -CN, (e) Fe_2/g -CN, (f) Co_2/g -CN, (g) Ni_2/g -CN, (h) Cu_2/g -CN, (i) Zn_2/g -CN, (j) Zr_2/g -CN, (k) Nb_2/g -CN, (l) Mo_2/g -CN, (m) Ru_2/g -CN, (n) Rh_2/g -CN, (o) Pd_2/g -CN, (p) Ag_2/g -CN, (q) Ta_2/g -CN, (r) W_2/g -CN, (s) Re_2/g -CN, (t) Os_2/g -CN, (u) Ir_2/g -CN, (v) Pt_2/g -CN, and (w) Au_2/g -CN.



Fig. S4. The stable configurations for O adsorption on (a) Ti_2/g -CN, (b) V_2/g -CN, (c) Cr_2/g -CN, (d) Mn_2/g -CN, (e) Fe_2/g -CN, (f) Co_2/g -CN, (g) Ni_2/g -CN, (h) Cu_2/g -CN, (i) Zn_2/g -CN, (j) Zr_2/g -CN, (k) Nb_2/g -CN, (l) Mo_2/g -CN, (m) Ru_2/g -CN, (n) Rh_2/g -CN, (o) Pd_2/g -CN, (p) Ag_2/g -CN, (q) Ta_2/g -CN, (r) W_2/g -CN, (s) Re_2/g -CN, (t) Os_2/g -CN, (u) Ir_2/g -CN, (v) Pt_2/g -CN, and (w) Au_2/g -CN.



Fig. S5. The stable configurations for OOH adsorption on (a) Ti_2/g -CN, (b) V_2/g -CN, (c) Cr_2/g -CN, (d) Mn_2/g -CN, (e) Fe_2/g -CN, (f) Co_2/g -CN, (g) Ni_2/g -CN, (h) Cu_2/g -CN, (i) Zn_2/g -CN, (j) Zr_2/g -CN, (k) Nb_2/g -CN, (l) Mo_2/g -CN, (m) Ru_2/g -CN, (n) Rh_2/g -CN, (o) Pd_2/g -CN, (p) Ag_2/g -CN, (q) Ta_2/g -CN, (r) W_2/g -CN, (s) Re_2/g -CN, (t) Os_2/g -CN,



Fig. S6. (a)-(w) The Gibbs free energy for O/OH/OOH adsorption on M_2 /g-CN (M = Ti, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt,

Au). (x) The relations between the d band center (ϵ_d) and $\Delta G_{O*}/\Delta G_{OH*}/\Delta G_{OOH*}$.



Fig. S7. The stable configurations for (a) $M_I M_{II}/g$ -CN ($M_I = Ti$, V, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; $M_{II} = V$, Cr, Mn, Fe, Co, Ni, Cu, Zn, Zr, Nb, Mo, Ru, Rh, Pd, Ag, Ta, W, Re, Os, Ir, Pt, Au; $M_I \neq M_{II}$; $M_I M_{II} \neq ZrCu$, NbCu, TaCu, RuZn, TaZn, ZnTi, NbZn, MoAg, ReAg, AgTi, ZrZn, ZrAg, AgNb, TaAg, WAg, MoZn, WZn, ReZn, OsZn), (b) $M_I M_{II}/g$ -CN ($M_I M_{II} = ZrCu$, NbCu, RuZn, TaZn), (c) $M_I M_{II}/g$ -CN ($M_I M_{II} = NbZn$, MoAg, ReAg), (d) $M_I M_{II}/g$ -CN ($M_I M_{II} = ZrZn$, ZrAg, TaAg, TaAg, WAg), (e) $M_I M_{II}/g$ -CN ($M_I M_{II} = ZnTi$, AgV), (f) $M_I M_{II}/g$ -CN ($M_I M_{II} = MoZn$, WZn, ReZn, OsZn), and (g) $M_I M_{II}/g$ -CN ($M_I M_{II} = AgTi$, AgNb).



Fig. S8. The stable configurations for H adsorption on (a) M_IM_{II}/g-CN (M_IM_{II} = TiPd, VPd, CrZr, ZnV, ZrNi, ZrPd, ZrAu, NbTi, NbNi, NbRh, NbPd, MoTi, MoMn, MoZr, MoRh, MoPd, MoIr, MoAu, RuTi, RuV, RuCu, RuZr, RuNb, RuAg, RuTa, RuAu, RhV, TaTi, TaNi, TaNb, TaRh, TaPd, WTi, WFe, WNi, WZr, WRu, WPd, WIr, WPt, WAu, ReTi, ReV, ReMn, ReFe, ReNi, ReCu, ReNb, ReRh, RePd, ReTa, ReOs, ReAu, OsTi, OsV, OsMn, OsRu, OsPd, OsAg, OsTa, OsAu, IrTi, IrV, IrNb, IrAg, IrTa, AuV,

AuMn, AuNi, AuZn, AuPd, AuAg), (b) $M_I M_{II}/g$ -CN ($M_I M_{II} = TiFe$, TiCo, TiNi, TiCu, TiPt, VTi, VFe, VCo, VNi, VCu, VPt, CrTi, CrV, CrMn, CrFe, CrCo, CrNi, CrCu, CrZn, CrNb, CrMo, CrRu, CrRh, CrPd, CrAg, CrTa, CrW, CrIr, CrPt, MnTi, MnV, MnFe, MnCo, MnNi, MnCu, MnPd, MnPt, FeCo, FeNi, FeCu, FePd, FePt, CoNi, CuCo, CuNi, CuPd, CuPt, ZnMn, ZnTi, ZnCo, ZnNi, ZnCu, ZnPd, ZnPt, ZrTi, ZrV, ZrMn, ZrFe, ZrCo, ZrPt, NbV, NbMn, NbFe, NbCo, NbPt, MoV, MoFe, MoCo, MoNi, MoCu, MoNb, MoTa, MoPt, RuMn, RuFe, RuCo, RuNi, RuRh, RuPd, RuIr, RuPt, RhTi, RhMn, RhFe, RhCo, RhNi, RHCu, RhZn, RhPd, RhPt, PdCo, PdNi, PdPt, AgMn, AgFe, AgCo, AgNi, AgCu, AgZn, AgRh, AgPd, AgPt, TaV, TaMn, TaFe, TaCo, TaPt, WV, WMn, WCo, WCu, WRh, ReCo, ReZr, ReRu, ReIr, RePt, OsFe, OsCo, OsNi, OsCu, OsZr, OsNb, OsRh, OsIr, OsPt, IrMn, IrFe, IrCo, IrNi, IrCu, IrZn, IrRh, IrPd, IrPt, PtCo, PtNi, AuPt), (c) $M_I M_{II}/g$ -CN ($M_I M_{II} = CrRe$, CrOs, CrAu, ZrNb, ZrRh, ZrTa, ZrIr, MoRu, MoW, MoRe, MoOs, AgNb, WNb, WTa, WRe, WOs, AuTi, AuFe, AuNb, AuTa, AuIr), (d) $M_I M_{II}/g$ -CN ($M_I M_{II} = ZrCu$, NbCu, RuZn, TaCu, TaZn), (e) $M_I M_{II}/g$ -CN ($M_I M_{II} = NbZn$, MoAg, ReAg), (f) $M_I M_{II}/g$ -CN ($M_I M_{II}$ = ZrZn, ZrAg, TaAg, WAg), (g) $M_I M_{II}/g$ -CN ($M_I M_{II} = ZnTi, AgV$), (h) $M_I M_{II}/g$ -CN $(M_IM_{II} = MoZn, WZn, ReZn, OsZn)$, and (i) M_IM_{II}/g -CN $(M_IM_{II} = AgTi, AgNb)$.



Fig. S9. The stable configurations for OH adsorption on (a) M_IM_{II}/g -CN ($M_IM_{II} =$ FePd, RuAu, IrAg), (b) M_IM_{II}/g -CN ($M_IM_{II} =$ AgCo, AgNi, AgRh, AgPd, AuFe, AuCo, AuNi, AuZn, AuRh, AuPd, AuIr, AuPt), and (c) M_IM_{II}/g -CN (FePt, CuNi, CuPd, CuPt, ZnNi, ZnPd, ZnPt, MoPt, RuPt, RhCu, RhPd, RhPt, PdCo, PdNi, PdPt, AgCu, AgPt, ReIr, IrCu, IrPd, IrPt, PtCo, PtNi, AuCu). The stable configurations for O adsorption on (d) M_IM_{II}/g -CN ($M_IM_{II} =$ MoPt, ReIr, IrAg), (e) AuFe/g-CN, and (f) M_IM_{II}/g -CN ($M_IM_{II} =$ FePd, FePt, CuNi, CuPd, CuPt, ZnNi, ZnPd, ZnPt, RePt, RuNi, CuPd, CuPt, ZnNi, ZnPd, ZnPt, RuPt, RuAu, RhCu, RhPd, RhPt, PdCo, PdNi, PdPt, AgCo, AgNi, AgCu, AgRh, AgPd, AgPt, IrCu, IrPd, IrPt, PtCo, PtNi, AuFe, AuCo, AuNi, AuCu, AuZn, AuRh, AuPd, AuIr, AuPt). The stable configurations for OOH adsorption on (g) M_IM_{II}/g -CN ($M_IM_{II} =$ FePd, KuAu, RhPd, AgPt, IrPd, IrAg, AuPd, AuPt), (h) M_IM_{II}/g -CN ($M_IM_{II} =$ PdPt, AgCo, AgNi, AgCu, AgRh, AgPd, AuFe, AuCo, AuNi, AuCu, AuZn, AuRh, AuIr), and (i) (FePt, CuNi, CuPt, ZnNi, ZnPd, ZnPt, RuPt, RhCu, RhPt, PdCo, PdNi, ReIr, IrCu, IrPt, PtCo, PtNi).



Fig. S10. Calculated free energy diagram of OER/ORR on (a) FePd/g-CN, (b) FePt/g-CN, (c) CuNi/g-CN, (d) CuPt/g-CN, (e) ZnNi/g-CN, (f) ZnPd/g-CN, (g) ZnPt/g-CN, (h) MoPt/g-CN, (i) RuPt/g-CN, (j) RuAu/g-CN, (k) RhCu/g-CN, (l) RhPd/g-CN, (m) RhPt/g-CN, (n) PdCo/g-CN, (o) PdNi/g-CN, (p) PdPt/g-CN, (q) AgCo/g-CN, (r) AgNi/g-CN, (s) AgRh/g-CN, (t) AgPt/g-CN, (u) ReIr/g-CN, (v) IrCu/g-CN, (w) IrPd/g-CN, (x) IrAg/g-CN, (y) IrPt/g-CN, (z) PtCo/g-CN, (a1) PtNi/g-CN, (b1) AuFe/g-CN, (c1) AuNi/g-CN, (d1) AuCu/g-CN, (e1) AuZn/g-CN, and (f1) AuPt/g-CN under U=0 V, wherein the PDS of the elementary reactions are denoted as red lines. (g1) The overpotentials of OER/ORR for $M_I M_{II}/g$ -CN ($M_I M_{II}$ = FePd, FePt,

CuNi, CuPt, ZnNi, ZnPd, ZnPt, MoPt, RuPt, RuAu, RhCu, RhPd, RhPt, PdCo, PdNi, PdPt, AgCo, AgNi, AgRh, AgPt, ReIr, IrCu, IrPd, IrAg, IrPt, PtCo, PtNi, AuFe, AuNi, AuCu, AuZn, AuPt).



Fig. S11. Calculated free energy diagram of ORR on the benchmark catalyst Pt(111).



Fig. S12. The band structures and DOS for (a) Ni₂/g-CN, (b) Zn₂/g-CN, (c) Pd₂/g-CN, (d) CuPd/g-CN, (e) NbCo/g-CN, (f) MoRu/g-CN, (g) MoOs/g-CN, (h) RuTi/g-CN, (i) RuV/g-CN, (j) RuTa/g-CN, (k) PdNi/g-CN, (l) AgPt/g-CN, (m) TaPt/g-CN, (n) WOs/g-CN, (o) OsV/g-CN, (p) OsTa/g-CN, (q) AuFe/g-CN, (r) AuCo/g-CN, (s) AuNi/g-CN, (t) AuRh/g-CN, (u) AgCu/g-CN, (v) AgPd/g-CN, (w) AuPd/g-CN, and (x) AuIr/g-CN.



Fig. S13. The surface pourbaix diagrams of (a) Ni₂/g-CN, (b) Pd₂/g-CN, (c) CuPd/g-CN, (d) PdNi/g-CN, (e) AgPt/g-CN, (f) AuFe/g-CN, (g) AuCo/g-CN, (h) AuNi/g-CN, (i) AuRh/g-CN, (j) AgPd/g-CN, (k) AuPd/g-CN, and (l) AuIr/g-CN. The black dashed lines represent the HER/OER/ORR at equilibrium ($U_{eq} = 1.23 V - 0.059 pH$). The red and blue dashed lines denote the oxidizing state for OER and ORR, respectively at an overpotential 0.5 V. The other solid lines represent the phase boundary of two oxidizing states.

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