

## Supporting Information

### **Colorimetric Sensor Array for Antioxidant Recognition based on Co<sub>3</sub>O<sub>4</sub> Dual Enzyme Activity**

**Pingping Hao<sup>a</sup>, Zhenchao Liu<sup>a</sup>, Zhiwei Wang<sup>a</sup>, Min Xie<sup>b,\*</sup> and Qingyun Liu<sup>a,\*</sup>**

*<sup>a</sup> College of Chemical and Biological Engineering, Shandong University of Science and Technology, Qingdao 266590, P R China*

*<sup>b</sup> Community Health Service Center (University Hospital), University of Science and Technology Beijing, Beijing 100083, P. R. China*

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**\* Corresponding Author**

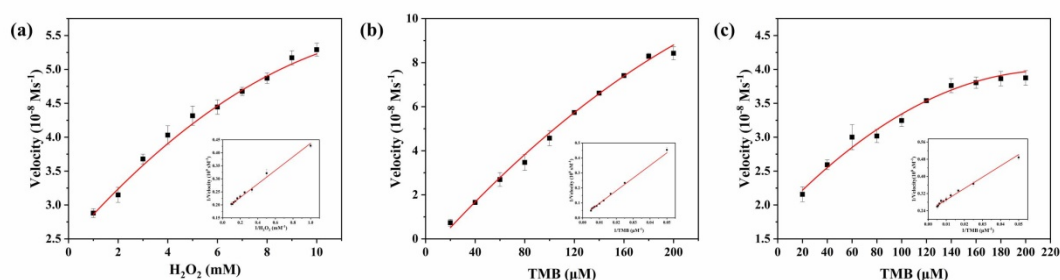
E-mail: min66@ustb.edu.cn; qyliu@sdust.edu.cn

Tel: +86 0532 86057757

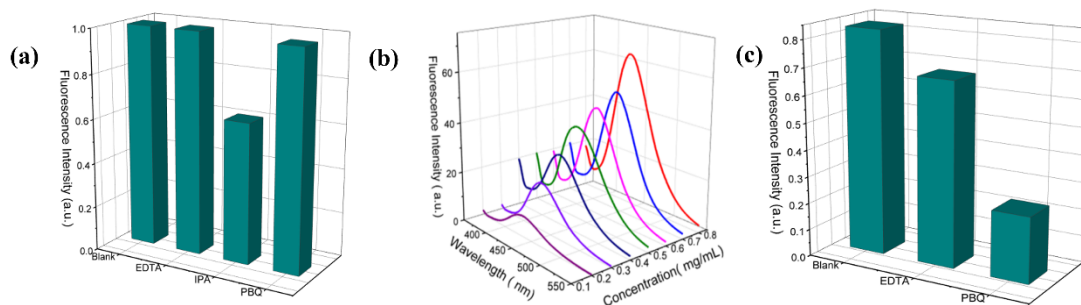
## Preparation

A solution containing cobalt nitrate hexahydrate ( $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ , 2.0 mmol), glycerol (16 mL) and isopropanol (60 mL) was placed in a Teflon-lined stainless-steel autoclave. The autoclave was then treated at 180 °C for 6 h. After cooling to room temperature naturally, the solution was centrifuged to recover the solid precipitate, which was washed several times with ethanol and dried in an oven at 80 °C to afford the precursor of Co.

A mixture containing the precursor (0.1 g) and deionized water (20 mL) was placed in a Teflon-lined stainless steel autoclave, and then treated at 160 °C for 3 h. After cooling to room temperature naturally, the solution was centrifuged to recover the solid precipitate. The finally obtained solid product is denoted as h-CoOH. The  $\text{Co}_3\text{O}_4$  samples were prepared by calcining h-CoOH in air at 400 °C for 2 h, respectively.



**Fig. S1** Kinetic analysis of the as-prepared  $\text{Co}_3\text{O}_4$  by the Michaelis–Menten model and double reciprocal plots, respectively.



**Fig. S2** (a), (c) Effects of various active scavengers during the catalysis of TMB with the aid of  $\text{Co}_3\text{O}_4$  and (b) Fluorescence intensity varies with the concentration of  $\text{Co}_3\text{O}_4$  at different wavelengths.

**Table S1** Comparison of peroxidase-like and oxidase-like kinetic parameters.

Catalyst	$K_m$ (mM)		$V_{max}$ ( $10^{-8}$ Ms $^{-1}$ )		Ref.
	$\text{H}_2\text{O}_2$	TMB	$\text{H}_2\text{O}_2$	TMB	
HRP	3.7	0.434	8.71	10	1
$\text{CeO}_2/\text{C}$	2.61	0.12	3.31	2.08	2
$\text{CoFe-LDH}/\text{CeO}_2$	10.82	0.419	\	\	3
$\text{CeO}_2$ NPs(OXD)	\	0.80	\	30.00	4
$\text{Co}_3\text{O}_4(\text{POD})$	1.3	0.221	5.62	71.22	<b>This work</b>
$\text{Co}_3\text{O}_4(\text{OXD})$	\	0.598	\	4.199	<b>This work</b>

**Table S2** Other methods used to detect DA.

<b>Linear range</b>	<b>LOD</b>	<b>Method</b>	<b>Ref.</b>
1-200 $\mu\text{M}$	0.07 $\mu\text{M}$	Fluorescence	5
0.5-25 $\mu\text{M}$	0.18 $\mu\text{M}$	Electrochemical	6
<b>0-1000 nM</b>	<b>8.26 nM</b>	<b>Colorimetric sensor array</b>	<b>This work</b>

**Table S3** Other methods used to detect AA.

<b>Linear range</b>	<b>LOD</b>	<b>Method</b>	<b>Ref.</b>
0-80 $\mu\text{M}$	0.026 $\mu\text{M}$	Colorimetric	7
10-250 $\mu\text{M}$	1.3 $\mu\text{M}$	Fluorescence	8
<b>0-1000 nM</b>	<b>5.42 nM</b>	<b>Colorimetric sensor array</b>	<b>This work</b>

**Table S4** Other methods used to detect GSH.

<b>Linear range</b>	<b>LOD</b>	<b>Method</b>	<b>Ref.</b>
1-10 $\mu\text{M}$	0.658 $\mu\text{M}$	Colorimetric	9
12.5-800 $\mu\text{M}$	0.7 $\mu\text{M}$	Fluorescence	10
<b>0-1000 nM</b>	<b>2.89 nM</b>	<b>Colorimetric sensor array</b>	<b>This work</b>

**Table S5** Other methods used to detect Cys.

<b>Linear range</b>	<b>LOD</b>	<b>Method</b>	<b>Ref.</b>
0-140 $\mu\text{M}$	11.1 nM	Fluorescence	11
0.05 -14.0 $\mu\text{M}$	20 nM	Colorimetric	12
<b>0-1000 nM</b>	<b>6.24 nM</b>	<b>Colorimetric sensor array</b>	<b>This work</b>

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