

## Supplementary material

to

### Development of Optical Sensor for Determination of Phenolic Compounds in Environmental Samples

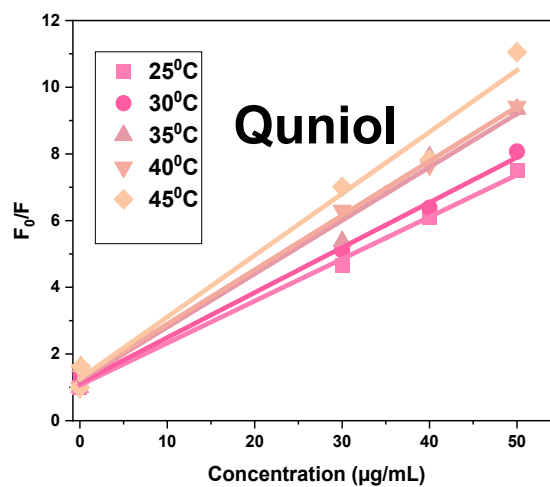
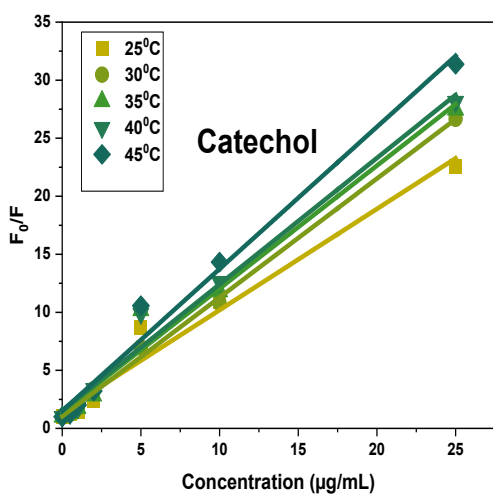
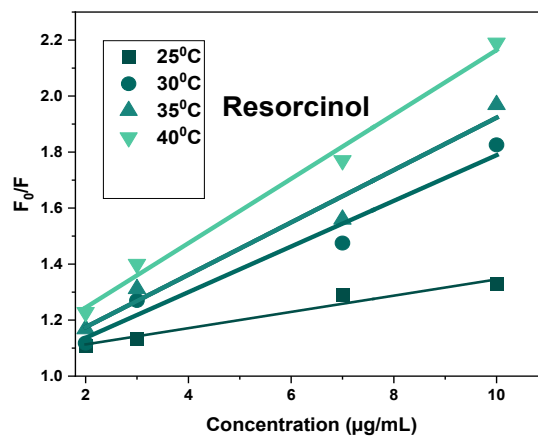
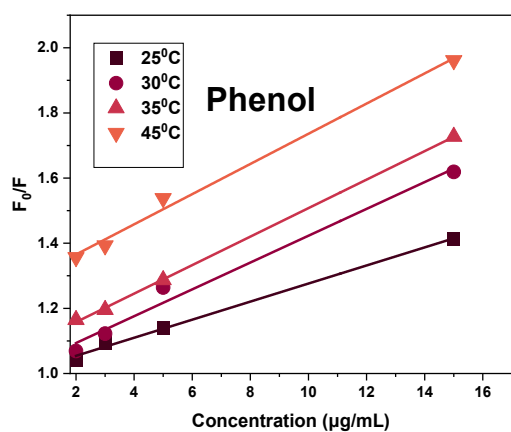
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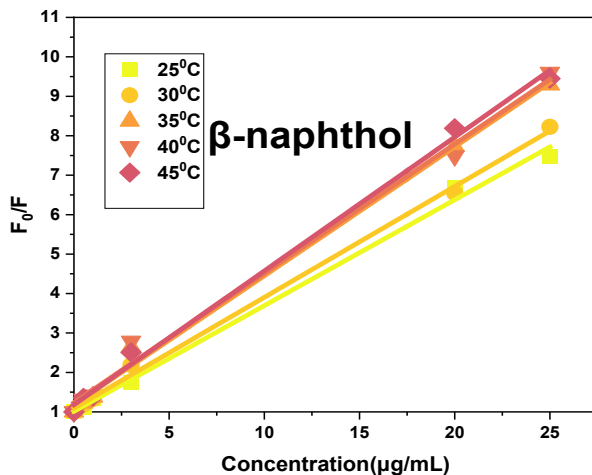
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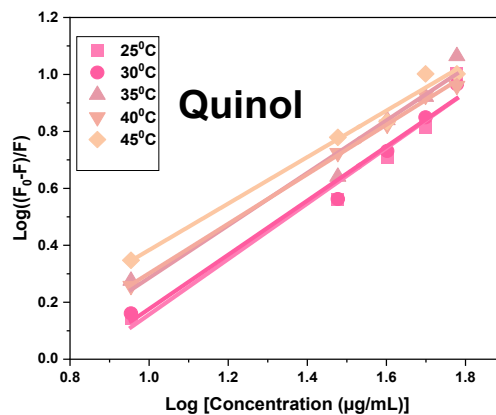
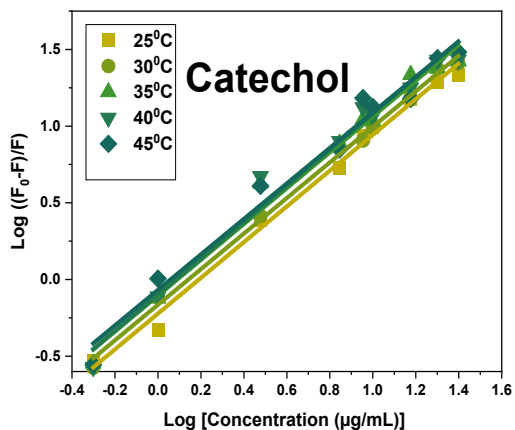
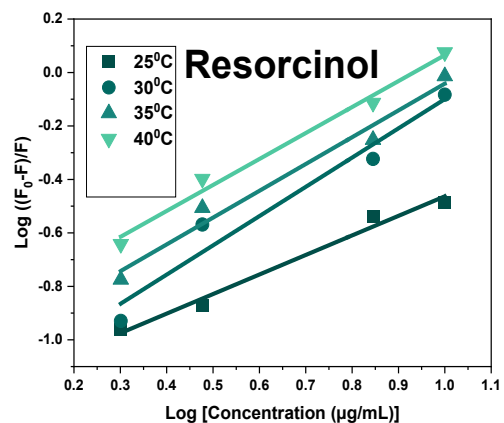
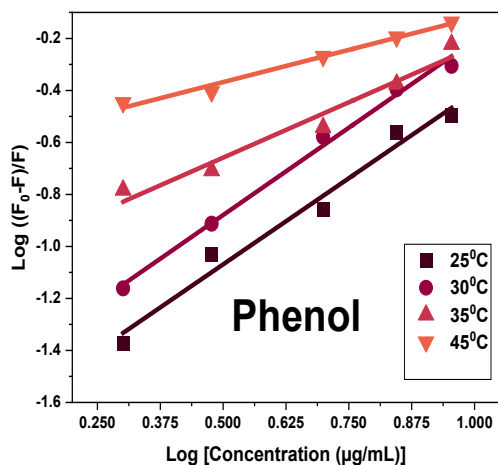
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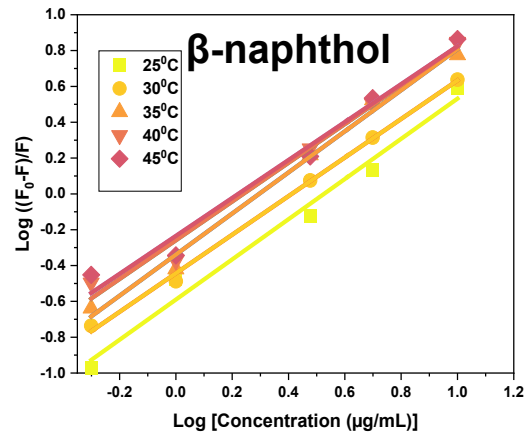
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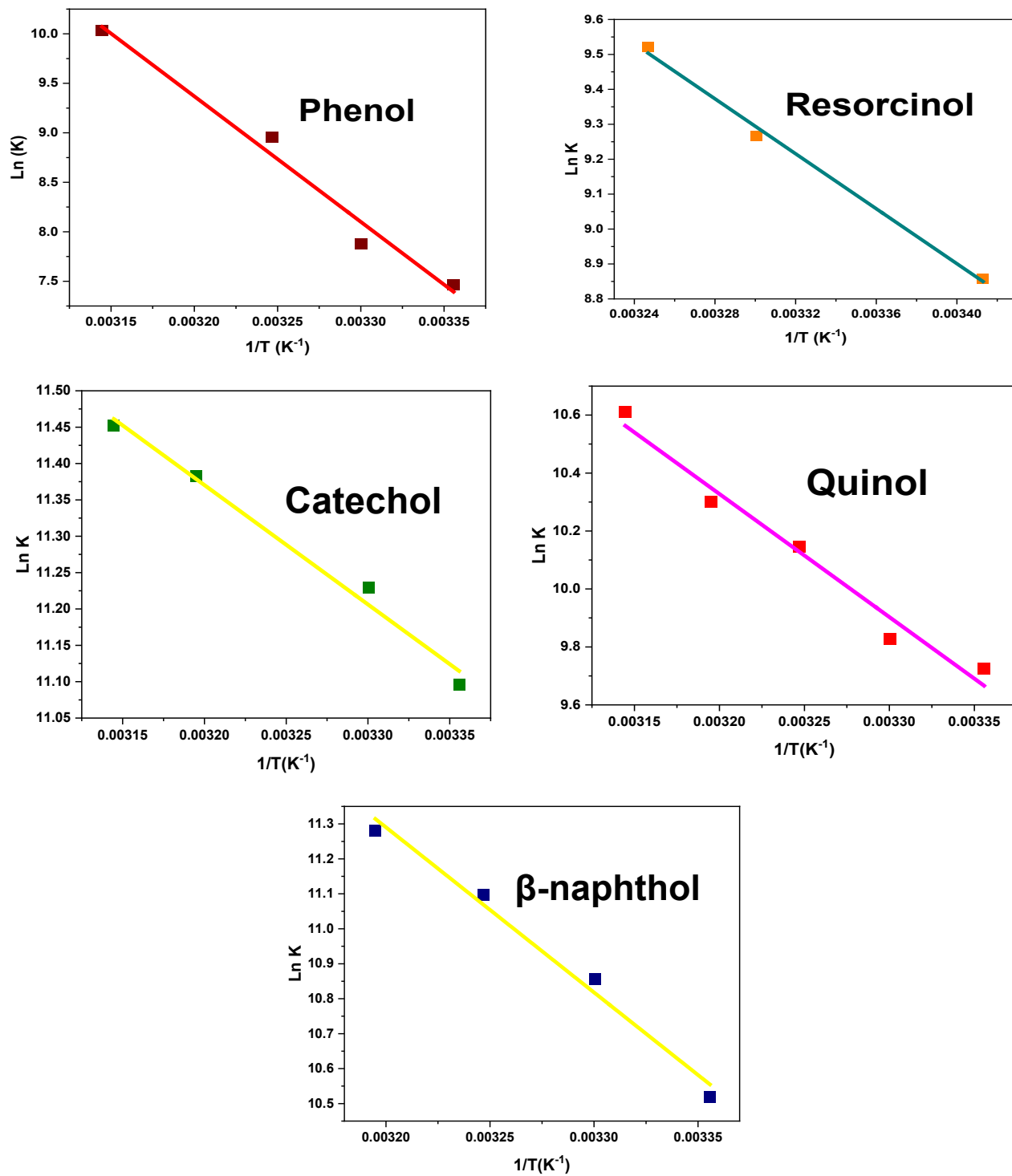


**Figure 1S.** Stern–Volmer plots for the reaction of Eu(III)-complex and phenolic compounds in PIPES buffer (pH=7.5) at different temperatures





**Figure 2S.** The modified Stern–Volmer plots between  $\log (F_0 - F)/F$  and  $\log$  of the concentration of phenolic compounds at different temperatures.



**Figure 3S.** The plot of Van't Hoff of Eu(III)-complex and phenolic compounds

**Table 1S.** Thermodynamic parameters for the interaction of Eu(III)-complex with the phenolic compounds under study

	Temperature (T)	$\Delta H^0$ (KJ/mol)	$\Delta S^0$ (J/mol K)	$\Delta G^0$ (KJ/mol)
<b>Phenol</b>	298 <sup>0</sup> K	<b>10.548</b>	<b>415.430</b>	-18.496
	303 <sup>0</sup> K			-19.853
	308 <sup>0</sup> K			-22.933
	318 <sup>0</sup> K			-26.531
<b>Resorcinol</b>	298 <sup>0</sup> K	<b>32.703</b>	<b>185.198</b>	-21.579
	303 <sup>0</sup> K			-21.951
	308 <sup>0</sup> K			-23.347
	313 <sup>0</sup> K			-24.381
<b>Catechol</b>	298 <sup>0</sup> K	<b>13.647</b>	<b>138.211</b>	-27.493
	303 <sup>0</sup> K			-28.290
	308 <sup>0</sup> K			-29.106
	313 <sup>0</sup> K			-29.622
	318 <sup>0</sup> K			-30.278
<b>Quinol</b>	298 <sup>0</sup> K	<b>35.325</b>	<b>198.912</b>	-24.097
	303 <sup>0</sup> K			-24.755
	308 <sup>0</sup> K			-25.980
	313 <sup>0</sup> K			-26.806
	318 <sup>0</sup> K			-28.057
<b><math>\beta</math>-naphthol</b>	298 <sup>0</sup> K	<b>39.264</b>	<b>219.520</b>	-26.063
	303 <sup>0</sup> K			-27.352
	308 <sup>0</sup> K			-28.418
	313 <sup>0</sup> K			-29.359
	318 <sup>0</sup> K			-29.975

**Table 2S.** Tolerance limit of interfering species affecting on detection of phenolic compounds using Eu(III)-complex.

Interfering species	Tolerance				
	Phenol	Resorcinol	Catechol	Quinol	$\beta$ -naphthol
Na <sup>+</sup>	0.68 <sup>Q</sup> $\mu\text{g/mL}$	0.84 <sup>Q</sup> $\mu\text{g/mL}$	1.41 <sup>Q</sup> $\mu\text{g/mL}$	1.41 <sup>Q</sup> $\mu\text{g/mL}$	1.69 <sup>Q</sup> $\mu\text{g/mL}$
K <sup>+</sup>	1.49 <sup>Q</sup> $\mu\text{g/mL}$	1.22 <sup>Q</sup> $\mu\text{g/mL}$	2.25 <sup>Q</sup> $\mu\text{g/mL}$	2.25 <sup>Q</sup> $\mu\text{g/mL}$	1.12 <sup>Q</sup> $\mu\text{g/mL}$
Ca <sup>2+</sup>	0.78 <sup>Q</sup> $\mu\text{g/mL}$	0.58 <sup>Q</sup> $\mu\text{g/mL}$	1.55 <sup>Q</sup> $\mu\text{g/mL}$	0.78 <sup>Q</sup> $\mu\text{g/mL}$	2.33 <sup>Q</sup> $\mu\text{g/mL}$
Cu <sup>2+</sup>	1.12 <sup>Q</sup> $\mu\text{g/mL}$	0.80 <sup>Q</sup> $\mu\text{g/mL}$	1.06 <sup>Q</sup> $\mu\text{g/mL}$	0.80 <sup>Q</sup> $\mu\text{g/mL}$	1.59 <sup>Q</sup> $\mu\text{g/mL}$
Pb <sup>2+</sup>	1.79 <sup>Q</sup> $\mu\text{g/mL}$	1.79 <sup>Q</sup> $\mu\text{g/mL}$	1.34 <sup>Q</sup> $\mu\text{g/mL}$	2.68 <sup>Q</sup> $\mu\text{g/mL}$	8.05 <sup>Q</sup> $\mu\text{g/mL}$
Ni <sup>2+</sup>	0.48 <sup>Q</sup> $\mu\text{g/mL}$	0.75 <sup>Q</sup> $\mu\text{g/mL}$	1.28 <sup>Q</sup> $\mu\text{g/mL}$	0.32 <sup>Q</sup> $\mu\text{g/mL}$	1.12 <sup>Q</sup> $\mu\text{g/mL}$
Mg <sup>2+</sup>	0.14 <sup>Q</sup> $\mu\text{g/mL}$	0.21 <sup>Q</sup> $\mu\text{g/mL}$	0.21 <sup>Q</sup> $\mu\text{g/mL}$	0.21 <sup>Q</sup> $\mu\text{g/mL}$	0.56 <sup>Q</sup> $\mu\text{g/mL}$
Cl <sup>-</sup>	1.36 <sup>Q</sup> $\mu\text{g/mL}$	1.30 <sup>Q</sup> $\mu\text{g/mL}$	2.17 <sup>Q</sup> $\mu\text{g/mL}$	2.14 <sup>Q</sup> $\mu\text{g/mL}$	2.61 <sup>Q</sup> $\mu\text{g/mL}$
SO <sub>4</sub> <sup>2-</sup>	0.78 <sup>Q</sup> $\mu\text{g/mL}$	1.22 <sup>Q</sup> $\mu\text{g/mL}$	2.09 <sup>Q</sup> $\mu\text{g/mL}$	0.81 <sup>Q</sup> $\mu\text{g/mL}$	2.17 <sup>Q</sup> $\mu\text{g/mL}$
NO <sub>3</sub> <sup>-</sup>	0.54 <sup>Q</sup> $\mu\text{g/mL}$	0.54 <sup>Q</sup> $\mu\text{g/mL}$	0.40 <sup>Q</sup> $\mu\text{g/mL}$	0.81 <sup>Q</sup> $\mu\text{g/mL}$	2.42 <sup>Q</sup> $\mu\text{g/mL}$
PO <sub>4</sub> <sup>3-</sup>	2.83 <sup>Q</sup> $\mu\text{g/mL}$	1.13 <sup>Q</sup> $\mu\text{g/mL}$	1.70 <sup>Q</sup> $\mu\text{g/mL}$	2.83 <sup>Q</sup> $\mu\text{g/mL}$	2.11 <sup>Q</sup> $\mu\text{g/mL}$
CO <sub>3</sub> <sup>2-</sup>	1.21 <sup>Q</sup> $\mu\text{g/mL}$	0.81 <sup>Q</sup> $\mu\text{g/mL}$	1.62 <sup>Q</sup> $\mu\text{g/mL}$	2.02 <sup>Q</sup> $\mu\text{g/mL}$	4.04 <sup>Q</sup> $\text{mg/L}$
p.Cresol	1.50 <sup>Q</sup> $\mu\text{g/mL}$	2.50 <sup>Q</sup> $\mu\text{g/mL}$	1.50 <sup>Q</sup> $\mu\text{g/mL}$	2.5 <sup>Q</sup> $\mu\text{g/mL}$	5 <sup>Q</sup> $\mu\text{g/mL}$
m.Cresol	2 <sup>Q</sup> $\mu\text{g/mL}$	3.50 <sup>Q</sup> $\mu\text{g/mL}$	2 <sup>Q</sup> $\mu\text{g/mL}$	2.5 <sup>Q</sup> $\mu\text{g/mL}$	3 <sup>Q</sup> $\mu\text{g/mL}$
$\alpha$ -naphthol	1 <sup>Q</sup> $\mu\text{g/mL}$	1.50 <sup>Q</sup> $\mu\text{g/mL}$	1.5 <sup>Q</sup> $\mu\text{g/mL}$	1.5 <sup>Q</sup> $\mu\text{g/mL}$	2 <sup>Q</sup> $\mu\text{g/mL}$
Pyrogallol	0.5 <sup>Q</sup> $\mu\text{g/mL}$	0.50 <sup>Q</sup> $\mu\text{g/mL}$	1 <sup>Q</sup> $\mu\text{g/mL}$	1 <sup>Q</sup> $\mu\text{g/mL}$	1 <sup>Q</sup> $\mu\text{g/mL}$

<sup>Q</sup> : quenching effect

**Table 3S.** Determination of Phenolic Compounds in different types of wastewater samples using Standard Addition method.

	<b>Added (<math>\mu\text{g/mL}</math>)</b>	<b>Found (<math>\mu\text{g/mL}</math>)</b>	<b>Recovery (%)</b>	<b>RSD (n=3) %</b>
<b>Sample (I)</b>	<b>10</b>	<b>10.8</b>	<b>108%</b>	<b>1.73%</b>
	<b>15</b>	<b>15.2</b>	<b>102%</b>	<b>1.10%</b>
	<b>20</b>	<b>18.9</b>	<b>95%</b>	<b>0.35%</b>
<b>Sample (II)</b>	<b>10</b>	<b>9.5</b>	<b>95%</b>	<b>1.82%</b>
	<b>15</b>	<b>13.7</b>	<b>91%</b>	<b>2.96%</b>
	<b>20</b>	<b>21.1</b>	<b>105%</b>	<b>1.92%</b>
<b>Sample (III)</b>	<b>10</b>	<b>10.4</b>	<b>104%</b>	<b>1.91%</b>
	<b>15</b>	<b>14.9</b>	<b>99.9%</b>	<b>2.51%</b>
	<b>20</b>	<b>19.6</b>	<b>98.0%</b>	<b>2.00%</b>