

Supplementary Information
for
**Controlled reduction of isocyanates to formamides using monomeric
magnesium**

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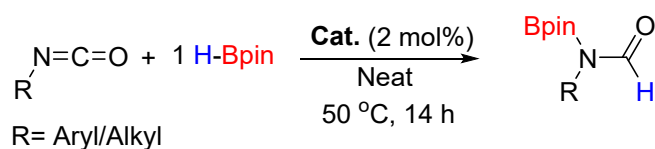
All reactions were carried out under argon atmosphere using Schlenk techniques or inside a MBraun glove box. Catalysts **1** was prepared according to our previous paper.^{S1}

Pinacolborane (HBpin), isocyanates were purchased from Sigma-Aldrich, TCI Chemicals and used without further purification. C₆D₆, toluene-d₈, CD₃CN and CDCl₃ were purchased from Sigma-Aldrich, were degassed by three freeze-pump-thaw cycles and stored over molecular sieves.

¹H, ¹³C{¹H} and ¹¹B NMR spectra were recorded on Bruker AV-200 MHz, AV-400 MHz and

AV-500 MHz and referenced to the resonances of the internal standard with respect to the solvent used. HRMS spectra were obtained using Thermo scientific Q-Exactive instrument.

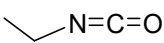
❖ **General catalytic procedure for the synthesis of N-boryl formamide:**

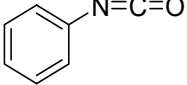
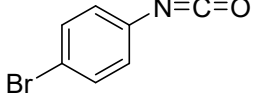
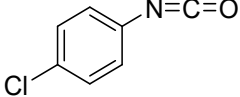
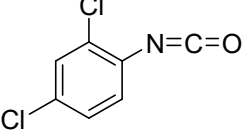
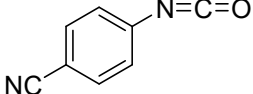
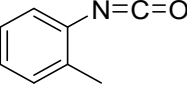
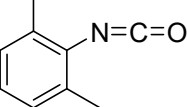
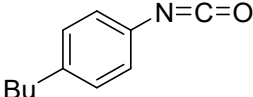
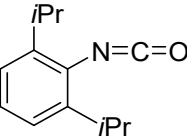
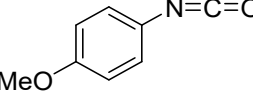
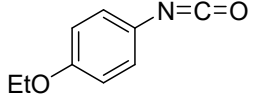


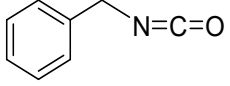
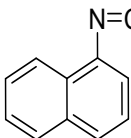
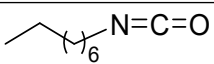
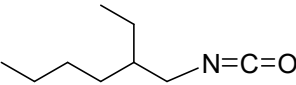
Scheme S1. General catalytic procedure for the synthesis of N-borylated formamide from phenyl isocyanate.

Ethyl isocyanate (0.250 mmol), pinacolborane (1.1 equiv., 0.275 mmol), catalyst (**1**, 2.0 mol%) were charged in a Schlenk tube/ nmr tube inside the glove box. The reaction mixture was allowed to heat at 50 °C for 14 h. Upon completion of the reaction, the volatiles were removed using vacuum in a Schlenk line and mesitylene (0.25 mmol) was added as the internal standard, while making the NMR in appropriate deuterated solvent. The progress of the reaction was monitored by the ¹H NMR spectroscopy, which indicated the completion of the reaction by the appearance of a new formamide -NCHO resonance.

❖ **Table S1. Isocyanate substrate scope for N-boryl formamide:**

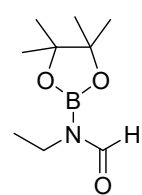
Entry	Substrate	Temperature (°C)	Time (h)	Catalyst 1 (mol%)	NMR Yield (%)	Product	Selectivity (Formamide: dihydroborated:amine)
1.	 N=C=O	50	14	2	93	2a	100:0:0

2.		50	14	2	20	2b	20:0:15
3.		50	14	2	58	2c	100:0:0
4.		50	14	2	55	2d	100:0:0
5.		50	14	2	30	2e	100:0:0
6.		50	14	2	41	2f	41:8:7
7.		50	14	2	69	2g	100:0:0
8.		50	14	2	43	2h	100:0:0
9.		50	14	2	62	2i	100:0:0
10.		50	14	2	26	2j	26:24:0
11.		50	14	2	52	2k	100:0:0
12.		50	14	2	56	2l	100:0:0

13.		50	14	2	59	2m	59:18:0
14.		50	14	2	70	2n	100:0:0
15.		50	14	2	59	2o	100:0:0
16.		50	14	2	53	2p	100:0:0

❖ Analytical data of N-boryl formamide of corresponding isocyanates with their NMR spectra:

N-ethyl-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2a): ^1H NMR (400 MHz,



CHLOROFORM-*d*) δ 8.66 (s, 1H, NCHO), 3.34-3.30 (t, 2H, NCH₂CH₃), 1.52-1.47 (q, 3H, NCH₂CH₃), 1.30 (s, 12H, CH₃) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl₃, 101 MHz, 298 K): δ 165.92 (s, 1 C), 84.01 (s, 1 C), 39.93 (s, 1 C), 29.51 (s, 1 C), 26.53 (s, 1 C), 24.79 (s, 1 C), 24.46 (s, 1 C), 24.41 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

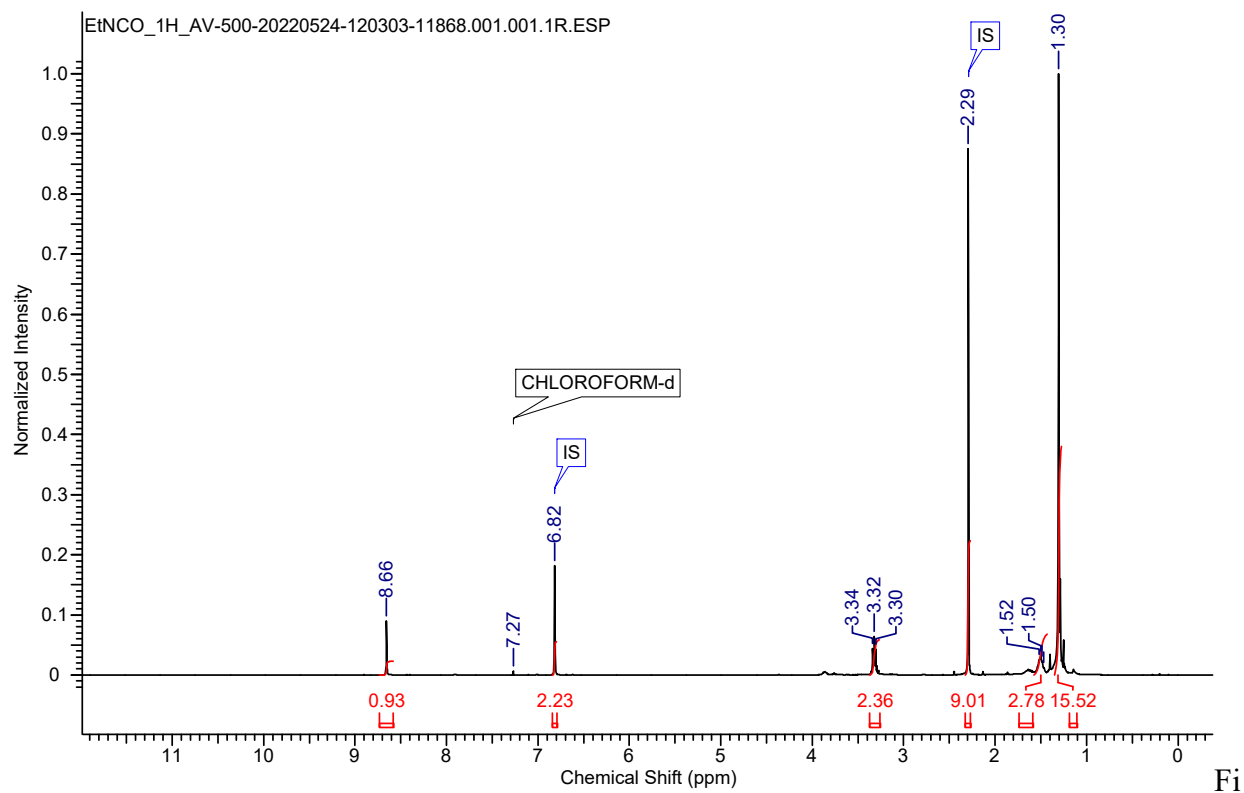


Figure S1. ^1H NMR spectrum of **2a** (CDCl_3 , 500 MHz, 298 K).

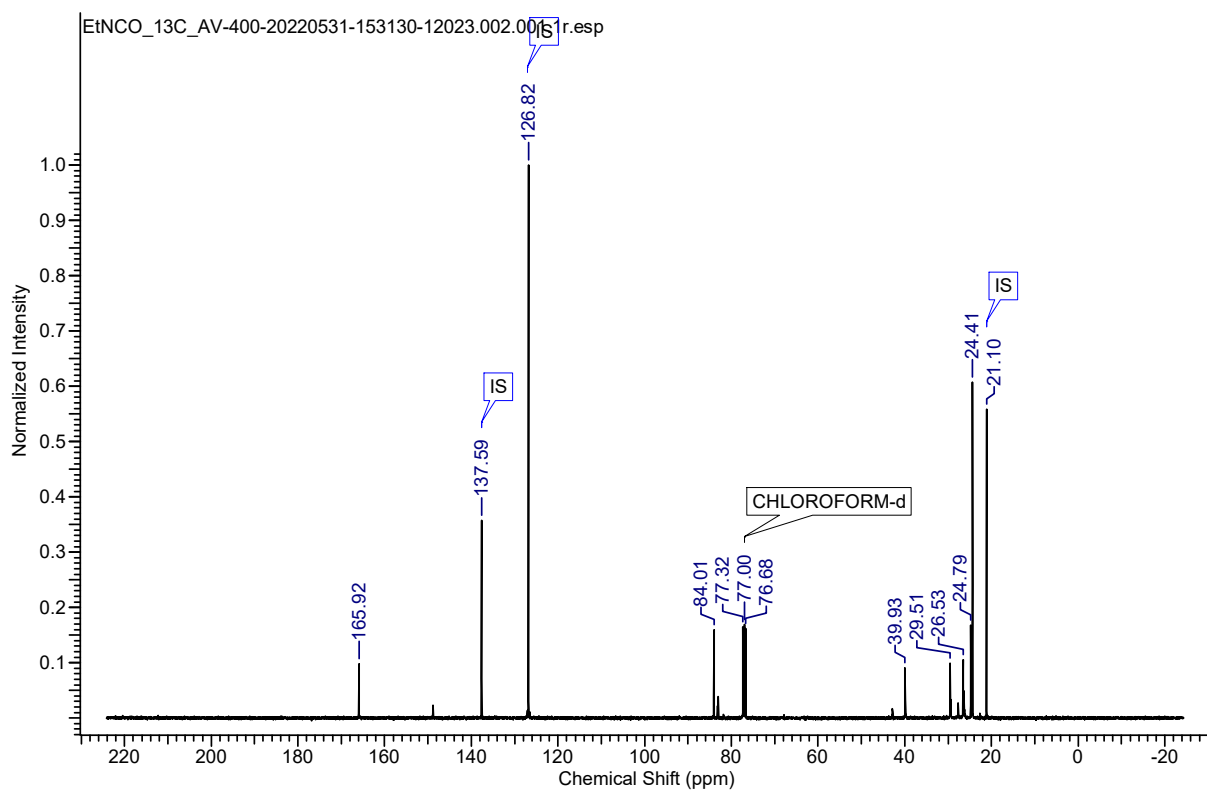
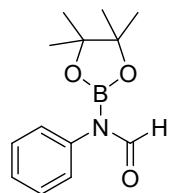


Figure S2. ^{13}C NMR spectrum of **2a** (CDCl_3 , 101 MHz, 298 K).

N-phenyl-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2b): ^1H NMR (400



MHz, CHLOROFORM-*d*) δ 8.71 (s, 1 H, NCHO), 7.29-7.20 (m, 5H, PhH), 2.87 (s, 3H, NCH₃), 1.08 (s, 12H) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl₃, 101 MHz, 298 K): δ 165.31 (s, 1 C), 148.58 (s, 1 C), 147.40 (s, 1 C), 133.58 (s, 1 C), 129.20 (s, 1 C), 128.84 (s, 1 C), 128.69 (s, 1 C), 128.40 (s, 1 C), 128.34 (s, 1 C), 128.24 (s, 1 C), 127.27 (s, 1 C), 127.05 (s, 1 C), 126.97 (s, 1 C), 120.51 (s, 1 C), 119.68 (s, 1 C), 118.74 (s, 1 C), 84.43 (s, 1 C), 82.98 (s, 1 C), 82.57 (s, 1 C), 81.78 (s, 1 C), 34.19 (s, 1 C), 24.55 (s, 1 C), 24.42 (s, 1 C), 24.34 (s, 1 C), 24.09 (s, 1 C), 22.60 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

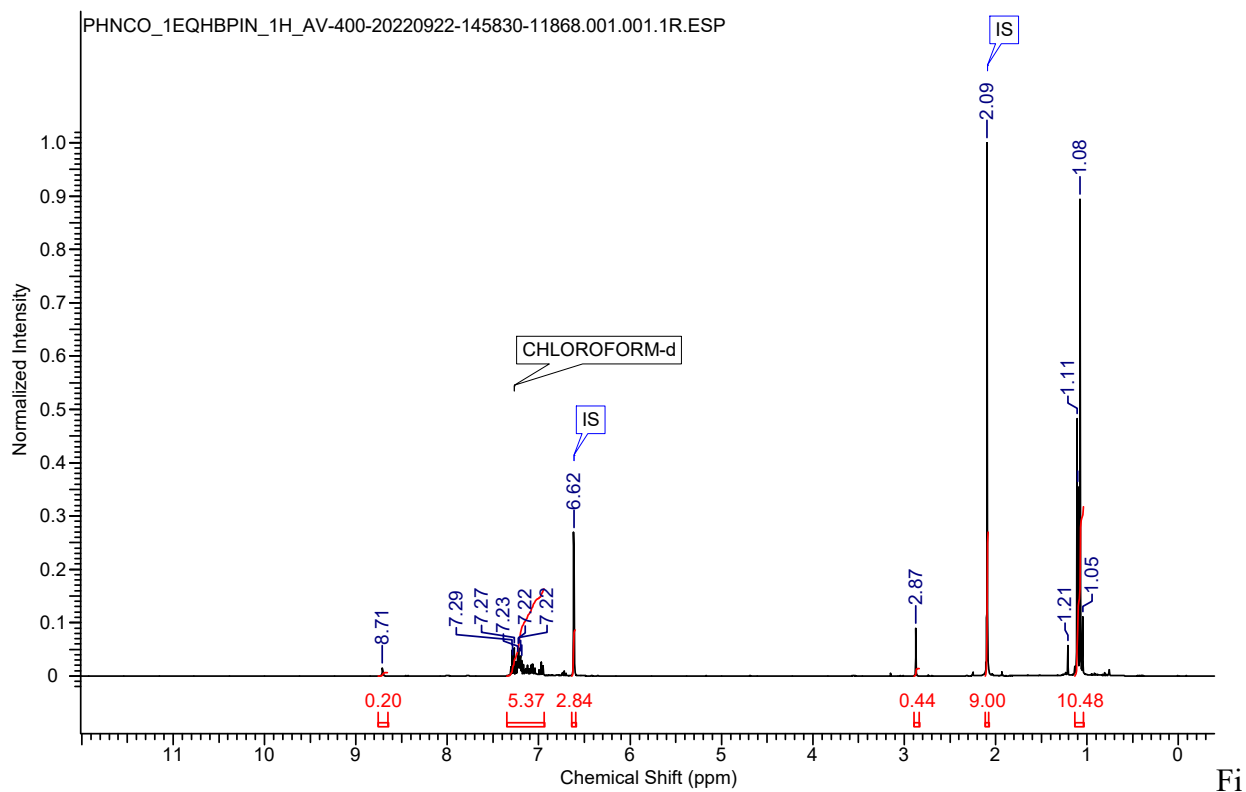


Figure S3. ^1H NMR spectrum of **2b** (CDCl₃, 400 MHz, 298 K).

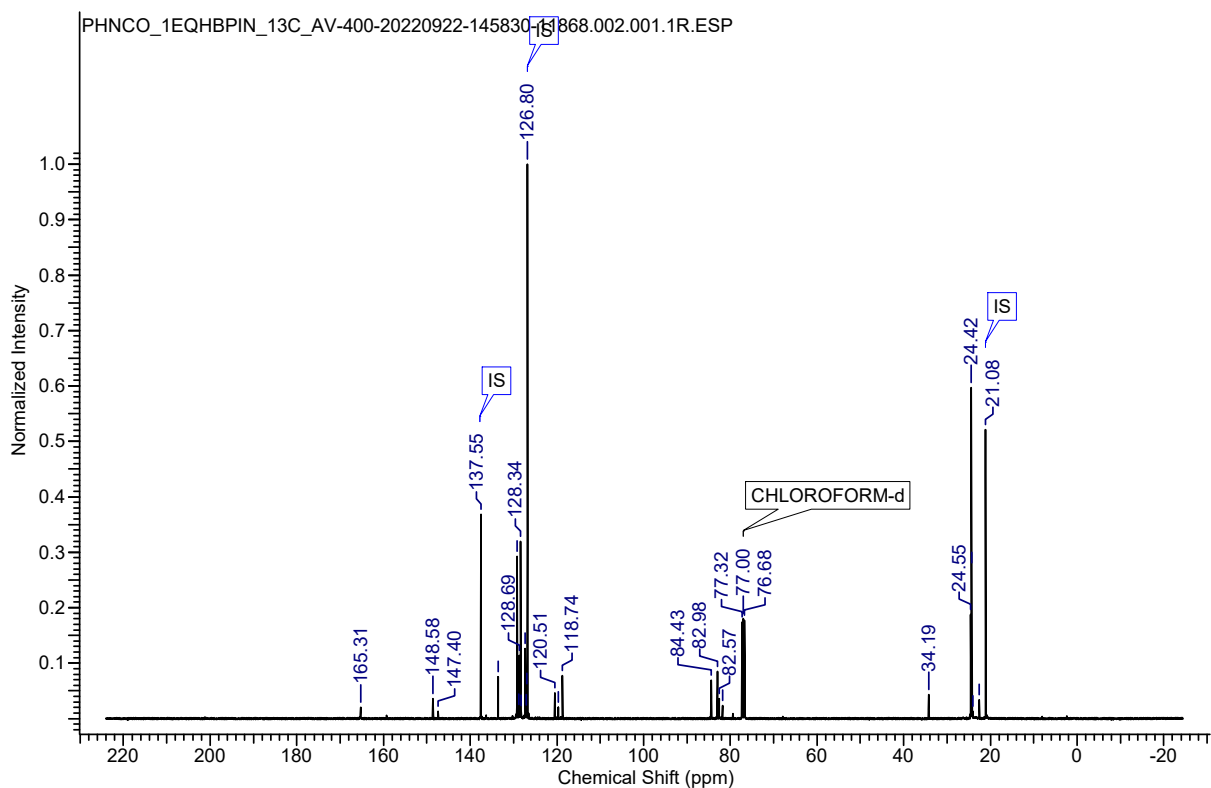
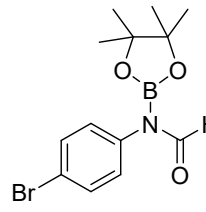


Figure S4. ^{13}C NMR spectrum of **2b** (CDCl_3 , 101 MHz, 298 K).


N-(4-bromophenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2c): ^1H NMR (CDCl_3 , 400 MHz, 298 K): δ 8.89 (s, 1H, NCHO), 7.53-7.51 (d, 2H, PhH), 7.08-7.06 (d, 2H, PhH), 1.34 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 50.28 MHz, 298 K): δ 165.05 (s, 1 C), 132.51 (s, 1 C), 131.86 (s, 1 C), 131.69 (s, 1 C), 130.03 (s, 1 C), 128.97 (s, 1 C), 120.67 (s, 1 C), 120.32 (s, 1 C), 84.70 (s, 1 C), 83.04 (s, 1 C), 24.56 (s, 1 C), 24.44 (s, 1 C), 24.39 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

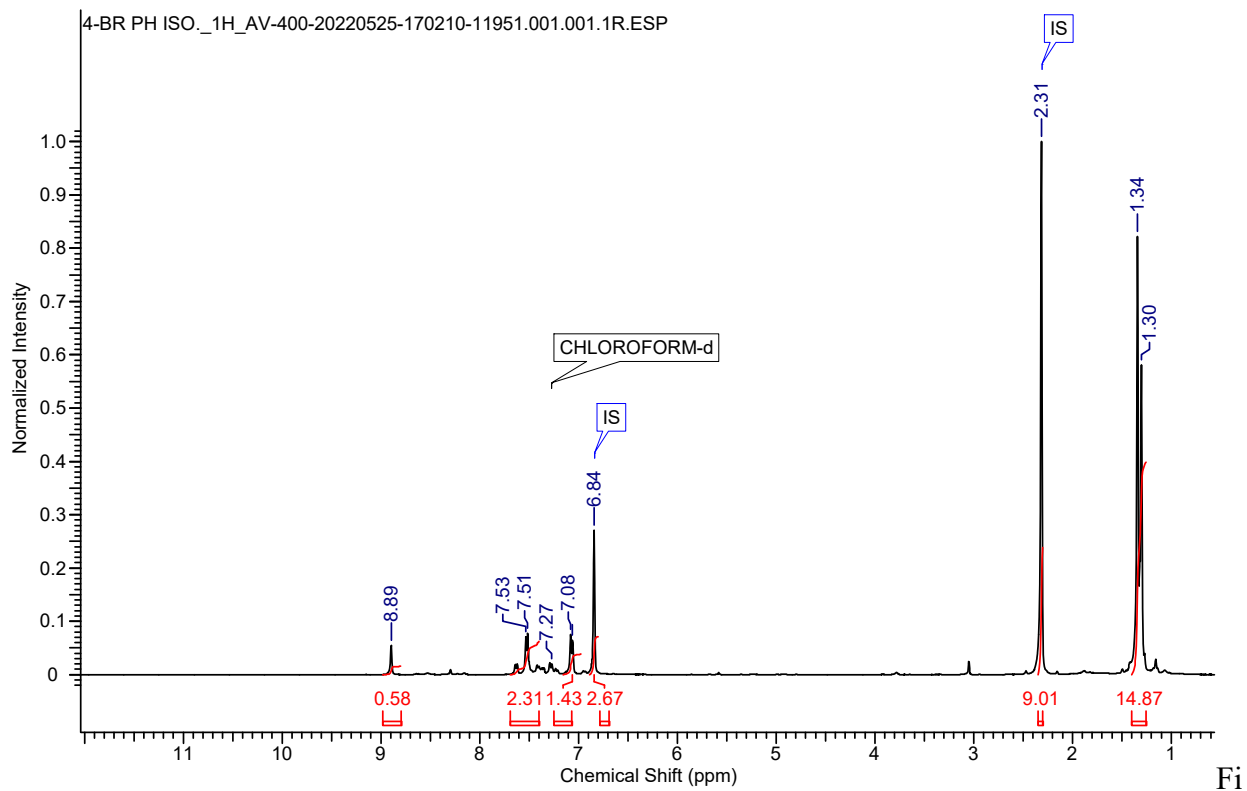


Figure S5. ^1H NMR spectrum of **2c** (CDCl_3 , 400 MHz, 298 K).

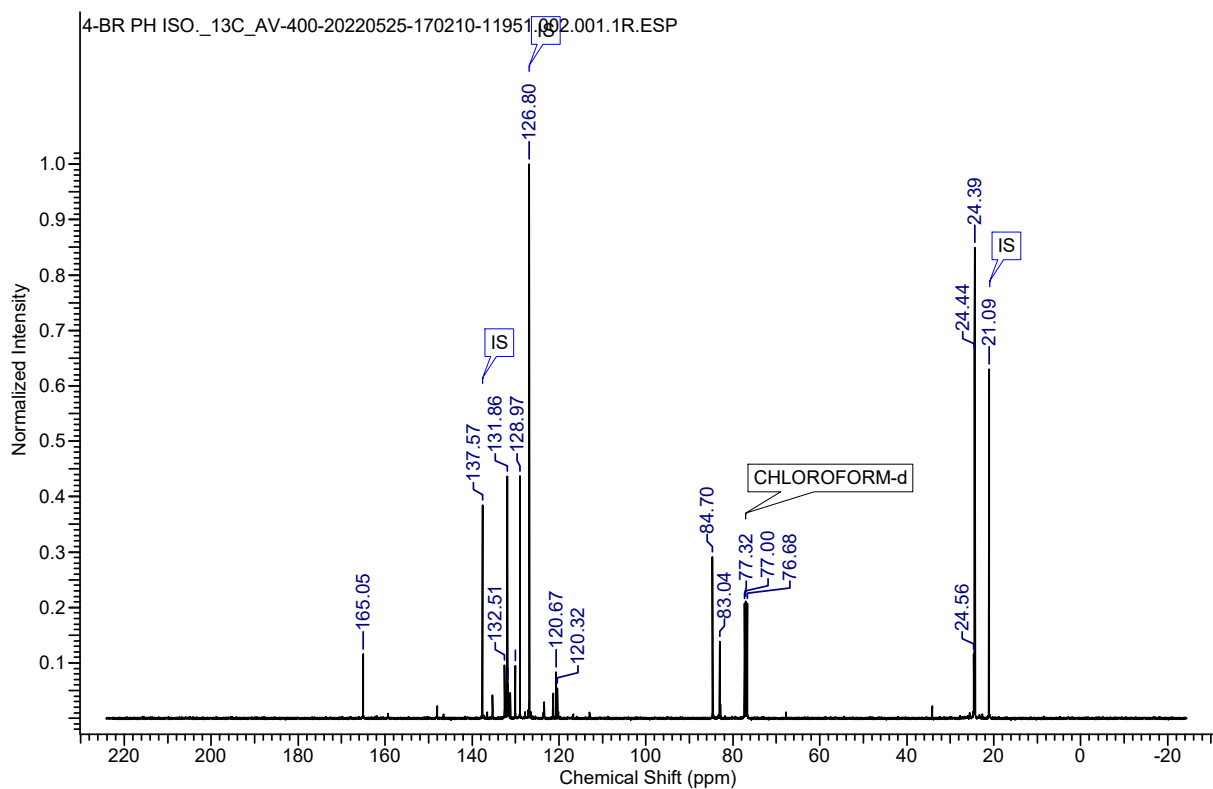
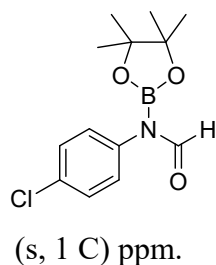


Figure S6. ^{13}C NMR spectrum of **2c** (CDCl_3 , 101 MHz, 298 K).

N-(4-chlorophenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2d): ^1H



NMR (CDCl_3 , 200 MHz, 298 K): δ 8.84 (s, 1H, NCHO), 7.33-7.30 (d, 2H, PhH), 7.08-7.06 (d, 2H, PhH), 1.29 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.10 (s, 1 C), 129.72 (s, 1 C), 129.51 (s, 1 C), 128.88 (s, 1 C), 128.60 (s, 1 C), 84.67 (s, 1 C), 83.01 (s, 1 C), 24.44 (s, 1 C), 24.37 (s, 1 C), 24.28 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

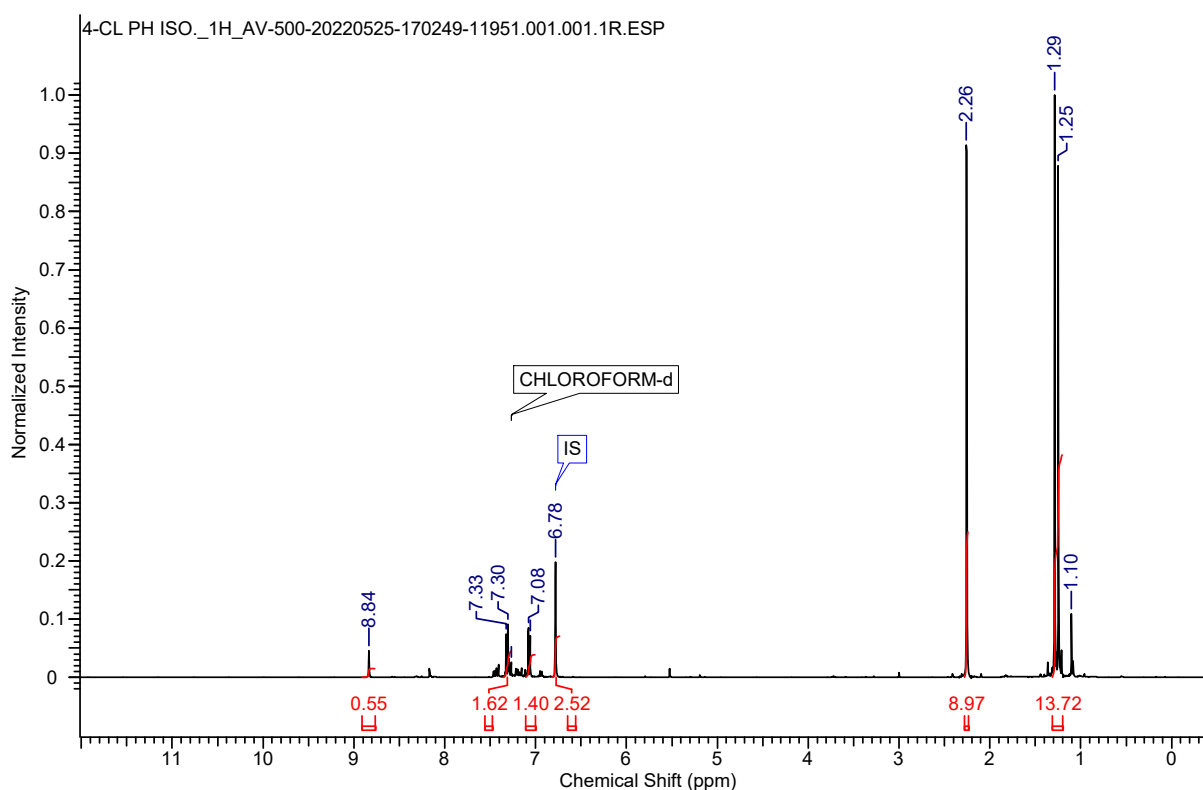


Figure S7. ^1H NMR spectrum of **2d** (CDCl_3 , 500 MHz, 298 K).

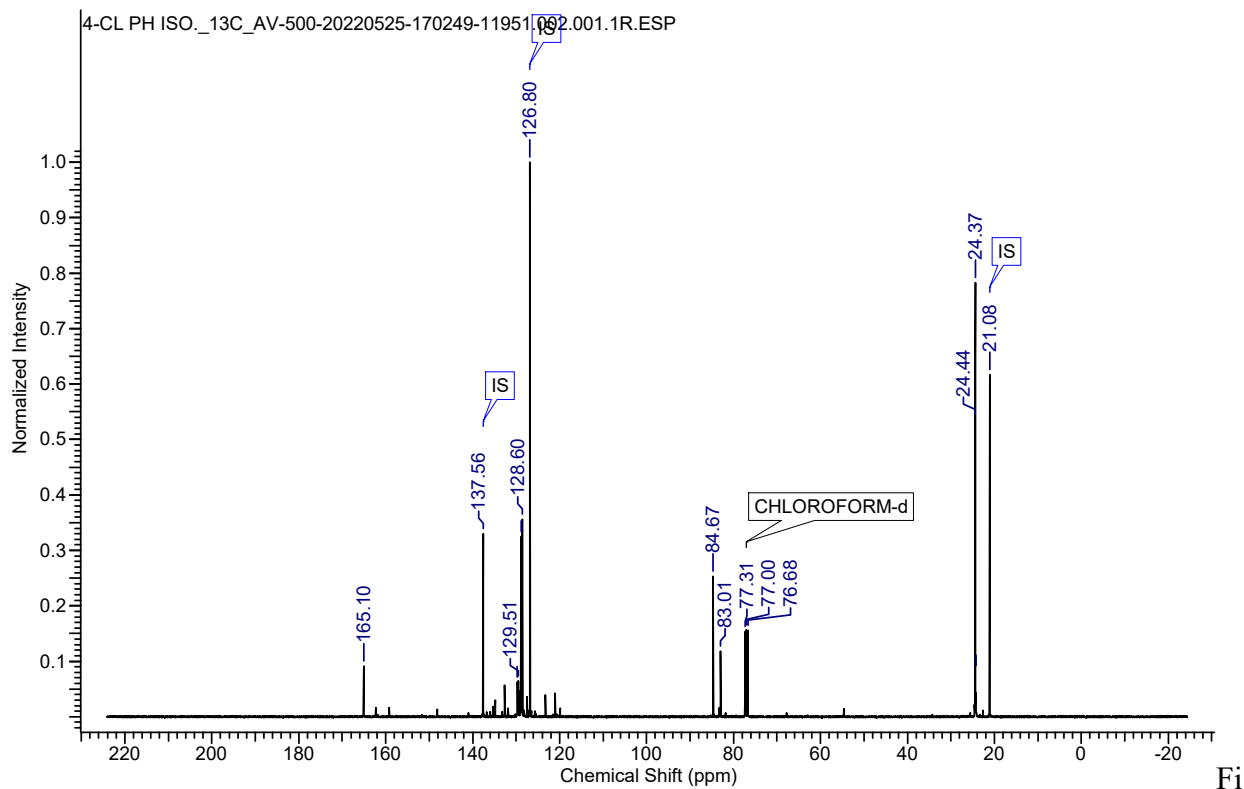
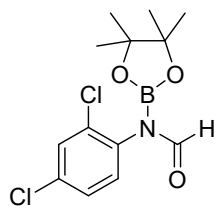


Figure S8. ^{13}C NMR spectrum of **2d** (CDCl_3 , 101 MHz, 298 K).

N-(2,4-dichlorophenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2e): ^1H



NMR (CDCl_3 , 200 MHz, 298 K): δ 8.91 (s, 1H, NCHO), 7.53-7.32 (m, 3H, PhH), 1.37 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.10 (s, 1 C), 129.72 (s, 1 C), 129.51 (s, 1 C), 128.88 (s, 1 C), 128.60 (s, 1 C), 84.67 (s, 1 C), 83.01 (s, 1 C), 24.44 (s, 1 C), 24.37 (s, 1 C), 24.28 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

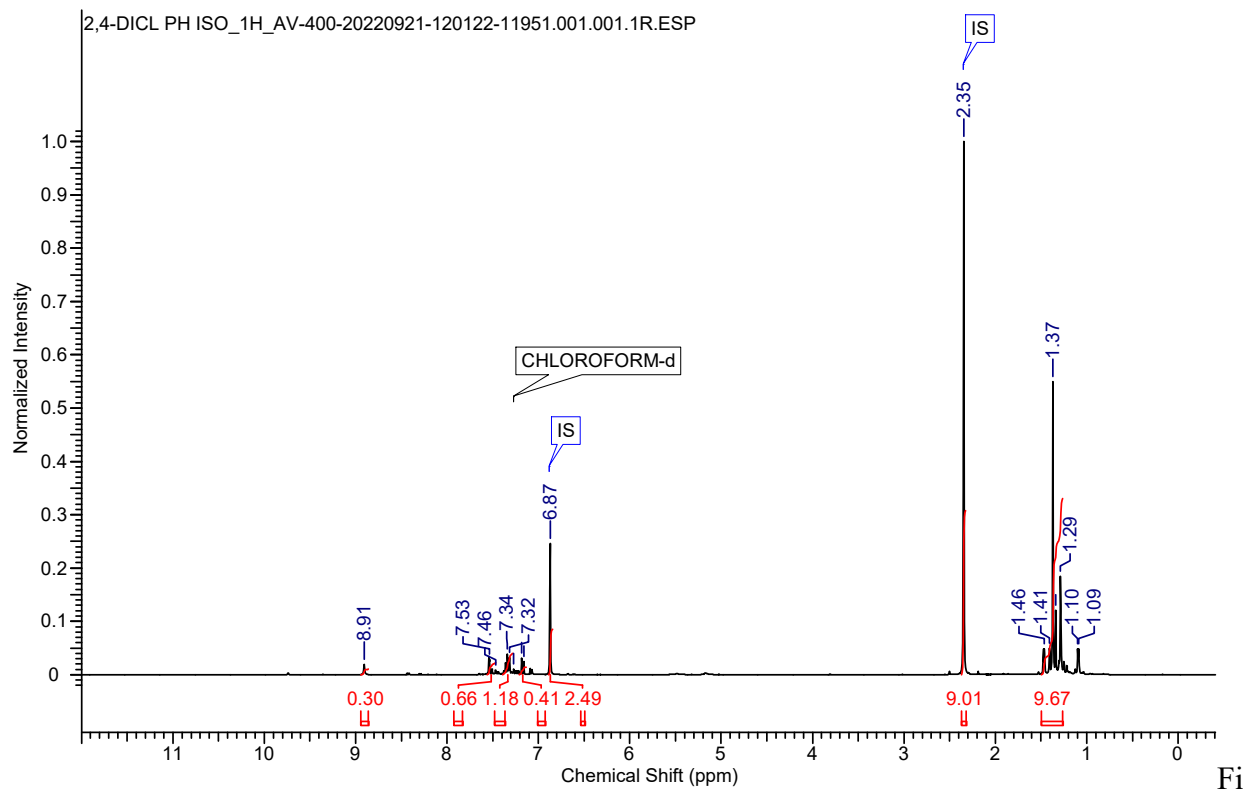


Figure S9. ^1H NMR spectrum of **2e** (CDCl_3 , 400 MHz, 298 K).

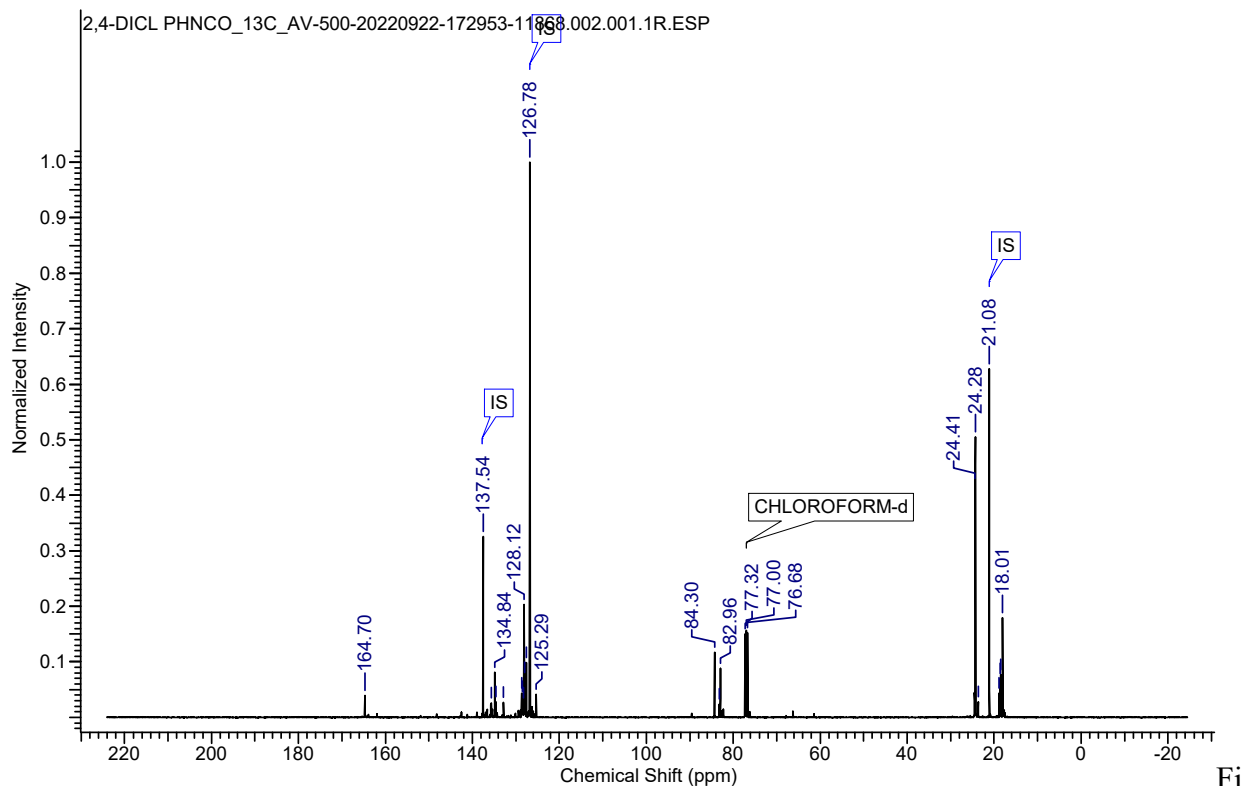


Figure S10. ^{13}C NMR spectrum of **2e** (CDCl_3 , 101 MHz, 298 K).

N-(4-cyanophenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2f): ^1H NMR (CDCl_3 , 500 MHz, 298 K): δ 8.74 (s, 1H, NCHO), 7.52-7.50 (d, 2H, PhH), 7.18-7.16 (d, 2H, PhH), 5.57 (s, 2H, $\text{NCH}_2\text{-OBpin}$), 2.92 (s, 3H, NCH_3), 1.13 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 164.54 (s, 1 C), 140.55 (s, 1 C), 128.05 (s, 1 C), 125.56 (s, 1 C), 120.95 (s, 1 C), 119.65 (s, 1 C), 118.32 (s, 1 C), 117.98 (s, 1 C), 110.51 (s, 1 C), 84.95 (s, 1 C), 83.93 (s, 1 C), 83.29 (s, 1 C), 82.98 (s, 1 C), 24.50 (s, 1 C), 24.38 (s, 1 C), 24.34 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

Figure S12. ^{13}C NMR spectrum of **2f** (CDCl_3 , 101 MHz, 298 K).

N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-N-(o-tolyl)formamide (2g): ^1H NMR (CDCl_3 , 400 MHz, 298 K): δ 8.96 (s, 1H, NCHO), 7.31-7.28 (m, 3H, PhH), 7.08-7.06 (m, 1H, PhH), 2.24 (s, 3H, NCH_3), 1.35 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.01 (s, 1 C), 134.90 (s, 1 C), 130.61 (s, 1 C), 128.12 (s, 1 C), 127.74 (s, 1 C), 126.52 (s, 1 C), 84.35 (s, 1 C), 82.99 (s, 1 C), 24.41 (s, 1 C), 24.25 (s, 1 C), 17.72 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

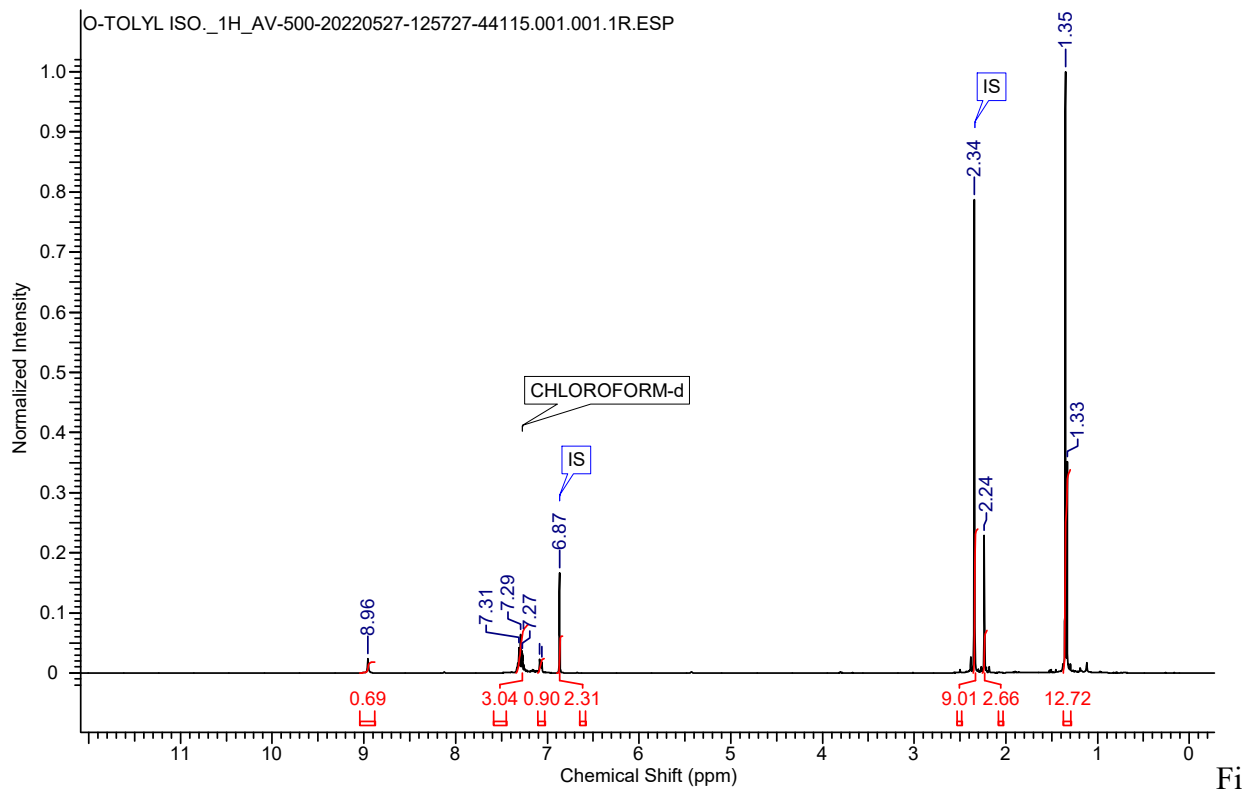


Figure S13. ^1H NMR spectrum of **2g** (CDCl_3 , 500 MHz, 298 K).

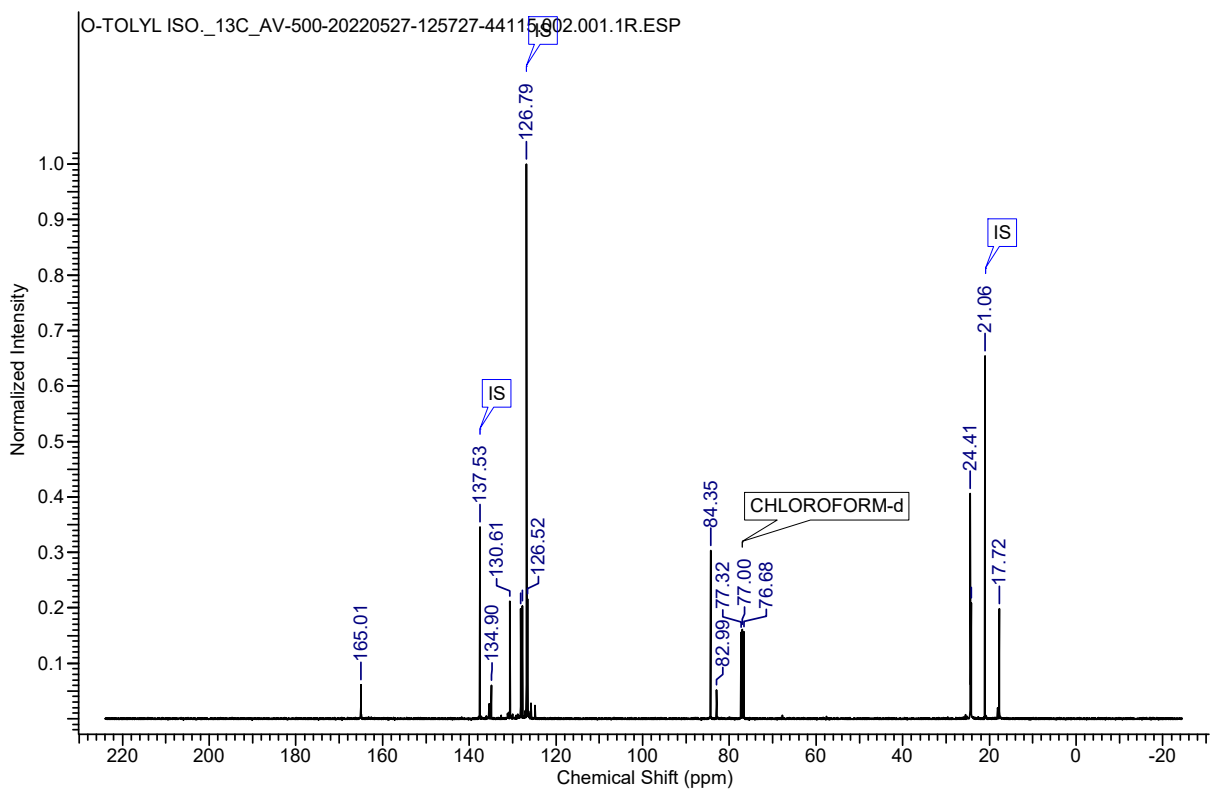
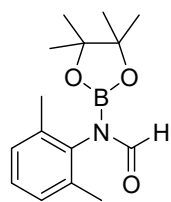


Figure S14. ^{13}C NMR spectrum of **2g** (CDCl_3 , 101 MHz, 298 K).

N-(2,6-dimethylphenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2h): ^1H



NMR (CDCl_3 , 500 MHz, 298 K): δ 8.99 (s, 1H, NCHO), 7.20-7.07 (m, 3H, PhH), 2.41 (s, 3H, PhCH₃), 2.23 (s, 3H, PhCH₃), 1.37 (s, 12H, CH₃) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 164.70 (s, 1 C), 134.84 (s, 1 C), 128.12 (s, 1 C), 127.94 (s, 1 C), 127.59 (s, 1 C), 84.30 (s, 1 C), 82.96 (s, 1 C), 24.41 (s, 1 C), 24.28 (s, 1 C), 18.59 (s, 1 C), 18.41 (s, 1 C), 18.01 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

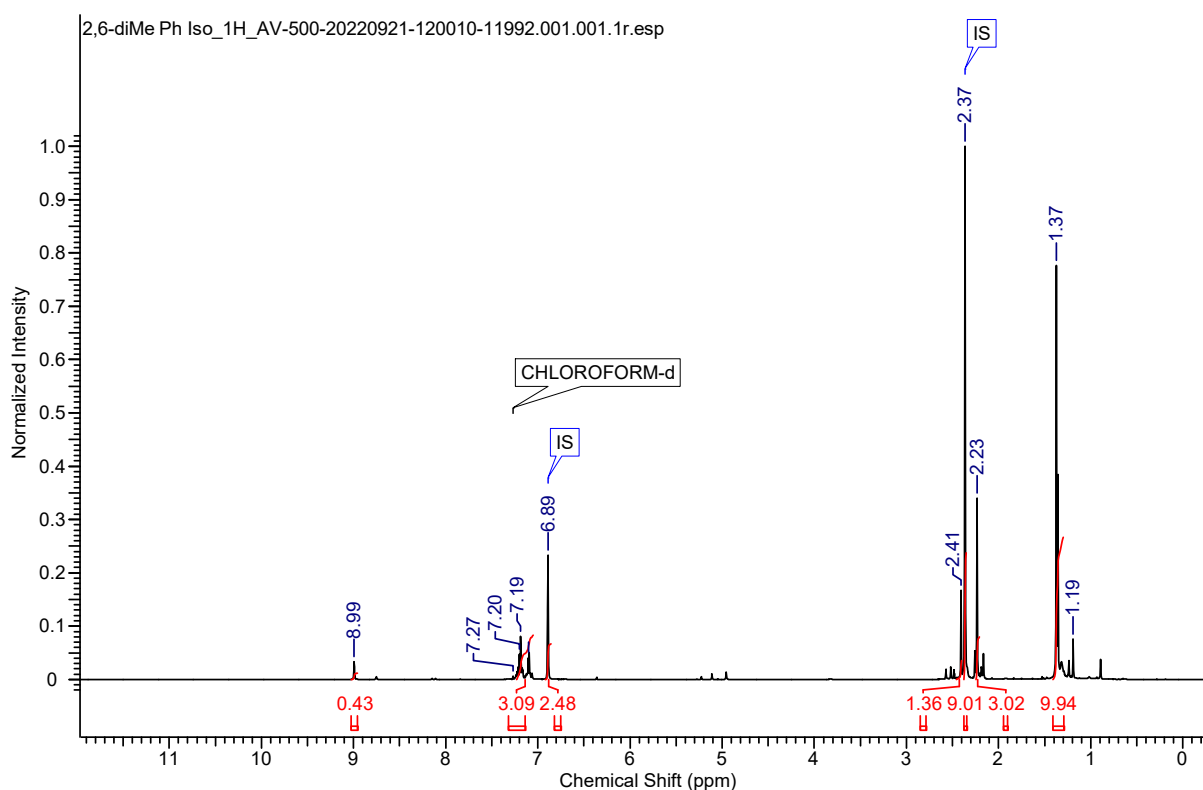


Figure S15. ^1H NMR spectrum of **2h** (CDCl_3 , 500 MHz, 298 K).

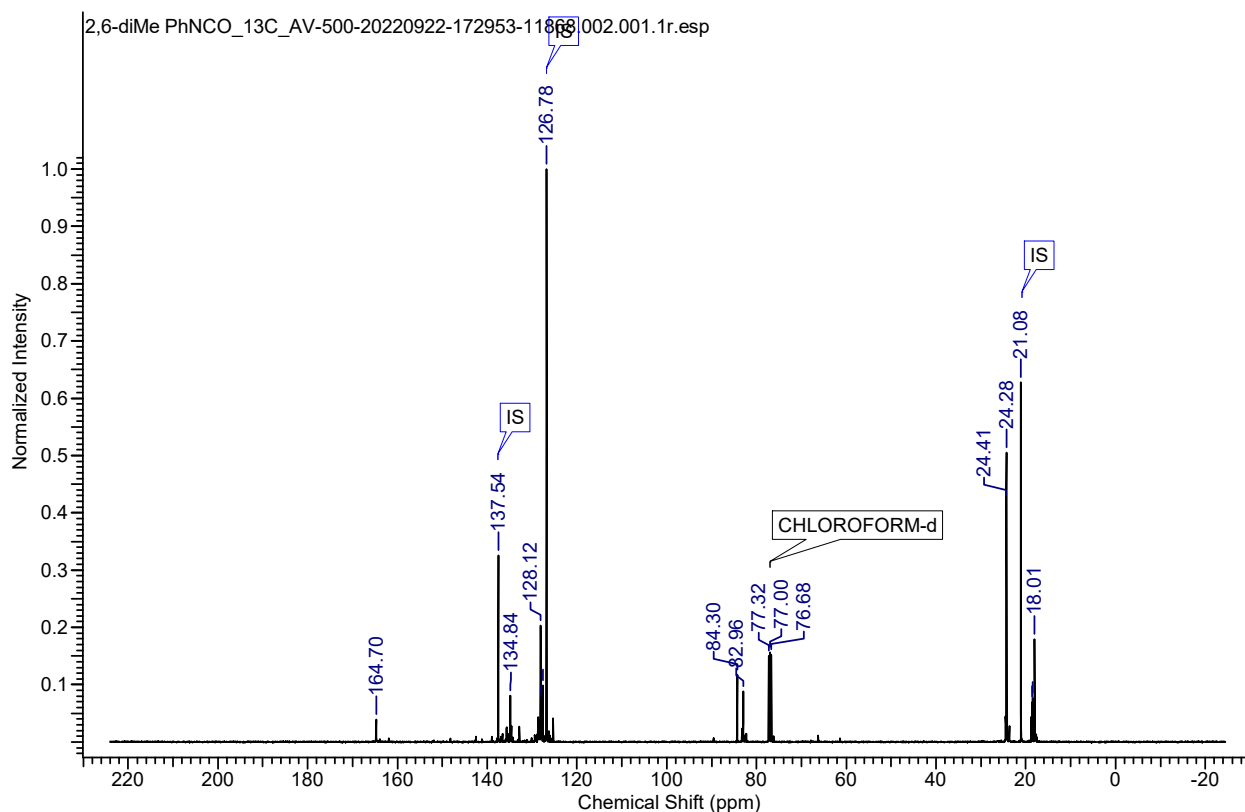


Figure S16. ^{13}C NMR spectrum of **2h** (CDCl_3 , 101 MHz, 298 K).

N-(4-butylphenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2i): ^1H NMR

(CDCl_3 , 100 MHz, 298 K): δ ppm 8.88 (s, 1H, NCHO), 7.19-7.17 (d, 2H, PhH), 7.05-7.03 (d, 2H, PhH), 2.62-2.58 (t, 2H, Bu), 1.62-1.59 (t, 2H, Bu), 1.39-1.35 (t, 2H, Bu), 1.30 (s, 12H, CH_3), 0.95-0.92 (t, 3H, Bu) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.48 (s, 1 C), 141.51 (s, 1 C), 128.67 (s, 1 C), 126.90 (s, 1 C), 84.36 (s, 1 C), 82.98 (s, 1 C), 35.16 (s, 1 C), 33.28 (s, 1 C), 24.54 (s, 1 C), 24.41 (s, 1 C), 24.34 (s, 1 C), 22.34 (s, 1 C), 13.83 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data. ^{S2}

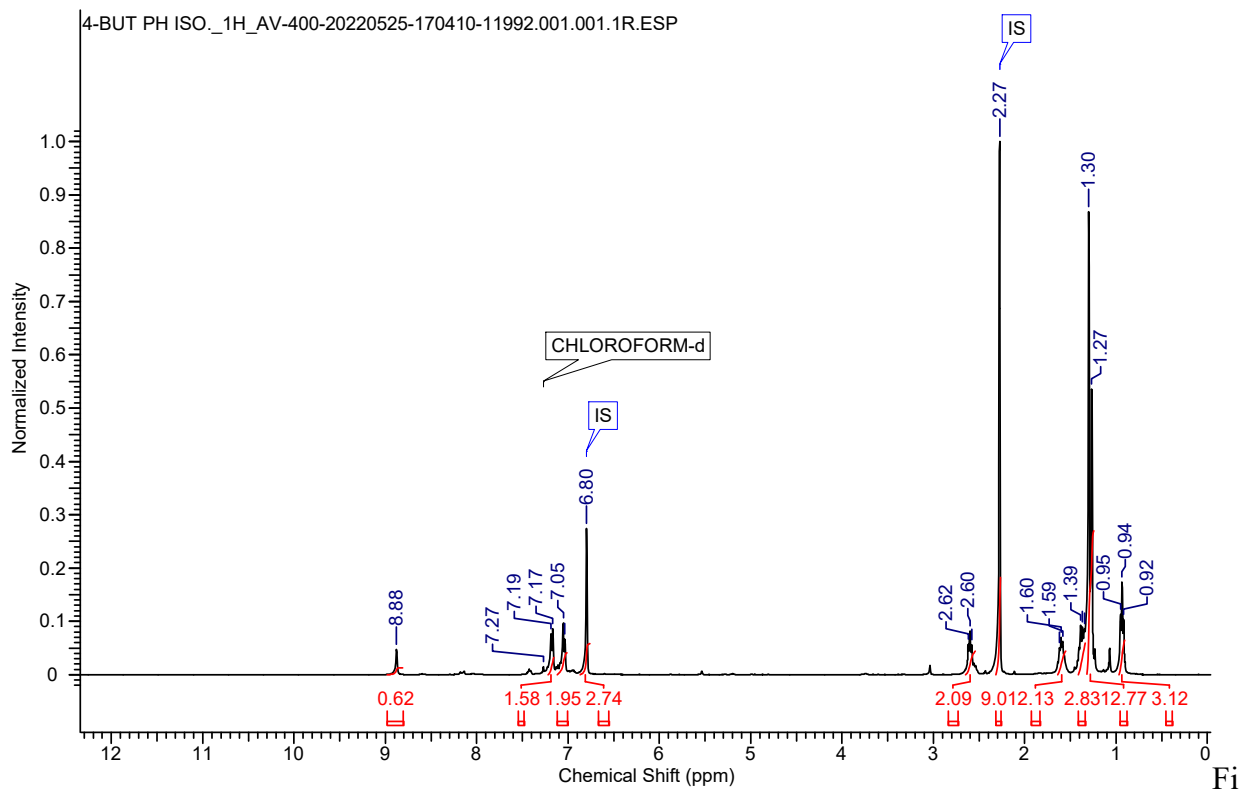


Figure S17. ^1H NMR spectrum of **2i** (CDCl_3 , 400 MHz, 298 K).

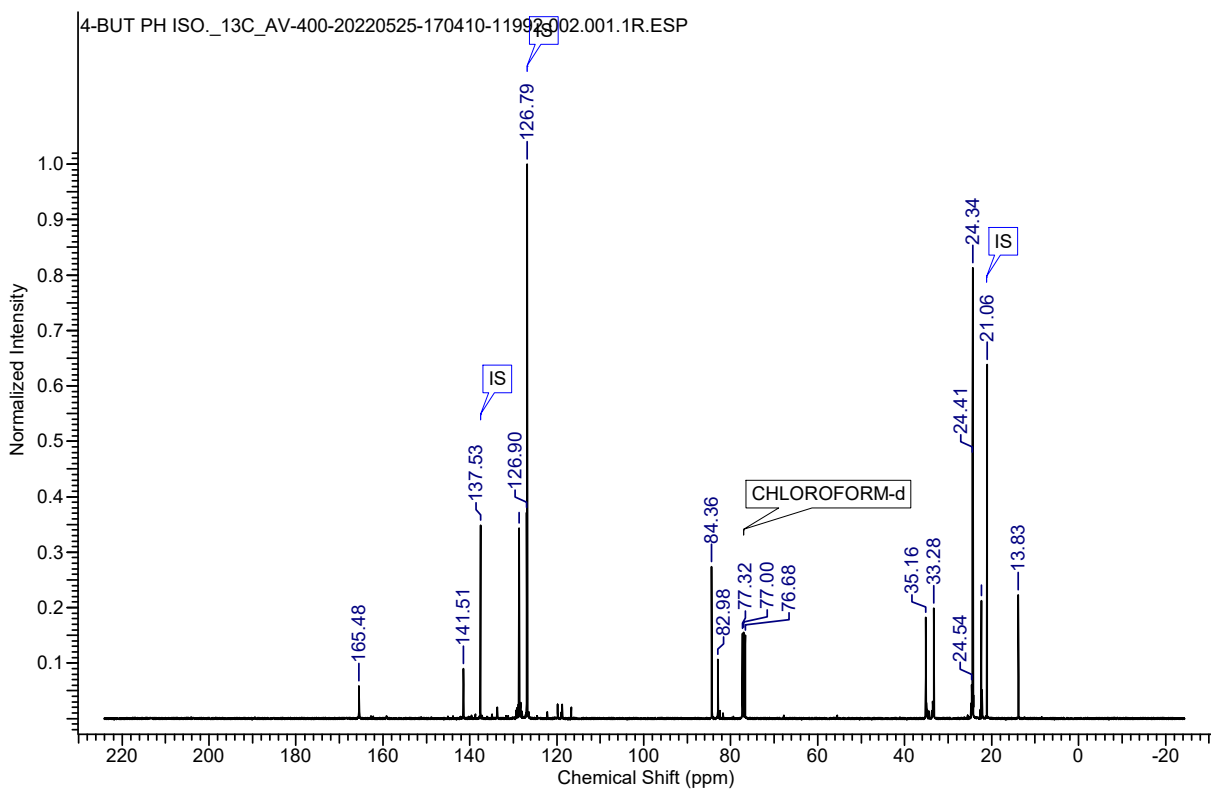
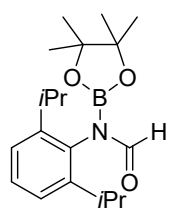


Figure S18. ^{13}C NMR spectrum of **2i** (CDCl_3 , 101 MHz, 298 K).

N-(2,6-diisopropylphenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2j):



^1H NMR (CDCl_3 , 500 MHz, 298 K): δ 8.91 (s, 1H, NCHO), 7.15-7.08 (m, 3H, PhH), 4.97 (s, 2H, NCH₂), 3.27-3.22 {q, 1H, CH-(CH₃)₂}, 2.83-2.78 (q, 1H, CH-(CH₃)₂), 1.22 (s, 12H, CH₃), 1.15-1.13 {d, 6H, CH-(CH₃)₂}, 1.04-1.02 {d, 6H, CH-(CH₃)₂} ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.45 (s, 1 C), 147.31 (s, 1 C), 145.22 (s, 1 C), 142.92 (s, 1 C), 128.37 (s, 1 C), 126.90 (s, 1 C), 123.46 (s, 1 C), 123.35 (s, 1 C), 123.28 (s, 1 C), 84.39 (s, 1 C), 82.73 (s, 1 C), 82.21 (s, 1 C), 77.07 (s, 1 C), 28.65 (s, 1 C), 27.98 (s, 1 C), 24.81 (s, 1 C), 24.45 (s, 1 C), 24.34 (s, 1 C), 24.33 (s, 1 C), 23.78 (s, 1 C), 23.57 (s, 1 C), 22.76 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

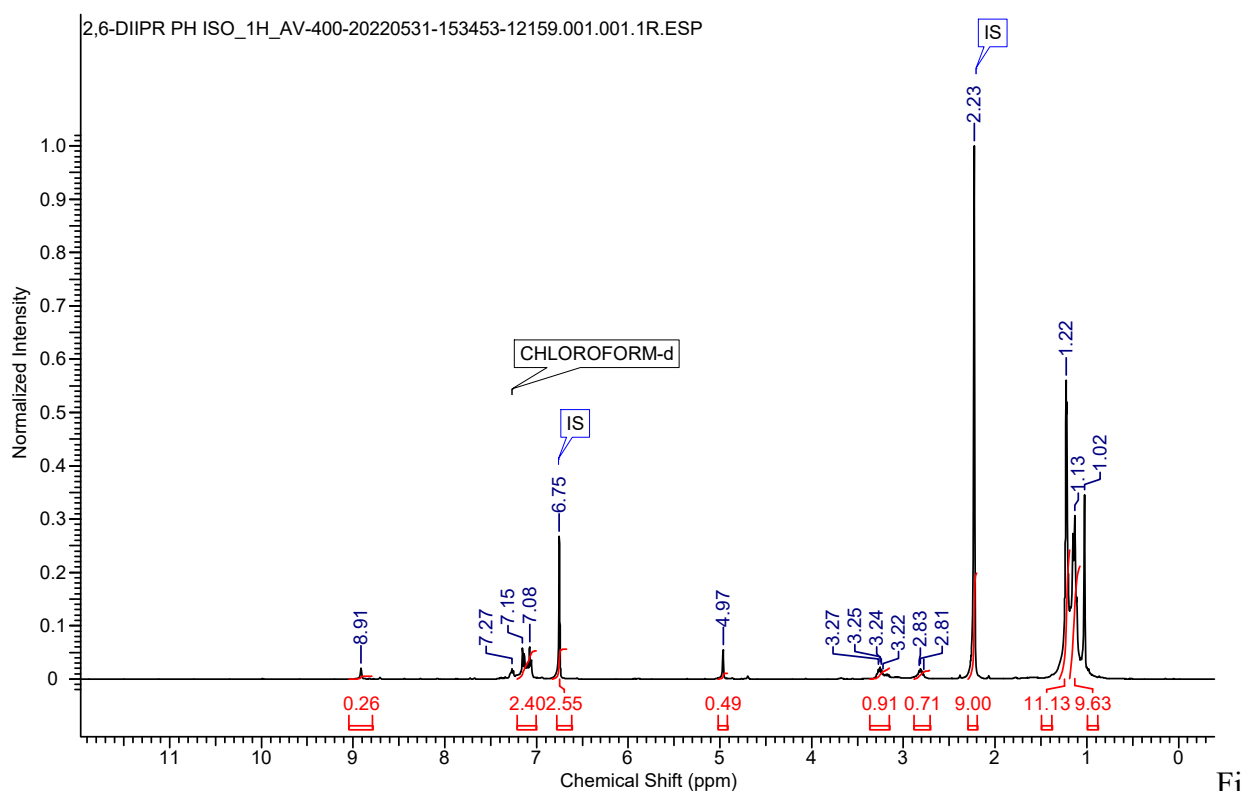


Figure S19. ^1H NMR spectrum of **2j** (CDCl_3 , 400 MHz, 298 K).

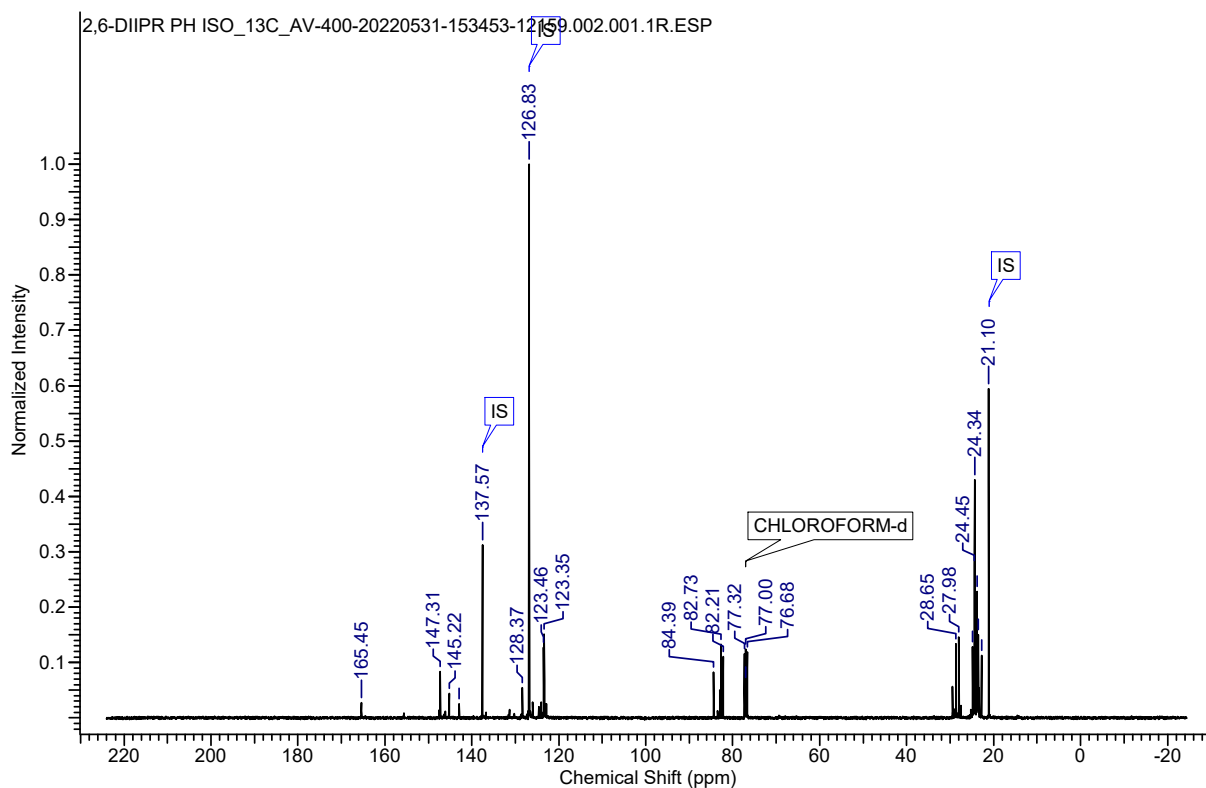
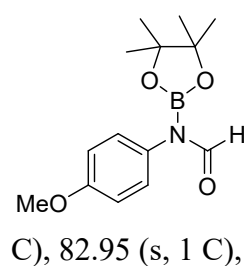


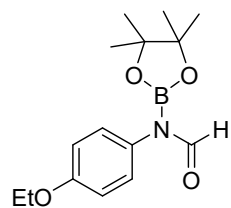
Figure S20. ^{13}C NMR spectrum of **2j** (CDCl_3 , 101 MHz, 298 K).

N-(4-methoxyphenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2k): ^1H

 NMR (CDCl_3 , 500 MHz, 298 K): δ 8.94 (s, 1H, NCHO), 6.99-6.96 (d, 2H, PhH), 6.82-6.80 (d, 2H, PhH), 3.69 (s, 3H, PhCH₃), 1.22 (s, 12H, CH₃) ppm;
 $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.56 (s, 1 C), 158.25 (s, 1 C), 128.18 (s, 1 C), 125.45 (s, 1 C), 124.04 (s, 1 C), 114.02 (s, 1 C), 84.37 (s, 1 C), 82.95 (s, 1 C), 55.16 (s, 1 C), 24.41 (s, 1 C), 24.34 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

N-(4-ethoxyphenyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2I): ^1H



NMR (CDCl_3 , 500 MHz, 298 K): δ 8.93 (s, 1H, NCHO), 7.11-7.09 (d, 2H, PhH), 6.95-6.93 (d, 2H, PhH), 4.08-4.03 (q, 2H, Ph- OCH_2CH_3), 1.47-1.44 (t, 3H, Ph- OCH_2CH_3), 1.36 (s, 12H, CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ ppm 165.58 (s, 1 C), 157.67 (s, 1 C), 129.30 (s, 1 C), 128.14 (s, 1 C), 118.66 (s, 1 C), 115.12 (s, 1 C), 114.86 (s, 1 C), 114.51 (s, 1 C), 84.35 (s, 1 C), 82.98 (s, 1 C), 63.59 (s, 1 C), 63.54 (s, 1 C), 63.36 (s, 1 C), 24.75 (s, 1 C), 24.53 (s, 1 C), 24.40 (s, 1 C), 24.34 (s, 1 C), 24.25 (s, 1 C), 14.71 (s, 1 C), 14.63 (s, 1 C), 14.59 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

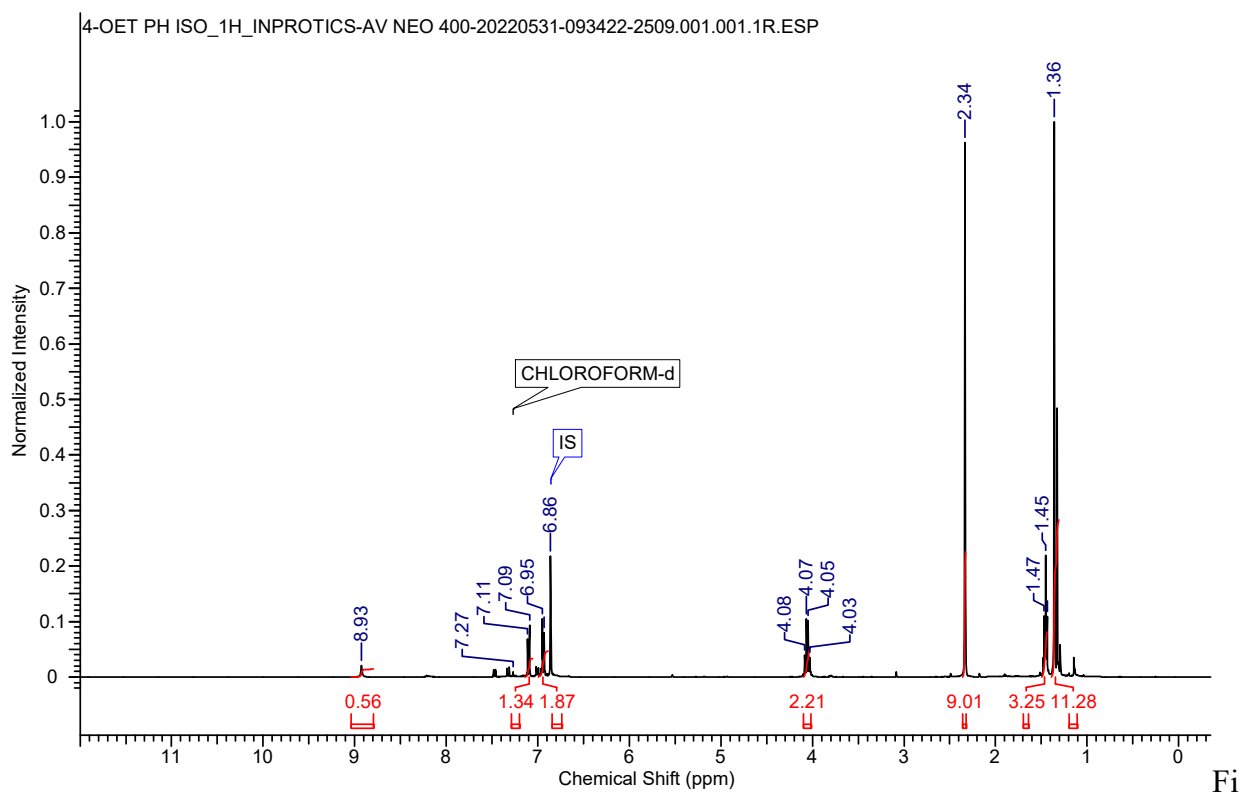


Figure S23. ^1H NMR spectrum of **2I** (CDCl_3 , 500 MHz, 298 K).

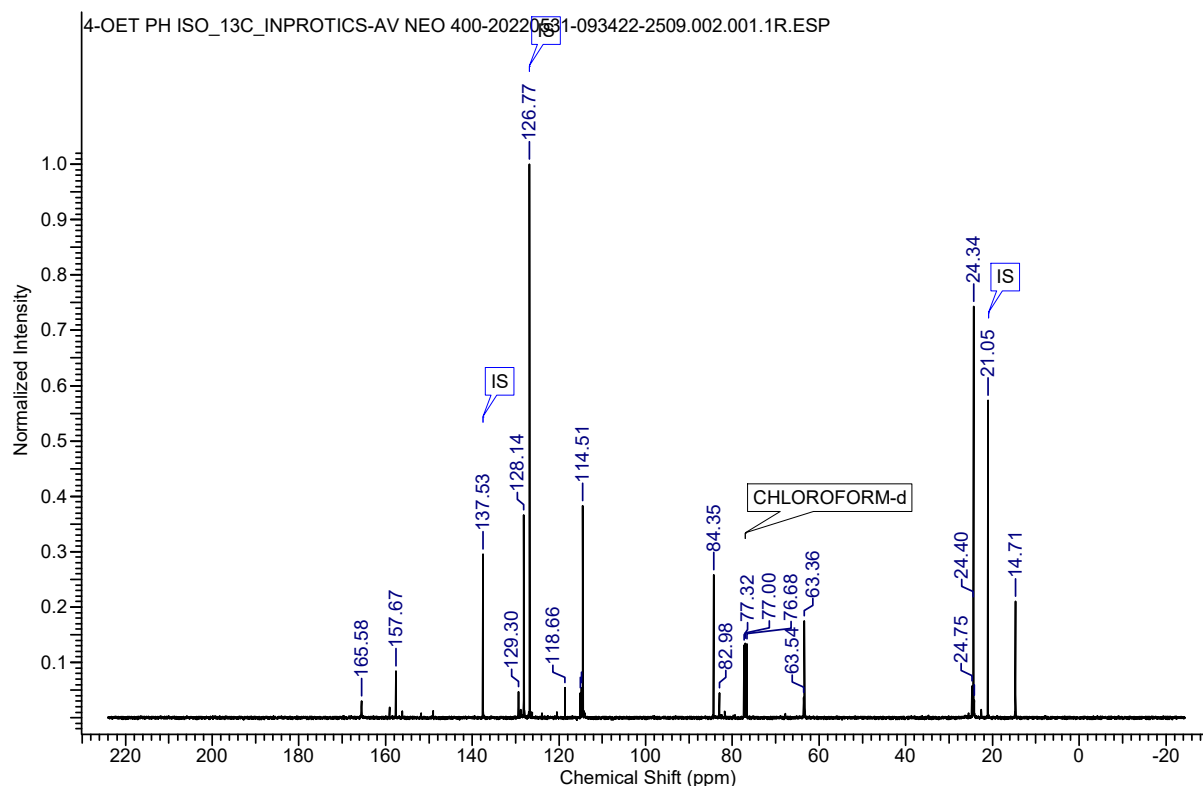
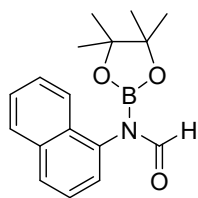


Figure S24. ^{13}C NMR spectrum of **2I** (CDCl_3 , 100.28 MHz, 298 K).

N-benzyl-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2I): ^1H NMR (CDCl_3 , 500 MHz, 298 K): δ 8.79 (s, 1H, NCHO), 7.43-7.29 (m, 5H, PhH), 5.10 (s, 2H, NCH₂), 4.58 (s, 2H, PhCH₂), 1.35 (s, 12H, CH₃) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 165.61 (s, 1 C), 148.98 (s, 1 C), 138.76 (s, 1 C), 135.71 (s, 1 C), 128.99 (s, 1 C), 128.79 (s, 1 C), 128.72 (s, 1 C), 128.57 (s, 1 C), 128.50 (s, 1 C), 128.33 (s, 1 C), 128.18 (s, 1 C), 128.09 (s, 1 C), 128.05 (s, 1 C), 127.83 (s, 1 C), 127.75 (s, 1 C), 127.19 (s, 1 C), 126.96 (s, 1 C), 126.60 (s, 1 C), 84.28 (s, 1 C), 82.99 (s, 1 C), 46.12 (s, 1 C), 44.29 (s, 1 C), 43.41 (s, 1 C), 24.53 (s, 1 C), 24.42 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

N-(naphthalen-1-yl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide(2n):



^1H NMR (CDCl_3 , 400 MHz, 298 K): δ ppm 9.20 (s, 1 H), 7.96-7.90 (m, 2 H), 7.81-7.79 (d, 1H), 7.60-7.55 (m, 3H), 7.37-7.35 (d, 1H), 1.36 (s, 12 H) ppm;

$^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ ppm 165.49 (s, 1 C), 134.27 (s, 1 C), 133.14 (s, 1 C), 130.07 (s, 1 C), 128.35 (s, 1 C), 128.18 (s, 1 C), 126.43 (s, 1 C), 125.96 (s, 1 C), 125.57 (s, 1 C), 125.37 (s, 1 C), 122.26 (s, 1 C), 84.46 (s, 1 C), 82.99 (s, 1 C), 24.43 (s, 1 C), 24.12 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data. ^{S2}

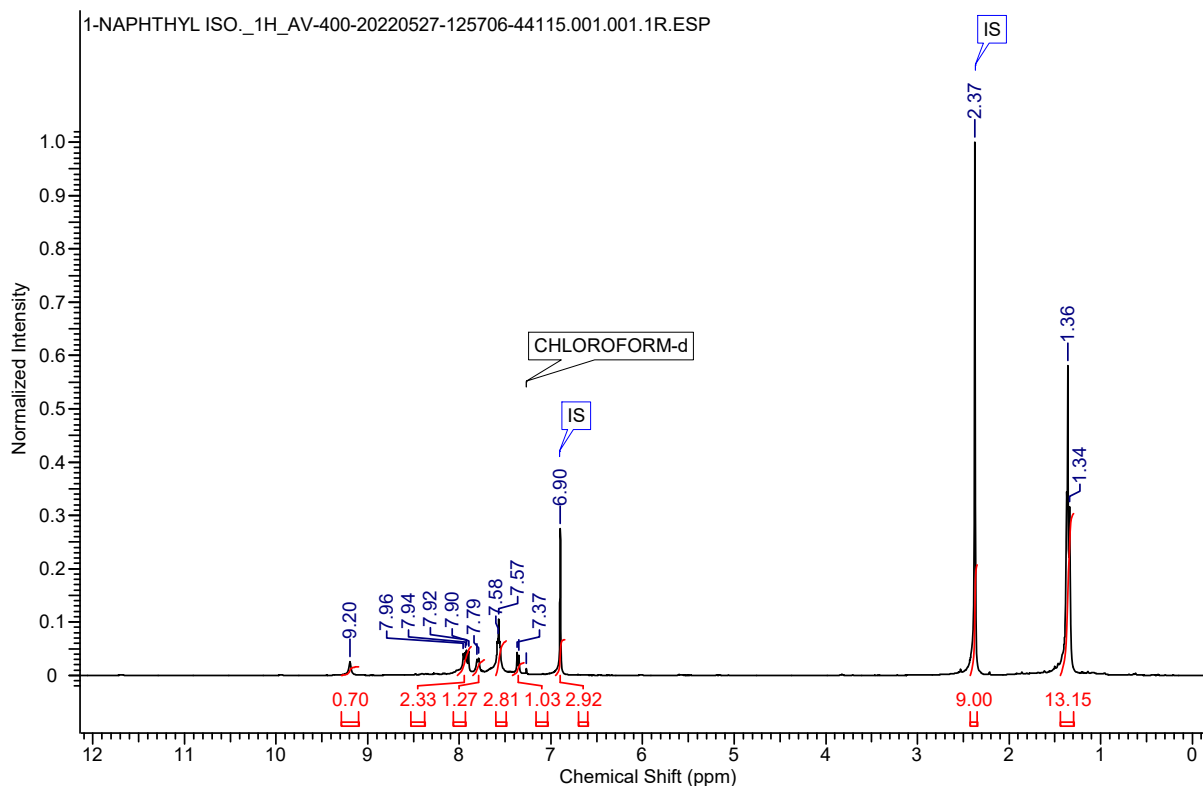


Figure S27. ^1H NMR spectrum of **2n** (CDCl_3 , 400 MHz, 298 K).

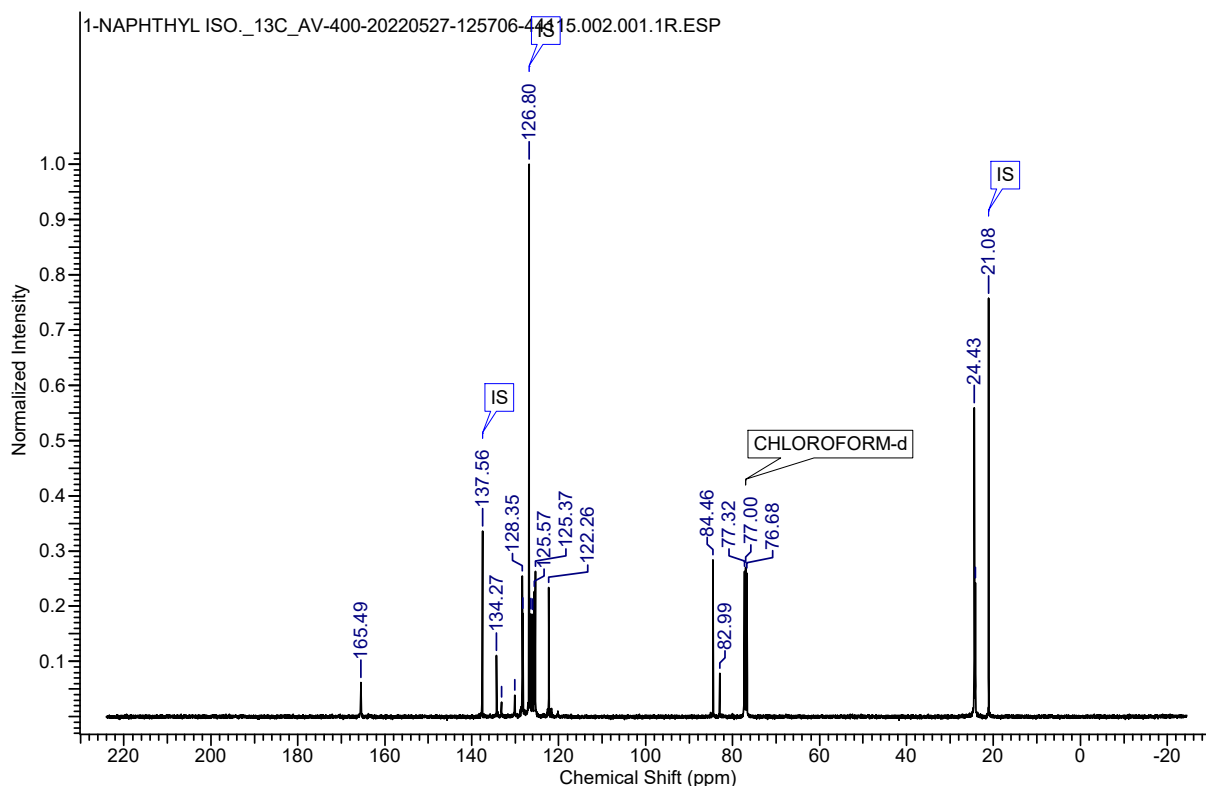


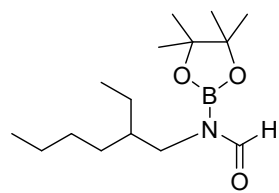
Figure S28. ^{13}C NMR spectrum of **2n** (CDCl_3 , 101 MHz, 298 K).

N-octyl-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2o): ^1H NMR (CDCl_3 ,

400 MHz, 298 K): δ 8.69 (s, 1H, *NCHO*), 3.37-3.34 (t, 2H, *Octyl*), 1.53 (bs, 2H, *Octyl*), 1.32 (s, 12H, CH_3) 1.29-1.25 (m, 12H, *Octyl*), 0.93 (t, 3H, *Octyl*) ppm;
 $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ ppm 165.91 (s, 1 C), 83.93 (s, 1 C), 82.93 (s, 1 C), 48.29 (s, 1 C), 42.86 (s, 1 C), 39.96 (s, 1 C), 31.71 (s, 1 C), 31.65 (s, 1 C), 29.49 (s, 1 C), 29.17 (s, 1 C), 29.14 (s, 1 C), 29.07 (s, 1 C), 26.73 (s, 1 C), 26.61 (s, 1 C), 24.43 (s, 1 C), 24.40 (s, 1 C), 24.34 (s, 1 C), 22.54 (s, 1 C), 22.51 (s, 1 C), 13.96 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

N-(2-ethylhexyl)-N-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)formamide (2p): ^1H NMR



(CDCl_3 , 400 MHz, 298 K): δ 8.70 (s, 1H, NCHO), 3.25-3.23 (t, 2H, NCH_2), 1.31 (s, 12H, CH_3) 1.29 (bs, 8H, CH_2), 0.94-0.89 (m, 7H, CH_2CH_3) ppm; $^{13}\text{C}\{^1\text{H}\}$ NMR (CDCl_3 , 101 MHz, 298 K): δ 166.16 (s, 1 C), 83.95 (s, 1 C), 82.95 (s, 1 C), 43.76 (s, 1 C), 40.71 (s, 1 C), 39.09 (s, 1 C), 38.27 (s, 1 C), 30.72 (s, 1 C), 30.35 (s, 1 C), 28.69 (s, 1 C), 28.42 (s, 1 C), 24.39 (s, 1 C), 24.33 (s, 1 C), 24.30 (s, 1 C), 23.92 (s, 1 C), 23.51 (s, 1 C), 22.98 (s, 1 C), 22.89 (s, 1 C), 13.97 (s, 1 C), 13.94 (s, 1 C), 10.65 (s, 1 C), 10.35 (s, 1 C) ppm.

The spectroscopic data is consistent with the literature data.^{S2}

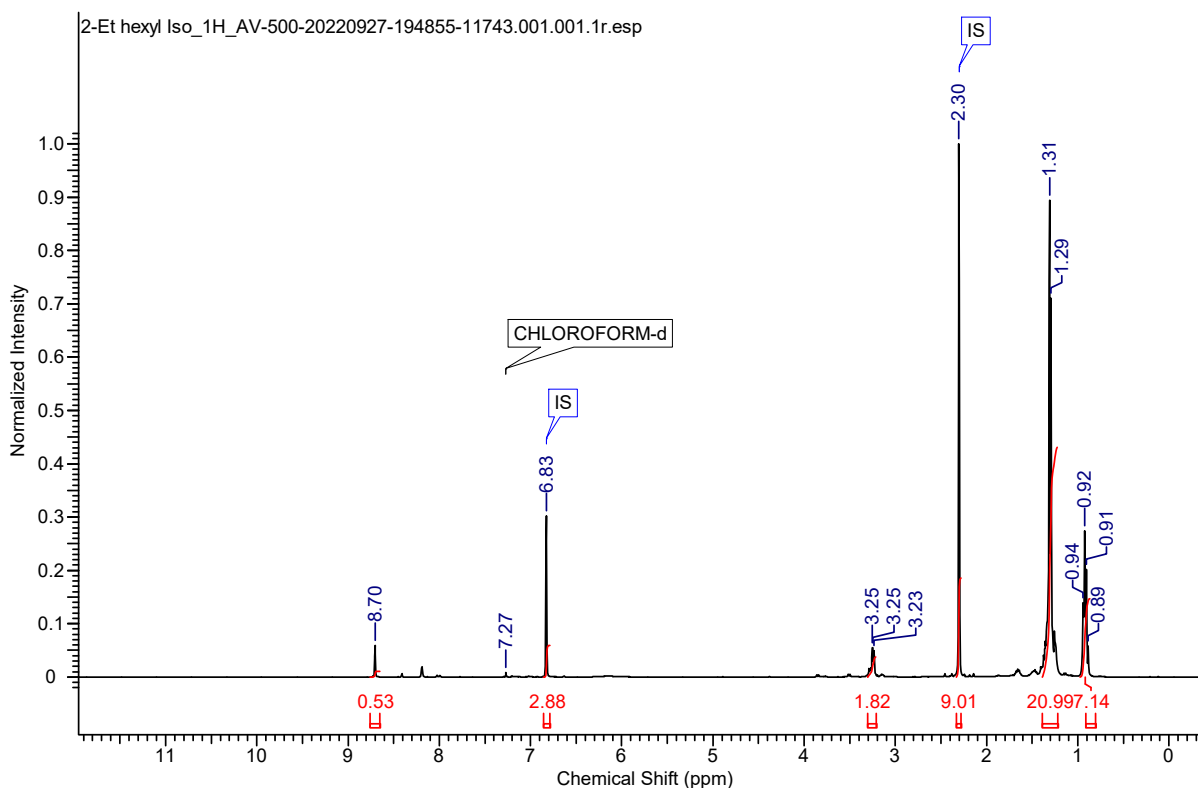


Figure S31. ^1H NMR spectrum of **2p** (CDCl_3 , 400 MHz, 298 K).

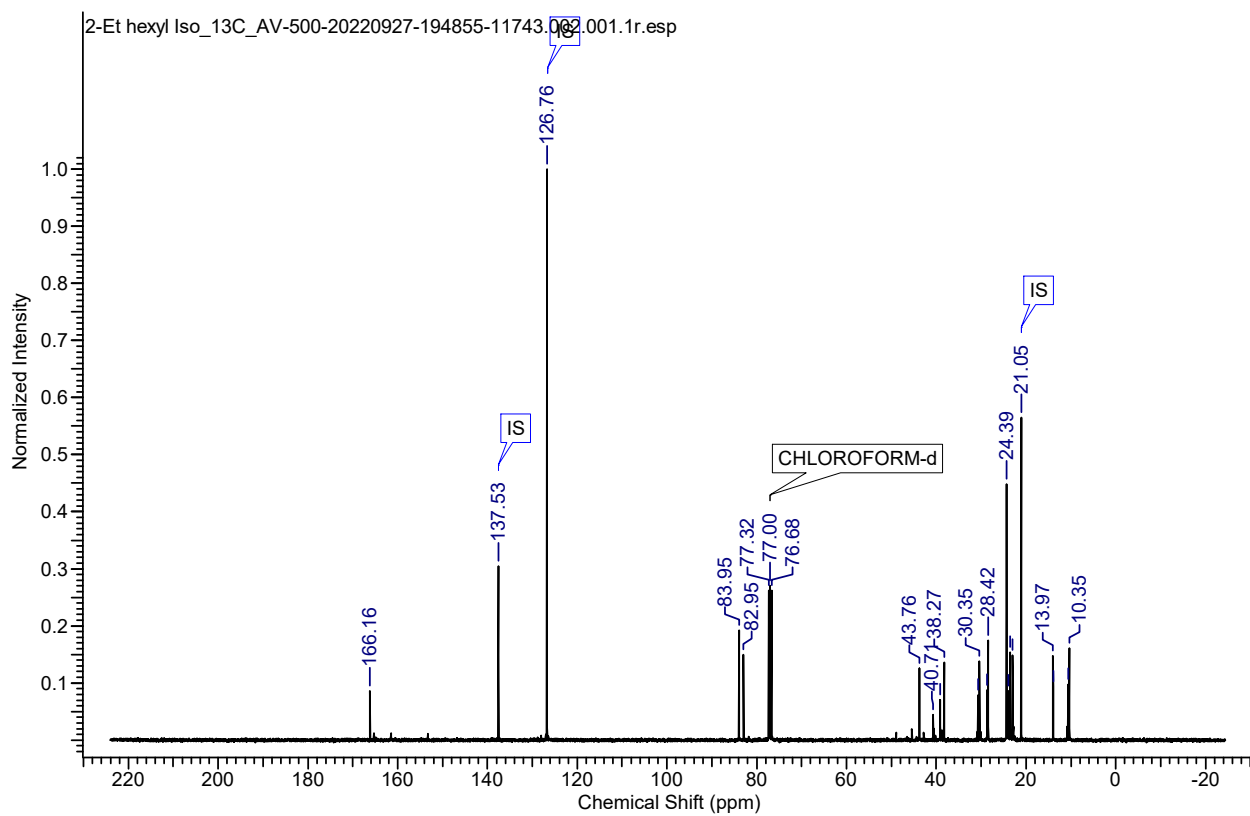
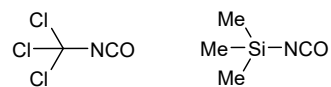


Figure S32. ¹³C NMR spectrum of **2p** (CDCl₃, 101 MHz, 298 K).

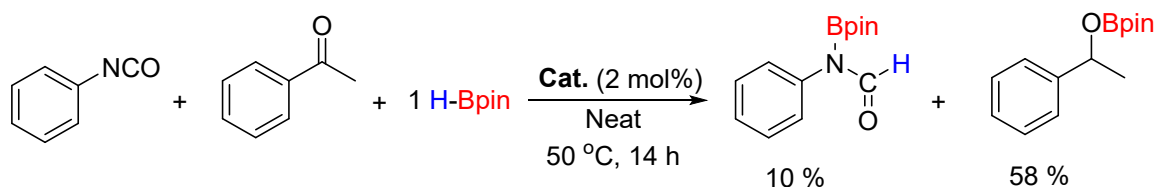
❖ **List of unsuccessful substrates for catalytic hydroboration:**



Very less yield due to unidentified side reactions

❖ **Competitive experiment for alkene/alkyne/nitrile/ketone/ester hydroboration- a selectivity study:**

1. Hydroboration of phenyl isocyanate and acetophenone in presence of 1 equiv. HBpin:



Scheme S2. Hydroboration of phenyl isocyanate and 4-fluoro acetophenone.

Phenyl isocyanate (0.25 mmol), acetophenone (0.25 mmol), pinacolborane (32 mg, 0.25 mmol), catalyst (2.0 mol% for **1**) were charged in a Schlenk tube inside the glove box. The reaction mixture was stirred for 14 hours at 50 °C after in neat conditions. Upon completion of the reaction, the progress of the reaction was monitored by ¹H NMR after addition of mesitylene (0.25 mmol) as an internal standard in CDCl₃. A sharp singlet resonance at $\delta = 8.89$ ppm indicates for the formation of hydroboration product from phenyl isocyanate and a quartet resonance at 5.29-5.24 ppm indicates the hydroboration product from acetophenone.

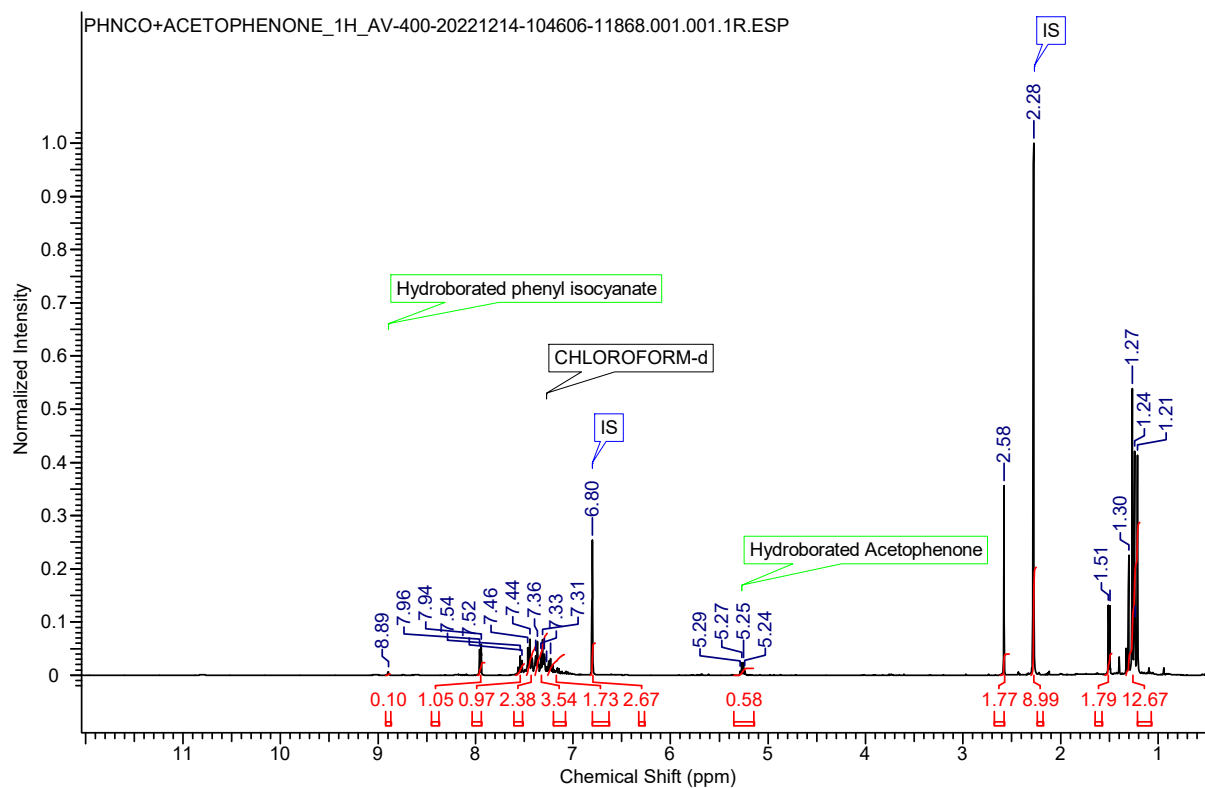
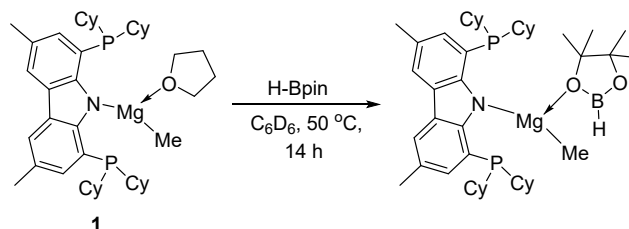


Figure S33. ^1H NMR spectrum of the reaction of phenyl isocyanate and acetophenone in presence of one equiv. HBpin (CDCl_3 , 400 MHz, 298 K).

❖ Mechanistic investigation:

Stoichiometric reaction of catalyst **1** and HBpin:



Scheme S3. Stoichiometric reaction of **1** with HBpin and the tentative product formation.

A solution of HBpin (20 mg, 0.14 mmol) in C₆D₆ was added drop by drop to the C₆D₆ solution of **1** (90 mg, 0.14 mmol) at room temperature or inside the glove box. The reaction mixture was stirred/heated for 14 hours at 50 °C. After that the reaction mixture was subjected for characterization. ¹H NMR (C₆D₆, 400 MHz, 298 K): ¹H NMR spectra shows a new peak at 2.29 ppm which can be a proton peak for HBPin. This supports the formation of O-coordinated magnesium methyl complex. The ¹¹B NMR also indicates that the Bpin moiety is coordinated to the Mg centre through one of the O-atoms. ¹¹B NMR (C₆D₆, 128 MHz, 298 K): δ 33.96 (s, Me-Bpin), 28.90 (d, for unreacted HBpin), 21.92 (s, Bpin coordinated catalyst) ppm.

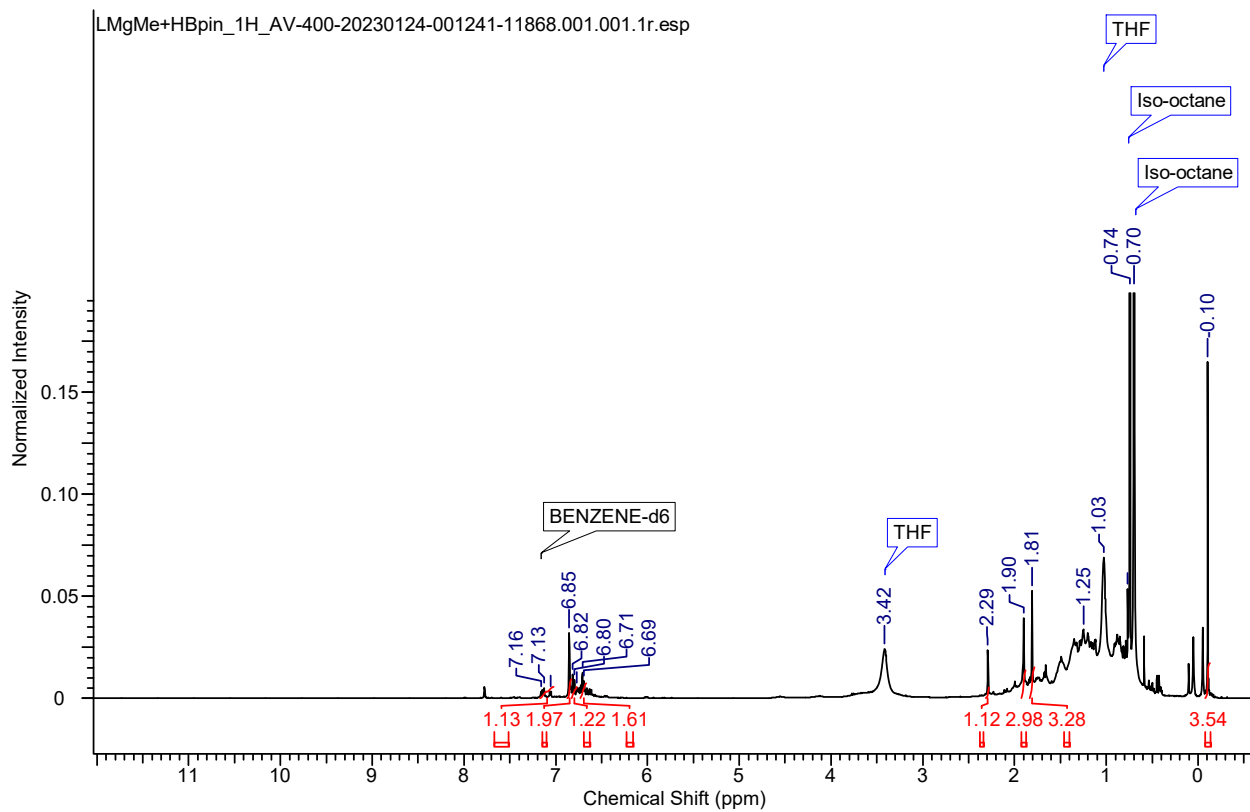


Figure S34. ^1H NMR spectrum of the reaction of **1** with HBpin (C_6D_6 , 400 MHz, 298 K).

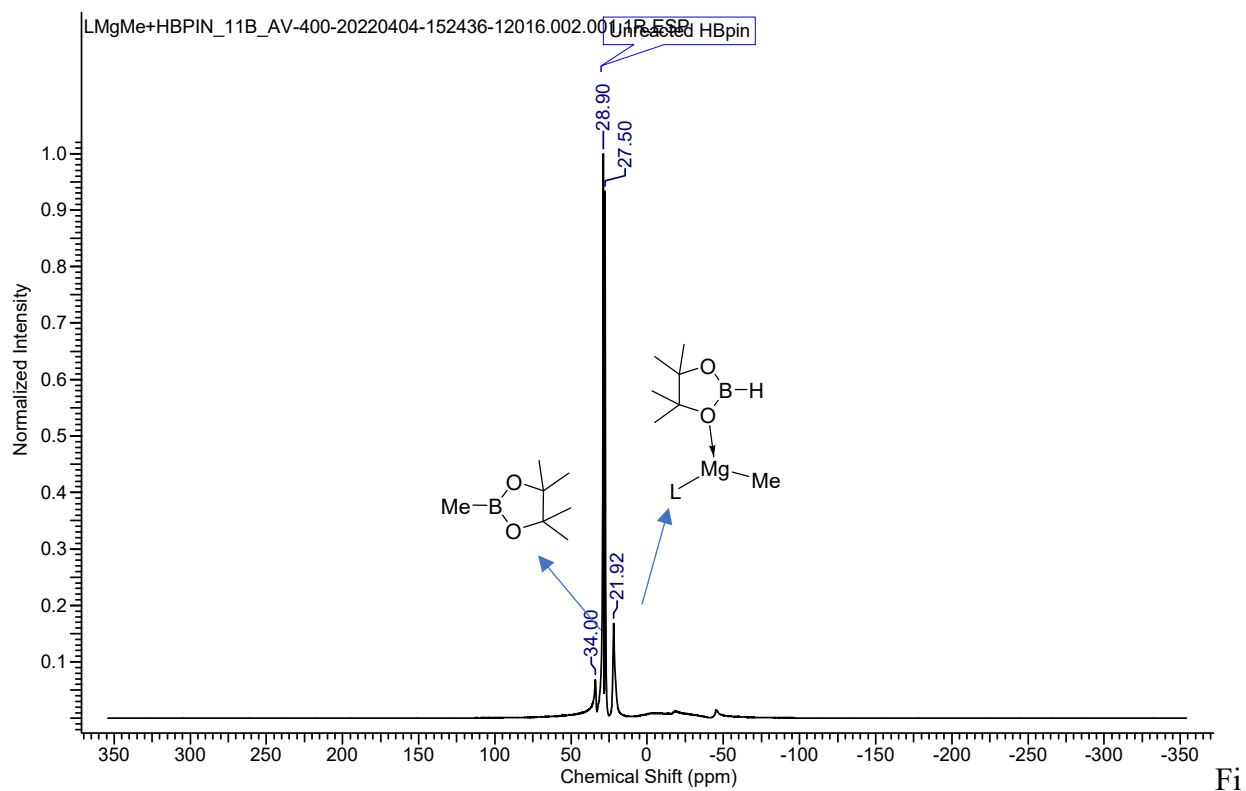


Figure S35. ^{11}B NMR spectrum of the reaction of **1** with HBpin (C_6D_6 , 128 MHz, 298 K).

Heating of the catalyst at 70 °C for 8 hours under vacuum:

We have heated the catalyst under vacuum at 70 °C for 8 hours but we did not observe the removal of THF confirmed by ^1H NMR.

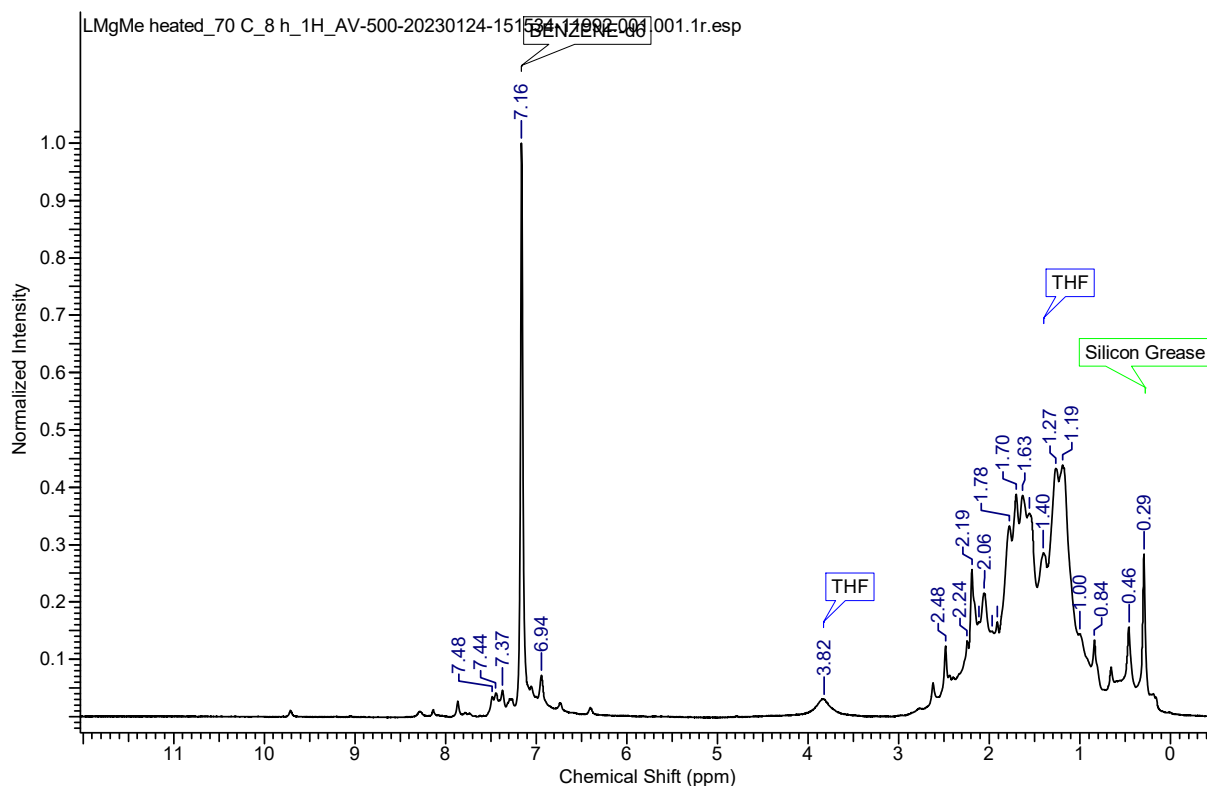
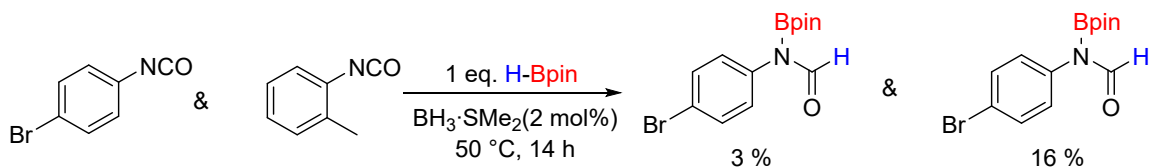


Figure S36. ^1H NMR spectrum heating of the catalyst at 70 °C for 8 hours under vacuum (C_6D_6 , 128 MHz, 298 K).

Stoichiometric reaction of **1** with phenyl isocyanate or ethyl isocyanate:

No reaction was observed when an equimolar amount of **1** was treated separately with phenyl isocyanate or ethyl isocyanate in THF/toluene solvent at room temperature and at 60-100 °C. Further heating at 80-100 °C overnight in neat conditions also does not indicate any appreciable changes in the ^1H NMR spectra.

Hydroboration of 4-bromo phenyl isocyanate and o-tolyl isocyanate with $\text{BH}_3 \cdot \text{SMe}_2$:



Scheme S4. Synthetic scheme for the reaction of 4-bromo phenyl isocyanate and o-tolyl isocyanate with $\text{BH}_3 \cdot \text{SMe}_2$.

To identify whether the BH_3 is behaving as hidden catalyst in the hydroboration of phenyl isocyanate or not, we performed the reaction with $\text{BH}_3 \cdot \text{SMe}_2$. A Screw cap NMR tube was charged with $\text{BH}_3 \cdot \text{SMe}_2$ (2 mol%) inside the argon-filled glovebox. Subsequently 4-bromo phenyl isocyanate (0.25 mmol, 1.0 equiv) and 2-methyl phenyl isocyanate were added to the reaction mixture in separate NMR tubes and heated at 50°C up to 14 hours in an oil bath. After completion of the reaction 0.25 mmol mesitylene was added as an internal standard prior to the NMR measurement in CDCl_3 solvent. The ^1H NMR suggests that there is very less product formation in both the reactions, i.e., N-boryl formamide was observed (Figure S35 and S36), which clearly suggest very less influence of BH_3 as an active catalyst for the hydroboration of isocyanate.

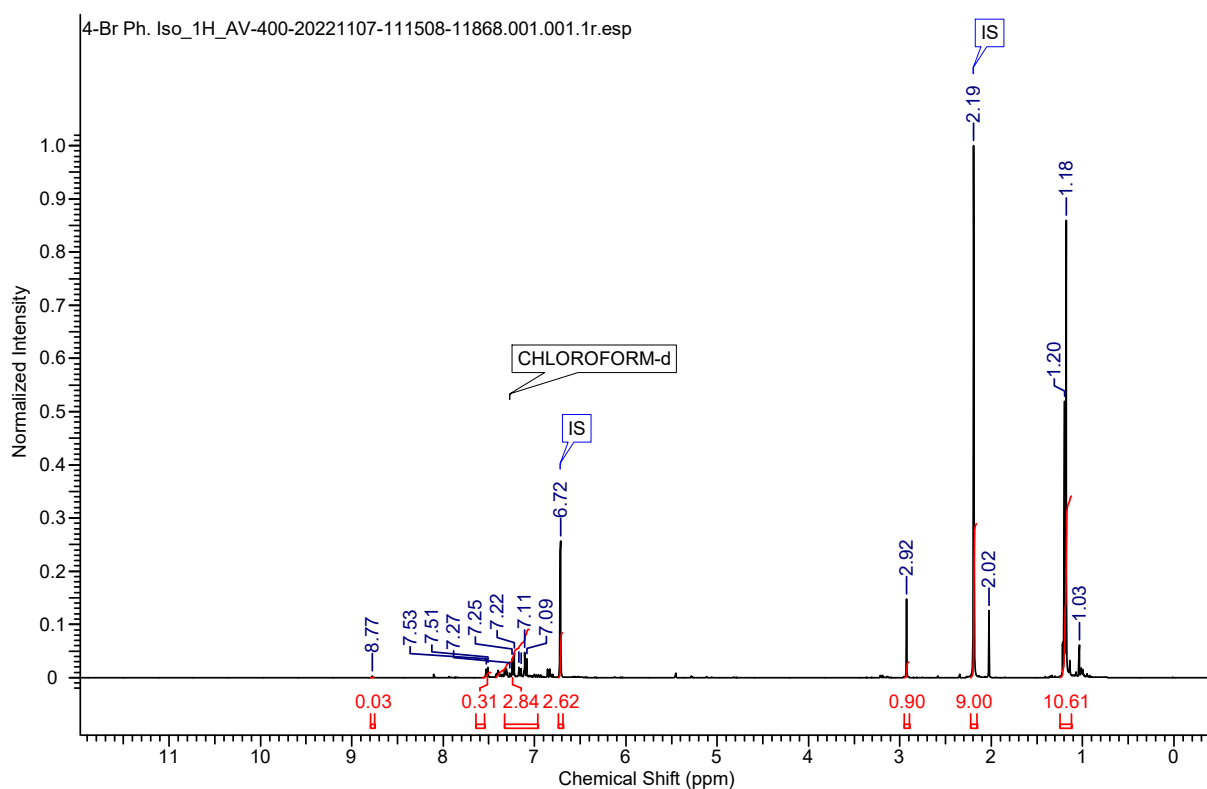


Figure S37. ^1H NMR spectrum of reaction of 4-bromo phenyl isocyanate with $\text{BH}_3 \cdot \text{SMe}_2$ at 50°C (CDCl_3 , 400 MHz, 298 K).

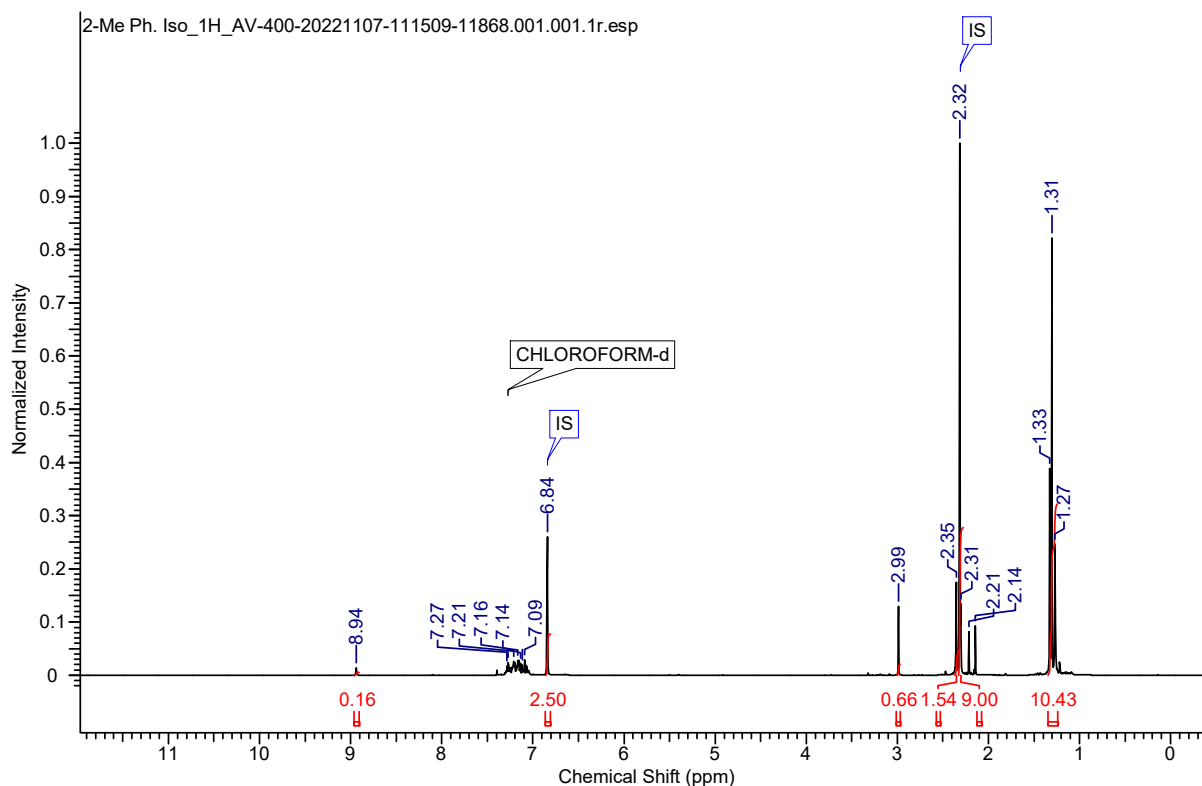


Figure S38. ^1H NMR spectrum of reaction of 2-methyl phenyl isocyanate with $\text{BH}_3 \cdot \text{SMe}_2$ at 50°C (CDCl_3 , 400 MHz, 298 K).

❖ Details of the theoretical calculations:

Full quantum calculations have been carried out with Density Functional theory(DFT)^{S3} using Turbomole 7.4. ^{S4} The TZVP basis set have been employed in all the calculations. ^{S5} Geometry optimizations were performed using the Perdew, Burke, and Ernzerhof functional (PBE).^{S6} Dispersion corrections ^{S7} have been included in all the calculations. The resolution of identity (RI) ^{S8} along with the multipole accelerated RI (marij) ^{S9} approximations have been used for an accurate and efficient treatment of the electronic Coulomb term in the DFT calculations. Solvent corrections have also been included in all the calculations using the cosmo model, ^{S10} with epsilon (ϵ) = 2.48, to model the HBpin which has been employed to study the reaction involved. Necessary care was taken to ensure that the obtained transition state structures possessed only one imaginary frequency corresponding to the correct normal mode, in order to obtain more reliable energy values for the investigated potential energy surface. In addition, intrinsic reaction coordinate (IRC)^{S11} calculations were done with all the transition states in order to further confirm that they were the

correct transition states, yielding the correct reactant and product structures. The values reported here are ΔG values, with zero-point energy, internal energy, and entropic contributions, with the temperature taken to be 323.15 K. This is because this is the temperature at which the experiments were done. The translational entropy term has been corrected by a free volume correction introduced by Mammen and co-workers.^{S12} The efficiency of the catalytic cycle for the reaction mechanism investigated has been calculated with the AUTOF^{S13-S14} program by employing the “Energetic Span Model” (ESM), developed by Shaik and co-workers.^{S14-S16} This has been done for all the free energy profiles discussed in the manuscript. The turnover frequency (TOF) calculations take into account the principal rate-determining transition state, as well as the potentially rate-influencing transition states and intermediates during the catalysis process. In most cases, the TOF is calculated from the TOF-determining transition state (TDTS), the TOF-determining intermediate (TDI), and from the reaction energy, ΔG_r as shown below:

$$\text{TOF} = \frac{KbT}{h} e^{-\delta E/RT}$$

$$\delta E = \text{TDTS} - \text{TDI} \quad \text{if TDTS appears after TDI}$$

$$\delta E = \text{TDTS} - \text{TDI} - \Delta G_r \quad \text{if TDTS appears before TDI}$$

$$\delta E = \text{TDTS} - \text{TDI} - \Delta G_r \quad \text{if TDTS appears before TDI}$$

This model has been employed to calculate the TOFs for the free energy profiles obtained for the mechanisms in the solvent phase discussed in the manuscript.

Discarded Mechanisms

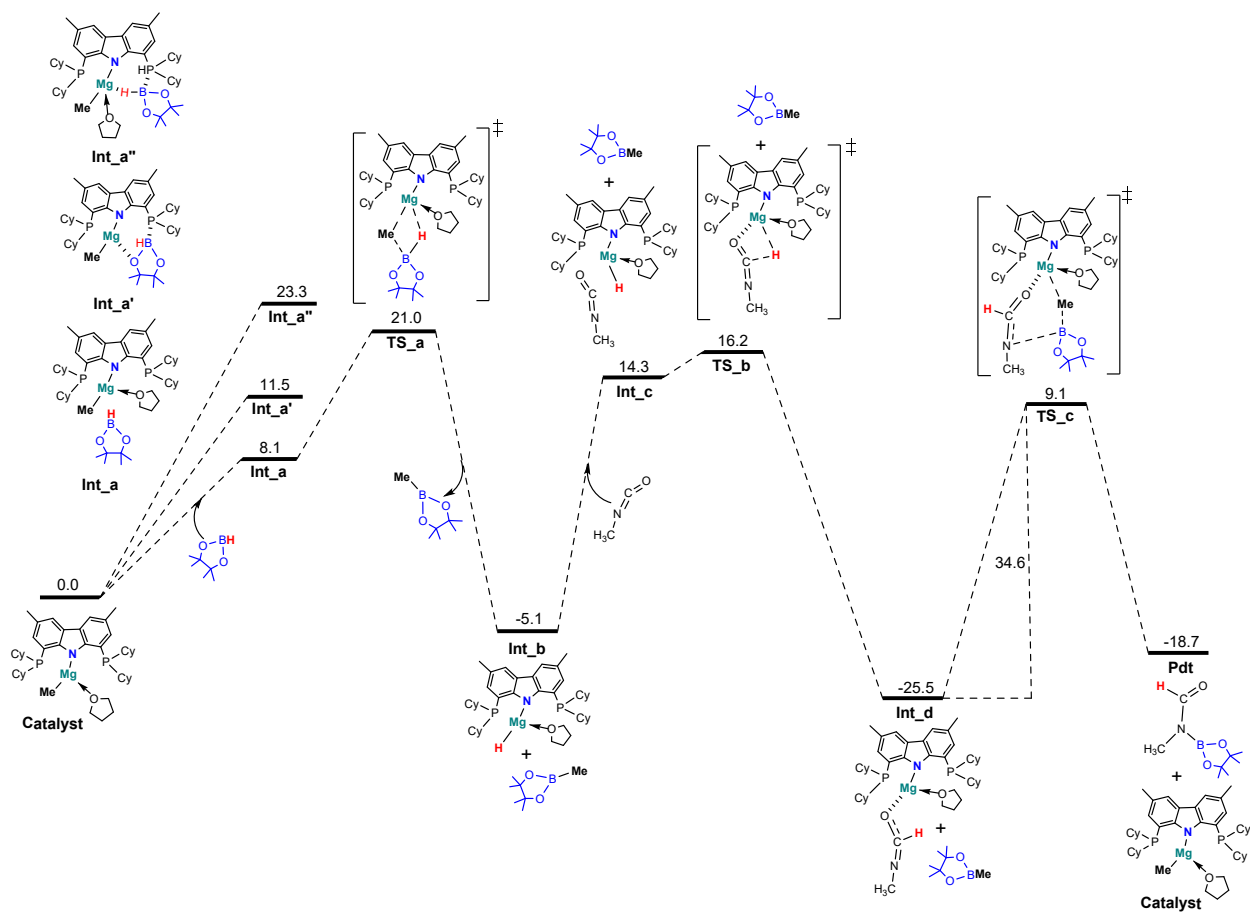


Figure S39. Reaction mechanism with DFT studies considering magnesium methyl as catalyst.

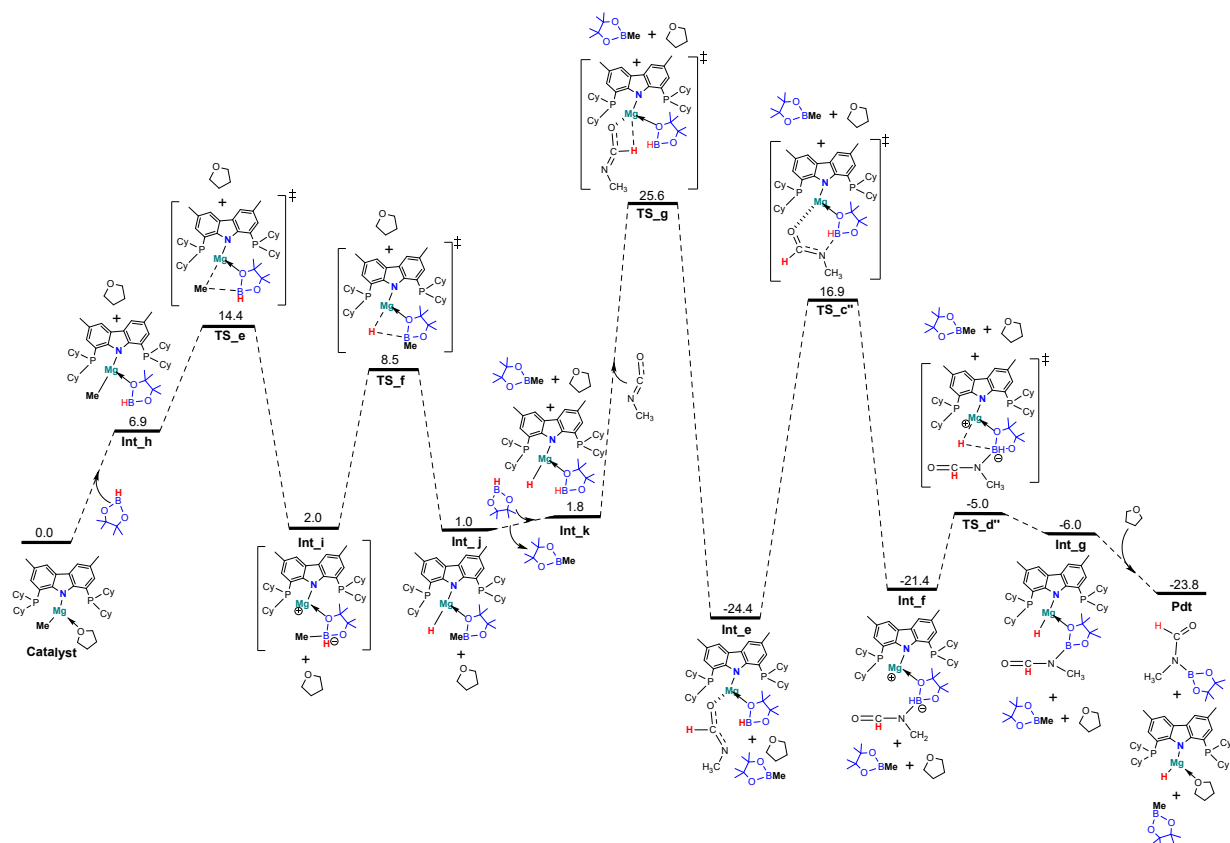


Figure S40. Mechanism via the formation of magnesium hydride

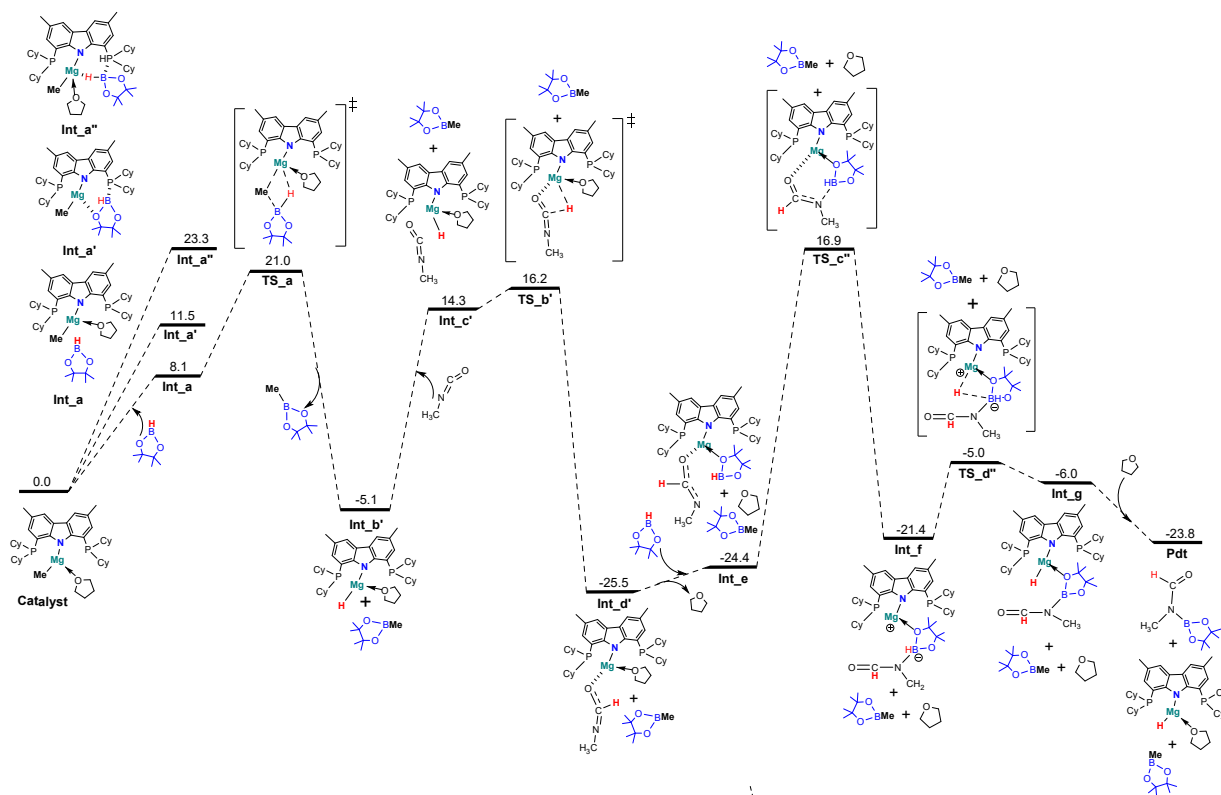


Figure S41. Alternative mechanism via the formation of magnesium hydride.

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❖ The XYZ coordinates of the optimized structures discussed in the manuscript

The optimized geometries of the structures reported in the manuscript (the atomic symbol followed by the three Cartesian coordinates, in Å).

Int_1			H	5.8279677	9.3919784	1.1644732	C	6.8987630	11.5028120	2.8076985	
P	4.8178000	9.5602086	3.4010279	C	6.6880395	7.7697315	2.3010177	H	6.1371121	12.1283238	2.3111811
P	1.2154317	9.6256753	7.5492842	H	7.5018517	8.3937568	2.7031737	H	7.3083623	10.8283531	2.0389762
Mg	3.2896158	8.9168997	5.7568746	H	6.3856680	7.0879777	3.1154826	C	-0.1678834	8.6923889	8.4053545
N	2.2276433	10.4568252	4.7829818	C	7.2019864	6.9456709	1.1120674	H	-0.8722996	9.4192696	8.8478247
C	3.5716435	10.7165680	2.6935821	H	7.6108523	7.6310824	0.3479727	C	-0.9188978	7.8524416	7.3581970
C	3.6030987	11.3342532	1.4391512	H	8.0357562	6.3022800	1.4362973	H	-0.1842276	7.2276821	6.8201678
H	4.4417284	11.1412704	0.7645274	C	6.0849992	6.1035115	0.4854074	H	-1.3725090	8.5128292	6.6033566
C	2.5783069	12.2053990	0.9959751	H	5.7428180	5.3502248	1.2189271	C	-1.9828693	6.9515452	7.9949772
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C	-4.8969947	-1.3219907	3.8896776	C	1.0192376	-1.6601024	-3.3705075	C	2.4616074	2.4290285	-0.4033097
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