Supporting Information

for

Highly selective α-hydrogermylation of alkynes catalyzed by

an *in-situ* **generated bulky NHC–cobalt complex**

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CONTENTS:

1. General methods and chemicals

All syntheses and catalytic tests were carried out under dry argon, using glove-box, standard Schlenk-line and vacuum techniques. ${}^{1}H$ NMR and ${}^{13}C$ NMR spectra were recorded in CDCl₃ on a Varian 400 operating at 402.6 and 101.2 MHz, respectively. GC-MS analyses were performed on a Varian Saturn 2100T equipped with a DB-1 capillary column (30 m in length and 0.25 mm in internal diameter) and an ion trap detector. Mass spectrum of product **P21** was obtained using Synapt Gs-S HDMS (Waters) mass spectrometer with Electrospray ion source and quadrupole-Time-offlight analyzer with resolving power FWMH 38000. Acetonitrile was utilized as the sample solvent. The Capillary Voltage was set to 4.5 kV, the sampling was set 40 and the source temperature was equal to 120°C. The most abundant ions in ESI-MS spectra was potassiated ions of desired product. Thin layer chromatography (TLC) was conducted on plates coated with a 250 μm thick silica gel layer and column chromatography was performed on silica gel 60 (70–230 mesh).

NHC carbene precursor was prepared according to literature procedures.^[S1] All the other reagents were commercially available and used as received. The solvents were dried over CaH₂ prior to use and stored over 4\AA molecular sieves under argon. Dichloromethane was additionally passed through an alumina column and degassed by repeated freeze-pump-thaw cycles.

2. Optimization studies

2.1. Solvent and temperature screening

Table S1. Optimization of solvent and temperature.

Reaction conditions: $[\textbf{1a}] : [\textbf{2a}] = 1 : 1$, $[Co_2(CO)_8] = 5.0$ mol%, $[IPr^{*OMe}] = 10.0$ mol%, argon. ^a Determined by GC-MS analysis; \overline{b} Determined by GC-MS analysis and confirmed by ¹H NMR spectroscopy of the crude reaction mixture.

3. Experimental procedures

3.1. General method for hydrogermylation of acetylenes – optimization studies

An oven-dried 5 mL glass reactor equipped with a magnetic stirring bar was charged with ligand (6.09 \times 10⁻⁶ mol), octacarbonyl dicobalt(0) (1.0 mg, 3.05 \times 10⁻⁶ mol) in the glovebox. Then solvent (0.5 mL) was added and reaction mixture was stirred for 30 minutes at room temperature. After this time, alkyne $(6.09 \times 10^{-5}$ mol), germane $(6.09\times10^{-5}$ mol) and internal standard (decane or dodecane, 15 μ L) were added. The reaction mixture was stirred at set temperature for 18h. Reaction course was monitored by GC-MS.

3.2. General method for hydrogermylation of acetylenes

An oven-dried 5 mL glass reactor equipped with a magnetic stirring bar was charged with NHC carbene (5.8 mg, 6.09×10^{-6} mol), octacarbonyl dicobalt(0) (1.0 mg, 3.05×10^{-6} mol) in the glovebox. Then DCE (0.5 mL) was added and reaction mixture was stirred for 30 minutes at room temperature. After this time, alkyne $(6.09 \times 10^{-5}$ mol), germane (6.09×10-5 mol) and internal standard (decane or dodecane, 15 μL) were added. The reaction mixture was stirred at 60 °C for 18h. Reaction course was monitored by GC-MS.

3.3. General procedure for the synthesis of α-vinylgermanes (P1-P24)

An oven-dried 5 mL glass reactor equipped with a magnetic stirring bar was charged with NHC carbene (28.8 mg, 3.05×10^{-5} mol), octacarbonyl dicobalt(0) (5.2 mg, 1.53×10^{-5} mol) in the glovebox. Then DCE (0.5 mL) was added and reaction mixture was stirred for 30 minutes at room temperature. After this time, alkyne (3.05×10⁻⁴ mol) and germane (3.05 \times 10⁻⁴ mol) were added. The reaction mixture was stirred at 60 °C for 18h. Reaction course was monitored by GC-MS. The products were purified by column chromatography on silica gel using hexane or a 10: 1 v/v mixture of *n*-hexane and diethyl ether as eluents. Evaporation of the solvents afforded analytically pure compounds. Purification of **P2**, **P7** – **P17**, **P20** was performed on preparative TLC plates with 10: 1 v/v mixture of *n*-hexane and diethyl ether as eluent.

4. Analytical data of isolated products

5. NMR spectra of isolated products

Figure S2.¹³C NMR (101 MHz, CDCl₃) of product P1

Figure S4.¹³C NMR (101 MHz, CDCl₃) of product P2

Figure S6. ¹³C NMR (101 MHz, CDCl3) of product **P3**

Figure S8. ¹³C NMR (101 MHz, CDCl₃) of product P4

Figure S10.¹³C NMR (101 MHz, CDCl₃) of product P5

Figure S12. ¹³C NMR (101 MHz, CDCl3) of product **P6**

Figure S14. ¹³C NMR (101 MHz, CDCl3) of product **P7**

Figure S15. ¹H NMR (400 MHz, CDCl₃) of product P8. Signal at 1.56 ppm derives from water from CDCl₃

Figure S16. ¹³C NMR (101 MHz, CDCl3) of product **P8**

Figure S18.¹³C NMR (101 MHz, CDCl₃) of product P9

Figure S20.¹³C NMR (101 MHz, CDCl₃) of product P10

Figure S22.¹³C NMR (101 MHz, CDCl₃) of product P11

Figure S24. ¹³C NMR (101 MHz, CDCl3) of product **P12**

Figure S26.¹³C NMR (101 MHz, CDCl₃) of product P13

Figure S28. ¹³C NMR (101 MHz, CDCl3) of product **P14**

Figure S32. ¹³C NMR (101 MHz, CDCl3) of product **P16**

Figure S34. ¹³C NMR (101 MHz, CDCl3) of product **P17**

Figure S36. ¹³C NMR (101 MHz, CDCl3) of product **P18**

Figure S38. ¹³C NMR (101 MHz, CDCl3) of product **P19**

Figure S40. ¹³C NMR (101 MHz, CDCl3) of product **P20**

Figure S42. ¹³C NMR (101 MHz, CDCl3) of product **P21**

Figure S44. ¹³C NMR (101 MHz, CDCl3) of product **P22**

Figure S46.¹³C NMR (101 MHz, CDCl₃) of product P23

Figure S48 ¹³C NMR (101 MHz, CDCl3) of product **P24**

6. X-ray analysis

Diffraction data were collected at 100K by the ω-scan technique on Rigaku Xcalibur four-circle diffractometer with Eos CCD detector and graphite-monochromated $M \circ K_{\alpha}$ radiation (λ=0.71069 Å). The data were corrected for Lorentz-polarization as well as for absorption effects^[S4]. Precise unit-cell parameters were determined by a least-squares fit of 4038 (**P17**) reflections of the highest intensity, chosen from the whole experiment. The structures were solved with SHELXT^[S5] and refined with the full-matrix leastsquares procedure on F² by SHELXL^[S6]. All non-hydrogen atoms were refined anisotropically. Hydrogen atoms H2A and H2B (bonded with C2 atom) were found in subsequent difference Fourier maps and refined isotropically. Rest of hydrogen atoms were placed in idealized positions and refined as 'riding model' with isotropic displacement parameters set at 1.2 times U_{eq} of appropriate carrier atoms.

Table S2 lists the relevant experimental data and refinement details. Crystallographic data (excluding structure factors) for the structural analysis has been deposited with the Cambridge Crystallographic Data Centre, Nos. CCDC-2277547 (**P17**). Copies of this information may be obtained free of charge from: The Director, CCDC, 12 Union Road, Cambridge, CB2 1EZ, UK. Fax: +44(1223)336-033, email:deposit@ccdc.cam.ac.uk, or www: www.ccdc.cam.ac.uk.

Figure S49. A perspective view of the molecule **P17**. Displacement ellipsoids are drawn at the 50% probability level.

Table S2. Crystal data, data collection and structure refinement

7. References

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