

# **Modulating surface electron density of Ni(OH)<sub>2</sub> nanosheets with longitudinal Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene nanosheets by Schottky effect toward enhanced hydrogen evolution reaction**

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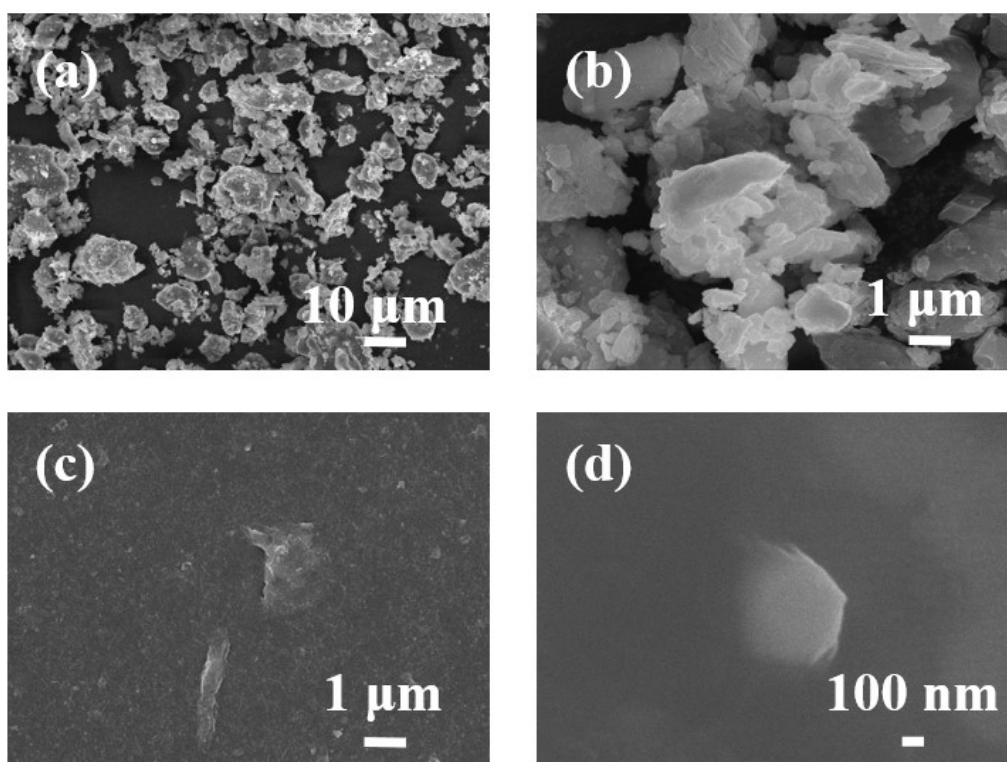
## **Physical Characterization**

X-ray diffraction (XRD) patterns were recorded on a Bruker D8 Advance instrument. Scanning electron microscopy (SEM) images were taken on a Carl Zeiss Ultra Plus. Transmission electron microscopy (TEM) images were taken on a JEOL JEM-2000FX and energy dispersive spectrometry (EDS) elemental mapping images were obtained using the same instrument. X-ray photoelectron spectroscopy (XPS) was conducted on a VG Escalab210 using Mg 300 W X as an X-ray source and the C 1 s peak at 285.0 eV as a reference binding energy calibration. Ultraviolet photoelectron spectroscopy (UPS) measurements were carried out on a Prevac spectrometer with a VG Scienta R3000 hemispherical electron energy analyzer. Photons with an energy of 21.22 eV generated by helium I were used for UPS spectrum generation. Specific surface areas of all

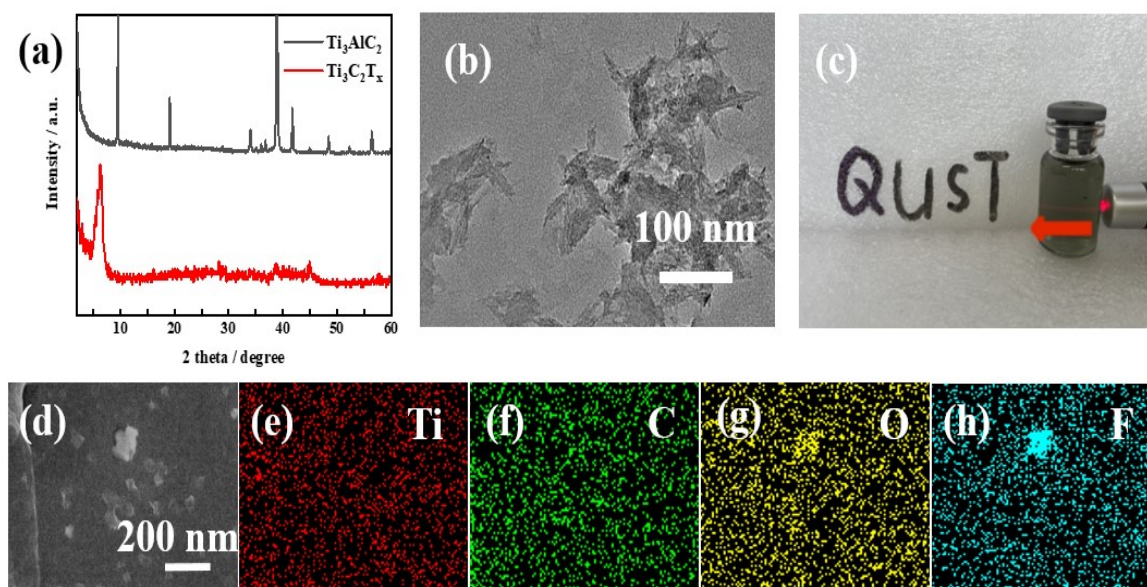
prepared samples were obtained according to Brunauer-Emmett-Teller (BET) method.

### **Electrochemical performance tests**

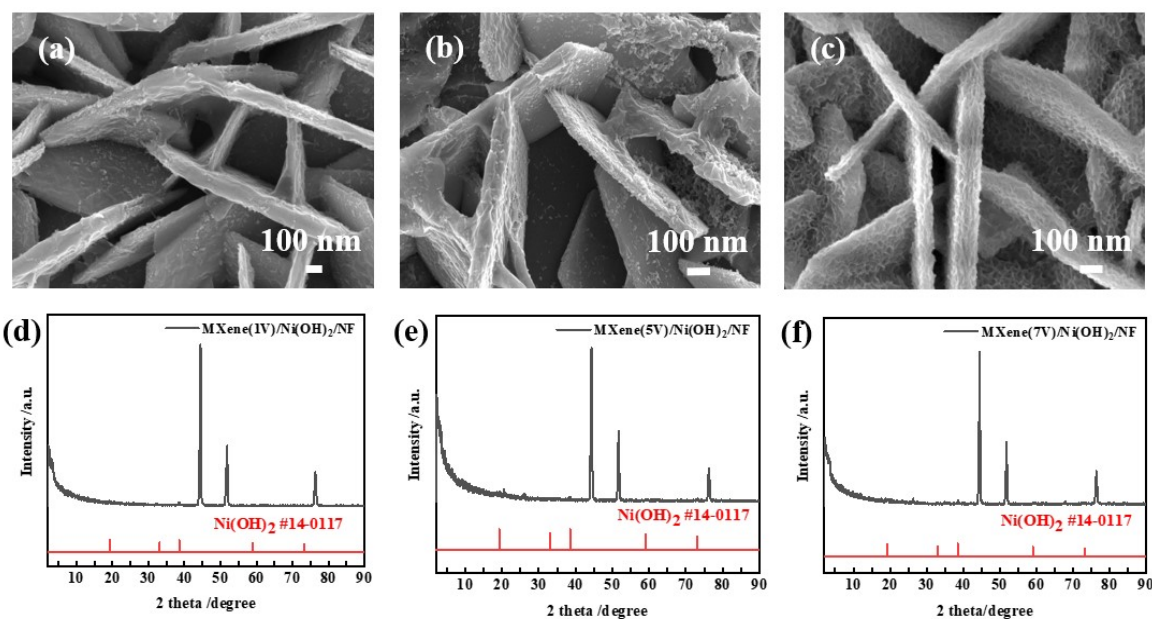
HER electrochemical tests were carried out in a three-electrode cell connected to the electrochemical workstation (CHI 660, CH Instruments) using as-prepared samples with a surface area of  $1\text{cm}^2$  as working electrodes and Hg/HgO, carbon rod and 1 M KOH as a reference electrode, counter electrode and electrolyte, respectively. The equation  $E_{\text{RHE}} = E_{\text{Hg/HgO}} + 0.059 \text{ pH} + 0.14 \text{ V}$  was used to establish the reversible hydrogen electrode (RHE) potential, and  $iR$  compensation was applied to all electrochemical experiments. Before tests, possible sample surface impurities were removed by scanning at  $-1 \sim -0.2 \text{ V}$  with a scan rate of  $50 \text{ mV s}^{-1}$  followed by linear sweep voltammetry (LSV) in 1 M KOH electrolyte at a scan rate of  $5 \text{ mV s}^{-1}$ . HER electrochemical impedance spectra (EIS) were obtained at potentials ranging from 0.1 Hz to 100 kHz. Tafel slopes of the samples were calculated using the Tafel equation  $\eta = b \log j + a$ , where  $b$  is the Tafel slope,  $\eta$  is the overpotential and  $j$  is the current density.



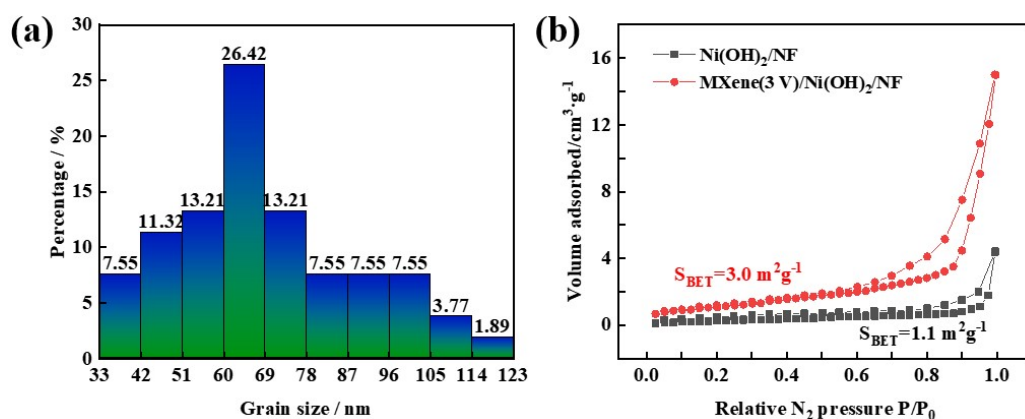
**Figure S1.** (a,b) SEM images of  $\text{Ti}_3\text{AlC}_2$  MAX; (c,d) SEM images of  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene.



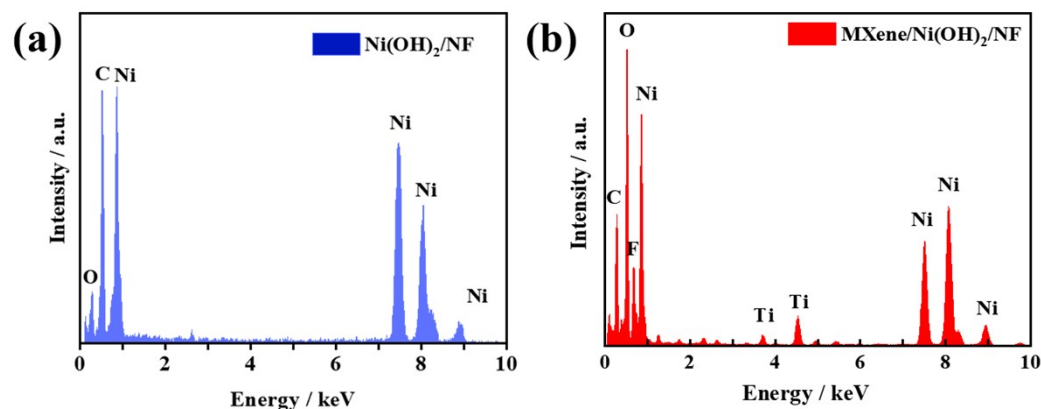
**Figure S2.** (a) XRD patterns of  $\text{Ti}_3\text{AlC}_2$  MAX and  $\text{Ti}_3\text{C}_2\text{T}_x$  MXenes; (b) TEM image of  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene; (c) Optical photograph of Tyndall phenomenon; (d)-(h) Elemental mapping images of Ti, C, O and F in  $\text{Ti}_3\text{C}_2$  MXene.



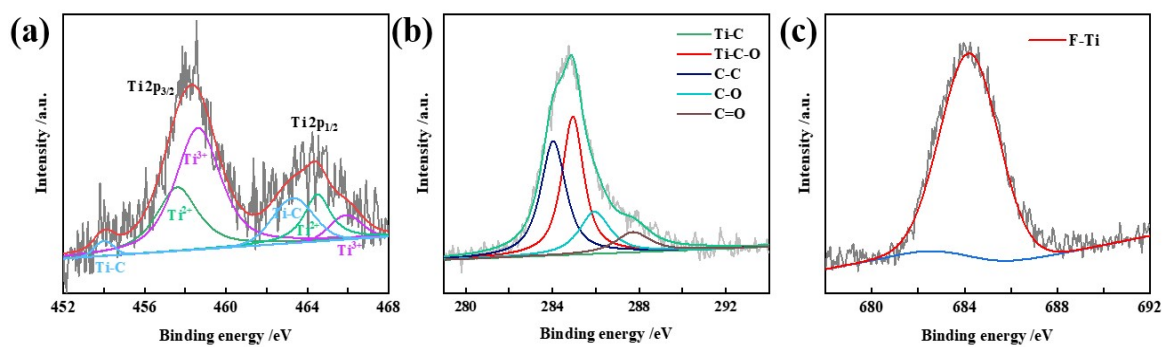
**Figure S3.** SEM images of (a) MXene(1 V) /Ni(OH)<sub>2</sub>/NF, (b) MXene(5 V) /Ni(OH)<sub>2</sub>/NF and (c) MXene(7 V) /Ni(OH)<sub>2</sub>/NF; XRD images of (d) MXene(1 V)/Ni(OH)<sub>2</sub>/NF, (e) MXene(5 V)/Ni(OH)<sub>2</sub>/NF and (f) MXene(7 V)/Ni(OH)<sub>2</sub>/NF.



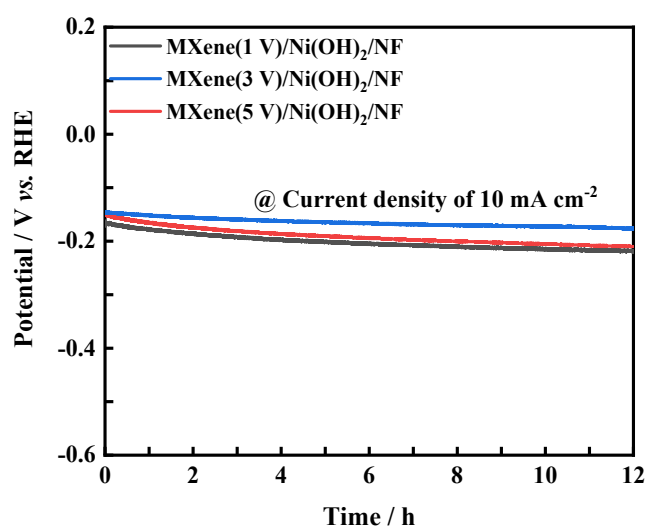
**Figure S4.** (a) Grain sizes of MXene in MXene/Ni(OH)<sub>2</sub>/NF; (b) N<sub>2</sub> adsorption-desorption isotherms of Ni(OH)<sub>2</sub>/NF and MXene/Ni(OH)<sub>2</sub>/NF.



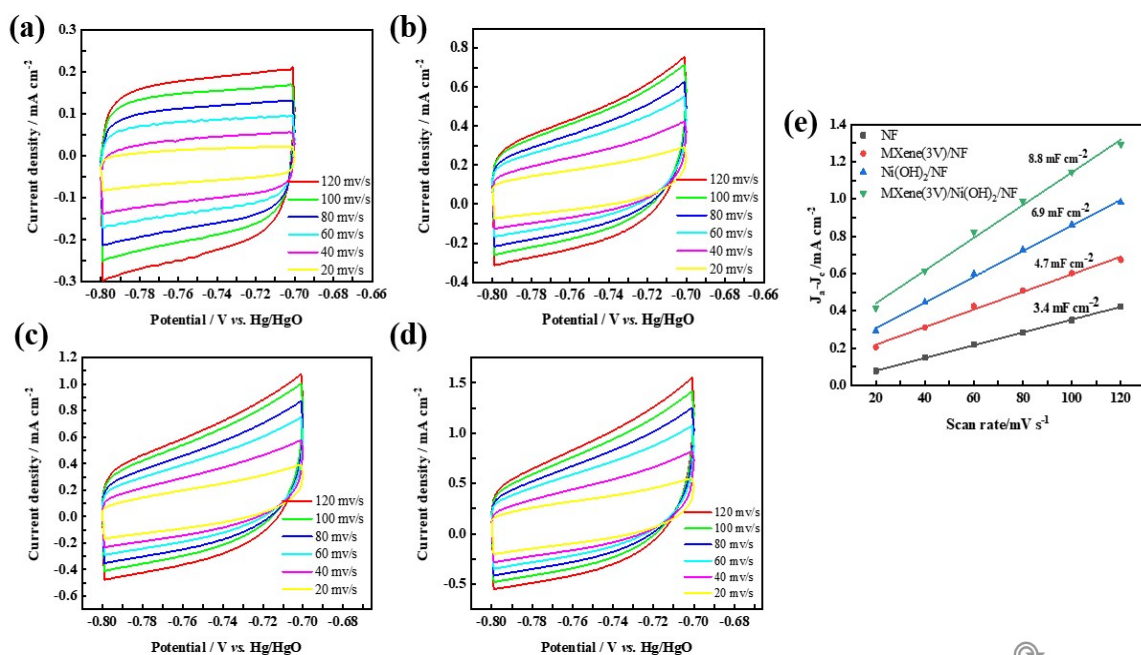
**Figure S5.** (a,b) EDS images of Ni(OH)<sub>2</sub>/NF and Ti<sub>3</sub>C<sub>2</sub>T<sub>x</sub> MXene/Ni(OH)<sub>2</sub>/NF.



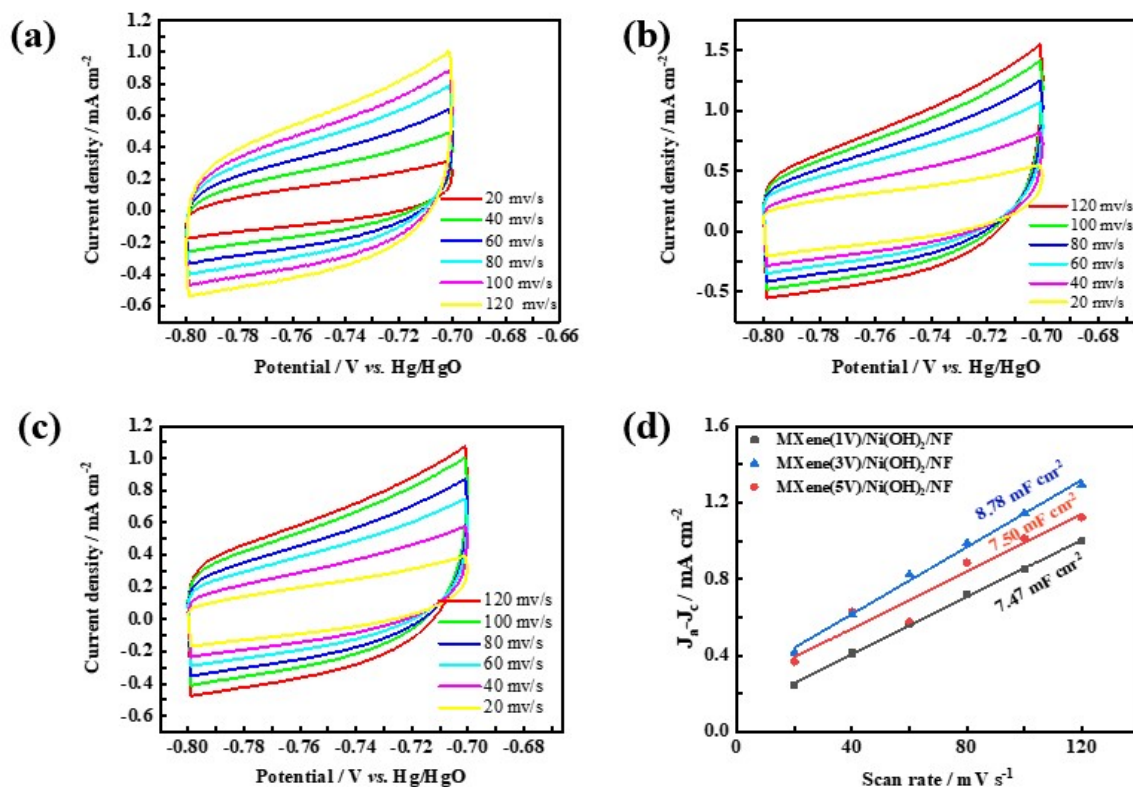
**Figure S6.** XPS of  $\text{Ti}_3\text{C}_2\text{T}_x$  MXene/ $\text{Ni}(\text{OH})_2/\text{NF}$ : (a) Ti 2p, (b) C 1s, (c) F 1s.



**Figure S7.** Chronopotential (CP) curve of the Mxene(1 V)/ $\text{Ni}(\text{OH})_2/\text{NF}$ , Mxene(3 V)/ $\text{Ni}(\text{OH})_2/\text{NF}$  and Mxene(5 V)/ $\text{Ni}(\text{OH})_2/\text{NF}$  samples at  $10 \text{ mA cm}^{-2}$ .



**Figure S8.** Cyclic voltammetry curves of (a) NF, (b) MXene(3 V)/NF, (c) Ni(OH)<sub>2</sub>/NF and (d) MXene(3 V)/Ni(OH)<sub>2</sub>/NF; (e) linear fitting of HER  $\Delta j$  vs. scan rates at +0.216 V vs. RHE.



**Figure S9.** Cyclic voltammetry curves of (a) MXene(1V)/Ni(OH)<sub>2</sub>/NF, (b) MXene(3 V)/Ni(OH)<sub>2</sub>/NF, (c) MXene(5 V)/Ni(OH)<sub>2</sub>/NF; (d) linear fitting of HER  $\Delta j$  vs. scan rates at +0.216 V vs. RHE.

**Table S1.** HER performance of MXene/Ni(OH)<sub>2</sub>/NF compared with previously reported electrocatalysts.

Catalysts	$\eta_{10}$ (mV)	Substrate	References
Ti <sub>3</sub> C <sub>2</sub> T <sub>x</sub> MXene/Ni(OH) <sub>2</sub> /NF	66	NF	This work
P-CoFe-LDH@MXene/NF	85	NF	Applied Catalysis B:Environmental 299 (2021) 120660
FeNi@Mo <sub>2</sub> TiC <sub>2</sub> T <sub>x</sub> @NF	165	NF	Nano Research 2021, 14(10):3474–3481
Ni <sub>0.7</sub> Fe <sub>0.3</sub> PS <sub>3</sub> @MXene	196	--	Adv. Energy Mater. 2018, 1801127
Ni <sub>2</sub> P/NF	116	NF	NanoResearch volume 13, pages2098–2105 (2020)
Ru-NiFeP/NF	105	NF	Appl. Surf. Sci. 2021, 536, 147952
NiFeLa-LDH/v-MXene/NF	38	NF	Journal of Energy Chemistry 70 (2022) 472–479
NiO@Ni <sub>2</sub> P/NF	79	NF	International Journal of Hydrogen Energy 47 (2022) 17097-17106
NiFeS@Ti <sub>3</sub> C <sub>2</sub> MXene/NF	180	NF	Applied Catalysis B:Environmental 321 (2023) 122039
TiVCT <sub>x</sub> @NF	151	NF	RSC Adv, 2022, 12, 23584–23594