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Supporting Information

Synthesis, Structure and Spectroscopic Properties of BODIPY Dyes Incorporating the Pentafluorosulfanylphenyl Group

Richard D. James, Fabio Cucinotta, Paul G. Waddell and Andrew C. Benniston

Chemistry-School of Natural and Environmental Sciences, Newcastle University,

Newcastle upon Tyne, NE1 7RU, UK.

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Figure S1. DFT calculated Kohn-Sham molecular orbitals using the B3LYP functional and the 6-311G(d,p) basis set for BD3.

Electronic Transition (oscillator strength)	Energy / eV (nm)	Orbitals	% Contribution
1 (0.6354)	2.85 (434)	$HOMO \rightarrow LUMO$	100
2 (0.0498))	3.87 (320)	HOMO-1 \rightarrow LUMO	100
3 (0.0002)	4.28 (290)	HOMO \rightarrow LUMO+1	100



Figure S2. DFT calculated Kohn-Sham molecular orbitals using the B3LYP functional and the 6-311G(d,p) basis set for BD4.

Table S2. TD-DFT calculated parameters for the carboxylic acid version of BD	4 in vacuo.
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Electronic Transition (oscillator strength)	Energy / eV (nm)	Orbitals	% Contribution
1 (1.1086)	2.12 (586)	$HOMO \rightarrow LUMO$	100
2 (2.1656)	3.54 (351)	HOMO-1 → LUMO HOMO → LUMO+2	79 21
3 (0.1365)	3.67 (338)	HOMO → LUMO+2 HOMO-1 → LUMO HOMO-1 → LUMO+3	74 22 4



Figure S3. DFT calculated Kohn-Sham molecular orbitals using the B3LYP functional and the 6-311G(d,p) basis set for BD5.

Electronic Transition (oscillator strength)	Energy / eV (nm)	Orbitals	% Contribution
1 (1.0604)	2.23 (556)	$HOMO \rightarrow LUMO$	100
2 (2.2145)	3.40 (365)	HOMO-1 → LUMO HOMO→ LUMO+1 HOMO-2 → LUMO HOMO-2 → LUMO+2 HOMO-1 → LUMO+2	86 5 3 3 3
3 (0.5933)	3.79 (327)	HOMO → LUMO+1 HOMO-1 → LUMO HOMO-1 → LUMO+4 HOMO-2 → LUMO+4	89 5 3 2

Table S3. TD-DFT calculated parameters for the carboxylic acid version of BD5 in vacuo.



Figure S4. DFT calculated Kohn-Sham molecular orbitals using the B3LYP functional and the 6-311G(d,p) basis set for **BD6**.

Electronic Transition (oscillator strength)	Energy / eV (nm)	Orbitals	% Contribution
1 (1.2547)	2.07 (599)	HOMO→LUMO HOMO-1→LUMO+1	97 3
2 (3.8114)	3.31 (375)	HOMO-1 \rightarrow LUMO HOMO \rightarrow LUMO+1 HOMO-2 \rightarrow LUMO+1 HOMO-1 \rightarrow LUMO+4 HOMO \rightarrow LUMO+6	47 40 5 4 4
3 (0.0999)	3.64 (341)	HOMO-3→LUMO HOMO→LUMO+4 HOMO-2→LUMO HOMO-1→LUMO+1 HOMO-1→LUMO+7	41 26 16 12 5



Figure S5. (A) Absorption and emission spectra recorded for **BD1** in CH₃CN at room temperature. (B) Fluorescence decay trace (top) and instrument response function (bottom) measured for **BD1** in CH₃CN using single-photon-counting.



Figure S6. (A) Absorption and emission spectra recorded for **BD2** in CH₃CN at room temperature. (B) Fluorescence decay trace (top) and instrument response function (bottom) measured for **BD2** in CH₃CN using single-photon-counting.



Figure S7. Absorption and emission spectra recorded for BD3 in CH₃CN at room temperature.



Figure S8. Absorption and emission spectra recorded for BD4 in CH_3CN at room temperature.



Figure S9. Absorption and emission spectra recorded for BD5 in CH_3CN at room temperature.



Figure S10. Absorption and emission spectra recorded for BD6 in CH₃CN at room temperature.



Figure S11. Absorption and emission spectra recorded for **BD4** (a and b), **BD5** (c and d) and **BD6** (e and f) in a range of solvents at room temperature.



Figure S12. Plots of absorption maxima and emission maxima versus the solvent polarizability function for BD4 (a and b), BD5 (c and d) and BD6 (e and f).



Figure S13. Catalán's Solvent Polarity (SPP) and Lippert Mataga plots for **BD4** (top), **BD5** (middle) and **BD6** (bottom).



Figure S14. (a) Excitation spectrum measured at 650 nm, (b) fluorescence spectrum using 480 nm excitation and (c) fluorescence lifetime decay for a crystalline sample of **BD1**.



Figure S15. (a) Excitation spectrum measured at 630 nm, (b) fluorescence spectrum using 400 nm excitation and (c) fluorescence lifetime decay for a crystalline sample of **BD2**.



Figure S16. Representation of planes, angles and distances with regard to the electronic transition dipoles (red) for **BD1** as obtained from the crystal packing diagram.



Figure S17. Representation of a dimer for BD1 and separation distance as obtained from the crystal packing diagram.



Figure S18. Representation of an offset face-to-face dimer for BD1 and separation distance as obtained from the crystal packing diagram.



Figure S19. Representation of a dimer for BD1 between separate stacks and the separation distance and angles as obtained from the crystal packing diagram.



Figure S20. Representation of an offset co-linear dimer for BD1 and separation distance and offset angle as obtained from the crystal packing diagram.



Figure S21. Representation of dimers for BD2 and separation distance as obtained from the crystal packing diagram.



Figure S22. Representation of pseudo co-linear dimers for BD2 and their separation distances as obtained from the crystal packing diagram. View is down the c-axis.



Figure S23. A 300 MHz ¹H NMR spectrum of 2 recorded in CDCl₃.



Figure S24. A 176 MHz ¹³C{¹H} NMR spectrum of 2 recorded in CDCI₃.



Figure S25. A 282 MHz ¹⁹F NMR spectrum of **2** recorded in CDCl₃.



Figure S26. A 700 MHz ¹H NMR spectrum of BD2 recorded in CDCl₃.



Figure S27. A 176 MHz ¹³C{¹H} NMR spectrum of BD2 recorded in CDCl₃.



Figure S28. A 300 MHz ¹H NMR spectrum of BD3 recorded in CDCl₃.



Figure S29. A 75 MHz $^{13}C\{^{1}H\}$ NMR spectrum of BD3 recorded in CDCl₃.



Figure S30. A 700 MHz ¹H NMR spectrum of BD5 recorded in CDCl₃.



Figure S31. A 176 MHz ¹³C{¹H} NMR spectrum of BD5 recorded in CDCl₃.



Figure S32. A 700 MHz ¹H NMR spectrum of **BD6** recorded in CDCI₃. X = trace solvent peaks from acetone and petrol.



Figure S33. A 176 MHz ¹³C{¹H} NMR spectrum of BD6 recorded in CDCl₃.



Figure S34. A 659 MHz ¹⁹F NMR spectrum of BD6 recorded in CDCl₃.



Figure S35. A 96 MHz ¹¹B NMR spectrum of BD6 recorded in CDCl₃.



Figure S36. TOF mass spectrum of BD3 and the comparison of calculated and observed data.



Figure S37. TOF mass spectrum of BD6 and the comparison of calculated and observed data.