

Development of a novel Cd(II) metal complex for solvent sensitive detection of Zn(II) and Mg(II) with the formation of Cd(II)-Zn(II)/Cd(II)-Mg(II) complexes and their application in effective schottky devices

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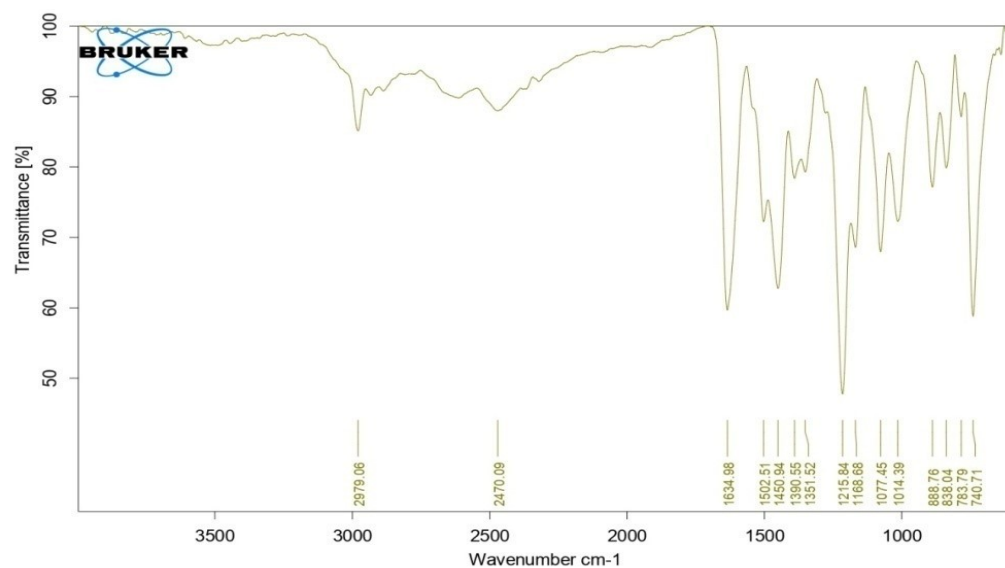


Fig. S1. FTIR of Complex **1**

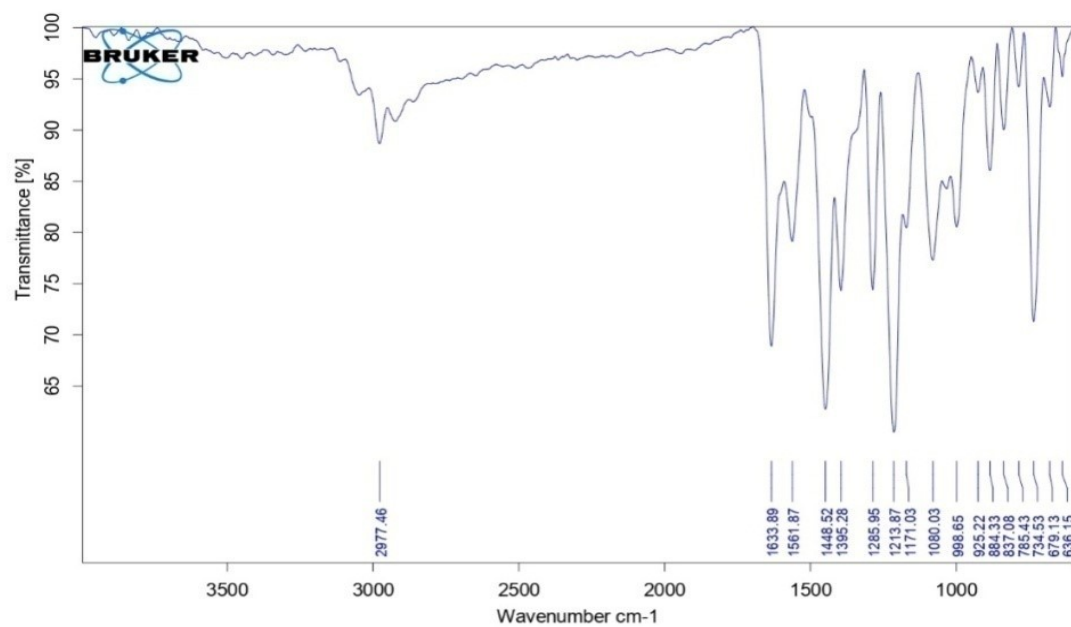


Fig. S2. FTIR of Complex **2**

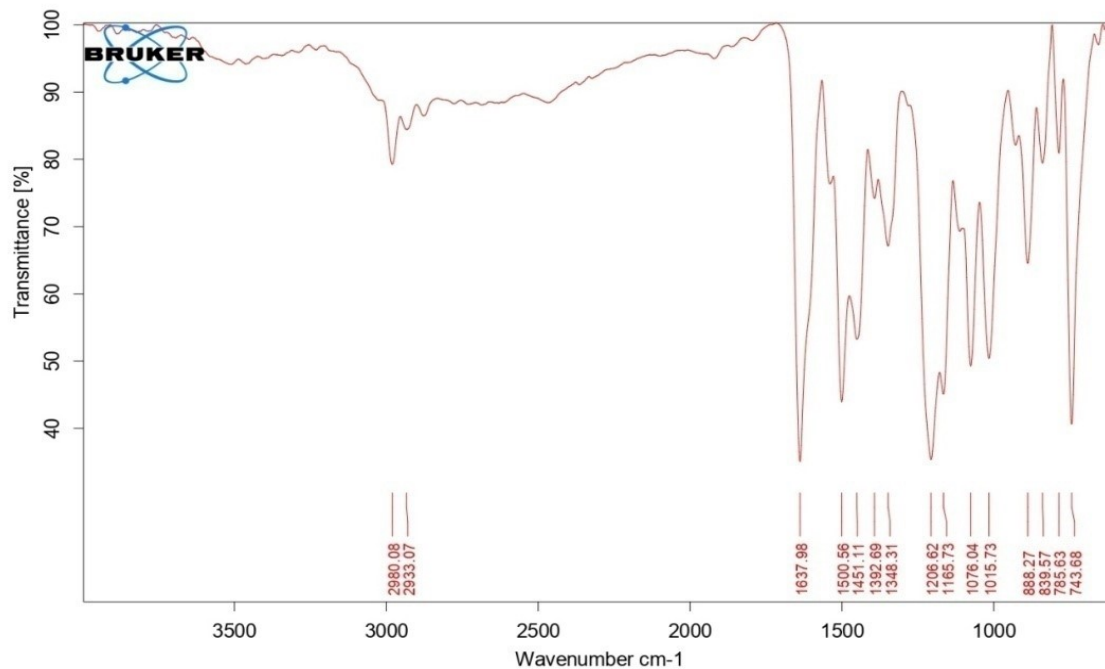


Fig. S3. FTIR of Complex 3

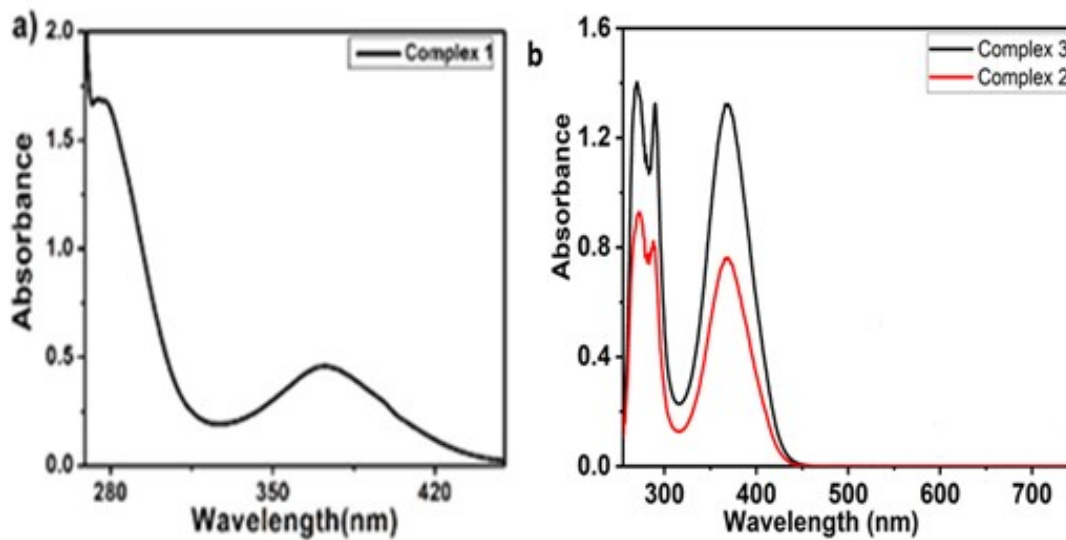


Fig. S4. Uv spectra of (a) Complex 1, (b) Complex 2 and 3

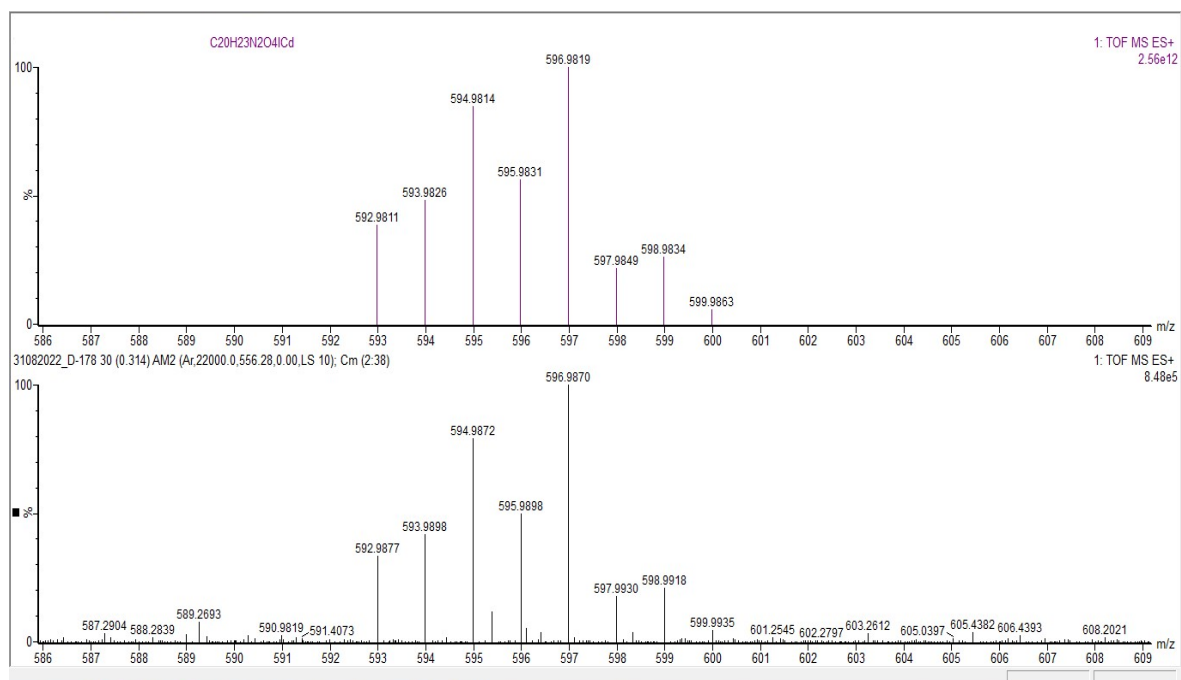


Fig. S5. Mass spectra of Complex 1

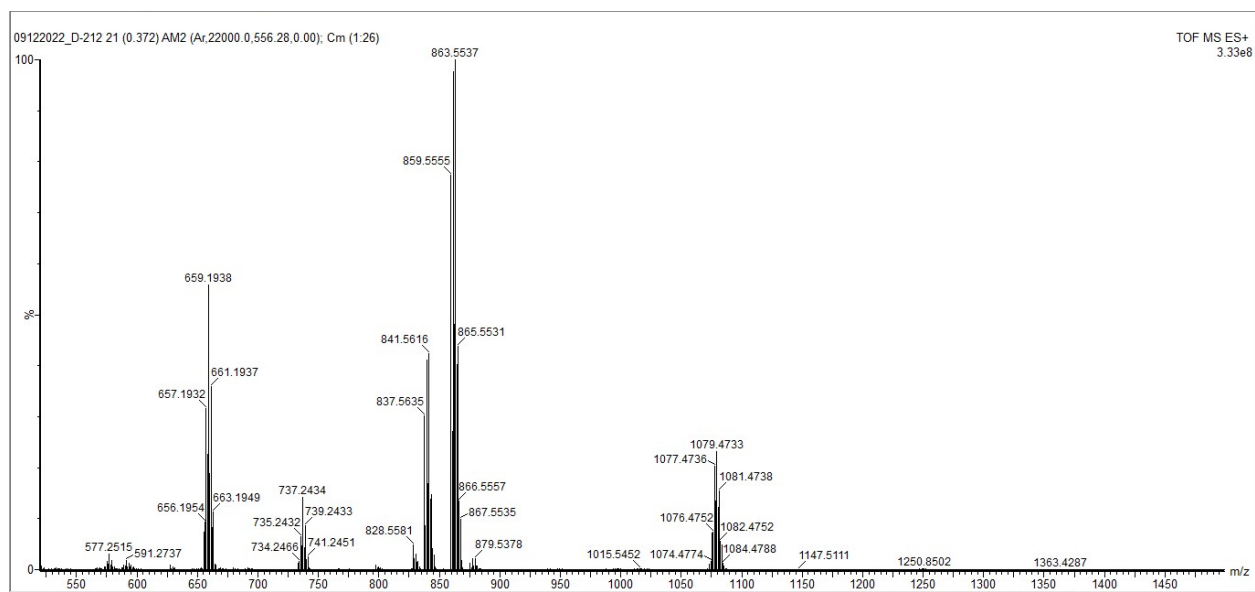


Fig. S6. Mass spectra of Complex 2

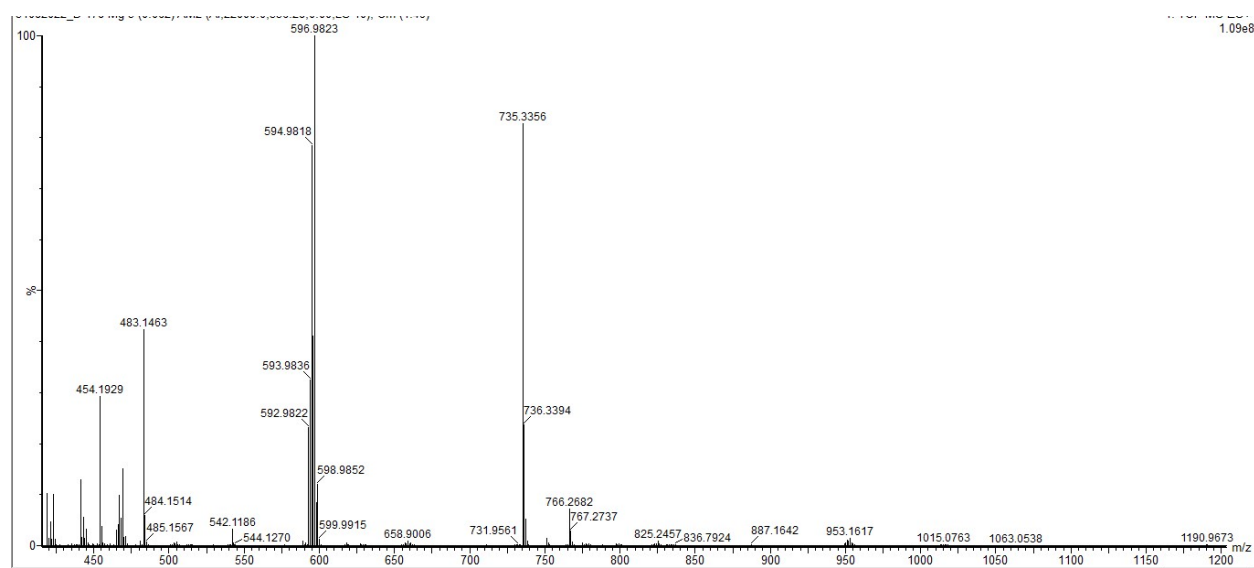


Fig. S7. Mass spectra of Complex 3

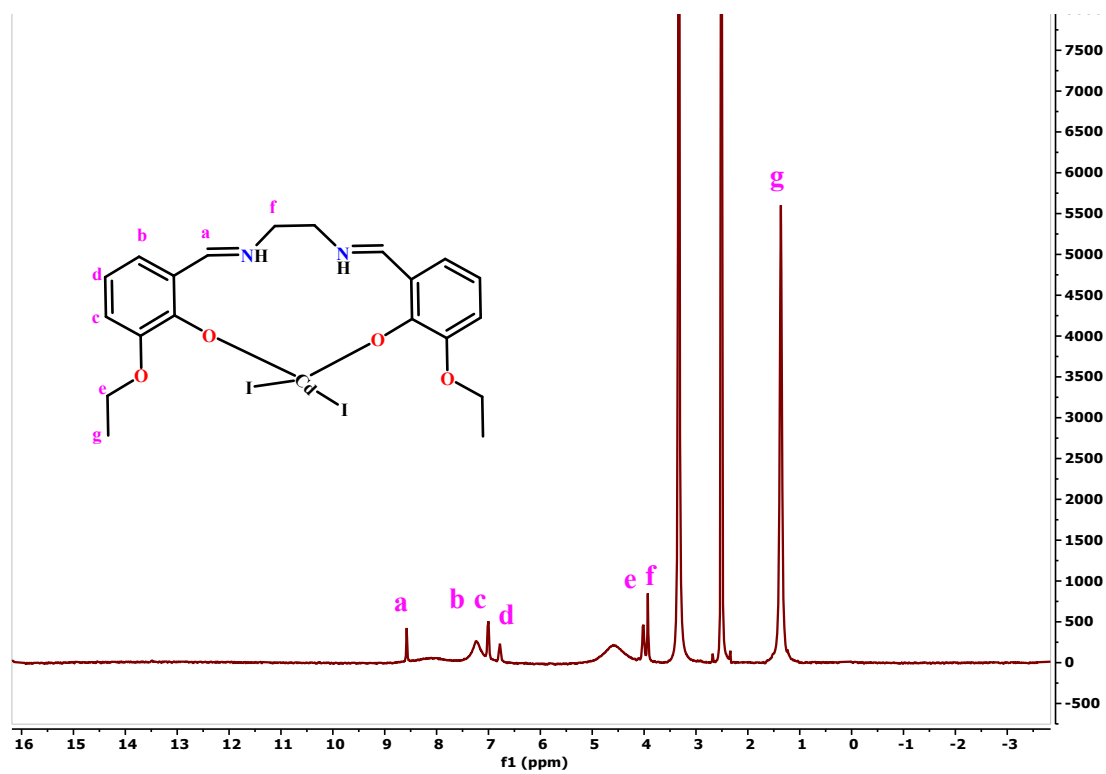


Fig. S8: ¹H NMR spectrum of complex 1

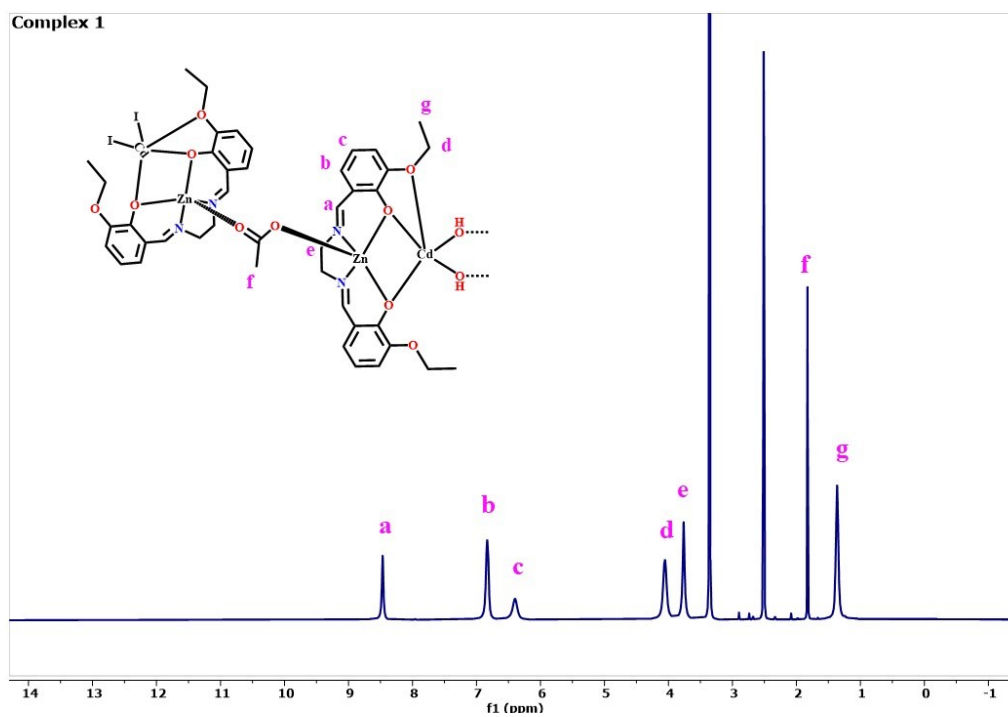


Fig. S9: NMR spectrum of complex 2

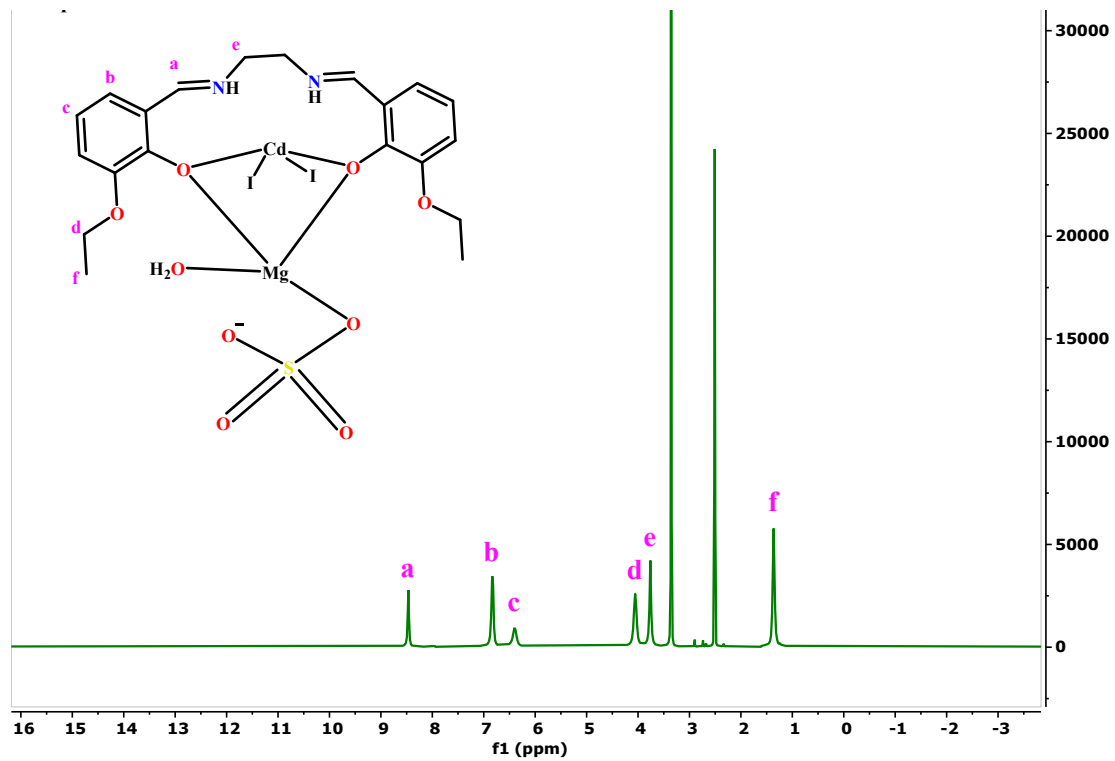


Fig S10: NMR spectrum of complex 3

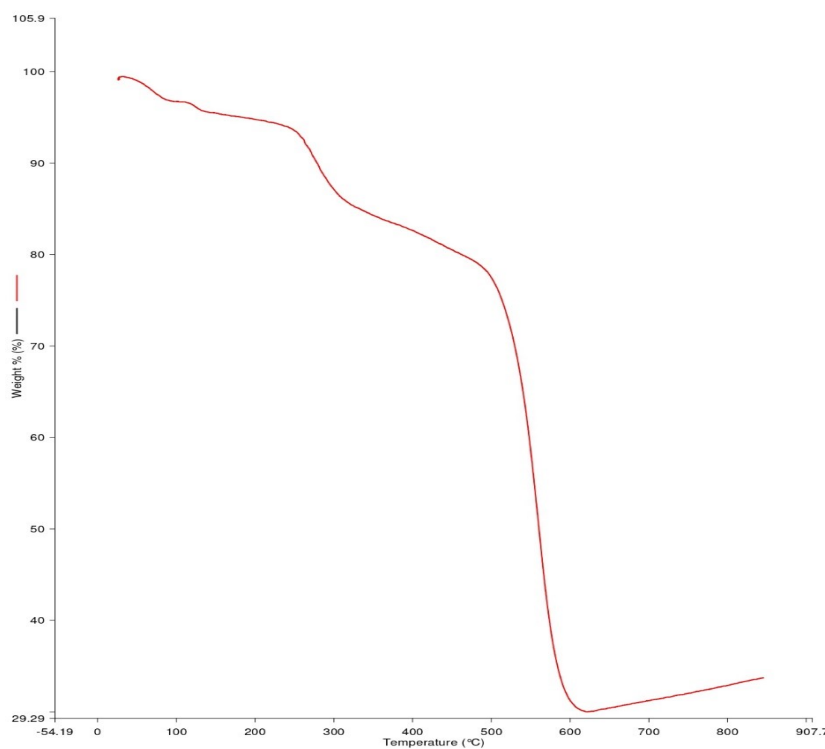


Fig. S11: The TGA graph for complex 2

TGA-DTA plot of complex 2

Weight loss calculation:

Monoclinic 'C 2/c' space group, Z = 4

Therefore FW = Unitcell contents/Z

= [4 Ligand + 2 CdI₂ + 6 Zn cations + 2 acetate (coordinated) + 2 oxygen atom bridged] + ~ 226 electron [SQUEEZE result]

= 2680 + (7 X 18)

= 2680 + 126 = 2806

Weight loss for 7 H₂O = 126/2806

= 4.5% (experimental 4.8 %)

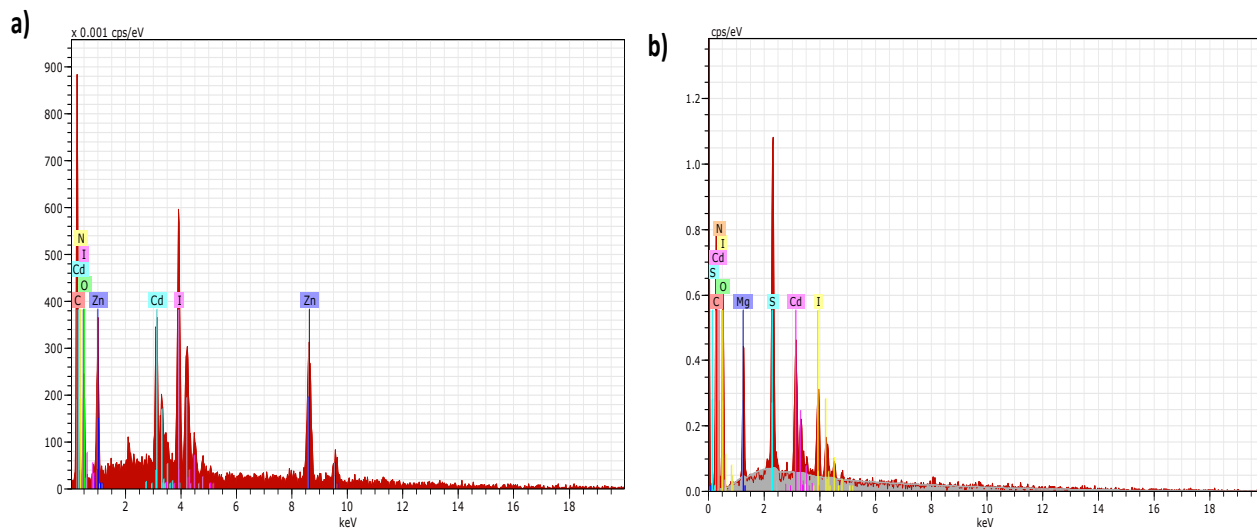


Figure S12: EDX result of a) complex 2 b) complex 3

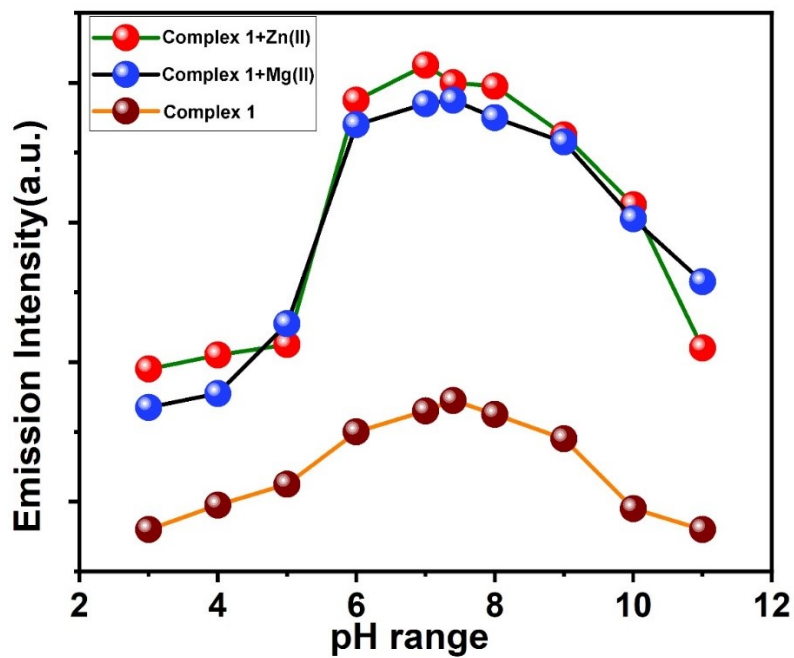


Figure. S13. Change of Emission Intensity of the complex 1, 2, 3 with the variation of pH

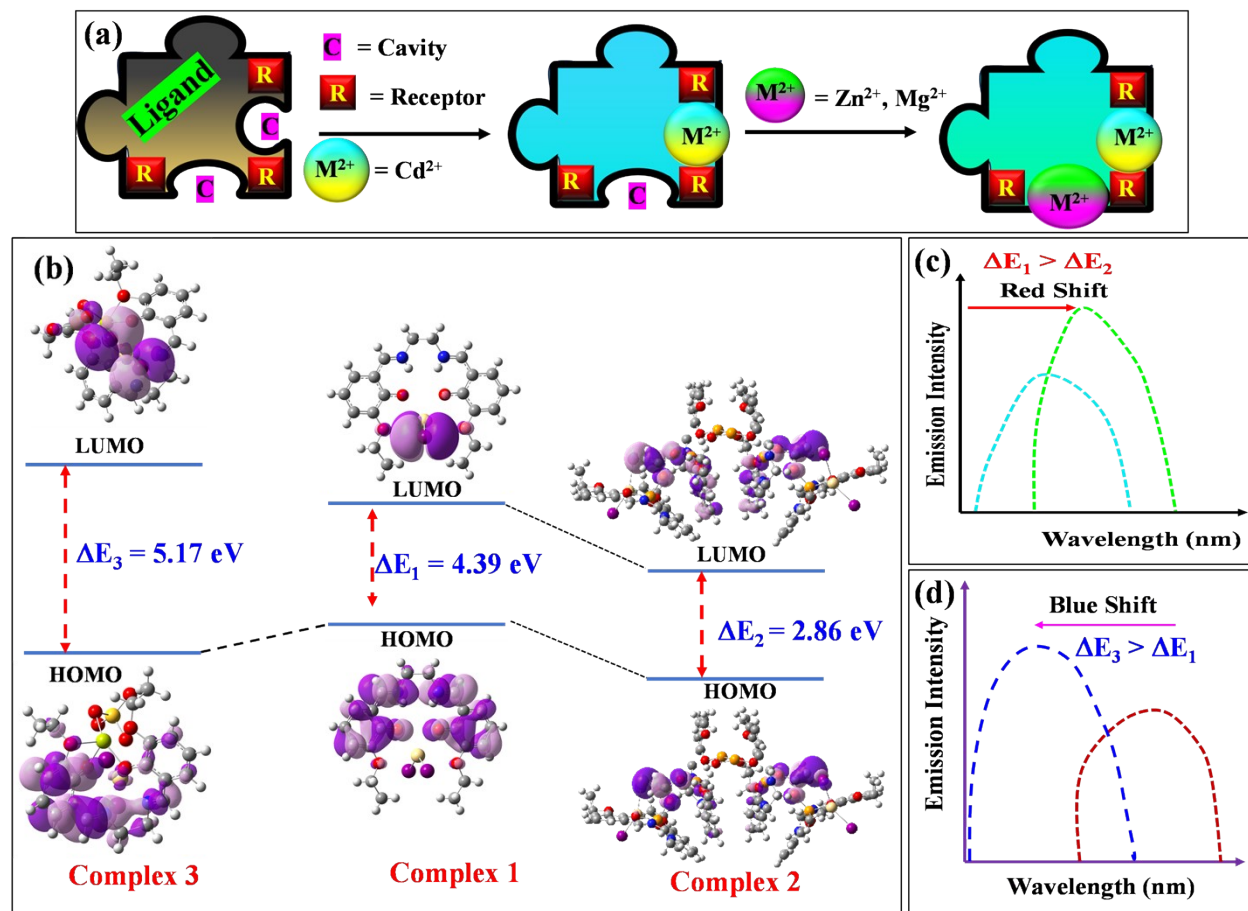


Fig S14: Schematic presentation of the mechanism of sensing phenomena of complex 1 towards Zn(II) and Mg(II).

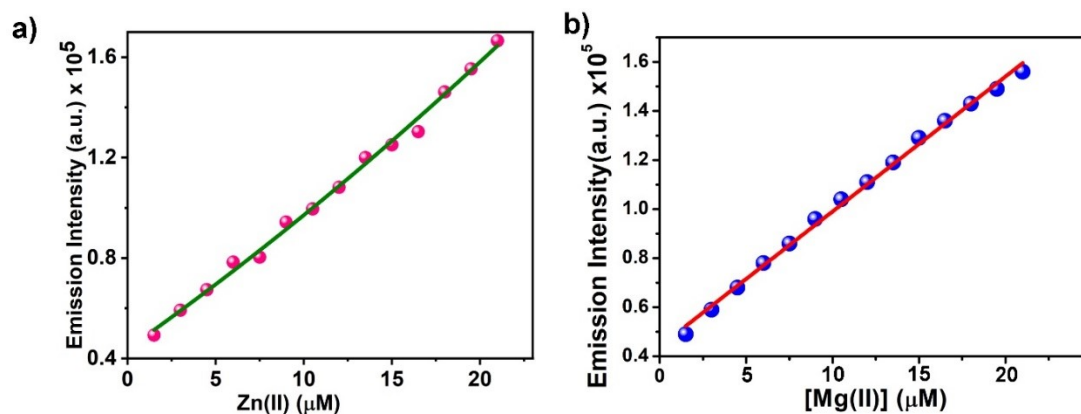


Fig. S15: Change of Fluorescence intensity of complex 1 ($3 \times 10^{-6} \text{ M}$) by varying the concentration of (a) Zn(II) ($1.0 - 25 \times 10^{-6} \text{ M}$) in 9:1 ethanol-water medium and (b) Mg(II) ($1.0 - 25 \times 10^{-6} \text{ M}$) in pure HEPES buffer solution for LOD calculation purpose.

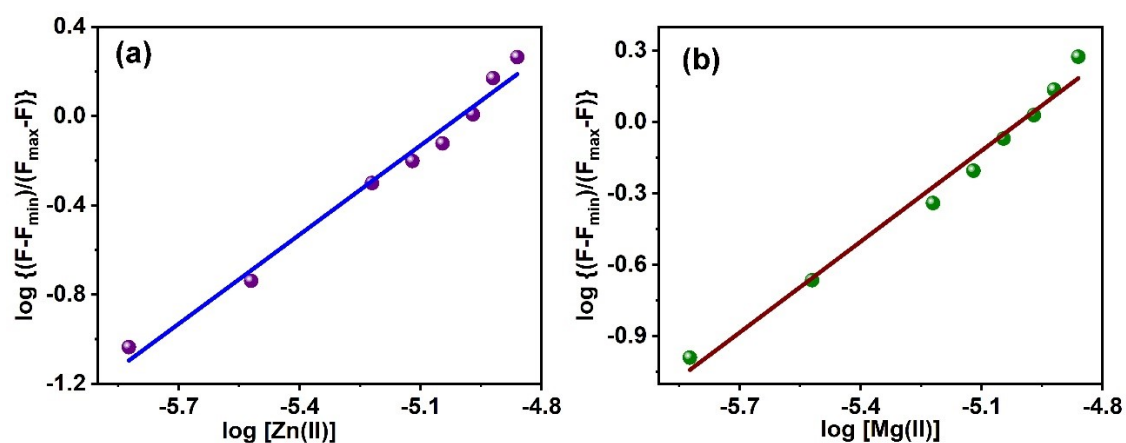


Fig. S16: Determination of Binding constant of (a) Complex 1 – Zn(II) and (b) Complex 1 – Mg(II) from Fluorescence titration results

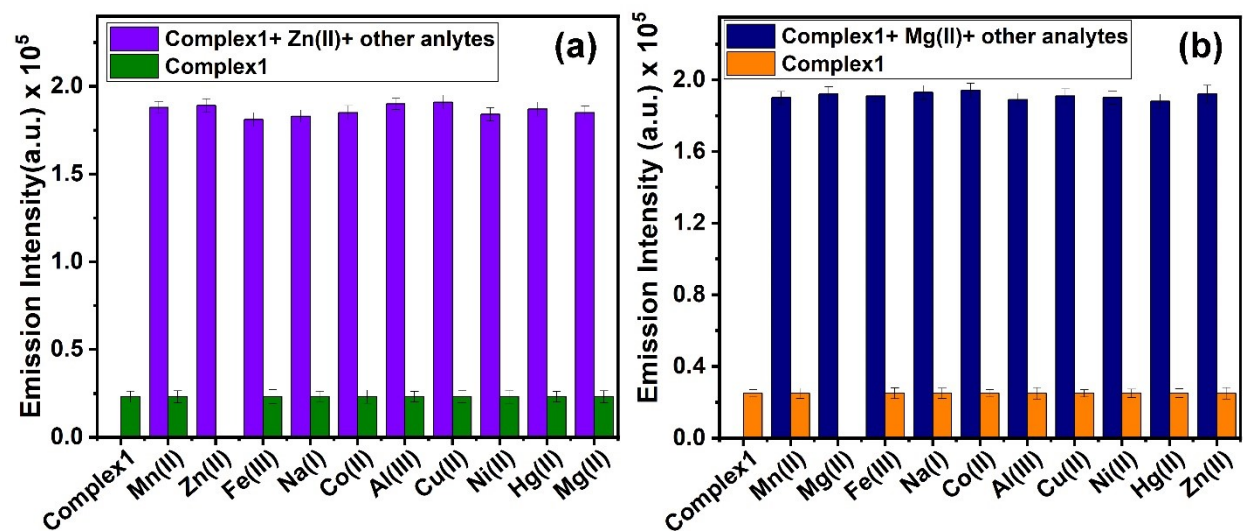


Fig. S17: Change of emission intensity of (a) complex1 – Zn(II) and (b) Complex1 – Mg(II) after the addition of several competitive metal ions.

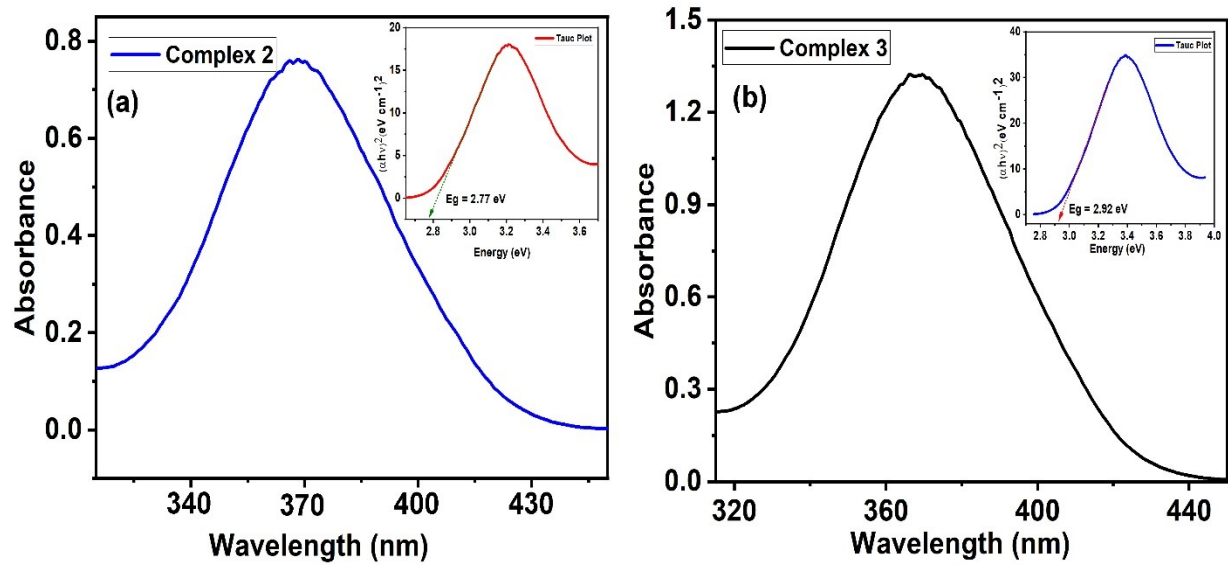


Fig S18: The band gap values for a) complex 2 b) complex 3 by using Tauc plot

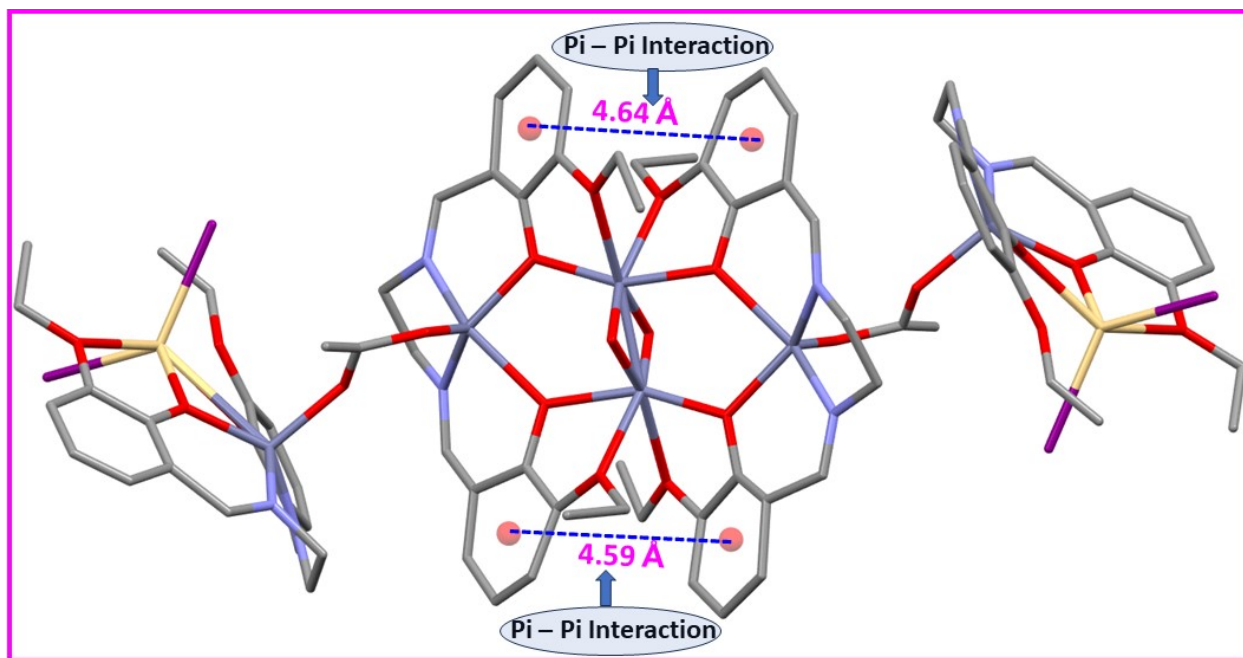


Fig. S19: Intramolecular π - π interaction in complex 2

Table S1: Intra-molecular Hydrogen bonding parameter of complex 1

D—H...A	D—H (Å)	H...A (Å)	D...A (Å)	D—H...A (°)	Symmetry operation for A
N(1)—H(1)...O2	0.900(19)	1.81(3)	2.565(3)	139(3)	x, y, z

Table S2: C—H... I interaction parameter of complex 2

C—H...I	C—H (Å)	H...I (Å)	C...I (Å)	C—H...I (°)	Symmetry operation for A
C(19)—H(19)...I2	0.9500	3.149	3.994	148.85	-x+1, -y+2, -z+1

Table S3: Crystal parameters of complex 1 and 2

Crystal parameters	Complex 1	Complex 2
CCDC No.	2246285	2246286
Empirical formula	C ₂₀ H ₂₄ N ₂ O ₄ I ₂ Cd	C _{82.99} H _{95.81} Cd ₂ I ₄ N ₈ O ₂₂ Zn ₆
Formula weight	722.61	2681.93
Crystal size/mm	0.14 X 0.09 X 0.04	0.095 X 0.059 X 0.032
Crystal system	Monoclinic	Monoclinic
Space group	'P 2/n'	'C 2/c'
a /Å	8.6378(19)	37.556(4)
b /Å	11.746(3)	12.9596(14)
c /Å	11.427(3)	27.018(3)
α/°	90	90
β/°	92.577(11)	125.815(3)

$\gamma/^\circ$	90	90
Volume/ \AA^3	1158.2(4)	10663.5(19)
Z	2	4
$D_{\text{calc}}/\text{gcm}^{-3}$	2.072	1.669
F(000)	688	5247
$\mu \text{ MoK}\alpha / \text{mm}^{-1}$	3.636	2.938
Temperature/K	296(2)	148.53
R_{int}	0.0290	0.1179
Range of h, k, l	-11/11, -16/14, -15/15	-42/53, -18/17, -38/31
$\theta_{\text{min/max}}/^\circ$	2.488/29.423	2.189/30.465
Reflections collected/unique/observed [$I > 2\sigma(I)$]	14233/3200/2631	64906/15869/7654
Data/restraints/ parameters	3200/1/137	15869/0/636
Goodness of fit on F^2	1.093	1.038
Final R indices	$R_1 = 0.0273$	$R_1 = 0.1379$
[$I > 2\sigma(I)$]	$wR_2 = 0.0396$	$wR_2 = 0.2328$
R indices (all data)	$R_1 = 0.0493$	$R_1 = 0.3471$
	$wR_2 = 0.0547$	$wR_2 = 0.4195$