

Influence of bridging atom in copper-based coordination polymers for enhancing urease inhibition activity

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Table S1 Crystallographic data for **Cu-CP-1** and **Cu-CP-2**.

Cu-CP	1	2
Empirical formula	C ₄₉ H ₅₅ ClCu ₂ N ₈ O ₁₆	C ₃₀ H ₂₄ Cu ₃ N ₄ O ₁₂
F_w	1174.54	823.15
Crystal system	Monoclinic	Triclinic
Space group	<i>C2/c</i>	<i>P-1</i>
a (Å)	14.4011(8)	8.3526(10)
b (Å)	16.9452(8)	8.6543(10)
c (Å)	22.2478(12)	10.7547(13)
α (°)	90	85.213(2)
β (°)	101.857(2)	75.719(2)
γ (°)	90	80.446(2)
V (Å ³), Z, T (K)	5313.3(5), 4, 296(2)	742.20(15), 1, 296(2)
$D_c/g\text{ cm}^{-3}$, F(000)	1.468, 2432	1.842, 415
Goodness-of-fit on F^2	1.030	1.019
Reflections collected	42125	3804
Unique data, R_{int}	6977, 0.0495	2586, 0.0107
θ Range (°)	2.23–28.92	1.96–25.00
R_1 ($I > 2\sigma(I)$) ^a	0.0437	0.0296
wR_2^b (all data) ^a	0.1256	0.0842

$$^a R_1 = \frac{\sum ||F_o| - |F_c||}{\sum |F_o|}, ^b wR_2 = \frac{\sum [w(F_o^2 - F_c^2)^2]}{\sum [w(F_o^2)^2]}^{1/2}.$$

Table S2 Selected bond distances (Å) and angles (°) for Cu-CP-1.

Cu(1)–O(3)	1.9646(16)	Cu(1)–O(1)	1.9850(17)
Cu(1)–N(1)	2.025(2)	Cu(1)–N(2)	2.055(2)
Cu(1)–Cl(1)	2.5447(7)	O(3)–Cu(1)–O(1)	159.83(8)
O(3)–Cu(1)–N(1)	89.05(8)	O(1)–Cu(1)–N(1)	87.76(8)
O(3)–Cu(1)–N(2)	90.64(7)	O(1)–Cu(1)–N(2)	90.56(7)
N(1)–Cu(1)–N(2)	174.25(8)	O(3)–Cu(1)–Cl(1)	111.19(6)
O(1)–Cu(1)–Cl(1)	88.96(6)	N(1)–Cu(1)–Cl(1)	95.81(6)
N(2)–Cu(1)–Cl(1)	89.65(6)		

Table S3 Selected bond distances (Å) and angles (°) for Cu-CP-2.

Cu(1)–O(6)	1.892(2)	Cu(1)–O(6)#1	1.892(2)
Cu(1)–O(1)#1	1.940(2)	Cu(1)–O(1)	1.940(2)
N(1)–Cu(2)	2.069(2)	Cu(2)–O(6)	1.880(2)
Cu(2)–O(3)#2	1.9641(19)	Cu(2)–O(2)	1.972(2)
O(6)–Cu(1)–O(6)#1	180	O(6)–Cu(1)–O(1)#1	86.95(9)
O(6)#1–Cu(1)–O(1)#1	93.05(9)	O(6)–Cu(1)–O(1)	93.05(9)
O(6)#1–Cu(1)–O(1)	86.94(9)	O(1)#1–Cu(1)–O(1)	180
O(6)–Cu(2)–O(3)#2	95.65(9)	O(6)–Cu(2)–O(2)	89.56(8)
O(3)#2–Cu(2)–O(2)	161.52(10)	O(6)–Cu(2)–N(1)	177.11(9)
O(3)#2–Cu(2)–N(1)	86.62(8)	O(2)–Cu(2)–N(1)	87.74(9)

Symmetry codes: #1 $-x + 1, -y, -z$; #2 $-x + 2, -y, -z$.**Table S4** Hydrogen bonding geometries (Å, °) of Cu-CP-1.

Cu-CP	D–H···A	D–H	H···A	D···A	D–H···A
1	N4–H4A···O2 ⁱ	0.86	2.10	2.8908	153

Symmetry code: ⁱ $1/2 - x, -1/2 + y, 1/2 - z$.

Table S5 Comparison of inhibition of urease by the Cu-CPs and other materials.

Materials	Percentage inhibition rate (100 μ M) ^a	IC ₅₀ \pm SEM (μ M) (3h) ^b	Reference
Cu ²⁺	87.50%	—	S1
3-dpye	—	—	—
HBCA	—	—	—
1,2-H ₂ BDC	—	—	—
DMSO	—	—	—
AHA	84.30%	37.20 \pm 4.00	S2
[2HL] ²⁺ ·[CuCl ₄] ²⁻			
L = N ¹ ,N ¹ ,N ² ,N ² -tetrakis(2-fluorobenzyl)ethane-1,2-diamine	96.70%	1.87 \pm 0.01	S3
[Cu(Hpz) ₄ Cl ₂]	99.70%	0.93 \pm 0.01	S4
Hpz = pyrazole			
tetrakis(m-bromo)-bis(m-4-chloro-2-((2-(piperidin-1-yl)ethyl)carboximidoyl)phenolato)-tri-copper(II)	—	1.00	S5
Cu(C ₁₉ H ₁₆ NO ₂) ₂	—	1.45	S6
Cu-CP-1	95.08	0.83 \pm 0.01	This work
Cu-CP-2	90.96	2.53 \pm 0.01	This work

^a The concentration of the tested material is 100 μ M.

^b IC₅₀ values represent mean \pm SD from three different experiments.

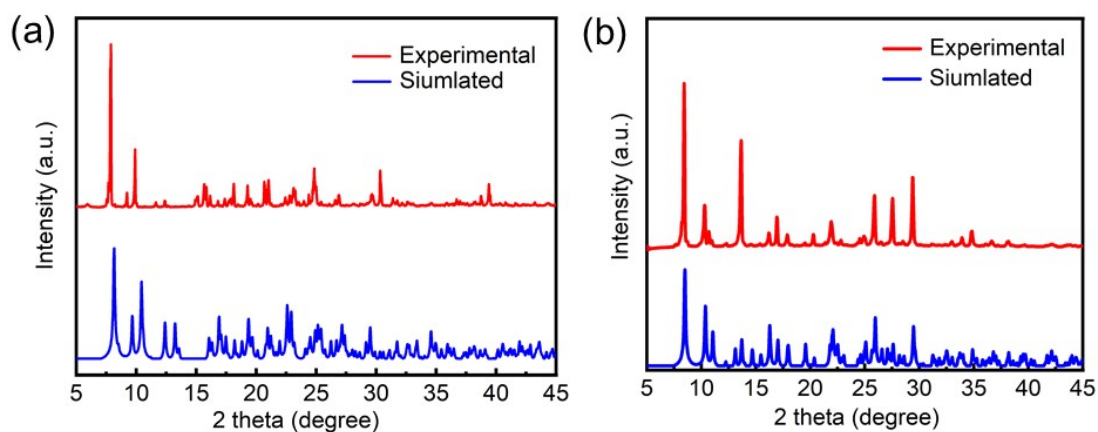


Fig. S1 Experimental (red) and simulated (blue) PXR D patterns of Cu-CP-1 (a) and Cu-CP-2 (b).

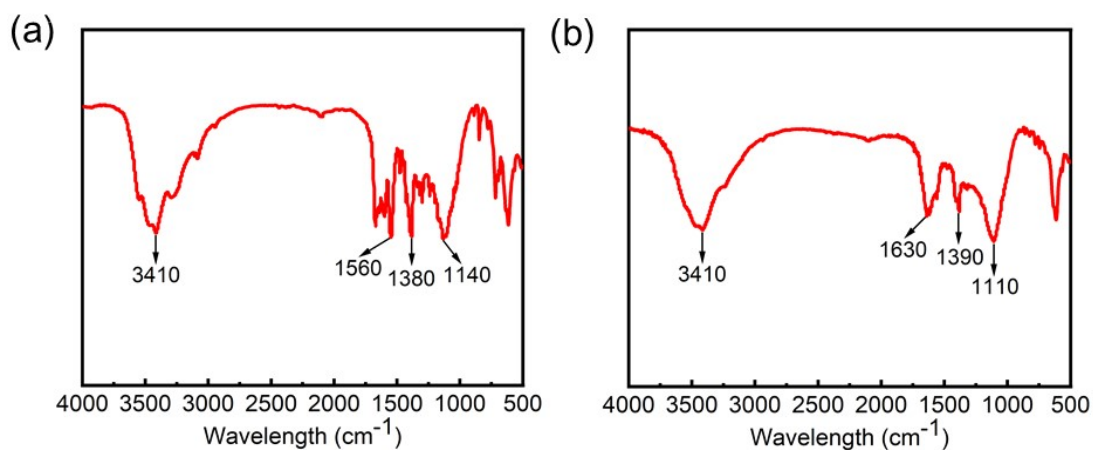


Fig. S2 IR spectra of Cu-CP-1 (a) and Cu-CP-2 (b).

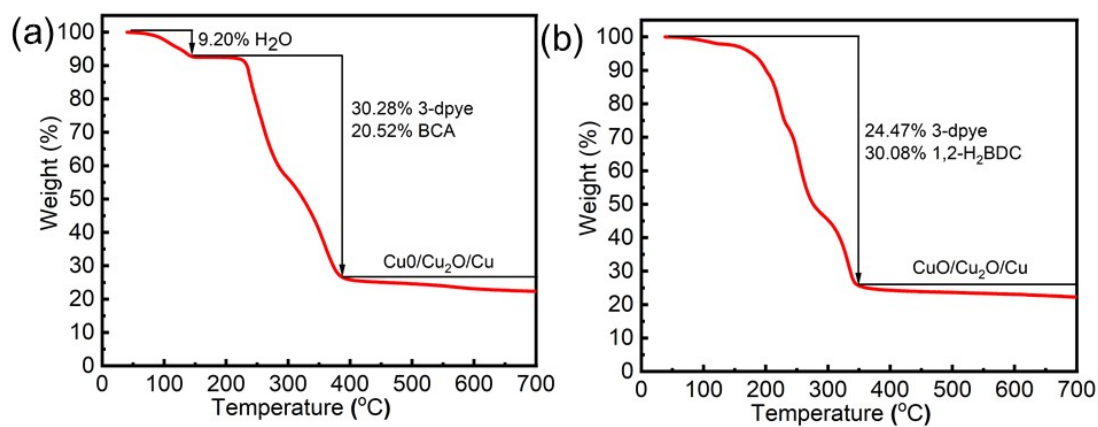


Fig. S3 TGA curves of Cu-CP-1 (a) and Cu-CP-2 (b).

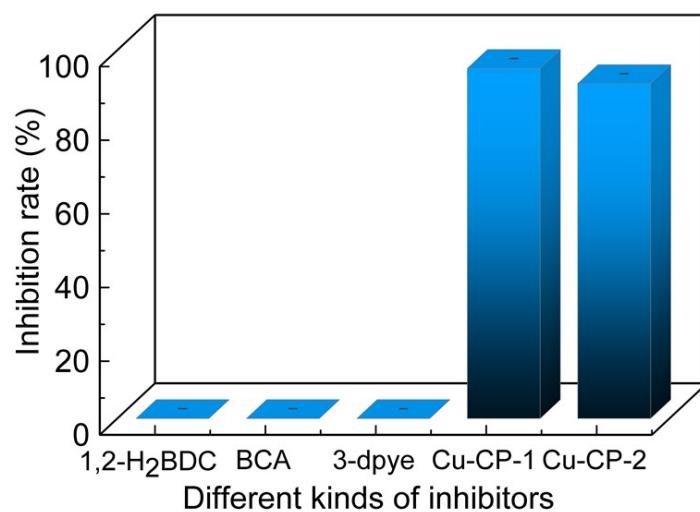


Fig. S4 Different kinds of inhibitors

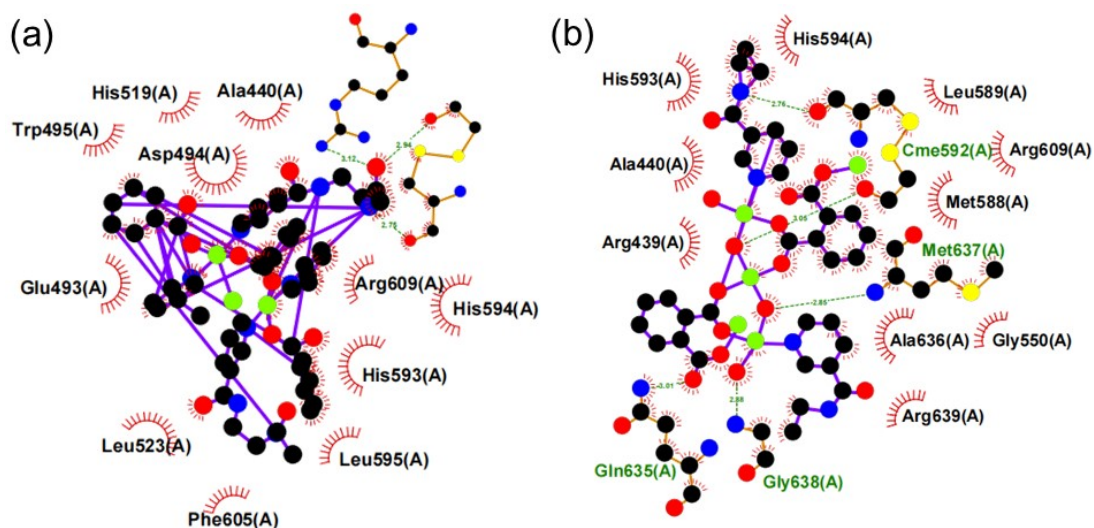


Fig. S5 Binding mode of Cu-CP-1 (a) and Cu-CP-2 (b) with urease in 2D representation. Hydrogen bonds are presented as light dotted lines.

References

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