

Supporting Information

Increasing Porosity in Hydrogen-bonded Organic Frameworks for Low- κ Interlayer Dielectric

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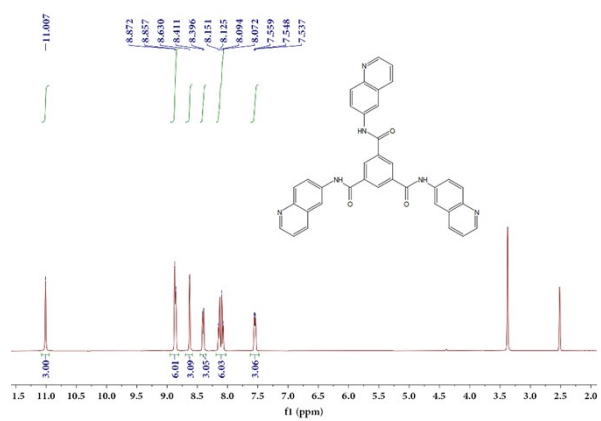


Fig. S1 ¹H NMR spectrum of N¹, N³, N⁵-tris (quinolin-6-yl) benzene-1,3,5-tricarboxamide.

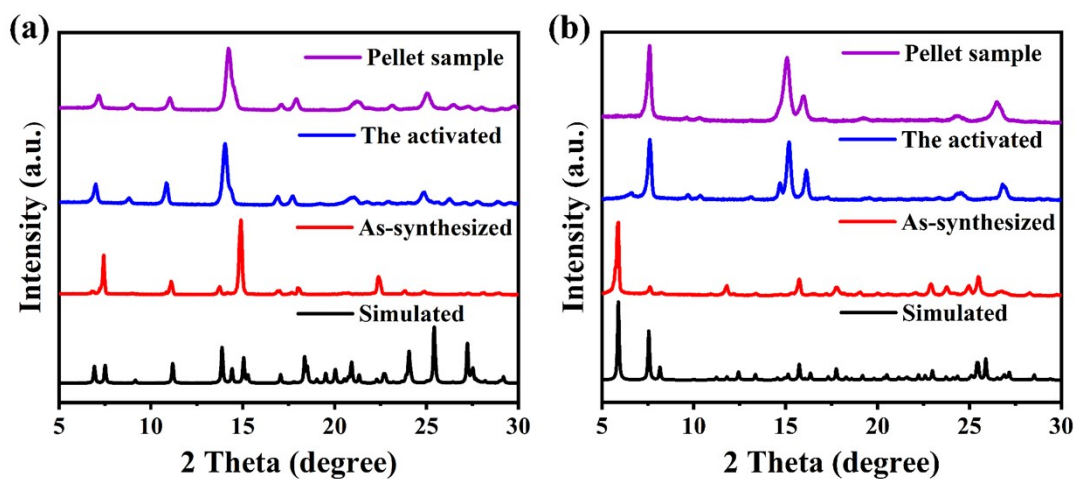


Fig. S2 The PXRD pattern of (a) **HOF-FJU-56** and (b) **HOF-FJU-57**.

The synthesized **HOF-FJU-56** and **HOF-FJU-57** are characterized by powder X-ray diffraction (PXRD) (Fig. S2) which show all the peak positions of the experimental PXRD patterns measured at room temperature are matching closely with simulated results transformed using single crystal data, thus further conforming the high purity of the synthesized samples. The PXRD of samples after activation (denoted as The activated) and pressing into pellets (denoted as Pellet sample) exhibits that both frameworks retained but with slightly shifted in 2θ values, indicating the flexible nature of the two frameworks.

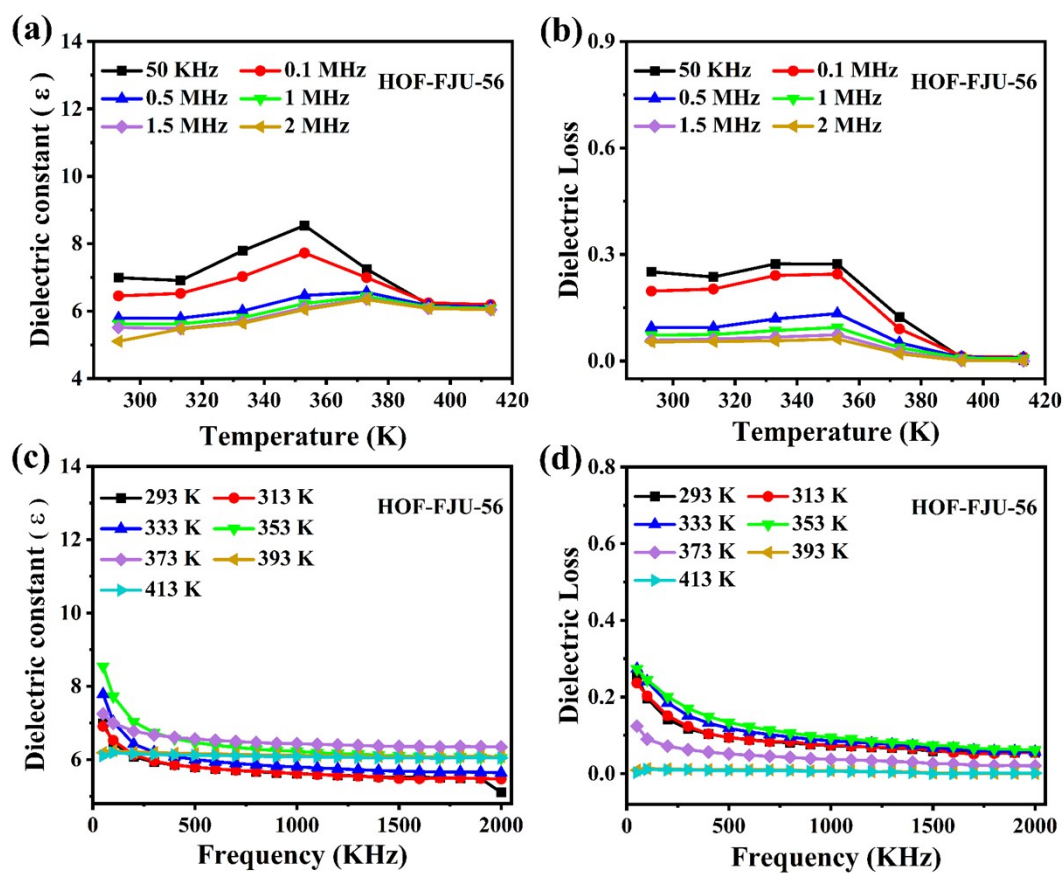


Fig. S3 Temperature-dependent dielectric constant (a) and dielectric loss (b) of **HOF-FJU-56** at different frequencies; Frequency-dependent dielectric constant (c) and dielectric loss (d) of **HOF-FJU-56** at different temperatures.

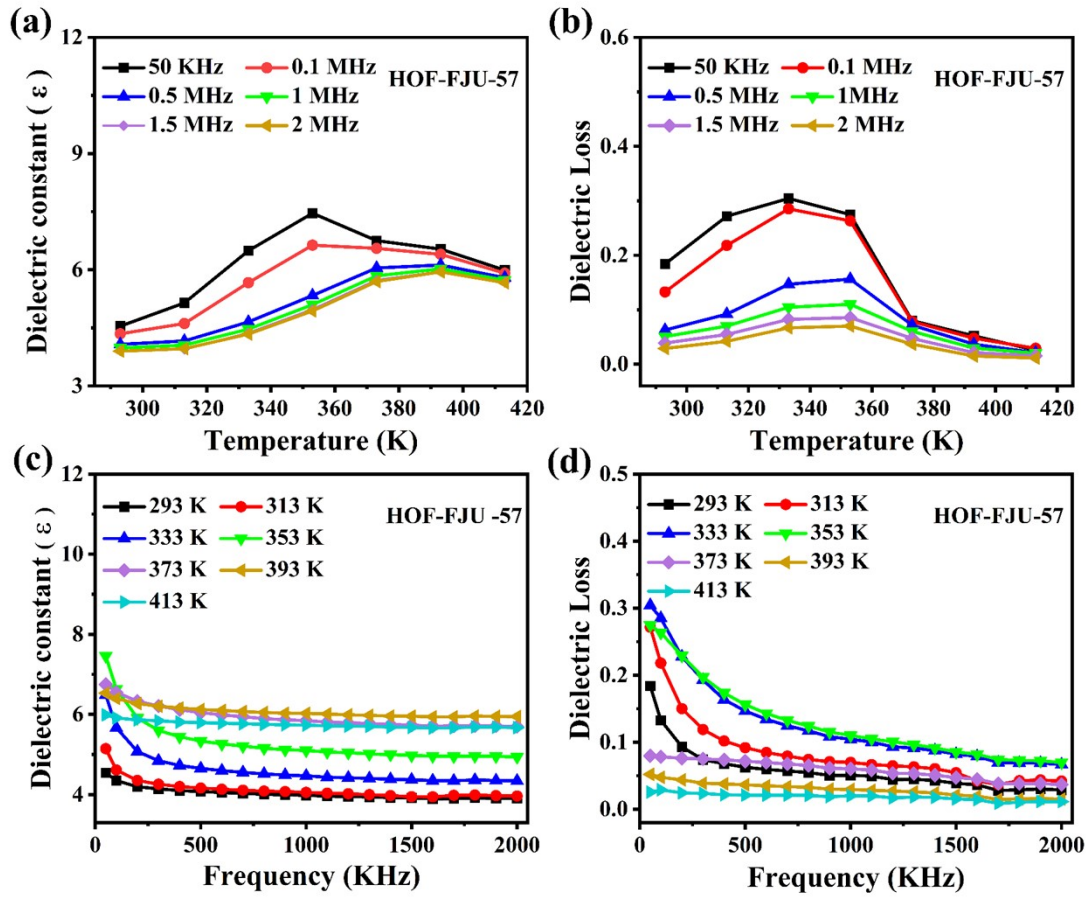


Fig. S4 Temperature-dependent dielectric constant (a) and dielectric loss (b) of HOF-FJU-57 at different frequencies; Frequency-dependent dielectric constant (c) and dielectric loss (d) of HOF-FJU-57 at different temperatures.

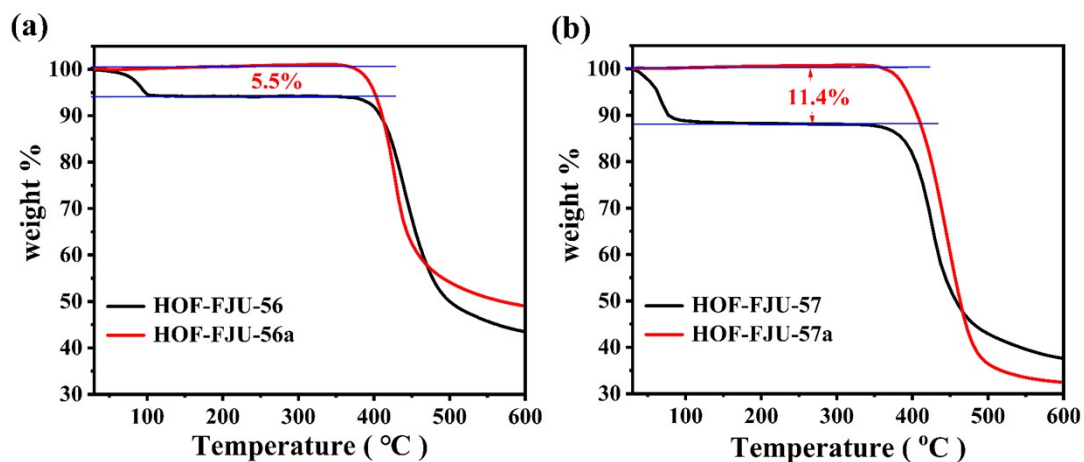


Fig. S5 Thermogravimetric plots for (a) **HOF-FJU-56, HOF-FJU-56a** and (b) **HOF-FJU-57, HOF-FJU-57a**.

To determine the thermal properties of these two crystals, thermogravimetric analysis (TGA) is performed from 30°C to 600°C under N₂ atmosphere. As shown in Fig. S5, the first weight loss rates of **HOF-FJU-56** and **HOF-FJU-57** are approximately 5.5% and 11.4%, respectively, which are consistent with the theoretically calculated solvent loss rates of 5.16% and 11.6% from their crystal data. The second weight loss of the two materials occurred at around 370 °C, which may be due to the decomposition of TQBTC.

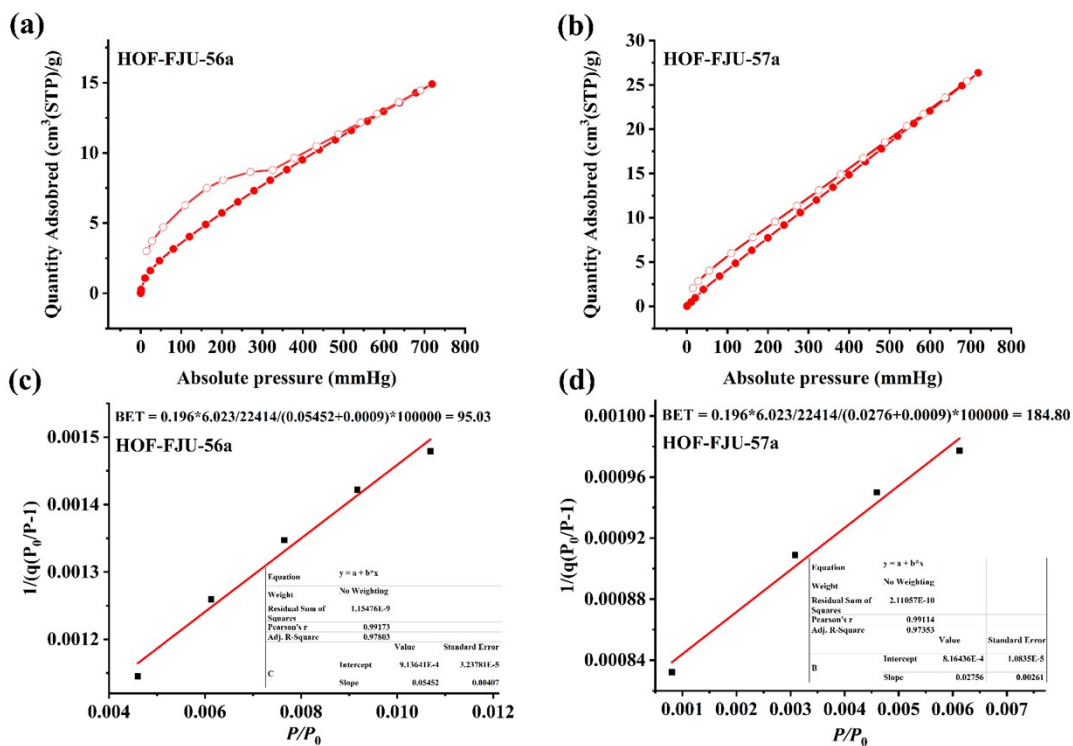


Fig. S6 CO₂ sorption isotherm at 196 K of **HOF-FJU-56a** (a) and **HOF-FJU-57a** (b).

The BET plots for **HOF-FJU-56a** (c) and **HOF-FJU-57a** (d).

$$S_{\text{BET}} = 0.196 * 6.023 / 22414 / (\text{slope} + \text{intercept}) * 100000$$

$$S_{\text{BET(HOF-FJU-56a)}} = 0.196 * 6.023 / 22414 / (0.05452 + 0.0009) * 100000 = 95.03 \text{ m}^2/\text{g}$$

$$S_{\text{BET(HOF-FJU-57a)}} = 0.196 * 6.023 / 22414 / (0.02760 + 0.0009) * 100000 = 184.80 \text{ m}^2/\text{g}$$

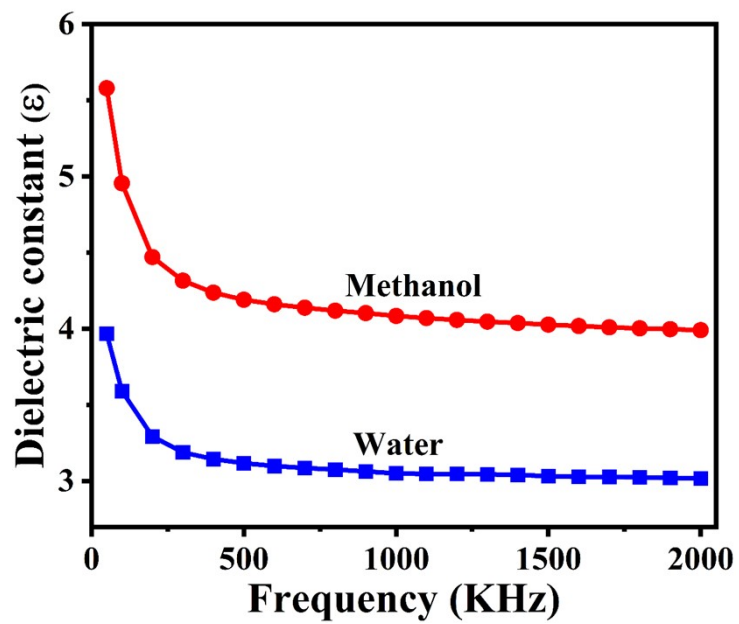


Fig. S7 Frequency-dependent dielectric constants of **HOF-FJU-57a** after immersion in methanol and water, respectively, at 298k.

Table S1 Crystallographic Data for **HOF-FJU-56** and **HOF-FJU-57**.

	HOF-FJU-56	HOF-FJU-57
CCDC	2241066	2245969
Empirical formula	C ₃₇ H ₂₇ N ₆ O ₄	C ₃₇ H ₃₃ N ₆ O _{6.5}
Formula weight	619.64	665.69
Temperature/K	150.00(10)	149.99(10)
Crystal system	triclinic	triclinic
Space group	<i>P1</i>	<i>P-1</i>
<i>a</i> /Å	4.8338(5)	8.9567(6)
<i>b</i> /Å	12.1052(8)	12.2826(7)
<i>c</i> /Å	13.0176(11)	15.4474(6)
<i>α</i> /°	101.200(7)	72.697(4)
<i>β</i> /°	90.866(7)	86.055(4)
<i>γ</i> /°	97.392(7)	88.054(5)
Volume/Å ³	740.37(11)	1618.48(16)
<i>Z</i>	1	2
$\rho_{\text{calc}}/\text{cm}^3$	1.39	1.366
μ/mm^{-1}	0.756	0.786
F(000)	323.0	698.0
Radiation	Cu K α ($\lambda = 1.54184$)	Cu K α ($\lambda = 1.54184$)
Data/restraints/parameters	4306/3/426	6143/6/459
Goodness-of-fit on F ²	1.085	1.005
Final R indexes [I $\geq 2\sigma$ (I)]	R ₁ = 0.0695, wR ₂ = 0.1776	R ₁ = 0.0887, wR ₂ = 0.2266
Final R indexes [all data] ^a	R ₁ = 0.0823, wR ₂ = 0.1919	R ₁ = 0.1346, wR ₂ = 0.2532

$$^a R_1 = \sum(|F_o| - |F_c|) / \sum|F_o|, \quad ^b wR_2 = [\sum w(F_o^2 - F_c^2)^2 / \sum w(F_o^2)^2]^{0.5}$$

Table S2. Compare the dielectric constant and dielectric loss of different materials at 1MHz, 298K.

Compound	Leakage current at 0.5 KV/cm	Dielectric constant	dielectric loss	The range of test temperature	Ref.
HOF-FJU-56a	1.06×10^{-10}	4.06	0.009	13-413 K	This work
HOF-FJU-57a	1.45×10^{-13}	2.12	0.067	13-413 K	This work
[Zn(CEIC) ₂ (H ₂ O)]·2DMF	/	3.90	0.067	303-363K	1
TAPB-TPOC ₆ -COF film	1.4×10^{-10}	1.25	0.010	293-433K	2
FMOF-1	/	1.28	/	298K	3
TmBPHF	/	2.09	0.001	298-573K	4
[Zn ₂ (Hbbim) ₂ (bbim)]n	/	3.05	0.019	3.5-350K	5
TmBPPA	/	2.23	0.009	293-573K	6
SiCOH	1.0×10^{-10}	3.48	0.009	298-523K	7
GD50	2×10^{-10}	3.16	/	298K	8
HGD50	7×10^{-11}	3.03	/	298K	8
Silica film	1.0×10^{-8}	2.0	0.020	298K	9
PSZ MFI low-κ film	1.35×10^{-8}	1.83	/	298K	10
MSQ film	6.0×10^{-12}	2.7	/	298K	11
BC:H	6.0×10^{-13}	3.3	/	298K	12
ZIF-8 film	3.0×10^{-12}	2.3	0.008	298-473K	13
ZIF-67 film	/	2.39	0.060	298K-423K	14

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