

Supporting Information For:

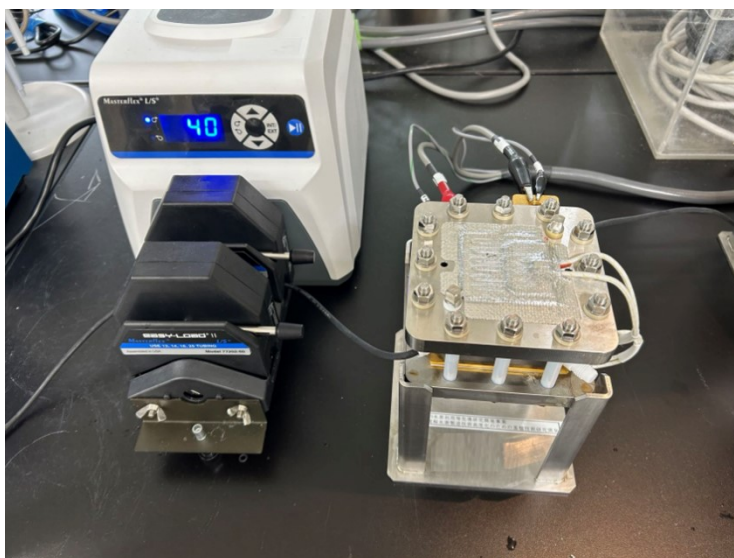
**Effects of Dissolved 3d-block Metal Ions on PEM Water Electrolysis Performance**

Shuang Kong<sup>\*a</sup>, Kazuna Fushimi<sup>a</sup>, Ailong Li<sup>a</sup>, Ryuhei Nakamura<sup>\*a,b</sup>

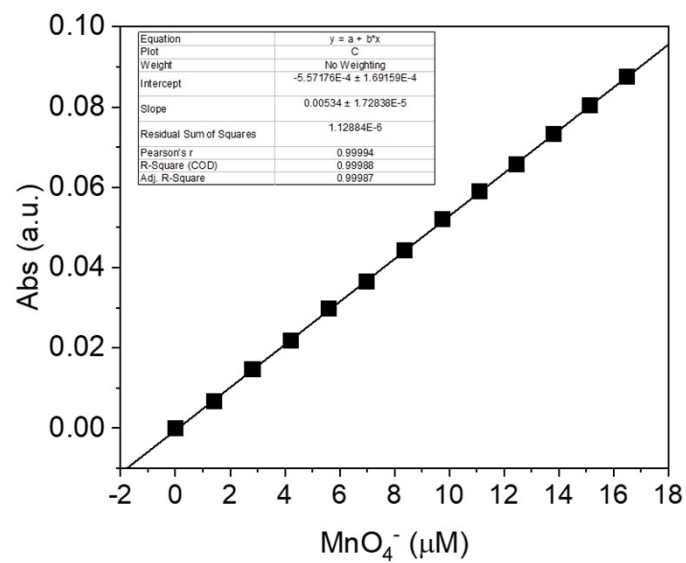
<sup>a</sup>Biofunctional Catalyst Research Team, RIKEN Center for Sustainable Resource Science (CSRS), 2-1 Hirosawa, Wako, Saitama, 351-0198, Japan

<sup>b</sup>Earth-Life Science Institute (ELSI), Tokyo Institute of Technology, 2-12-1-I7E Ookayama, Meguro-ku, Tokyo, 152-8550, Japan

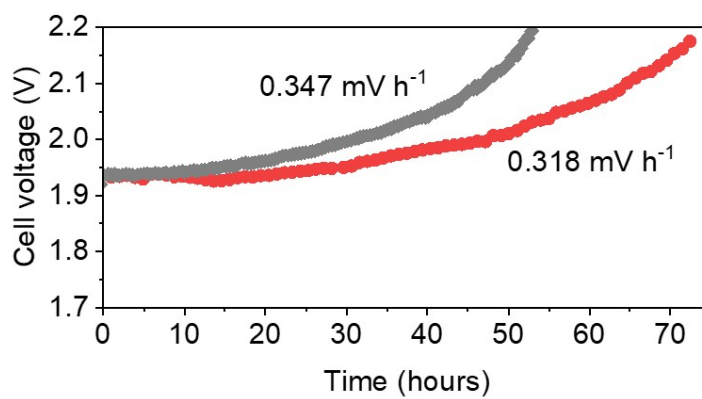
\*Corresponding authors: [shuang.kong@riken.jp](mailto:shuang.kong@riken.jp) and [ryuhei.nakamura@riken.jp](mailto:ryuhei.nakamura@riken.jp)



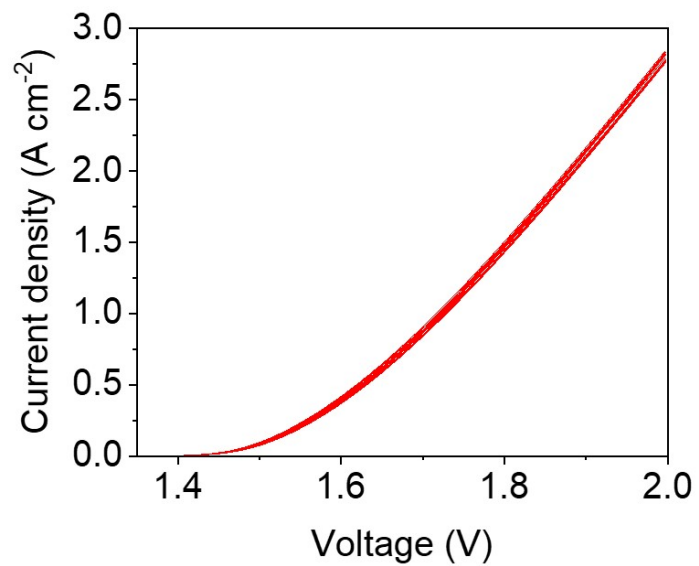
**Fig. S1.** Set-up of the PEM electrolyzer used in this study.



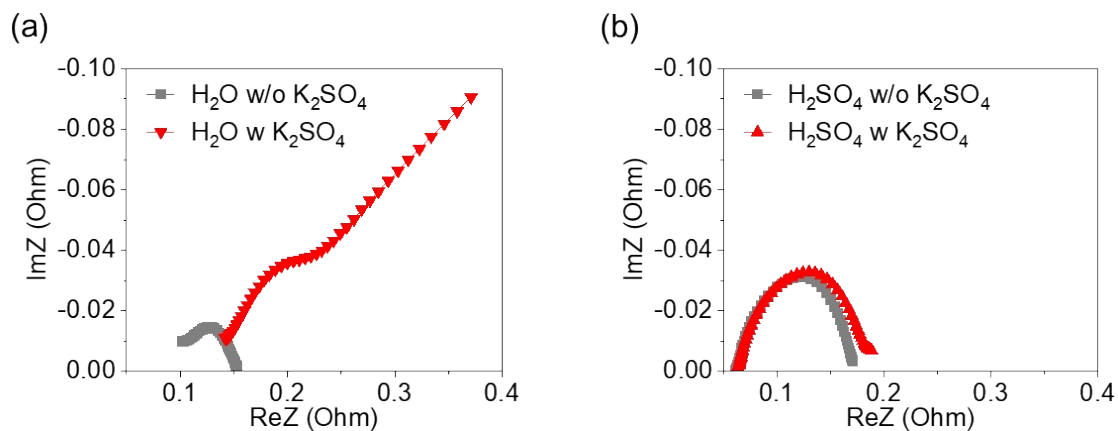
**Fig. S2.** Calibration curve of absorbance versus  $\text{MnO}_4^-$  concentration in 1 M  $\text{H}_2\text{SO}_4$ . The inset table shows the parameters used for the linear fit.



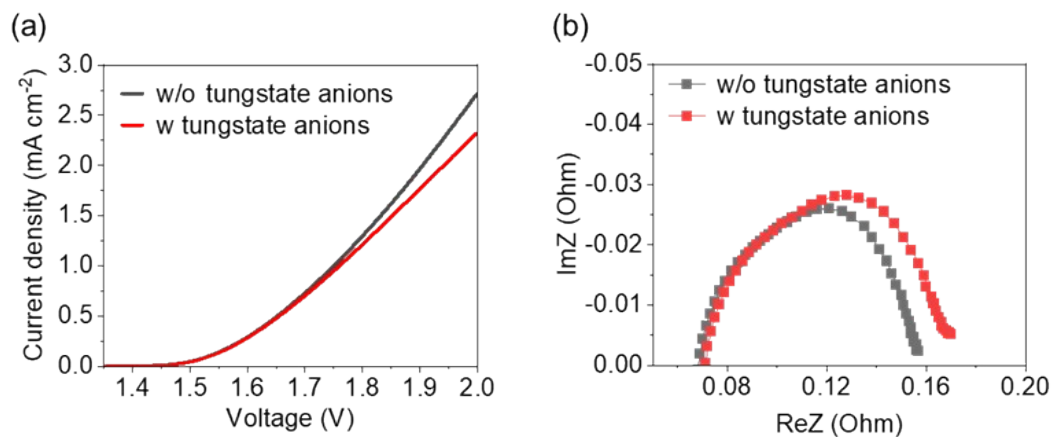
**Fig. S3.** Long-term stability test of PEM water electrolysis at  $1 \text{ A cm}^{-2}$  in the presence of  $\text{MnO}_4^-$  (grey line) and in the absence of  $\text{MnO}_4^-$  (red line).  $\gamma\text{-MnO}_2$  was used as the anode catalyst. The temperature is  $80 \text{ }^\circ\text{C}$ . The deviation in the degradation rate is attributable to the decomposition of  $\text{MnO}_4^-$  into  $\text{MnO}_2$  accompanied by oxygen release.



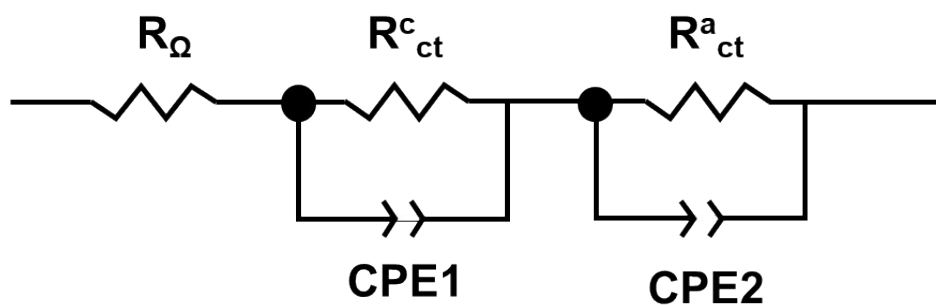
**Fig. S4.** 20 Linear sweep voltammetry scans performed during PEM water electrolysis using a 0.08 mg/cm<sup>2</sup> Ir catalyst on the anode side.



**Fig. S5.** (a) Electrochemical impedance spectra at 1.8 V measured in the presence and absence of K<sub>2</sub>SO<sub>4</sub> in PEM water electrolysis. (b) Electrochemical impedance spectra at 1.8 V measured in the presence and absence of K<sub>2</sub>SO<sub>4</sub> in 0.05 M H<sub>2</sub>SO<sub>4</sub> in PEM water electrolysis.

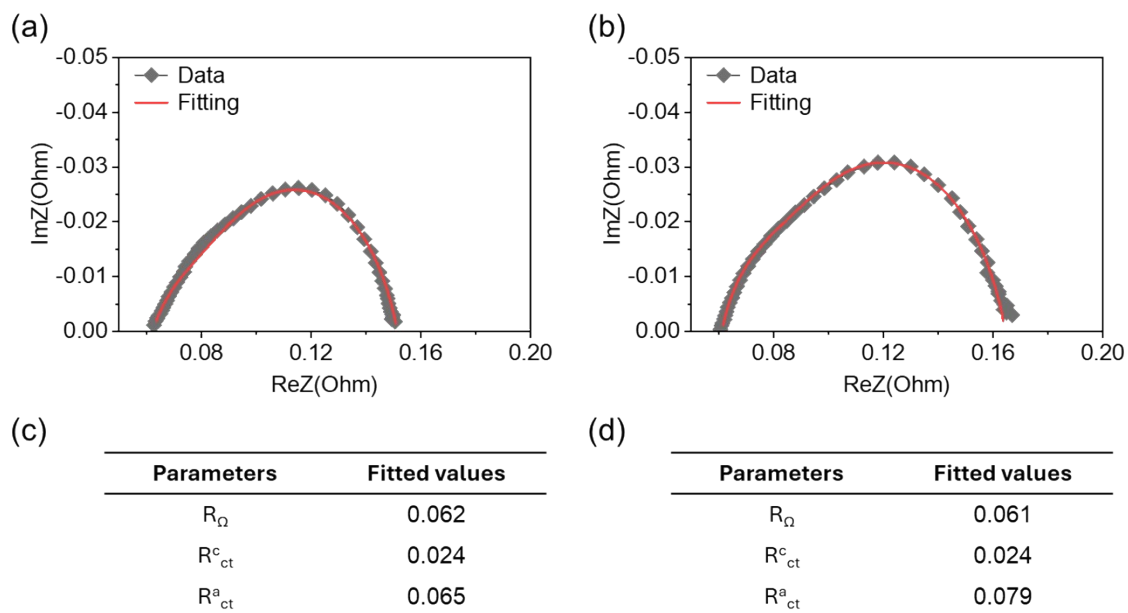


**Fig. S6.** (a) Linear sweep voltammetry scans (LSV) were performed during PEM water electrolysis with and without a direct feed of Na<sub>2</sub>WO<sub>4</sub>. (b) Electrochemical impedance spectra at 1.8 V were measured in the presence and absence of Na<sub>2</sub>WO<sub>4</sub> over a frequency range of 100 kHz to 1 Hz.

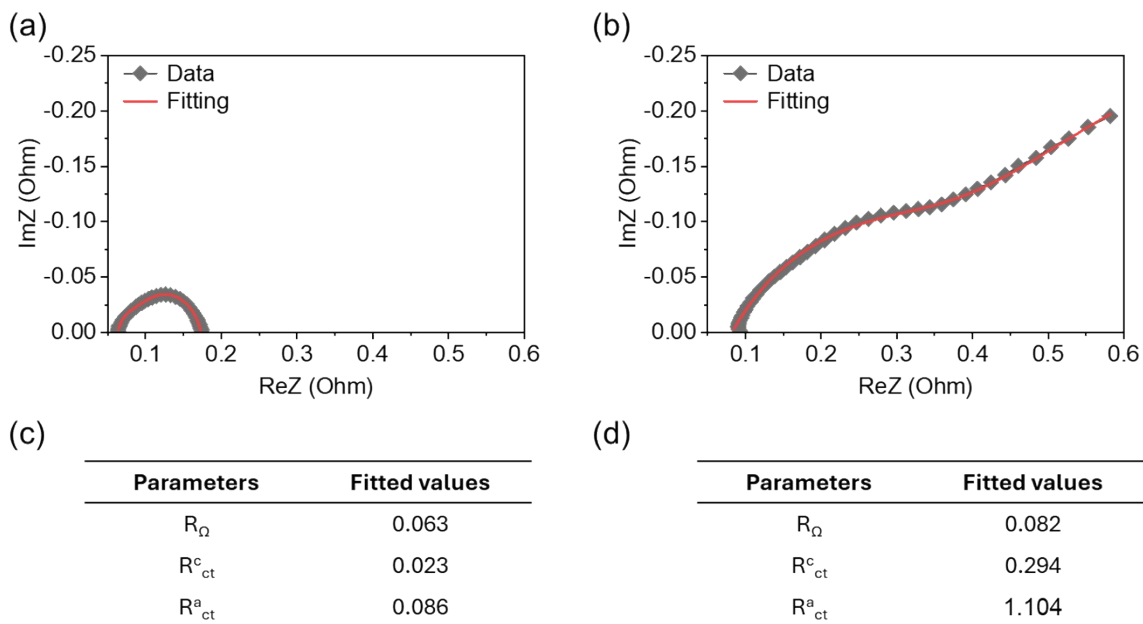


**Fig. S7.** Diagram of the electrical circuits used to model impedance spectra. An Ohmic resistor ( $R_{\Omega}$ ) was connected in series with two parallel components, which consisted of a resistor ( $R^a_{ct}$  or  $R^c_{ct}$ ) and a constant phase element (CPE).

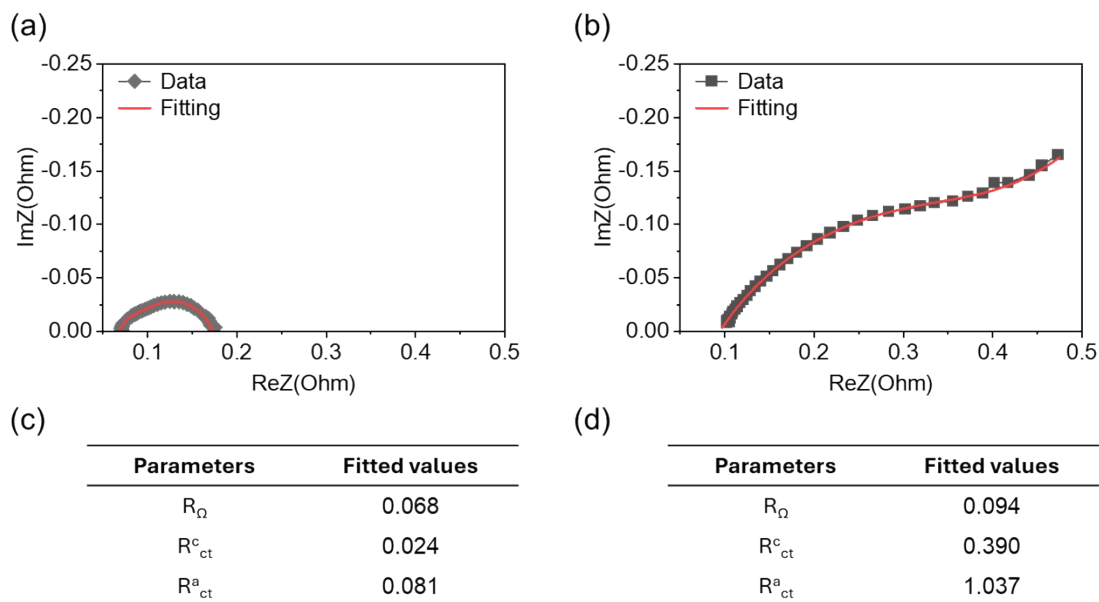




**Fig. S8.** Impedance spectra measured at 1.8 V during PEM electrolysis conducted with a feed solution of 0.05 M H<sub>2</sub>SO<sub>4</sub> (a) or with 1 mM KMnO<sub>4</sub> in 0.05 M H<sub>2</sub>SO<sub>4</sub> (b). Dotted lines represent the raw data and solid curves show the fitted results. The fitting parameters are displayed in (c) and (d), respectively.



**Fig. S9.** Impedance spectra measured at 1.8 V during PEM electrolysis conducted with a feed solution of 0.05 M H<sub>2</sub>SO<sub>4</sub> (a) or with 1 mM Ni<sup>2+</sup> in 0.05 M H<sub>2</sub>SO<sub>4</sub> (b). Dotted lines represent the raw data and solid curves show the fitted results. The fitting parameters are displayed in (c) and (d), respectively.



**Fig. S10.** Impedance spectra measured at 1.8 V during PEM electrolysis conducted with a feed solution 0.05 M  $H_2SO_4$  (a) or with 1 mM  $Co^{2+}$  in 0.05 M  $H_2SO_4$  (b). Dotted lines represent the raw data and solid curves show the fitted results. The fitting parameters are displayed in (c) and (d), respectively.