Electronic Supplementary Information for

Detection and disaggregation of amyloid fibrils by luminescent amphiphilic platinum(II) complexes

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Materials and general methods

All reactions were carried out under Ar unless specifically noted. All commercially available materials, reagents and solvents were used as supplied unless otherwise noted. K_2PtCl_4 was purchased from Strem Chemicals, Inc., hen egg-white lysozyme (HEWL, lyophilized powder) was purchased from Sigma-Aldrich, and beta amyloid 1-42 (A β_{42} , HFIP-treated) was purchased from rPeptide. SH-SY5Y neuroblastoma cells were obtained from ATCC. Reported chemical yields are isolated yields. Proton (1H) and carbon (13C) NMR were collected on Bruker NMR spectrometers at 600, 500, or 400 MHz for 1H and 150, 125, or 100 MHz for 13C. Chemical shifts (δ) are reported in parts per million (ppm) referenced to undeuerated residual solvent. High resolution mass spectra were collected on a Thermo Q-Exactive (Orbitrap) mass spectrometer. Absorption spectra were obtained using an Agilent Cary 100 UV-Vis spectrophotometer. Emission spectra were obtained using either an Agilent Cary Eclipse fluorescence spectrometer or a FluoroMax Plus fluorimeter (HORIBA Scientific, Ltd.). All luminescence spectrometer was performed on samples open to air. Dynamic light scattering (DLS) experiments were performed using a Malvern Zetasizer Nano ZS instrument.

Synthesis of Ligands (L2, L3, L4, L6-L8)

Synthesis of L2



3,4-Dimethoxystyrene was prepared according to literature reports.¹ 2-(4-Bromophenyl)pyridine (0.56 g, 2.37 mmol) was dissolved in 3.3 mL Et₃N and the solution was stirred and purged with Ar for 15 mins. 3,4-Dimethoxystyrene (0.43 g, 2.62 mmol), Pd(OAc)₂ (0.011 g, 0.047 mmol) and tri(o-tolyl)phosphine (0.043 g, 0.14 mmol) were added. The reaction mixture was stirred at 95 °C for 24 hours. The resulting mixture was allowed to cool to room temperature. Brine was added and the reaction mixture was extracted with DCM. The combined extracts were dried over Na₂SO₄, filtered, and concentrated under vacuum. The crude product was further purified by silica gel column chromatography, eluting with 30% EtOAc/70% hexane, to give **L2** in 88% yield (0.67 g, 2.09 mmol). 1H NMR (500 MHz, CDCl₃) δ 8.67 (dt, J = 4.7, 1.3 Hz, 1H), 7.99-7.97 (m, 2H), 7.70-7.65 (m, 2H), 7.58-7.56 (m, 2H), 7.18-7.15 (m, 1H), 7.10 (d, J = 16.3 Hz, 1H), 7.07-7.02 (m, 2H), 6.99 (d, J = 16.3 Hz, 1H), 6.82 (d, J = 8.3 Hz, 1H), 3.91 (s, 3H), 3.86 (s, 3H); 13C NMR (125 MHz, CDCl₃) δ 157.0, 149.8, 149.3, 149.2, 138.3, 138.2, 136.8, 130.5, 129.2, 127.2, 126.8, 126.4, 122.1, 120.4, 120.2, 111.4, 109.0, 56.0, 56.0; HRMS (ESI) calcd for C₂₁H₂₀O₂N [M+H]⁺ 318.1489, found 318.1483.

Synthesis of L3



3,4-Dimethoxybenzoic acid (2.01 g, 10.97 mmol) was dissolved in 40 mL DCM. 4-(2-Pyridinyl)benzenamine (2.02 g, 12.07 mmol), EDC (hydrochloride, 3.15 g, 16.46 mmol), HOBt (monohydrate, 2.52 g, 16.46 mmol), and Et₃N (1.66 g, 16.46 mmol) were added. The reaction mixture was stirred at room temperature for 24 h. The solvent was removed under vacuum. Ethanol (200 mL) was added to the crude sample and the resulting mixture was stirred at 78 °C for 10 mins. Insoluble **L3** was collected by vacuum filtration and dried under vacuum. NMR indicated this material was of sufficient purity for use in subsequent reactions. Yield – 1.45 g (4.33 mmol), 40%. 1H NMR (600 MHz, CDCl₃) δ 8.61-8.60 (m, 1H), 7.96-7.94 (m, 2H) 7.90 (br s, 1H), 7.70-7.65 (m, 4H), 7.45 (d, J = 2.0 Hz, 1H), 7.35 (dd, J= 8.4, 2.1 Hz, 1H), 7.15-7.13 (m, 1H), 6.84-6.83 (m, 1H), 3.88 (s, 3H), 3.87 (s, 3H); 13C NMR (150 MHz, CDCl₃) δ 165.2, 156.7, 152.1, 149.6, 149.2, 138.9, 136.8, 135.2, 127.6, 127.4, 121.9, 120.1, 120.0, 119.4, 110.7, 110.3, 56.0; HRMS (ESI) calcd for C₂₀H₁₉O₃N₂ [M+H]⁺ 335.1390, found 335.1385.

Synthesis of L4



L4 was prepared using a method reported in the literature.² L3 (0.21 g, 0.59 mmol) was dissolved in 6 mL toluene and tetramethylammonium fluoride (TMAF, 0.14 g, 1.43 mmol) was added. The reaction mixture was allowed to stir at 100 °C for 16 hours. The mixture was allowed to cool to room temperature and the solvent was removed under vacuum. The crude product was purified by silica gel column chromatography, eluting with hexane with increasing amount of EtOAc from 50% hexane/50% EtOAc to 30% hexane/70%EtOAc, to give L4 in 72% yield (0.15 g, 0.43mmol). 1H NMR (400 MHz, CDCl₃) δ 8.65 (dt, J = 4.7, 0.8 Hz, 1H), 7.91-7.88 (m, 2H), 7.72 (td, J = 7.6 Hz, 1.7Hz, 1H), 7.65 (d, J = 7.9Hz, 1H), 7.23-7.20 (m, 1H), 7.16-7.14 (m, 2H), 6.96-6.91 (m, 2H), 6.61 (d, J = 8.3 Hz, 1H), 3.78 (s, 3H), 3.67 (s, 3H), 3.53 (s, 3H); 13C NMR (100 MHz, CDCl₃) δ 171.0, 170.0, 156.1, 150.2, 149.6, 148.0, 146.1, 137.1, 136.8, 127.7, 127.6, 126.8, 122.7, 122.3, 120.2, 112.3, 109.9, 55.7, 55.6, 38.4, 14.1; HRMS (ESI) calcd for C₂₁H₂₁O₃N₂ [M+H]⁺ 349.1547, found 349.1539.



L6-L8 were prepared using modifications of a previously reported procedure.³ In general, K_2CO_3 (0.1 equivalent) was added to acetylacetone (acac, neat, 5 equivalents) and stirred at room temperature. The appropriate acrylate (1 equivalent) was then added and the reaction mixture was heated to 70 °C and maintained with stirring for 24 h. Insoluble material from the reaction mixture was removed by vacuum filtration and washed with DCM and MeOH. The filtrate was then concentrated under vacuum to give crude product. The acrylates were prepared using previously reported methods.^{4, 5}



L6 was prepared according to the general procedure from acac and 2-methoxyethyl acrylate (0.78 g, 5.99 mmol). Crude **L6** obtained from the reaction (93%, 1.28 g, 5.55 mmol) was used directly in subsequent reactions. 1H NMR (600 MHz, CDCl₃), mixture of tautomers: δ 4.27-4.25 (m, 2H), 3.8 (t, J = 7.0 Hz, 0.6H), 3.63-3.61 (m, 2H), 3.41-3.40 (m, 3H), 2.65-2.62 (m, 0.6H), 2.50-2.47 (m, 0.6H), 2.39-2.38 (t, J = 7.2 Hz, 1.4H), 2.23 (s, 4H), 2.19-2.14 (m, 3H); 13C NMR (150 MHz, CDCl₃), mixture of tautomers δ 203.6, 191.2, 172.3, 172.3, 108.5, 70.2, 70.1, 66.6, 63.4, 63.4, 58.7, 58.7, 34.4, 31.3, 29.2, 22.7, 22.7, 22.6; HRMS (ESI) calcd for C₁₁H₁₉O₅ [M+H]⁺ 231.1227, found 231.1225.



L7 was prepared according to the general procedure from acac and N-(2-(2-hydroxyethyloxy)ethyl)acrylamide (1.75 g, 10.99 mmol). The crude product obtained directly from the reaction mixture was further purified by silica gel column chromatography, eluting with DCM with increasing amount of MeOH from 99% DCM/1% MeOH to 94% DCM/6% MeOH, to give L7 in 17% yield (0.50 g, 1.94 mmol). 1H NMR (400 MHz, CDCl₃), mixture of tautomers: δ 3.79-3.75 (m, 2H), 3.61-3.54 (m, 4H), 3.48-3.43 (m,4H), 2.65-2.61 (m, 0.5H), 2.29-2.25 (m, 0.5H), 2.21 (s, 4H), 2.17-2.15 (m, 4H);13C NMR (100 MHz, CDCl₃), mixture of tautomers δ 204.6, 191.6, 172.2, 172.0, 170.8, 146.2, 117.5, 109.1, 72.5, 72.2, 69.7, 69.6, 69.3, 66.8, 61.5, 41.3, 39.3, 39.2, 36.9, 33.3, 31.2, 29.8, 29.4, 23.5, 23.4, 22.9, 22.1, 16.3; HRMS (ESI) calcd for C₁₂H₂₂O₅N [M+H]⁺ 260.1492, found 260.1491.



L8 was prepared according to the general procedure from acac and N,N-bis[2-(2-hydroxy)ethyl]-2-propenamide (0.53 g, 2.16 mmol). The crude product was further purified by silica gel column chromatography, eluting with DCM with increasing amount of MeOH from 95% DCM/5% MeOH to 92% DCM/ 8% MeOH, to give **L8** in 54% yield (0.40 g, 1.16 mmol). 1H NMR (400 MHz, CDCl₃), mixture of tautomers: δ 3.81 (t, J = 6.8 Hz, 1H), 3.72-3.53 (m, 21H), 2.62-2.58 (m, 0.5H), 2.50-2.46 (m, 0.5H), 2.43 (t, J = 7.1 Hz, 2H), 2.22 (s, 5H), 2.18-2.11 (m, 3.5H); 13C NMR (100 MHz, CDCl₃), mixture of tautomers: δ 205.0, 191.3, 172.7, 172.6, 128.1, 127.8, 109.1, 72.9, 72.8, 72.4, 72.3, 69.2, 69.1, 69.1, 69.0, 66.8, 61.6, 61.5, 61.5, 48.8, 48.7, 46.6, 46.4, 33.6, 30.2, 29.4, 23.0, 23.0, 22.9; HRMS (ESI) calcd for C₁₆H₃₀O₇N [M+H]⁺ 348.2017, found 348.2020.



Complexes 7-14 were prepared via modifications to the two-step method previously reported for the preparation of 6.⁶ In general, K_2PtCl_4 (1 equivalent) was suspended in a mixture of $H_2O/2$ -ethoxyethanol (~1:3 – 1:4, v/v) and the 2-phenylpyridine ligand (ppy, 1.0 – 3.5 equivalents, L1-L4) was added. The reaction mixture was heated with stirring to 100 °C for 48 h. The reaction was allowed to cool to room temperature and H_2O added until precipitation of the intermediate chloride-bridged Pt dimer was complete. The precipitate was collected by vacuum filtration, washed with H_2O , and dried under vacuum. These materials were used directly in the next step without further purification.

For the second step, the chloride-bridged dimer obtained in the first step (1 equivalent) was suspended in 2-ethoxyethanol or 1,4-dioxane. The acetonylacetone (acac) ligand L5-L8 (3 equivalents) and base (3 – 10 equivalents of Na₂CO₃ or NaH) were added. The reaction mixture was heated to 100 °C for 24 h. The reaction mixture was then allowed to cool to room temperature and MeOH was added. The solvent was then removed under vacuum to give the crude product. Overall yields were calculated based on amount of K₂PtCl₄ used.



Complex 6

Complex **6** was prepared according to the general procedure using **L1** (0.11 g, 0.96 mmol) and K₂PtCl₄ (0.30 g, 0.96 mmol) in 10 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The chloride-bridged dimer obtained from this reaction (0.22 g, 0.28 mmol) was suspended in 6 mL 1,4-dioxane and combined with **L5** (0.087 g, 0.87 mmol) and Na₂CO₃ (0.30 g, 2.90 mmol). The crude product obtained from the reaction was further purified by silica gel column chromatography, eluting with 20% EtOAc/80% hexane, to give **6** in 37% overall yield (0.12 g, 0.27 mmol). 1H NMR (400 MHz, CDCl₃) δ 9.02-8.93 (m, 1H), 7.77 (td, J = 7.8, 1.5 Hz, 1H), 7.62-7.58 (m, 2H), 7.42 (dd, J = 7.8, 1.0 Hz, 1H), 7.19 (td, J = 7.49, 1.3 Hz, 1H), 7.10-7.06 (m, 2H), 5.46 (s, 1H), 1.99 (s, 3H), 1.99(s, 3H); 13C NMR (100 MHz, CDCl₃) δ 185.7, 184.1, 168.3,

147.3, 144.6, 138.9, 138.1, 130.5, 129.2, 123.5, 123.0, 121.1, 118.3, 102.5, 28.3, 27.1; HRMS (ESI) calcd for C₁₆H₁₆O₂NPt [M+H]⁺ 449.0823, found 449.0820.



Complex 7

Complex 7 was prepared according to the general procedure except that all operations were performed in the dark (aluminum foil-wrapped glassware) to prevent light-induced isomerization of the stilbene alkene. Ligand L2 (3.5 equivalents, 0.57 g, 1.81 mmol) and K₂PtCl₄ (0.21 g, 0.51 mmol) were combined in 12 mL H₂O/2-ethoxyethanol (1:3, v/v) for the first step. Platinum dimer obtained from this reaction (0.40 g, 0.36 mmol) was suspended in 3 mL 2-ethoxyethanol. Acetylacetone (L5, 0.11 g, 1.12 mmol) and Na₂CO₃ (0.39 g, 3.74 mmol) were added and the reaction performed as described in the general procedure. The crude product was further purified by silica gel column chromatography, eluting with 20% hexane/80% DCM, to give 7 in 51% overall yield (0.12 g, 0.27 mmol). X-ray quality single crystals of 7 were grown from CH₂Cl₂/MeOH solution. 1H NMR (600 MHz, CDCl₃) δ 9.03-8.93 (m, 1H), 7.78 (td, J = 7.7, 1.5 Hz, 1H), 7.70 (d, J = 1.6 Hz, 1H), 7.61-7.57 (m, 1H), 7.42 (d, J = 8.0 Hz, 1H), 7.28 (dd, J = 8.0, 1.6 Hz, 1H), 7.15 (d, J = 16.1 Hz, 1H), 7.12-7.06 (m, 3H), 7.04 (d, J = 16.1 Hz, 1H), 6.86 (d, J = 8.2 Hz, 1H), 5.49 (s, 1H), 3.96 (s, 3H), 3.91 (s, 3H), 2.06 (s, 3H), 2.01 (s, 3H);). 13C NMR (150 MHz, CDCl₃) δ 185.8, 184.1, 167.8, 149.0, 148.9, 147.3, 143.8, 138.7, 138.2, 138.0, 130.8, 129.0, 128.4, 127.6, 123.2, 121.6, 120.8, 120.1, 118.3, 111.1, 108.8, 102.5, 55.9, 55.9, 28.2, 27.2; HRMS (ESI) calcd for C₂₆H₂₆O₄NPt [M+H]⁺ 611.1504, found 611.1499.



Complex 8

Complex 8 was prepared according to the general procedure using L3 (0.30 g, 0.91 mmol) and K_2PtCl_4 (0.37 g, 0.91 mmol) in 10 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The resulting chloride-bridged dimer (0.45 g, 0.40 mmol) was suspended into 10 mL 1,4-dioxane, and combined with L5 (0.12 g, 1.20 mmol) and Na₂CO₃ (0.42 g, 4.01 mmol). The crude product was further purified by silica gel column chromatography, eluting with 4% MeOH/96% DCM, to

give **8** in 26% overall yield (0.15g, 0.24 mmol). X-ray quality single crystals of **8** were grown from CH₂Cl₂/MeOH solution. 1H NMR (500 MHz, CDCl₃) δ 9.02-8.88 (m, 1H), 8.02 (dd, J = 8.3, 1.7 Hz, 1H), 7.87 (s, 1H), 7.79-7.76 (m, 1H), 7.57 (d, J = 7.8 Hz, 1H), 7.53-7.501 (m, 1H), 7.48 (d, J = 8.3 Hz, 1H), 7.42 (dd, J = 8.1, 1.8 Hz, 1H), 7.27-7.26 (m, 1H), 7.08-7.06 (m, 1H), 6.92 (d, J = 8.3 Hz, 1H), 5.48 (s, 1H), 3.96 (s, 3H), 3.94 (s, 3H), 2.03 (s, 3H), 2.01 (s, 3H); 13C NMR (125 MHz, CDCl₃) δ 186.1, 184.1, 167.9, 165.3, 152.1, 149.3, 147.2, 141.1, 139.5, 138.7, 138.3, 128.2, 124.3, 120.7, 120.4, 119.5, 118.3, 115.4, 111.0, 110.5, 102.8, 56.2, 56.2, 28.3, 27.3; HRMS (ESI) calcd for C₂₅H₂₅O₅N₂Pt [M+H]⁺ 628.1411, found 628.1408.



Complex 9

Complex **9** was prepared according to the general procedure using **L4** (0.22 g, 0.64 mmol) and K₂PtCl₄ (0.26 g, 0.64 mmol) in 10 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The resulting chloride-bridged dimer (0.13 g, 0.11 mmol) was suspended in 7 mL 1,4-dioxane and combined with **L5** (0.034 g, 0.34 mmol) and Na2CO3 (0.12 g, 1.13 mmol). The crude sample was further purified by silica gel column chromatography, eluting with hexane with increasing amount of EtOAc from 50% hexane/50% EtOAc to 30% hexane/70% EtOAc, to give **9** in 23% overall yield (0.092 g, 0.14 mmol). X-ray quality single crystals of **9** were grown from CH₂Cl₂/MeOH solution. 1H NMR (400 MHz, CDCl₃) δ 9.00-8.92 (m, 1H), 7.80-7.75 (m, 1H), 7.48 (d, J = 8.0 Hz, 1H), 7.43 (d, J = 2.1 Hz, 1H), 7.20 (d, J = 8.2 Hz, 1H), 7.12-7.09 (m, 1H), 7.05-7.00 (m, 2H), 6.62-6.57 (m, 2H), 5.48 (s, 1H), 3.78 (s, 3H), 3.69 (s, 3H), 3.55(s, 3H), 2.01 (s, 6H); 13C NMR (100 MHz, CDCl₃) δ 185.9, 184.2, 170.0, 167.1, 150.0, 147.8, 147.3, 145.6, 142.7, 140.2, 138.3, 128.2, 127.5, 123.3, 122.8, 122.7, 121.3, 118.5, 112.5, 109.7, 102.6, 55.7, 55.7, 38.5, 28.2, 21.1; HRMS (ESI) calcd for C₂₆H₂₇O₅N₂Pt [M+H]⁺ 642.1562, found 642.1559.



Complex 10

Complex 10 was prepared according to the general procedure except that all operations were performed in the dark (aluminum foil-wrapped glassware) to prevent light-induced isomerization of the stilbene alkene. Two equivalents L2 (0.88 g, 2.79 mmol) and K₂PtCl₄ (0.56 g, 1.38 mmol) were reacted in 100 mL H₂O/2-ethoxyethanol (1:3, v/v) for the first step. The crude dimer intermediate (0.83 g, 0.76 mmol) was suspended in 50 mL 1,4-dioxane and combined with L6 (0.53 g, 2.29 mmol) and NaH (60% wt, 0.092 g, 2.29 mmol). The crude product was further purified by silica gel column chromatography, eluting with 60% hexane/40% EtOAc, to give 10 in 29% overall yield (0.29g, 0.34 mmol). 1H NMR (600 MHz, CDCl₃) & 8.95-8.89 (m, 1H), 7.79-7.76 (m, 1H), 7.68 (d, J = 1.6 Hz, 1H), 7.58 (d, J = 8.1 Hz, 1H), 7.41 (d, J = 8.0 Hz, 1H), 7.28-7.26 (m, 1H), 7.15 (d, J = 16.1 Hz, 1H), 7.11-7.06 (m, 3H), 7.03 (d, J = 16.1 Hz, 1H), 6.87 (d, J = 8.2 Hz, 1H), 4.28-4.25 (m, 2H), 3.96 (s, 3H), 3.91 (s, 3H), 3.63-3.61 (m, 2H), 3.40 (s, 3H), 2.77-2.74 (m, 2H), 2.52-2.49 (m, 2H), 2.21 (s, 3H), 2.17 (s, 3H); 13C NMR (150 MHz, CDCl₃) § 184.8, 182.7, 173.0, 168.0, 149.0, 148.9, 147.1, 143.9, 139.4, 138.2, 138.0, 130.7, 129.0, 128.5, 127.6, 123.2, 121.6, 120.8, 120.1, 118.3, 111.1, 110.0, 108.8, 70.4, 63.6, 59.0, 55.9, 55.9, 35.2, 27.8, 27.1, 26.4; HRMS (ESI) calcd for C₃₂H₃₆O₇NPt [M+H]⁺ 741.2134, found 741.2130.



Complex 11

Complex **11** was prepared according to the general procedure using 2 equivalents of **L3** (1.27 g, 3.81 mmol) and K₂PtCl₄ (0.79 g, 1.90 mmol) in 100 mL H₂O/2-ethoxyethanol (1:3, v/v) for the first step. The dimer intermediate (1.05 g, 0.93 mmol) was suspended in 50 mL 1,4-dioxane and combined with **L6** (0.65 g, 2.82 mmol) and NaH (60% wt, 0.11 g, 2.82 mmol). The crude

product was further purified by silica gel column chromatography, eluting with 40% hexane/60% EtOAc, to give **11** in 32% overall yield (0.47 g, 0.62 mmol). 1H NMR (500 MHz, CDCl₃) δ 8.87 (d, J = 5.5 Hz, 1H), 7.97-7.93 (m, 2H), 7.79-7.72 (m, 1H), 7.56-7.46 (m, 3H), 7.41 (dd, J = 8.4, 2.0 Hz, 1H), 7.32-7.29 (m, 1H), 7.05 (t, J = 6.6 Hz, 1H), 6.91-6.89 (m, 1H), 4.26-4.24 (m, 2H), 3.95 (s, 3H), 3.93 (s, 3H), 3.62-6.60 (m, 2H), 3.39 (s, 3H), 2.75-2.71 (m, 2H), 2.50-2.46 (m, 2H), 2.16 (s, 6H); 13C NMR (125 MHz, CDCl₃) δ 185.0, 182.7, 173.1, 167.9, 165.3, 152.0, 149.2, 147.0, 141.1, 140.2, 138.7, 138.2, 128.0, 124.2, 120.6, 120.6, 119.5, 118.3, 115.3, 110.9, 110.4, 110.2, 70.5, 63.7, 59.1, 56.2, 56.1, 35.3, 27.9, 27.1, 26.5; HRMS (ESI) calcd for C₃₁H₃₅O₈N₂Pt [M+H]⁺ 758.2036, found 758.2036.



Complex 12

Complex **12** was prepared according to the general procedure using **L3** (0.24 g, 0.72 mmol) and K₂PtCl₄ (0.30 g, 0.72 mmol) in 10 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The resulting chloride-bridged dimer (0.35 g, 0.31 mmol) was suspended in 15 mL 1,4-dioxane and combined with **L7** (0.24 g, 0.93 mmol) and NaH (60% wt, 0.075 g, 1.87 mmol). The crude sample was further purified by silica gel column chromatography, eluting with DCM with increasing amount of MeOH from 98% DCM/2% MeOH to 95% DCM/5% MeOH, to give **12** in 17% overall yield (0.078 g, 0.099 mmol). 1H NMR (400 MHz, CDCl₃) δ 7.78 (d, J = 5.5 Hz, 1H), 8.58 (s, 1H), 7.81-7.73 (m, 1H), 7.65 (t, J = 7.6 Hz, 1H), 7.52-7.44 (m, 4H), 7.38 (d, J = 8.5 Hz, 1H), 7.02-6.94 (m, 1H), 6.93-6.88 (m, 1H), 6.79 (d, J = 8.3 Hz, 1H), 3.86 (s, 3H), 3.85 (s, 3H), 3.67-3.59 (m, 2H), 3.51-3.43 (m,4H), 3.38-3.31 (m, 2H), 2.66-2.55(m, 2H), 2.24-2.15 (m, 2H), 2.02 (s, 3H), 1.98 (s, 3H); 13C NMR (100 MHz, CDCl₃) δ 184.9, 182.5, 173.0, 167.6, 165.8, 151.9, 148.9, 146.7, 141.0, 140.6, 138.6, 138.1, 127.6, 124.0, 121.3, 120.6, 120.2, 118.3, 115.8, 110.9, 110.5, 110.3, 72.3, 69.8, 61.5, 56.0, 56.0, 39.3, 37.2, 27.9, 27.0, 26.8; HRMS (ESI) calcd for C₃₂H₃₈O₈N₃Pt [M+H]⁺ 787.2301, found 787.2296.



Complex **13** was prepared according to the general procedure using **L3** (0.15 g, 0.45 mmol) and K₂PtCl₄ (0.18 g, 0.45 mmol) in 5 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The crude dimer intermediate (0.089 g, 0.08 mmol) was suspended in 8 mL 1,4-dioxane and combined with **L8** (0.083 g, 0.24 mmol) and NaH (60% wt, 0.028 g, 0.71 mmol). The crude sample was further purified by silica gel column chromatography, eluting with DCM with increasing amount of MeOH from 99% DCM/1% MeOH to 96% DCM/4% MeOH, to give **13** in 45% overall yield (0.17 g, 0.19 mmol). 1H NMR (400 MHz, CDCl₃) δ 8.88-8.86 (m, 1H), 8.03 (s, 1H), 7.94-7.89 (m, 1H), 7.80-7.73 (m, 1H), 7.55 (d, J = 8.1 Hz, 1H), 7.51 (d, J = 1.7 Hz, 1H), 7.48-7.41 (m, 2H), 7.39-7.31 (m, 1H), 7.09-7.02 (m, 1H), 6.90 (d, J = 8.3 Hz, 1H), 3.94 (s, 3H), 3.93 (s, 3H), 3.75-3.49 (m, 16H), 2.79-2.68 (m, 2H), 2.56-2.47 (m, 2H), 2.16 (s, 3H), 2.15(s, 3H); 13C NMR (100 MHz, CDCl₃) δ 185.1, 182.7, 173.3, 167.9, 165.4, 152.1, 149.2, 147.0, 141.1, 140.3, 138.7, 138.2, 128.0, 124.2, 120.7, 120.7, 119.6, 118.3, 115.4, 110.9, 110.6, 110.4, 72.9, 72.4, 69.5, 69.2, 61.8, 61.7, 56.2, 56.1, 49.0, 46.9, 34.4, 28.0, 27.2, 26.6; HRMS (ESI) calcd for C₃₆H₄₆O₁₀N₃Pt [M+H]⁺ 875.2825, found 875.2815.



Complex 14

Complex 14 was prepared according to the general procedure using L1 (0.11g, 0.72 mmol) and K_2PtCl_4 (0.30 g, 0.72 mmol) in 5 mL H₂O/2-ethoxyethanol (1:4, v/v) for the first step. The dimer intermediate (0.18 g, 0.23 mmol) was suspended in 6 mL 1,4-dioxane and combined with L8 (0.21 g, 0.61 mmol) and NaH (60% wt, 0.088 g, 2.22 mmol). The crude sample was further purified by silica gel column chromatography, eluting with DCM with increasing amount of MeOH from 98% DCM/2% MeOH to 94% DCM/6% MeOH, to give 14 in 33% overall yield

(0.16 g, 0.23 mmol). 1H NMR (400 MHz, CDCl₃) δ 8.99-8.86 (m, 1H), 7.81-7.74 (m, 1H), 7.66-7.54 (m, 2H), 7.43 (dd, J = 7.8, 1.0 Hz, 1H) 7.19 (td, J = 7.4, 1.3 Hz, 1H), 7.12-7.05 (m, 2H), 3.71-3.54 (m, 16H), 2.78-2.72 (m, 2H), 2.54-2.48 (m, 2H), 2.16 (s, 3H), 2.14 (s, 3H); 13C NMR (100 MHz, CDCl₃) δ 185.0, 182.8, 173.3, 168.5, 147.2, 144.8, 139.8, 138.2, 130.7, 129.3, 123.6, 123.1, 121.2, 118.4, 110.5, 72.9, 72.4, 69.5, 69.3, 61.8, 61.7, 49.0, 46.9, 34.4, 28.0, 27.1, 26.6; HRMS (ESI) calcd for C₂₇H₃₇O₇N₂Pt [M+H]⁺ 696.2243, found 696.2236.

Absorption and Emission Spectra



Fig. S1. Absorption spectra of 6 in various solvents. [6] = $10 \mu M$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S2. Absorption spectra of 7 in various solvents. $[7] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S3. Absorption spectra of 8 in various solvents. [8] = 10 μ M with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S4. Absorption spectra of 9 in various solvents. $[9] = 10 \mu M$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S5. Absorption spectra of 10 in various solvents. $[10] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S6. Absorption spectra of 11 in various solvents. $[11] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S7. Absorption spectra of 12 in various solvents. $[12] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S8. Absorption spectra of 13 in various solvents. $[13] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S9. Absorption spectra of 14 in various solvents. $[14] = 10 \ \mu\text{M}$ with 0.5% (v/v) DMSO in the indicated solvents.



Fig. S10. Emission spectra of 7 in different solvents (containing 0.5% (v/v) DMSO); [7] = 10 μ M; $\lambda_{ex} = 410$ nm.



Fig. S11. Emission spectra of 10 in different solvents containing 0.5% (v/v) DMSO; $[10] = 10 \mu$ M; $\lambda_{ex} = 410 \text{ nm}$.



Fig. S12. Emission spectra of 6 in different solvents containing 0.5% (v/v) DMSO; [6] = 10 μ M; $\lambda_{ex} = 410$ nm.



Fig. S13. Emission spectra of 9 in different solvents containing 0.5% (v/v) DMSO; [9] = 10 μ M; $\lambda_{ex} = 410$ nm.



Fig. S14. Emission spectra of 11 in different solvents containing 0.5% (v/v) DMSO; [11] = 10 μ M; $\lambda_{ex} = 410$ nm.



Fig. S15. Emission spectra of 12 in different solvents containing 0.5% (v/v) DMSO; $[12] = 10 \mu$ M; $\lambda_{ex} = 410 \text{ nm}$.



Fig. S16. Emission spectra of 14 in different solvents containing 0.5% (v/v) DMSO; [14] = 10 μ M; $\lambda_{ex} = 410$ nm.



Fig. S17. ThT fluorescence assay of HEWL fibrillization. $\lambda_{ex} = 440 \text{ nm}$, $\lambda_{em} = 480 \text{ nm}$, error bar n = 3, $I_o = ThT$ emission at T = 0.



Fig. S18. Plot of emission enhancement (I/I₀) for complex **13** (cyan line) and **14** (black line) as a function of HEWL aggregation. $\lambda_{ex} = 410$ nm for both complexes, $\lambda_{em} = 490$ nm for **14** and 510 nm for **13**, error bar n = 3.



Fig. S19. Saturation binding isotherm generated by plotting maximum emission of 8 at 510 nm in the presence of 5 μ M HEWL fibrils versus [8] (0-35 μ M). Data analysis performed using GraphPad Prism, error bars reflect results of 3 trials.



Fig. S20. Saturation binding isotherm generated by plotting maximum emission of 13 at 510 nm in the presence of 5 μ M HEWL fibrils versus [13] (0-35 μ M). Data analysis performed using GraphPad Prism, error bars reflect results of 3 trials.



Fig. S21. Relative luminescence response (I/I_o) of **13** (20 μ M) at 505 nm in the presence of select biomolecules (10 μ M) in PBS (0.5% DMSO); I_o = emission intensity in presence of HEWL monomer; $\lambda_{ex} = 410$ nm, error bar n = 3.

Preparation of HEWL oligomers/fibrils

HEWL oligomers/fibrils were prepared according to literature procedures.^{7, 8} Specifically, lyophilized HEWL was dissolved at 5 mg/mL (350 μ M) in 10 mM pH 3 sodium citrate buffer with 0.1 M NaCl. The solution was sonicated for 1 minute and then incubated in a 70 °C oil bath and magnetically stirred at 80 rpm for 8 hours. During this time aliquots were withdrawn at different time intervals and immediately diluted with pH 7.4 PBS (phosphate buffer saline) to stop further fibrillization. Samples were then refrigerated

until use in spectroscopic experiments. Fibrillization was complete after 8 h of incubation according to ThT fluorescence assay (see Fig. S17) and AFM imaging.

Preparation of Aβ₄₂ fibrils

 $A\beta_{42}$ fibrils were prepared according to literature methods.⁹ Specifically, HFIP-treated $A\beta_{42}$ (0.5 mg) was solubilized in 22 µL dry DMSO (5 mM) and sonicated for 2 minutes. The $A\beta_{42}$ solution was then diluted to 100 µM with 1085 µL 7.4 PBS buffer. The solution was incubated at 37 °C for 21 hours, after which time fibrillization was complete as determined by ThT fluorescence assay and AFM imaging.

General methods for luminescence assays

All luminescence assays involving ThT and Pt(II) complexes 6-14 in the presence of HEWL or A β_{42} aggregates were performed in 96-well plates using a SpectraMax® M3 plate reader (Molecular Device). Emission spectra of Pt complexes were obtained using $\lambda_{ex} = 410$ nm and monitoring emission wavelength from 440-800 nm. ThT assays were performed using $\lambda_{ex} = 440$ nm and monitoring emission at 480 nm.

Luminescence spectra of Pt(II) complexes in the presence of HEWL aggregates and fibrils

A 350 μ M HEWL solution in pH 3 sodium citrate buffer was prepared as described above and aliquots were withdrawn and diluted to 10 μ M in pH 7.4 PBS at 1 h time intervals throughout the fibrillization reaction (0 to 8 hours). Stock solutions of each Pt(II) complex (2 mM in DMSO) were prepared and aliquots from stock solutions were added to 10 μ M HEWL solutions in pH 7.4 PBS to give final Pt(II) complex concentrations of 20 μ M. The HEWL-Pt(II) complex solutions were allowed to stand at room temperature for 2 min and then the emission spectra were recorded.

Luminescence measurements of Pt(II) complexes 12 and 13 in the presence of $A\beta_{42}$ fibrils

Aliquots from a 100 μ M solution of A β_{42} fibrils (prepared as described above) were diluted to 10 μ M in pH 7.4 PBS. Aliquots from stock solutions of platinum(II) complexes **12** and **13** (0.4 mM in DMSO) were added to 10 μ M A β_{42} fibril solutions to give a final Pt complex concentration of 20 μ M. The Ab₄₂-Pt(II) complex solutions were allowed to stand at room temperature for 2 min, and then the emission spectra were recorded.

ThT fluorescence assays

ThT fluorescence assays were performed by adding aliquots of a ThT stock solution (2 mM in DMSO for HEWL experiments and 0.4 mM in DMSO for A β_{42} experiments) to 10 μ M amyloid solutions in pH 7.4 PBS to achieve a final ThT concentration of 20 μ M. The respective ThT-amyloid mixtures were allowed to stand at room temperature for 2 min and then emission spectra were recorded.

Determination of K_d for binding of 8 and 13 to HEWL fibrils

Equilibrium dissociation constants (K_d) for binding of Pt complexes 8 and 13 to HEWL fibrils:

 $K_d = [Pt(II) \text{ probe}][HEWL \text{ fibrils}]/[Pt(II) \text{ probe-HEWL complex}]$

were determined using luminescence titration experiments. Specifically, different concentrations of **8** and **13** (1-35 μ M) were titrated against a fixed concentration of HEWL fibrils (5 μ M) in 7.4 PBS. Emission spectra of Pt complexes were recorded with excitation at 410 nm and emission at 510 nm after incubation with HEWL fibrils for 2 mins at room temperature.

Changes in emission intensity were plotted versus probe concentration to yield binding isotherms (Figs. S19-S20). Experiments were performed in triplicate and results are given as the mean \pm SD. The corresponding isotherms were analyzed by GraphPad Prism (GraphPad Software, Inc., San Diego, CA, USA) by non-linear regression using one-site-binding equation. The K_d values were calculated using following equation:¹⁰

 $Y = B_{max} * X(K_d + X)$

where X is concentration of Pt(II) complex, Y is change in emission intensity, B_{max} is the maximum specific binding, which refers to the total number of receptors, and K_d is the equilibrium binding constant. It should be noted that while the presence of multiple binding sites in amyloid fibrils is possible,¹¹ application of a one-site binding model greatly simplifies data analysis.

AFM imaging experiments – general materials and methods

AFM height imaging was carried at room temperature and ambient pressure using a Molecular Force Probe 3D AFM system (Asylum Research, Santa Barbara, CA). Si_3N_4 AFM probes (MikroMasch, San Jose, CA) were used for all AFM experiments with a nominal spring constant of 0.4 N/m and a typical tip radius of curvature of 8 nm. AFM height images were acquired in intermittent contact mode (AC mode) at a typical scan rate of 1 Hz.

Preparation of Pt(II) complex – HEWL fibril samples for AFM imaging

HEWL fibrils were prepared by incubation of HEWL (350 μ M) in pH 3 sodium citrate buffer for 8 hours as described above. Aliquots were removed from the incubation solution and diluted to 10 μ M in pH 7.4 PBS. Aliquots from 2 mM DMSO stock solutions of platinum(II) complexes (8, 13, and 14) were added to diluted fibril solutions to give final Pt(II) complex concentrations of 20 μ M. The HEWL fibril/Pt complex mixtures were allowed to stand at room temperature for 24 h. After this time, 70 μ L aliquots from each solution were transferred onto freshly cleaved mica and placed in a room temperature dust-free environment for 20 min to allow physisorption to occur. The mica substrates were then washed with 70 μ L of HPLC-grade H₂O three times and airdried in a dust-free environment overnight prior to AFM imaging. A control sample of HEWL fibrils was prepared as described above by mixing diluted HEWL fibril solutions with an equivalent amount of DMSO instead of Pt(II) complex stock solution.

Preparation of $A\beta_{42}$ fibril samples for AFM imaging

A solution of $A\beta_{42}$ fibrils (100 μ M) was prepared as described above. A 10 μ L aliquot of this solution was withdrawn and transferred directly onto freshly cleaved mica. The sample was placed in a room temperature dust-free environment for 20 min. for physisorption to occur and then washed with 20 μ L HPLC-grade H₂O three times followed by air-drying overnight in a dust-free environment prior to AFM imaging.

MTT assays

SH-SY5Y cells (purchased from ATCC) were treated with Pt(II) complexes **8** and **11-13** at concentrations ranging from 1 μ M to 100 μ M for 24 h. Cells were also treated with the cytotoxic agent Rotenone (1 μ M, 5 μ M and/or 10 μ M) as a negative control. Cell viability was assessed using the colorimetric reagent, 3-[4,5-dimethylthiazol-2-yl]-2,5-diphenyl tetrazolium bromide (MTT) (Sigma Aldrich). Cultures were incubated in HBSS/glucose with 2 mg/mL MTT for 2 h at 37 °C. Following incubation, 0.4 mL DMSO was added to each well to solubilize the formazan product. Reduced MTT was measured on a microplate reader (Molecular Devices Spectra Max 190) at 570 nm with a reference of 650 nm.

X-Ray crystallography

General data collection

Data were collected on a Bruker D8 VENTURE DUO diffractometer equipped with a IµS 3.0 microfocus source operated at 75 W (50kV, 1.5 mA) to generate Mo K α radiation ($\lambda = 0.71073$ Å) and a PHOTON III detector. Crystals were transferred from the vial and placed on a glass slide in type NVH immersion oil by Cargille. A Zeiss Stemi 305 microscope was used to identify a suitable specimen for X-ray diffraction from a representative sample of the material. The crystal and a small amount of the oil were collected on a MiTeGen 100 micron MicroLoop and transferred to the instrument where it was placed under a cold nitrogen stream (Oxford 800 series) maintained at 100K throughout the duration of each experiment. Samples were optically centered with the aid of a video camera to ensure that no translations were observed as crystals were rotated through all positions. A unit cell collection was then carried out. After it was determined that the unit cell was not present in the CCDC database a data collection strategy was calculated by *APEX4*.¹² The crystal was measured for size, morphology, and color.

Refinement details

After data collection, the unit cell was re-determined using a subset of the full data collection. Intensity data were corrected for Lorentz, polarization, and background effects using the *APEX4*.¹² A numerical absorption correction was applied based on a Gaussian

integration over a multifaceted crystal and followed by a semi-empirical correction for adsorption applied using *SADABS*.¹² The program *SHELXT*¹³ was used for the initial structure solution and *SHELXL*¹⁴ was used for the refinement of the structure. Both programs were utilized within the OLEX2 software.¹⁵ Hydrogen atoms bound to carbon atoms were located in the difference Fourier map and were geometrically constrained using the appropriate AFIX commands.

CCDC number	2322597
Empirical formula	$C_{26}H_{25}NO_4Pt$
Formula weight	610.56
Temperature/K	150.15
Crystal system	monoclinic
Space group	C/c
a/Å	15.5673(16)
b/Å	21.225(2)
c/Å	6.6731(7)
α/°	90
β/°	95.563(5)
γ/°	90
Volume/Å ³	2194.5(4)
Ζ	4
$\rho_{calc}g/cm^3$	1.848
μ/mm ⁻¹	6.428
F(000)	1192.0
Crystal size/mm ³	$0.220 \times 0.220 \times 0.145$
Radiation	MoKα (λ = 0.71073)
2Θ range for data collection/°	5.258 to 54.010
Index ranges	$-19 \le h \le 19, -27 \le k \le 27, -8 \le l \le 8$
Reflections collected	12942
Independent reflections	4529 [$R_{int} = 0.0199, R_{sigma} = 0.0379$]
Data/restraints/parameters	4529/2/293
Goodness-of-fit on F ²	0.609
Final R indexes $[I \ge 2\sigma(I)]$	$R_1 = 0.0181, wR_2 = 0.0518$
Final R indexes [all data]	$R_1 = 0.02\overline{12}, wR_2 = 0.0715$
Largest diff. peak/hole / e Å ⁻³	1.225/-0.519

 Table S1. Crystallographic data for 7.



Fig. S22. ORTEP plot of Pt(II) complex 7 with thermal ellipsoids represented at 50% probability. Carbon, nitrogen, oxygen, and platinum atoms are represented by gray, blue, red, and light blue ellipsoids, respectively. Platinum(II) bond distances and angles are typical for a cyclometallated square-planar Pt(II) acac complex. Pt1…N1 1.996 Å; Pt1…C7 1.978 Å; Pt1…O3 2.011 Å; Pt1…O4 2.087 Å; N1-Pt1-C7 81.61°; C7-Pt1-O3 93.58°; O3-Pt1-O4 91.98°; O4-Pt1-N1 92.81°; N1-Pt1-O3 175.19°; C7-Pt1-O4 174.12°. $\tau_4 = 0.08.^{16}$

CCDC number	2322598
Empirical formula	$C_{26}H_{28}N_2O_6Pt$
Formula weight	659.59
Temperature/K	100.00
Crystal system	monoclinic
Space group	P2 ₁ /c
a/Å	11.5142(6)
b/Å	12.2748(6)
c/Å	17.2202(10)
α/°	90
β/°	96.978(2)
$\gamma/^{\circ}$	90
Volume/Å ³	2415.8(2)
Ζ	4
$\rho_{calc}g/cm^3$	1.814
µ/mm ⁻¹	5.853
F(000)	1296.0
Crystal size/mm ³	$0.238 \times 0.139 \times 0.118$
Radiation	MoKα (λ = 0.71073)
2Θ range for data collection/°	4.766 to 72.734
Index ranges	$-19 \le h \le 18, -20 \le k \le 17, -28 \le l \le 28$
Reflections collected	99722
Independent reflections	11747 [$R_{int} = 0.0769, R_{sigma} = 0.0463$]
Data/restraints/parameters	11747/0/323
Goodness-of-fit on F ²	1.032
Final R indexes [I>= 2σ (I)]	$R_1 = 0.0340, wR_2 = 0.0668$
Final R indexes [all data]	$R_1 = 0.0505, wR_2 = 0.0726$
Largest diff. peak/hole / e Å ⁻³	2.17/-1.57

Table S2. Crystallographic data for 8·CH₃OH.



Fig. S23. ORTEP plot of Pt(II) complex 8·CH₃OH with thermal ellipsoids represented at 50% probability. Carbon, nitrogen, oxygen, and platinum atoms are represented by gray, blue, red, and light blue ellipsoids, respectively. Platinum(II) bond distances and angles are typical for a cyclometallated square-planar Pt(II) acac complex. Pt1…N1 1.991 Å; Pt1…C7 1.971 Å; Pt1…O5 2.003 Å; Pt1…O4 2.081 Å; N1-Pt1-C7 81.66°; C7-Pt1-O5 93.69°; O5-Pt1-O4 91.79°; O4-Pt1-N1 92.85°; N1-Pt1-O5 175.34°; C7-Pt1-O4 174.51°. $\tau_4 = 0.07.^{16}$

CCDC number	2322599
Empirical formula	C ₂₆ H ₂₆ N ₂ O ₅ Pt
Formula weight	641.58
Temperature/K	100.00
Crystal system	monoclinic
Space group	$P2_1/c$
a/Å	14.1430(6)
b/Å	12.3571(7)
c/Å	14.0848(8)
α/°	90
β/°	107.375(2)
$\gamma/^{\circ}$	90
Volume/Å ³	2349.2(2)
Ζ	4
$\rho_{calc}g/cm^3$	1.814
μ/mm ⁻¹	6.013
F(000)	1256.0
Crystal size/mm ³	$0.16 \times 0.137 \times 0.132$
Radiation	MoKα (λ = 0.71073)
20 range for data collection/°	4.468 to 61.04
Index ranges	$-18 \le h \le 20, -17 \le k \le 17, -20 \le l \le 20$
Reflections collected	52908
Independent reflections	7167 [$R_{int} = 0.0685$, $R_{sigma} = 0.0431$]
Data/restraints/parameters	7167/0/313
Goodness-of-fit on F ²	1.041
Final R indexes [I>= 2σ (I)]	$R_1 = 0.0278, wR_2 = 0.0530$
Final R indexes [all data]	$R_1 = 0.0379, wR_2 = 0.0564$
Largest diff. peak/hole $\overline{/ e \text{ Å}^{-3}}$	1.22/-1.12

 Table S3. Crystallographic data for 9.



Fig. S24. ORTEP plot of Pt(II) complex **9** with thermal ellipsoids represented at 50% probability. Carbon, nitrogen, oxygen, and platinum atoms are represented by gray, blue, red, and light blue ellipsoids, respectively. Platinum(II) bond distances and angles are typical for a cyclometallated square-planar Pt(II) acac complex. Pt1…N1 1.991 Å; Pt1…C7 1.967 Å; Pt1…O5 2.008 Å; Pt1…O4 2.087 Å; N1-Pt1-C7 81.67°; C7-Pt1-O5 92.81°; O5-Pt1-O4 91.81°; O4-Pt1-N1 93.74°; N1-Pt1-O5 174.10°; C7-Pt1-O4 175.32°. $\tau_4 = 0.08.^{16}$

References

- 1. E. K. Aratikatla, T. R. Valkute, S. K. Puri, K. Srivastava and A. K. Bhattacharya, *Eur. J. Med. Chem.*, 2017, **138**, 1089-1105.
- 2. H.-G. Cheng, M. Pu, G. Kundu and F. Schoenebeck, Org. Lett., 2019, 22, 331-334.
- 3. J. Suttil, J. Kucharyson, I. Escalante-Garcia, P. Cabrera, B. James, R. Savinell, M. Sanford and L. Thompson, *J. Mater. Chem. A*, 2015, **3**, 7929-7938.
- 4. M. Fu, L. Chen, Y. Jiang, Z.-X. Jiang and Z. Yang, Org. Lett., 2016, 18, 348-351.
- 5. L. J. O'Driscoll, D. J. Welsh, S. W. Bailey, D. Visontai, H. Frampton, M. R. Bryce and C. J. Lambert, *Chem-Eur. J.*, 2015, **21**, 3891-3894.
- 6. J. Brooks, Y. Babayan, S. Lamansky, P. I. Djurovich, I. Tsyba, R. Bau and M. E. Thompson, *Inorg. Chem.*, 2002, **41**, 3055-3066.
- P. L. Donabedian, T. K. Pham, D. G. Whitten and E. Y. Chi, ACS Chem. Neurosci., 2015, 6, 1526-1535.
- 8. M. Mulaj, J. Foley and M. Muschol, J. Am. Chem. Soc., 2014, 136, 8947-8956.
- 9. L. Sun, H.-J. Cho, S. Sen, A. S. Arango, T. T. Huynh, Y. Huang, N. Bandara, B. E. Rogers, E. Tajkhorshid and L. M. Mirica, *J. Am. Chem. Soc.*, 2021, **143**, 10462-10476.
- 10. S. Dhein, F. W. Mohr and M. Delmar, *Practical methods in cardiovascular research*, Springer, 2005.
- B. Jiang, U. Umezaki, A. Augustine, V. M. Jayasinghe-Arachchige, L. F. Serafim, Z. M. S. He, K. M. Wyss, R. Prabhakar and A. A. Martí, *Chemical Science*, 2023, 14, 1072-1081.
- 12. Bruker (2021). APEX4, SAINT and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.
- 13. G. M. Sheldrick, Acta Crystallogr., 2015, A71, 3-8.
- 14. G. M. Sheldrick, Acta Crystallogr., 2015, C71, 3-8.
- 15. O. V. Dolomanov, L. J. Bourhis, R. J. Gildea, J. A. K. Howard, and H. Puschmann, J. *Appl. Cryst.*, 2009, **42**, 339-341.
- 16. L. Yang, D. R. Powell, and R. P. Houser, *Dalton Trans.* 2007, 955-964.

















































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