

Supporting information

Design and development of copper-manganese oxide nanocatalyst for the sustainable synthesis of 1,4-disubstituted 1,2,3-triazoles

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Electronic Supplementary Information (ESI) includes PXRD and TEM image of recycled CuO@MnO₂ nanocatalyst; Green chemistry metric calculations, ¹H NMR and ¹³C NMR spectra of all compounds.

Total No of Pages: 19, Total No of Tables: 0, Total number of Figures: 2.

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1. General Remarks

Chemicals and solvents were purchased from Sigma Adrich, Alfa-Aesar and Merck India Pvt. ^1H and ^{13}C spectra were recorded on a Jeol Spectrospin spectrometer at 400 MHz and 100 MHz respectively by keeping TMS as internal standard. Chemical shift values were recorded in terms of δ and coupling constants (J) are in hertz (Hz). FT-IR spectra were recorded on Perkin Elmer. Field Emission Scanning Electron microscopy (FESEM) measurement was performed on a Zeiss Gemini SEM 500 thermal field emission type with acceleration voltage 0.02 – 30kV at USIC (University Science Instrument Centre), University of Delhi, Delhi, India. Field Emission Gun -Transmission electron microscope 200 KV (HR-TEM) was obtained on a JEOL JEM 2100F at Centre for Research in Nanotechnology and Science (CRNTS), IIT Bombay, Bombay, India. X-ray diffraction (XRD) patterns were recorded on Rigaku Rotaflex spectrometer at 2θ range of $10\text{--}90^\circ$ with Cu $K\alpha$ radiation. The elemental composition and electronic structure analysis were obtained from X-ray photoelectron spectra (XPS) of PHI Versa Probe III instrument equipped with a monochromatic Al $K\alpha$ source at ISM dhanbad. ICP-MS was measured on Agilent ICP-MS 7900 with UHMI at the ICP-MS Lab, (CRF Adopted Facility), IIT Delhi, INDIA. BET experiment was performed on BET surface area analyzer, Quantachrome Instruments, ASI-CI-11 at USIC (University Science Instrument Centre), University of Delhi, Delhi, India.

2. PXRD and TEM images of recycled CuO@MnO₂ nanocomposites

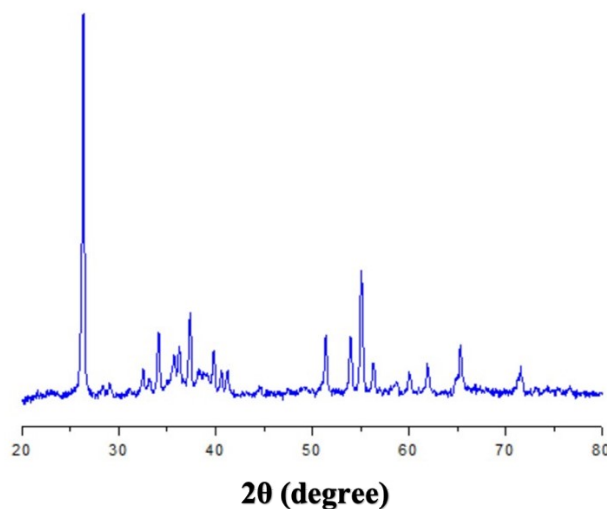


Figure S1: PXRD of recycled CuO@MnO₂ nanocomposites.

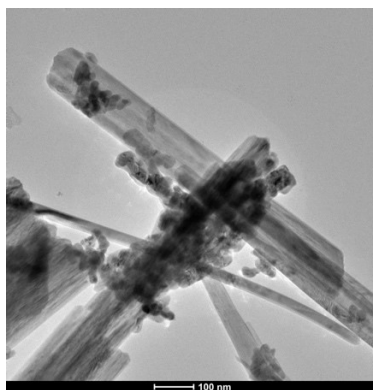
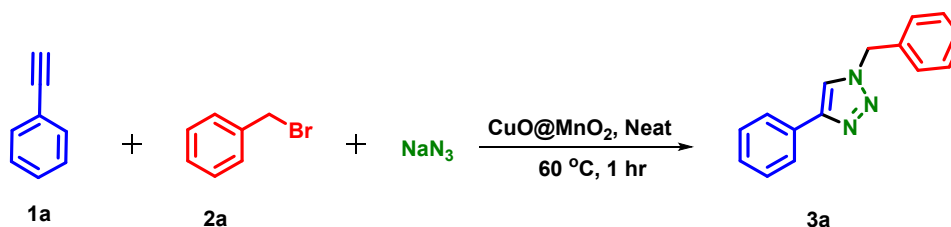


Figure S2: TEM image of recycled CuO@MnO₂ nanocomposites.

3. Calculation of E-factor and atom economy of compound 3a



FW	102.13 g/mol	171.04 g/mol	65.01g/mol	235.29 g/mol
m.mol	0.50	0.50	1	0.47
amount	51 mg	86 mg	65 mg	110 mg

E-factor:

The ideal value of E-factor is zero.

E-factor = [total mass of raw materials - the total mass of product]/ mass of product.

$$\begin{aligned} \text{E-factor of } \mathbf{3a} &= [(51 + 86 + 65) - 110] / 110 \\ &= 0.84. \end{aligned}$$

Process mass intensity (PMI):

$$\begin{aligned} \text{PMI} &= \sum (\text{mass of stoichiometric reactants}) / [\text{mass of product}] \\ &= (51 + 86 + 65) / 110 \\ &= 1.84. \end{aligned}$$

Reaction mass efficiency (RME):

$$\begin{aligned} \text{RME} &= [\text{mass of product} / \sum (\text{mass of stoichiometric reactants})] \times 100 \\ &= [110 / (51 + 86 + 65)] \times 100 \\ &= 54.45\% \end{aligned}$$

Carbon efficiency (CE):

CE denotes the percentage of carbon in the reactants that remains in the product.

$$\begin{aligned} \text{CE} &= [\text{Amount of carbon in product} / \text{Total carbon present in reactants}] \times 100 \\ &= [\text{no. of moles of product} \times \text{no. of carbons in product} / (\text{moles of 1a} \times \text{carbons in 1a} + \text{moles of 2} \times \text{carbons in 2})] \times 100 \\ &= [0.46 \times 15 / (0.50 \times 7 + 0.50 \times 8)] \times 100 \\ &= [6.9 / (3.5 + 4)] \times 100 \\ &= 92\% \end{aligned}$$

Turnover number (TON): The number of moles of substrate that converted by a mole of catalyst before its inactivation. TON is calculated for 5 cycles of the reaction with 25 mg of CuO@MnO₂.

$$\text{TON} = [\text{No. of moles of product}] \div [\text{No. of moles of catalyst}]$$

$$\text{TON of compound 3a: } [0.47 (1) + 0.46 (2) + 0.46 (3) + 0.44 (4) + 0.44 (5)] / [0.00252]$$

$$\text{TON of compound 3a: } 901$$

4. ^1H and ^{13}C Spectra of compounds

