Supporting Information

for

The role of interactions between cationic backbone and basic anions on green and ultra-selective catalytic synthesis of ethyl methyl carbonate in tunable ionized frameworks

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Fig. S1 FTIR spectra of BbIm, Trbb, CDI, and [CPIL-M]₄ [Br].

S2 Supplementary Figures



Fig. S2 Bond length of O-H on the hydroxyl of (A) pure ethanol, (B) ethanol interacted with [CPIL] [PhO], and (C) ethanol interacted with [CPIL-M][PhO].



Fig. S3 Swelling properties of $[CPIL-M]_2[PhO]$, $[CPIL-M]_4[PhO]$ and $[CPIL-M]_6[PhO]$ in the DMC/ethanol mixture.

S3 Supplementary Table

Table. S1 The basicity strength of [CPIL-M]₄[Br] and [CPIL-M]_n[PhO].

Catalyst	Basicity strength (mmol g ⁻¹)
[CPIL-M]₄[Br]	-1.333 ±0.428ª
[CPIL-M]₂[PhO]	1.538 ±0.064
[CPIL-M]₃[PhO]	1.959 ±0.084
[CPIL-M]4[PhO]	2.591 ±0.101
[CPIL-M]₅[PhO]	1.221 ±0.047
^a The [CPIL-M] ₄ [Br] could perform an ionic exchange with OH, thus showing an acidic	
feature with a minus basicity strength.	

The basicity strength of $[CPIL-M]_4[Br]$ and $[CPIL-M]_{1-6}[PhO]$ was assessed through acid-base titration. Specifically, 0.1 g of $[CPIL-M]_4[Br]$ and $[CPIL-M]_{1-6}[PhO]$ were each dissolved in 10 mL of 0.1 mol L⁻¹ HCl and allowed to react fully for 5 hours at 298 K. Subsequently, 3 mL of the supernatant liquid and 3 drops of phenolphthalein indicator were titrated with 0.01 mol L⁻¹ NaOH until the color of the solution turned pink. The basicity strength of these PILs were calculated by the following formula,

Basicity strength =[(10mL/3mL)*(3 mL*0.1 mol L⁻¹ -V_{NaOH} mL*0.01 mol L⁻¹)]/0.1 g

Catalyst	PMI _{RRC}
	(g g ⁻¹)
[CPIL-M]1[PhO]	3.825
[CPIL-M] ₂ [PhO]	3.791
[CPIL-M]₃[PhO]	3.426
[CPIL-M]₄[PhO]	2.768
[CPIL-M]₅[PhO]	3.903
[CPIL-M] ₆ [PhO]	3.831

Table. S2 The PMI_{RRC}^{1} of [CPIL-M]₁₋₆[PhO] under the reaction condition of reaction temperature of 80 °C, reaction temperature of 6 h, ethanol/DMC molar ratio of 1:1, catalyst dosage of 1.7%.

1.*PMI_{RRC}*: Process mass intensity: reactants, reagents, catalyst.