

## Supporting Information

### **Ionic liquid-based chemodosimeter probe for selective detection and removal of bisulfite in pure aqueous system, with potential uses in biosensing†**

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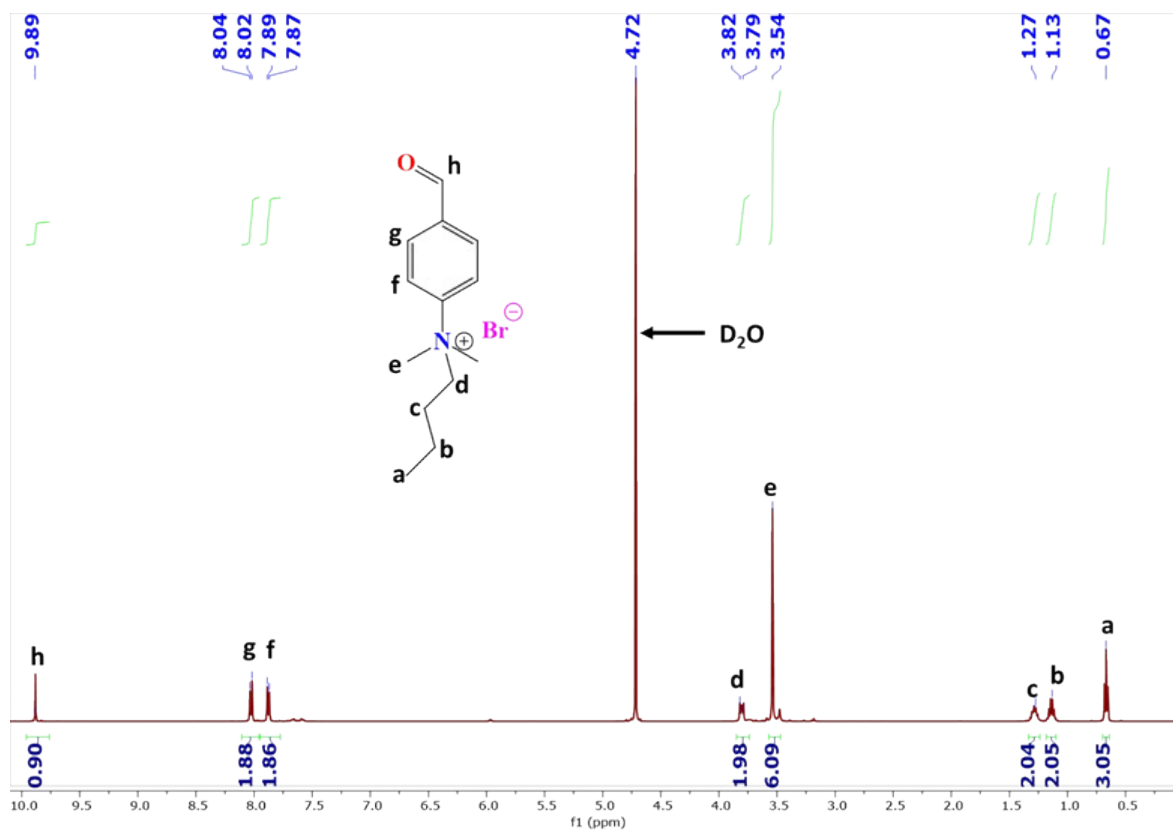


Fig. S1  $^1\text{H}$  NMR of hydrophilic TSIL-1 at 500 MHz in  $\text{D}_2\text{O}$  solvent.

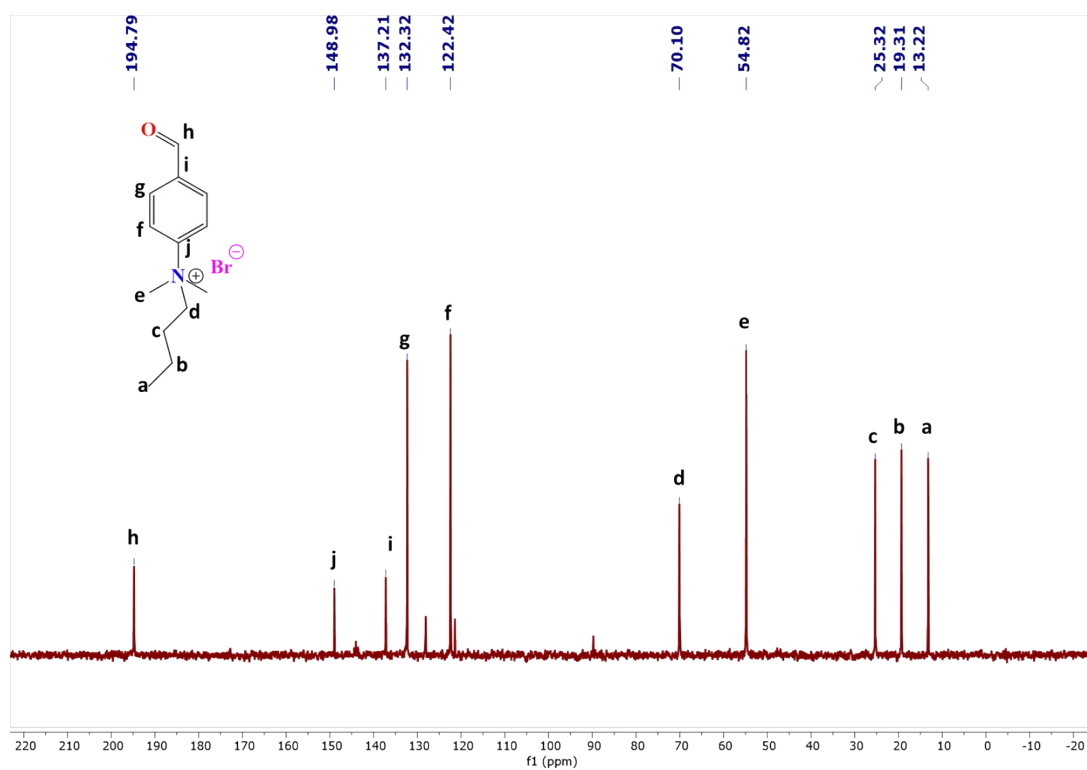
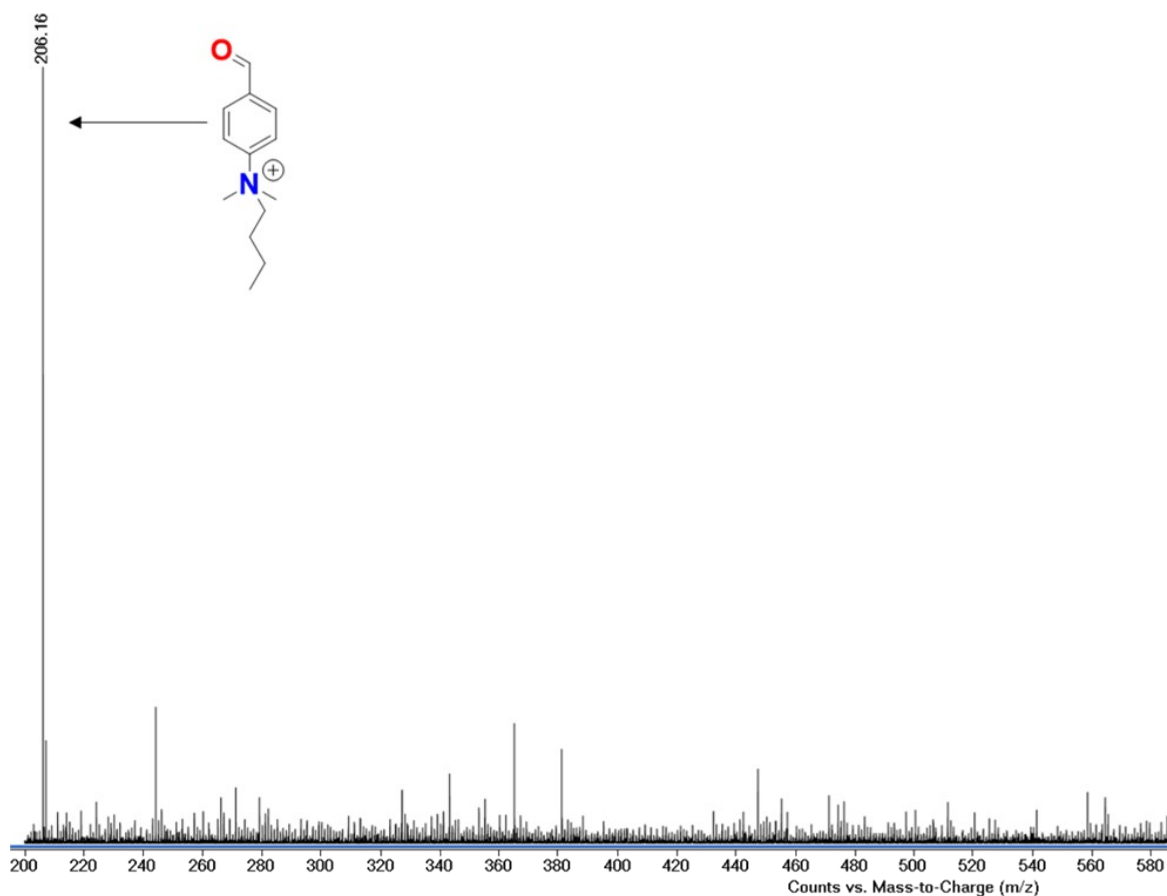
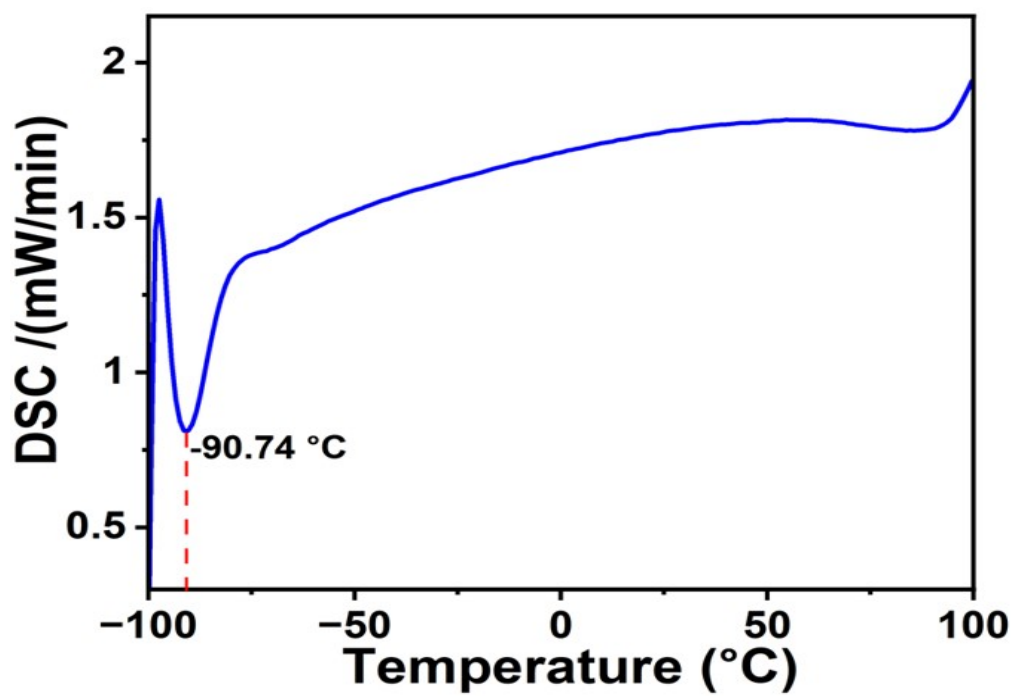


Fig. S2  $^{13}\text{C}$  NMR of hydrophilic TSIL-1 at 500 MHz in  $\text{D}_2\text{O}$  solvent.



**Fig. S3** ESI-MS of TSIL-1 (+ve Scan).



**Fig. S4** DSC profile of TSIL-1 showing room temperature ionic liquid behaviour (mp: -90 °C).

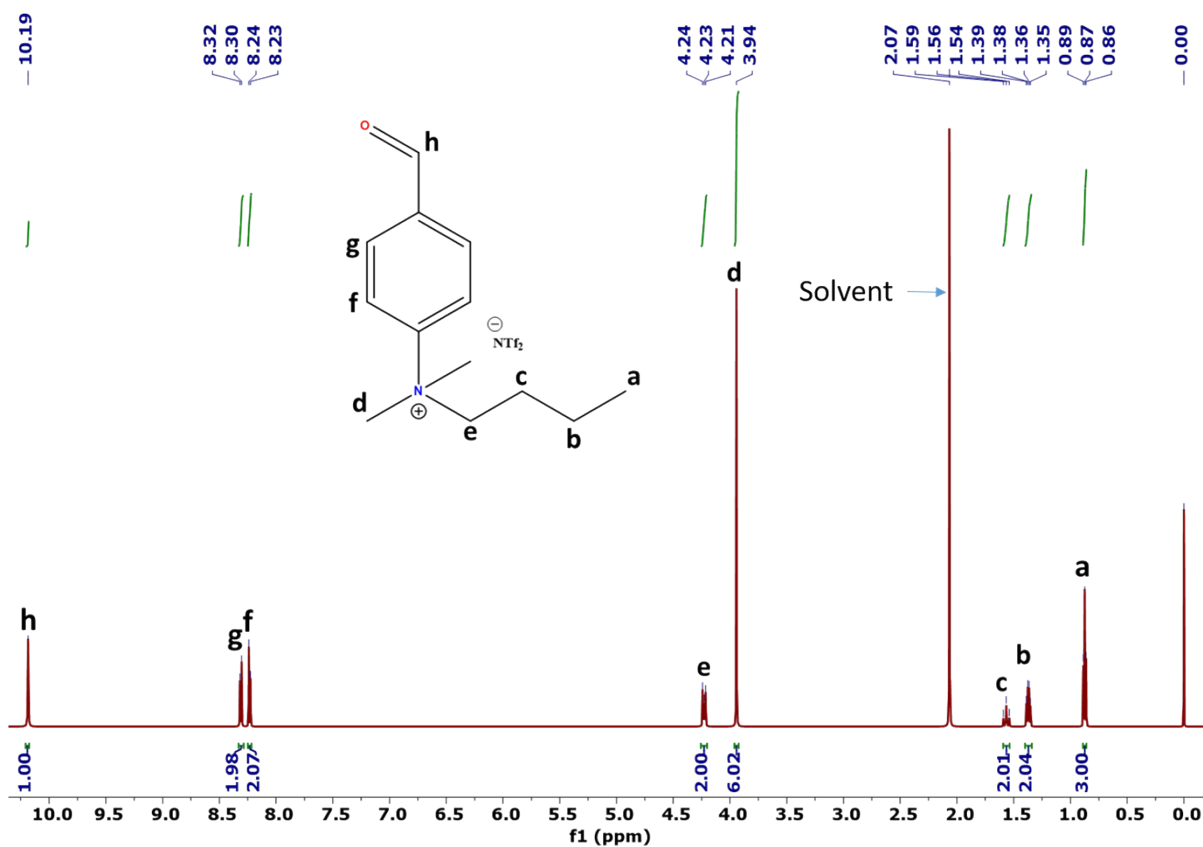


Fig. S5  $^1\text{H}$  NMR of hydrophobic TSIL-2 at 600 MHz in Acetone- $\text{d}_6$  solvent.

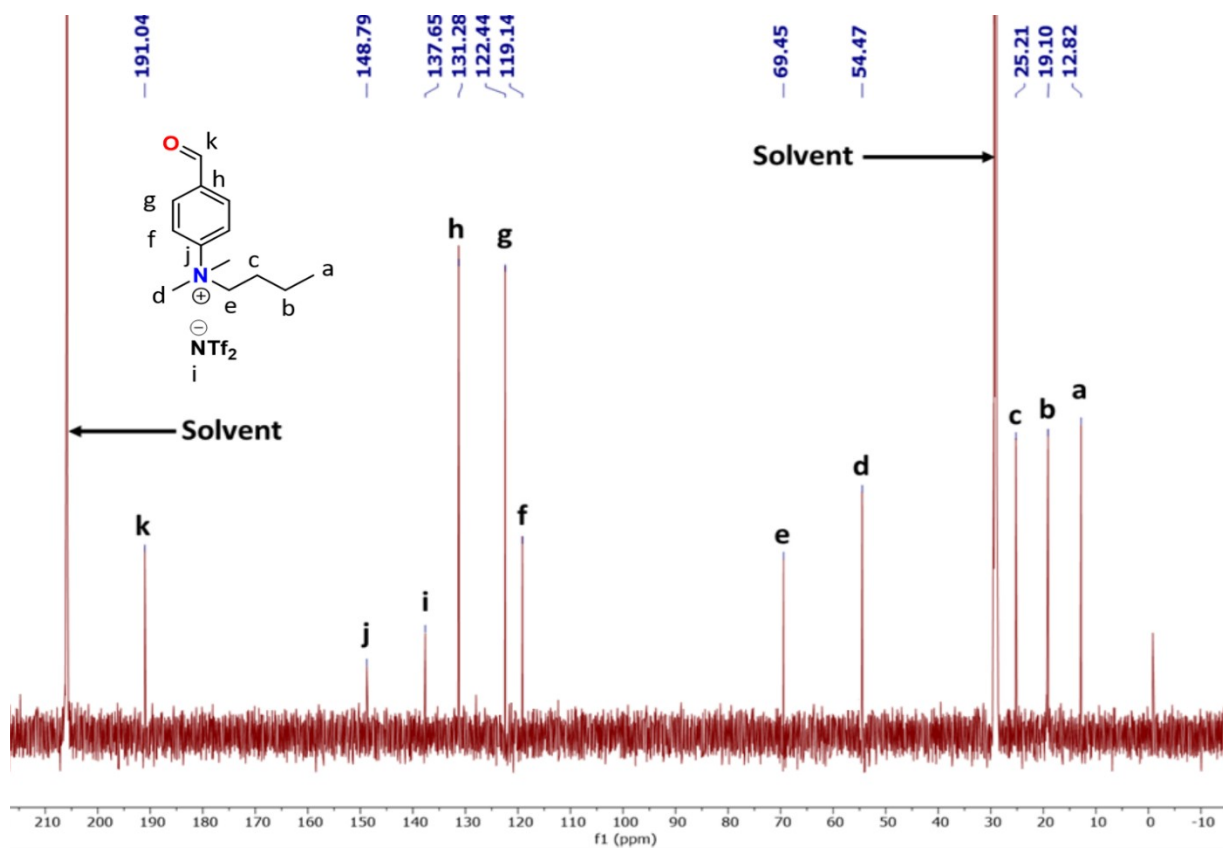
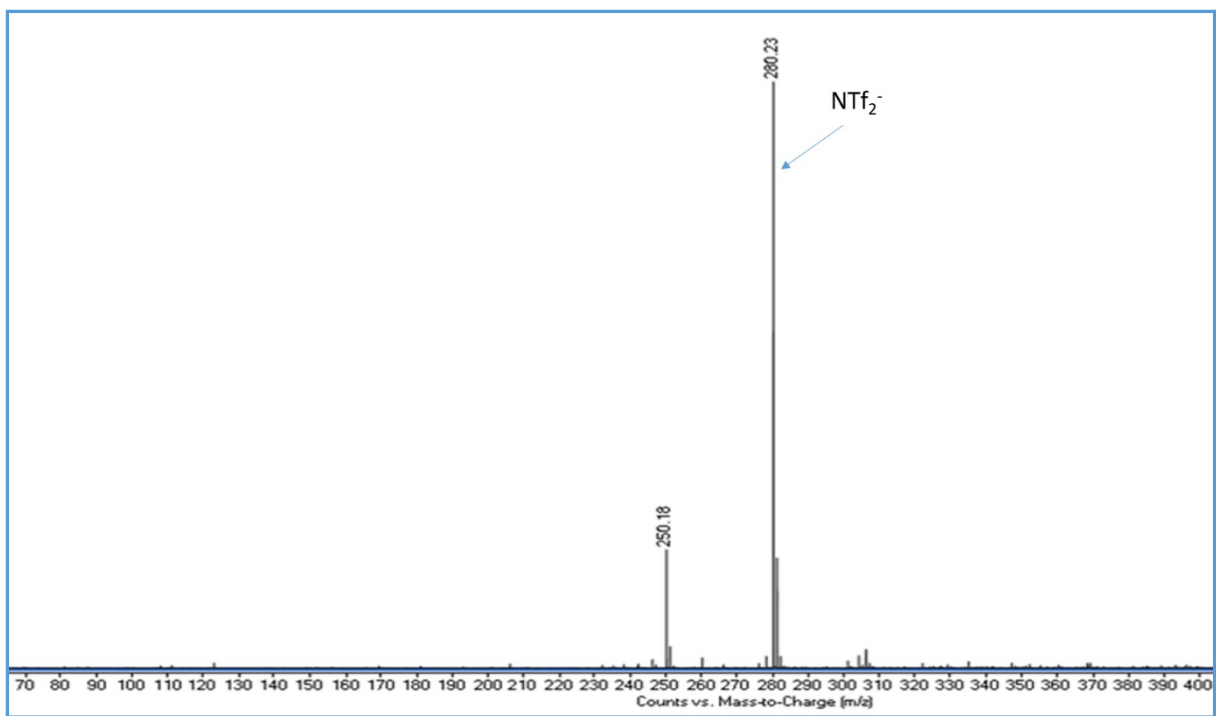
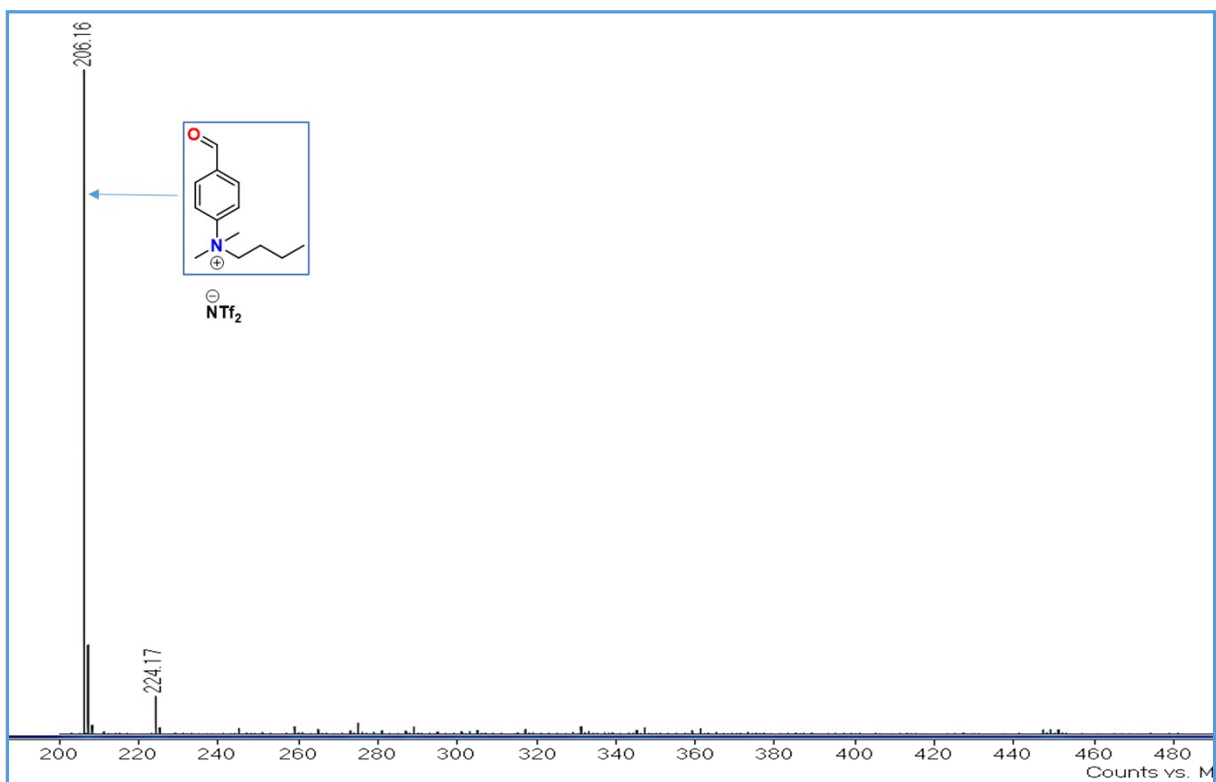


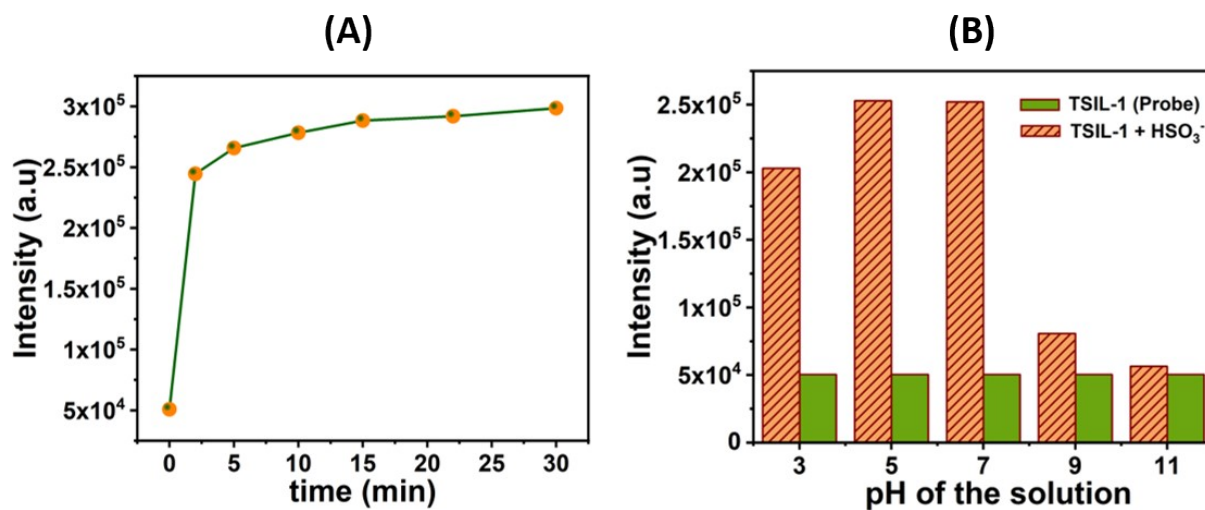
Fig. S6  $^{13}\text{C}$  NMR of hydrophobic TSIL-2 at 600 MHz in Acetone- $\text{d}_6$  solvent.



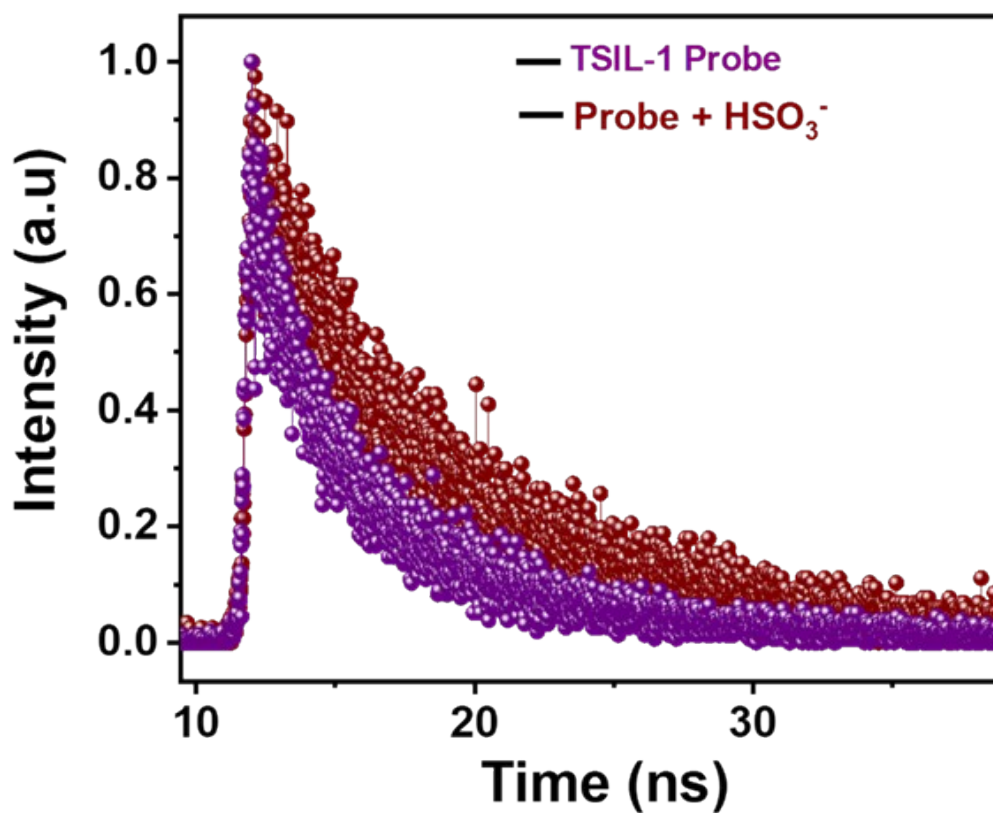
**Fig. S7** ESI-MS of TSIL-2 (-ve Scan).



**Fig. S8** ESI-MS of TSIL-2 (+ve Scan).



**Fig. S9** (A) The fluorescence intensity of the probe (TSIL-1) with addition of bisulfite over time (190  $\mu\text{M}$  bisulfite, pH  $\sim 7.2$ ); (B) The fluorescence emission of the probe (TSIL-1) with and without bisulfite at different pH.



**Fig. S10** The time-correlated single photon counting (TCSPC) experiment of the probe material (TSIL-1) before and after bisulfite addition.

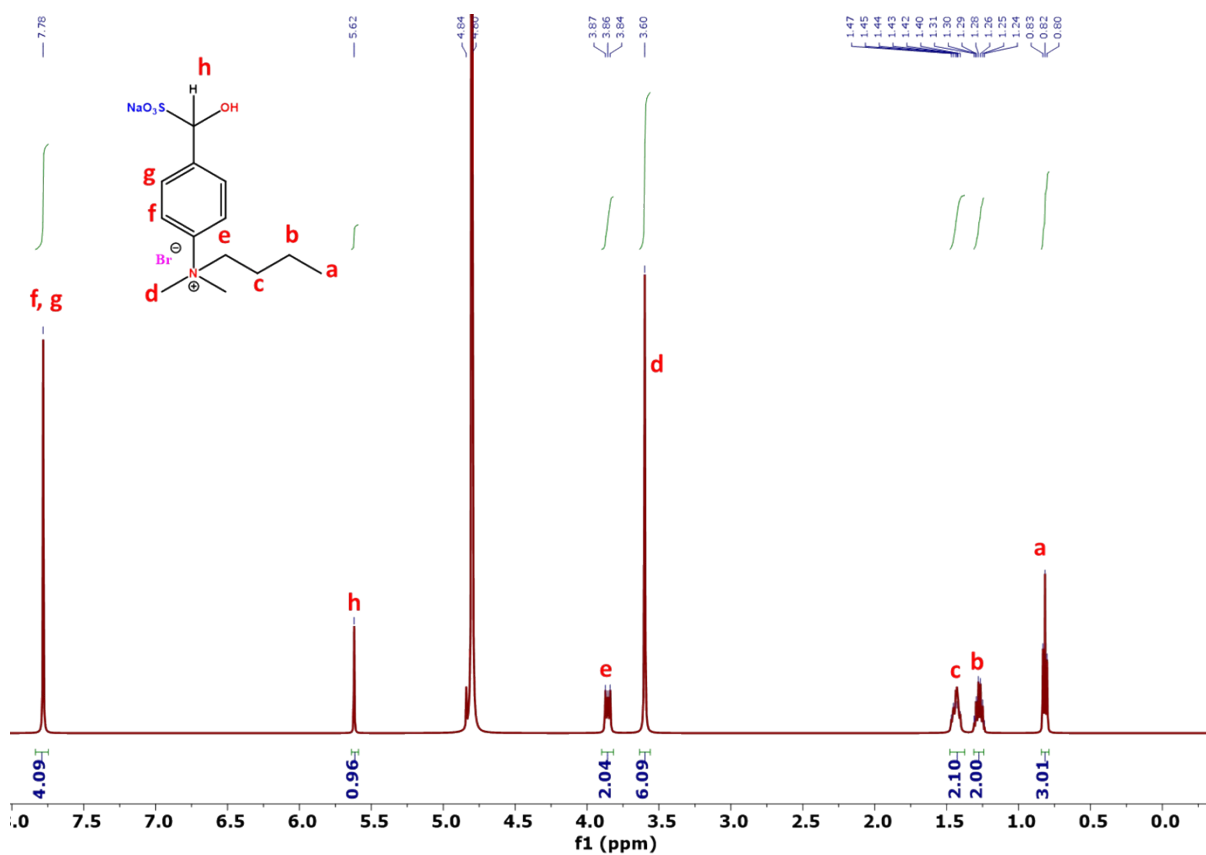
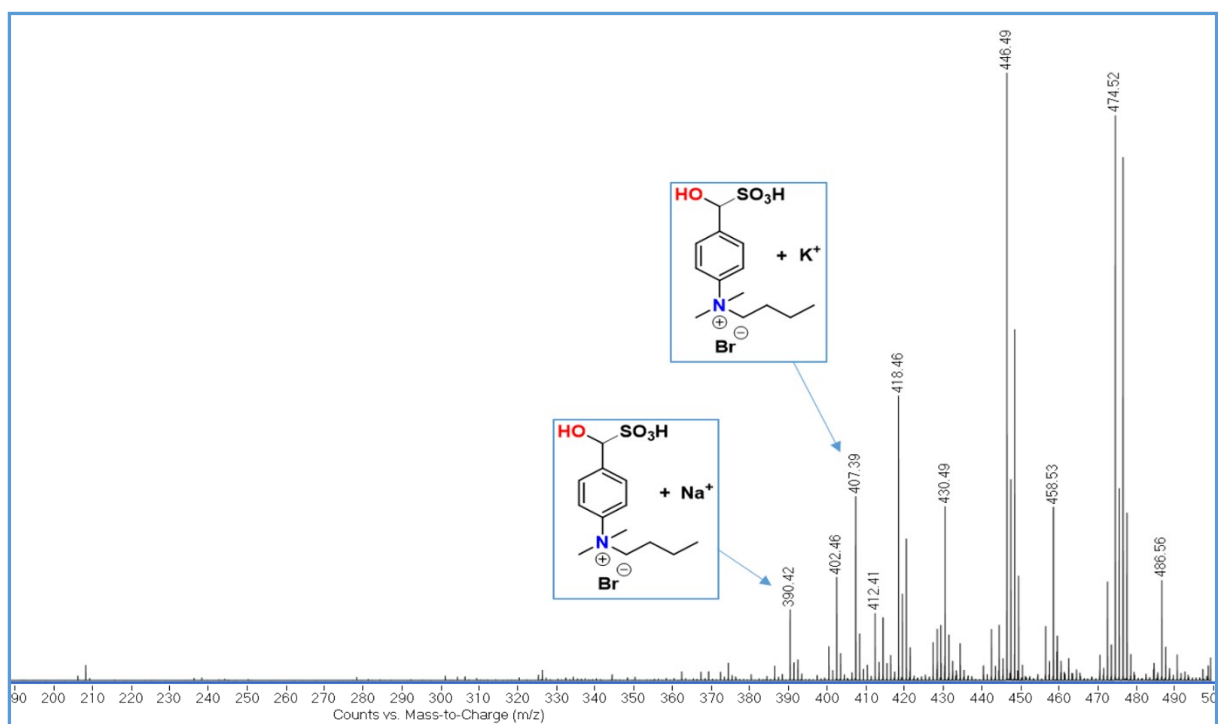
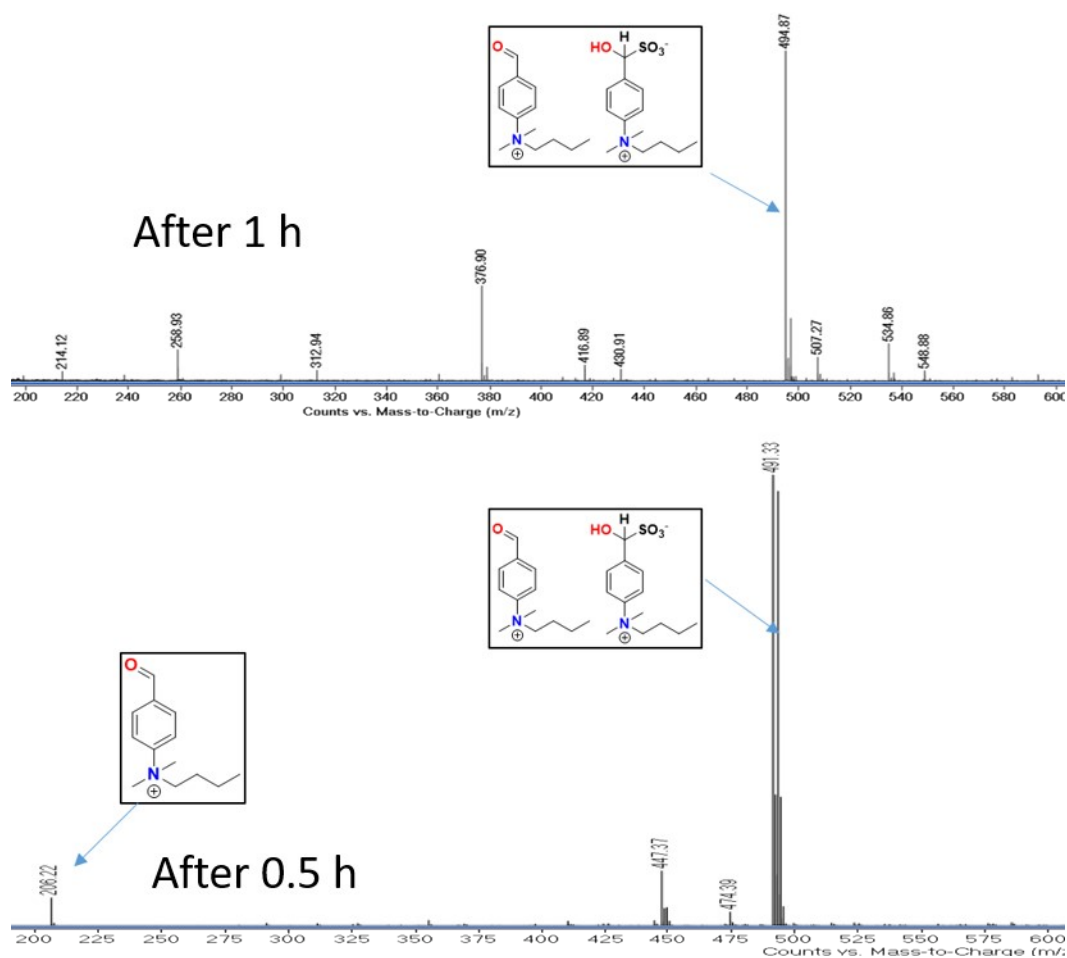


Fig. S11 <sup>1</sup>H NMR spectrum of TSIL-1 with NaHSO<sub>3</sub> in D<sub>2</sub>O.

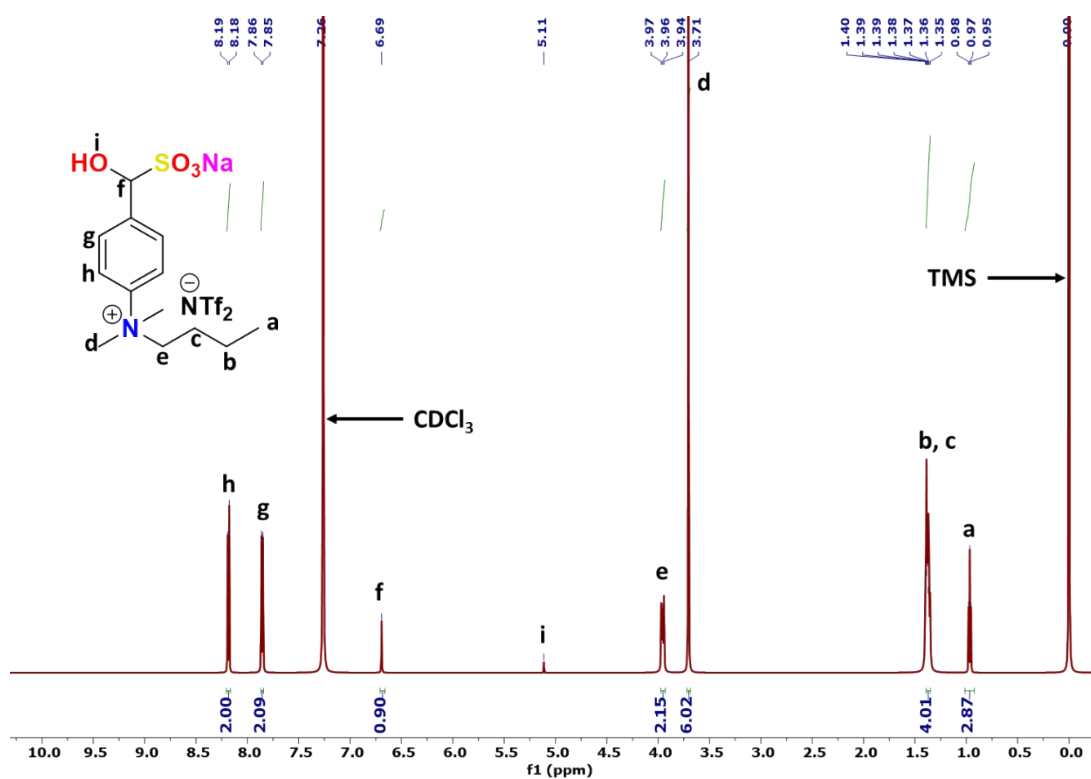


**Fig. S12** ESI-MS of TSIL-1 with NaHSO<sub>3</sub> in D<sub>2</sub>O (+ve Scan).

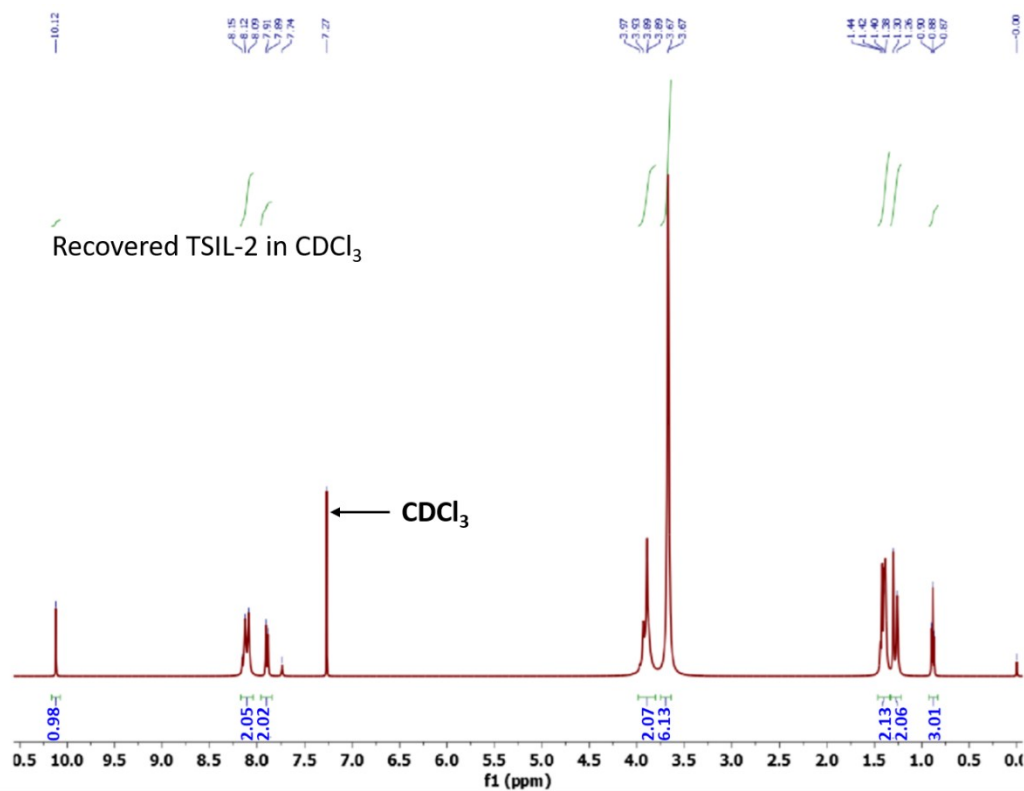


**Fig. S13** The mass spectrum of organic phase (TSIL-2 in ethylacetate) showing formation of bisulfite adduct with time.

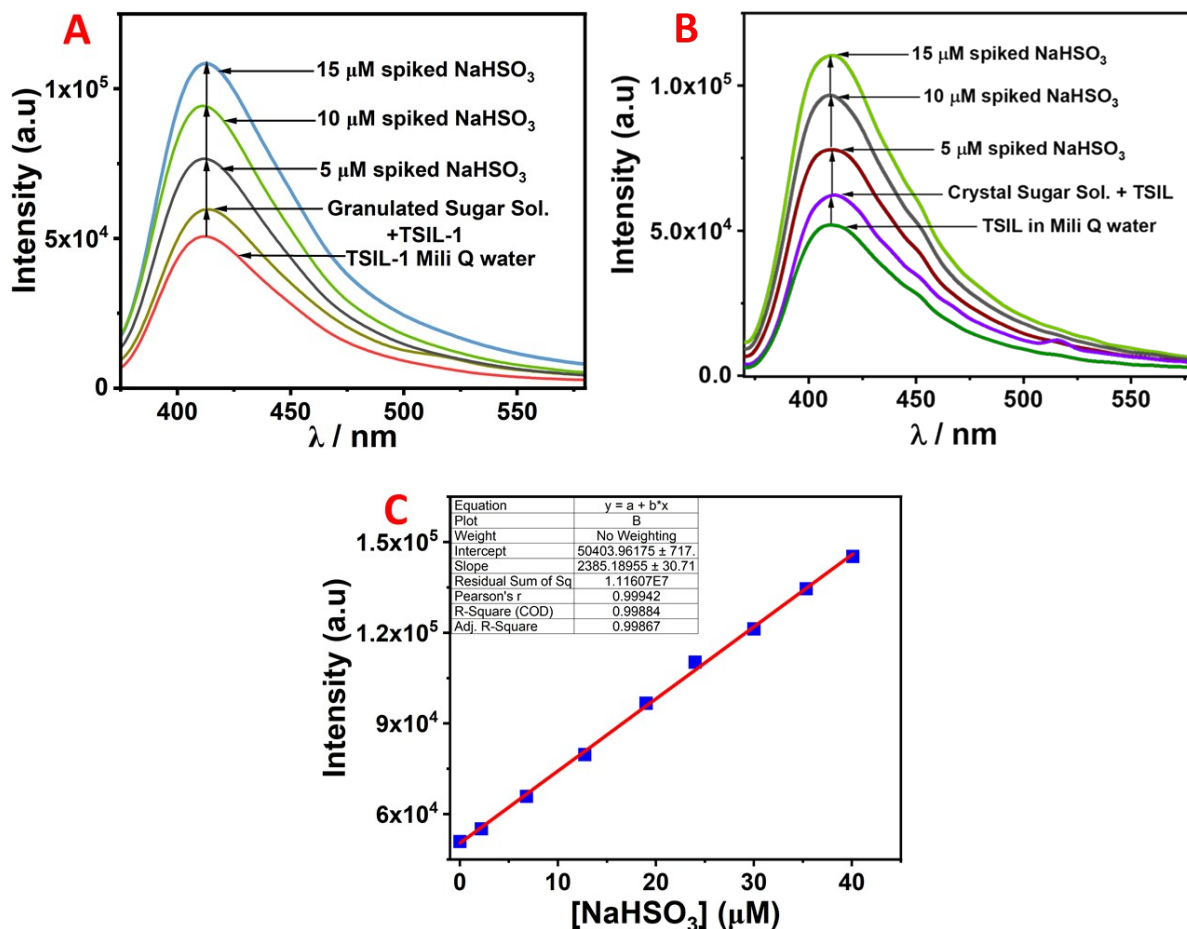




**Fig. S14** The dried organic phase from liquid-liquid extraction experiment showing formation of bisulfite adduct in CDCl<sub>3</sub>.



**Fig. S15** The dried organic phase from the third cycle of recycling experiment from liquid-liquid extraction experiment showing formation of bisulfite adduct in CDCl<sub>3</sub>.



**Fig. S16** Real sample quantification of (A) Granulated sugar and (B) Crystal sugar; (C) Calibration plot.

**Table S1** Performance comparison of the probe with some literature reports.

S.No	Probe	Type of response	Sensing Phase	Removal Studies	L.O.D Value	Ref
1	Functionalized silica	Turn off	Aqueous	yes	64 ppb	[8]
2	Organic Probe	NIR	Aqueous + DMSO	No	24 nM	[10]
3	Benzopyranium Salt	Ratiometric	Aqueous + EtOH	No	0.017 $\mu$ M	[2]
4	Organic Probe	Ratiometric	Aqueous + DMF	No	12.6 nM	[16]
5	Dicyanoisophorone-quinolinium-based	Ratiometric	Aqueous + DMSO	No	2.5 $\mu$ M	[18]

6	Organic Salt	Turn off	Aqueous + DMSO	No	2.1 $\mu$ M	[19]
7	Coumarin-Benzopyran derivative	Turn on	Aqueous	No	177 nM	[20]
8	Triphenylamine-Benzopyrylium based	Turn on	Aqueous	No	12.7 nM	[21]
9	Unsymmetrical azine	Turn off	Aqueous + organic	No	25 nM	[22]
10	Organic probe	Ratiometric	Aqueous + DMSO	No	58 $\mu$ M	[29]
11	Organic Probe	Turn on	Aqueous + Acetonitrile	No	10 $\mu$ M	[30]
<b>12</b>	<b>Ionic liquid (This work)</b>	<b>Turn on</b>	<b>Aqueous</b>	<b>Yes</b>	<b>91 nM</b>	<b>This work</b>