

Supporting information

Tunable Heteroassembly of 2D CoNi LDH and Ti₃C₂ Nanosheets with Enhanced Electrocatalytic Activity for Oxygen Evolution

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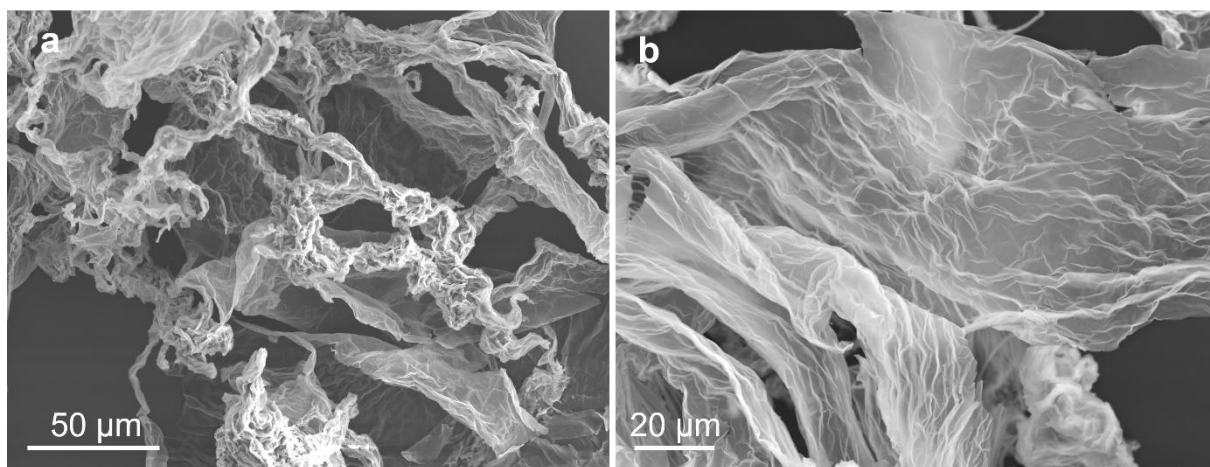


Figure S1. SEM images of CoNi LDH nanosheets after freeze drying.

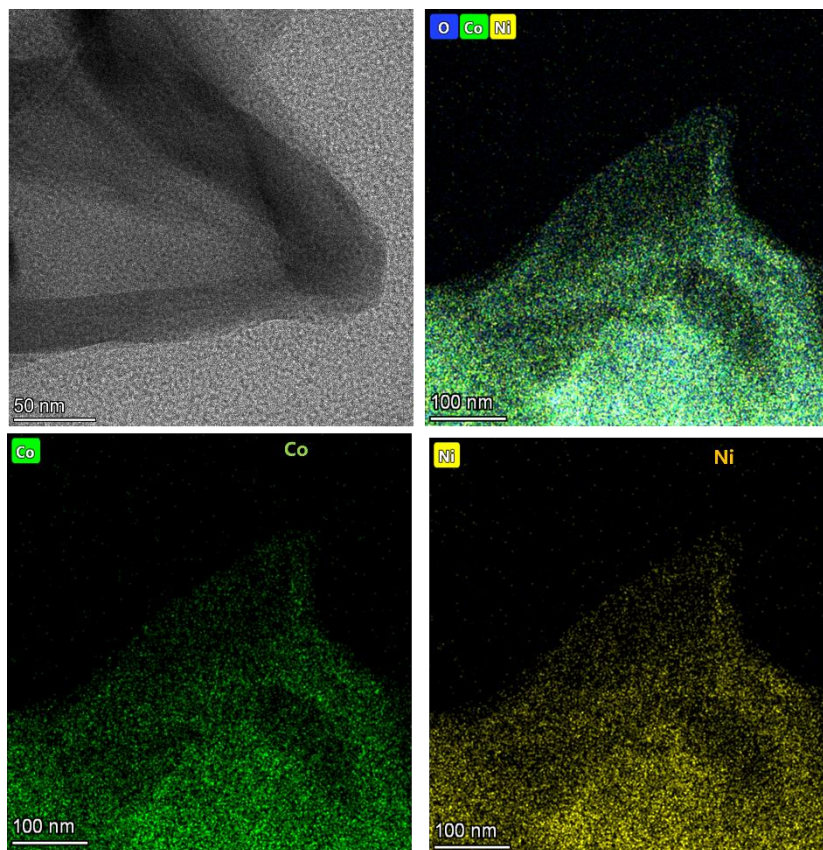


Figure S2. TEM image and elemental mappings of CoNi LDH nanosheets after exfoliation.

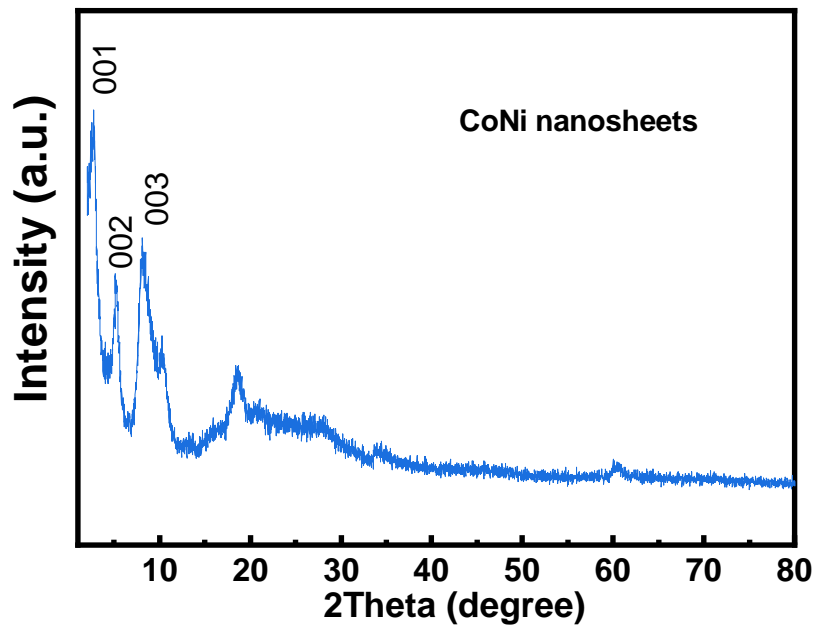


Figure S3. XRD pattern of CoNi LDH nanosheets after freeze drying.

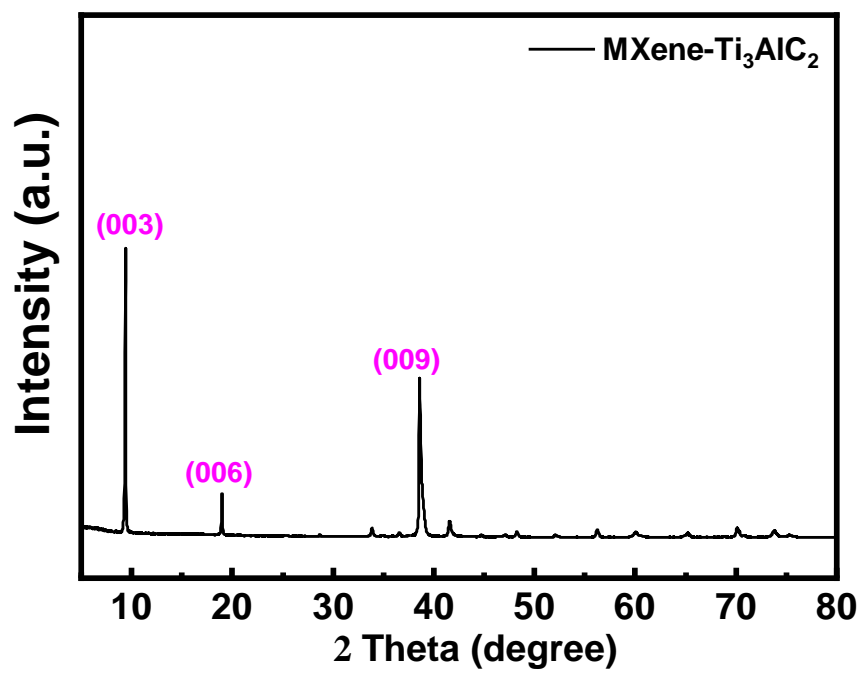


Figure S4. XRD pattern of the multilayered Ti₃AlC₂ MAX.

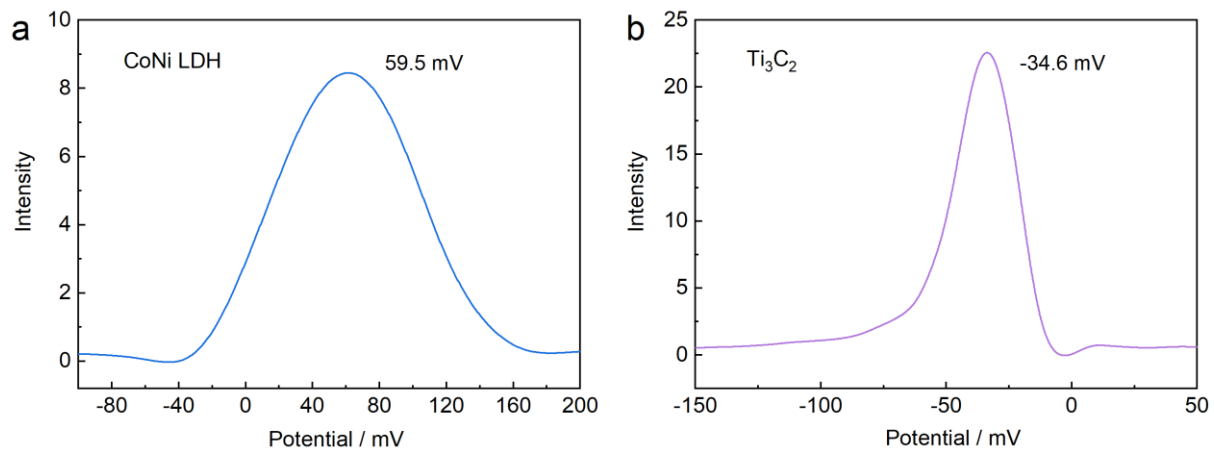
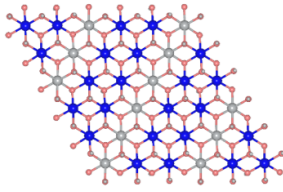


Figure S5. Zeta potential profiles of CoNi LDH and Ti₃C₂ nanosheet dispersion.

CoNi nanosheets



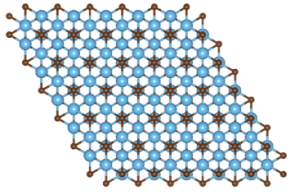
2D weight density:

$$W_{(\text{LDH})} = M_{(\text{LDH})} / (a \times a \times \sin 120^\circ \times N_A)$$

$$\text{unit cell } a = 0.307 \text{ nm} \quad N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$$

$$M_{(\text{LDH})} = \text{Co}_{0.67}\text{Ni}_{0.33}(\text{OH})_{0.8}\text{DS}_{0.2} \cdot 0.76 \text{ H}_2\text{O} = 154.4 \text{ g mol}^{-1}$$

Ti₃C₂ nanosheets



$$W_{(\text{Ti}_3\text{C}_2\text{T}_x)} = M_{(\text{Ti}_3\text{C}_2\text{T}_x)} / (a \times a \times \sin 120^\circ \times N_A)$$

$$a = 0.305 \text{ nm} \quad N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$$

$$M_{(\text{Ti}_3\text{C}_2\text{T}_x)} = 167 \text{ g mol}^{-1}$$

$$m_{(\text{LDH})} / m_{(\text{Ti}_3\text{C}_2\text{T}_x)} = W_{(\text{LDH})} / W_{(\text{Ti}_3\text{C}_2\text{T}_x)} \approx 0.92$$

Figure S6. Crystal structure of monolayer CoNi LDH and Ti₃C₂ nanosheets and the area match model for design of the heterostructures.

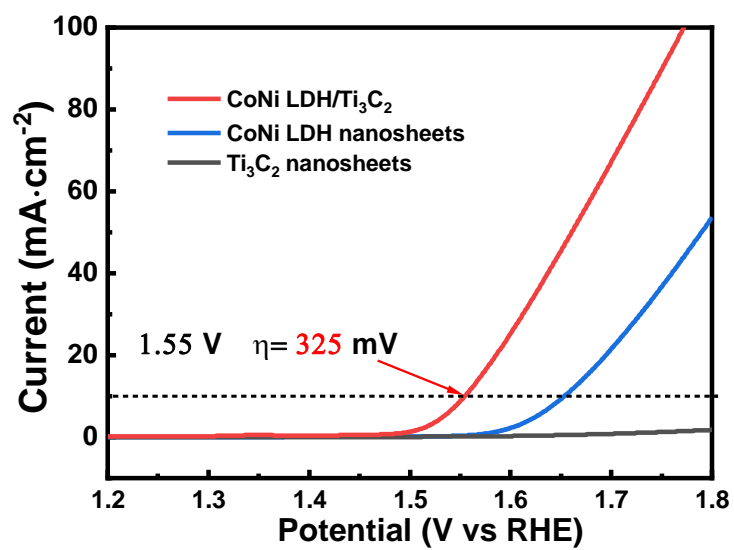


Figure S7. LSV curves for CoNi LDH, Ti₃C₂ nanosheets and CoNi LDH/Ti₃C₂ materials.

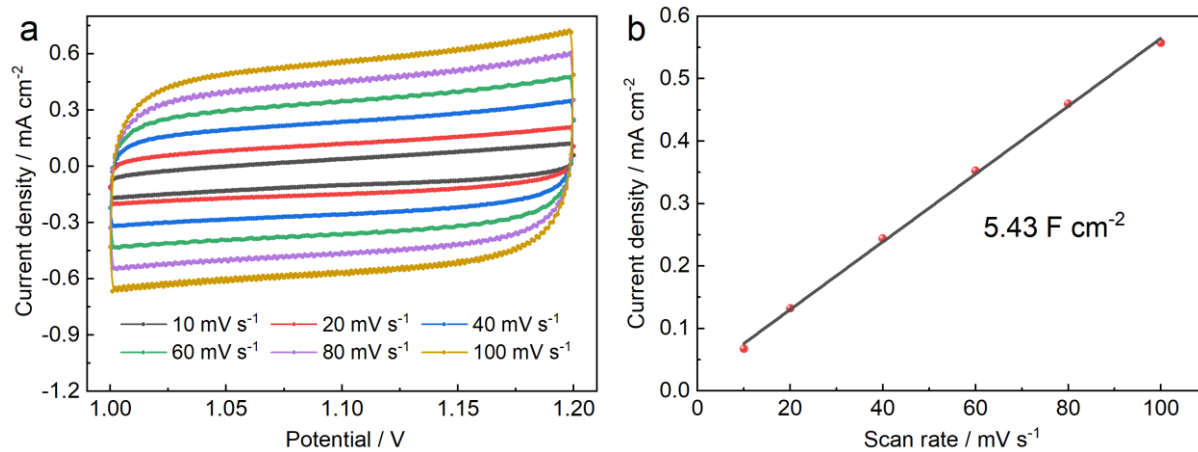


Figure S8. CV profiles of CoNi LDH/Ti₃C₂ at various scan rates and the capacitive currents at 1.1 V plotted as a function of scan rate.

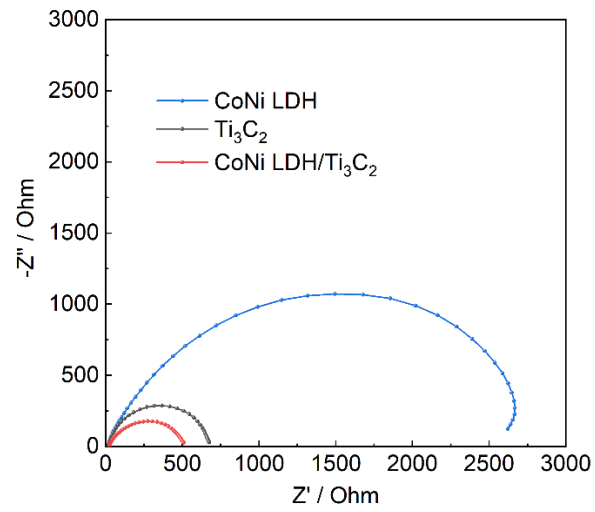


Figure S9. EIS spectra of CoNi LDH, Ti₃C₂ and CoNi LDH/Ti₃C₂ superlattice heterostructures.

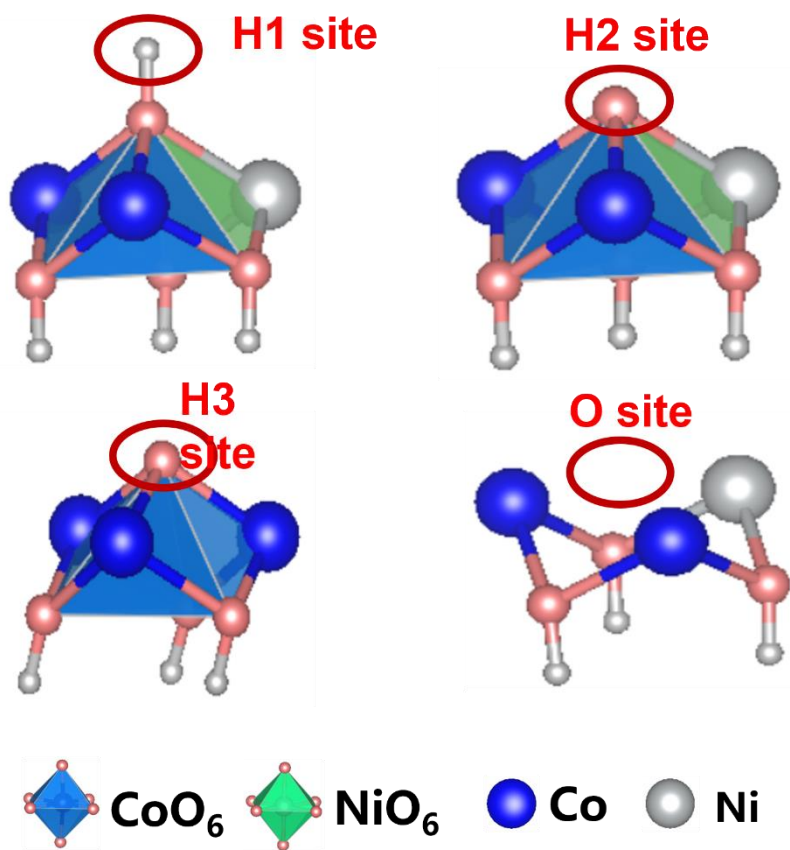


Figure S10. Different active site positions at the CoNi LDH/Ti₃C₂ interface.

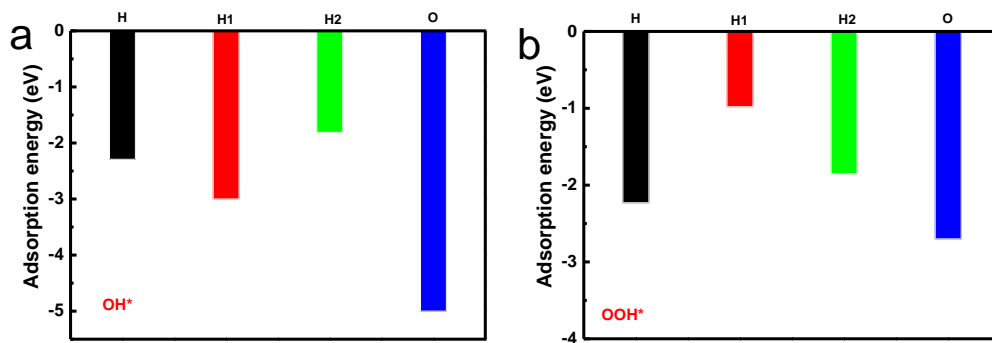


Figure S11. Adsorption energies of different active sites for the intermediates (OH* and OOH) of the OER reaction.

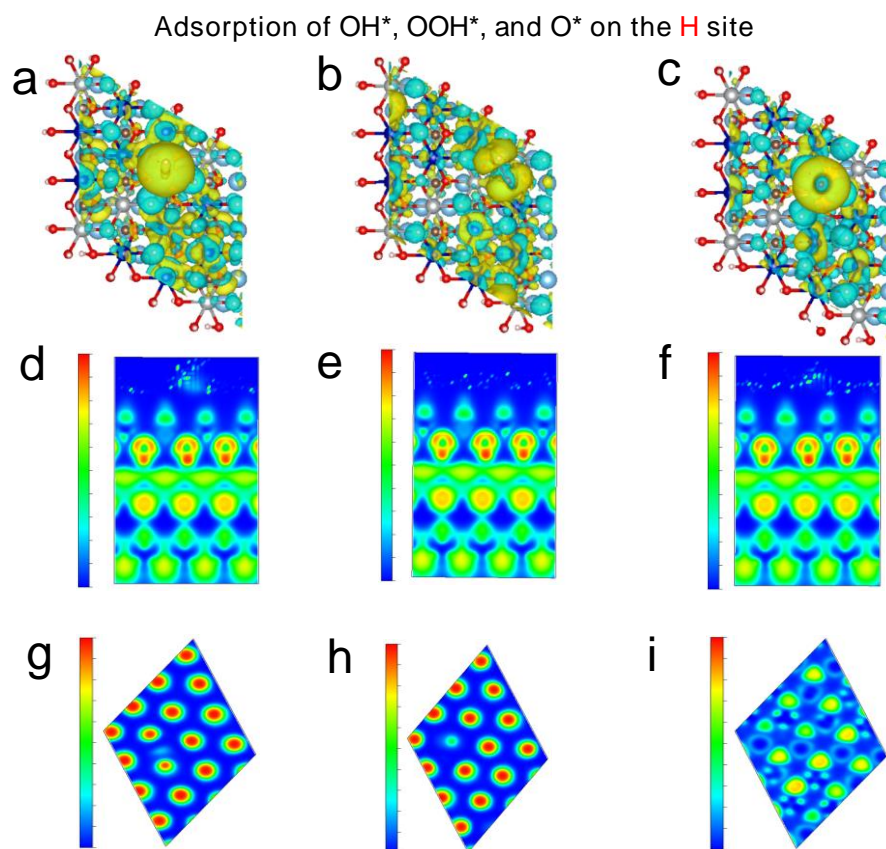


Figure S12. Electron localization function (a-c) and charge density difference (d-i) diagrams for the adsorption of OH* (a, d, f), OOH* (b, e, h), and O* (c, f, i) on the H site.

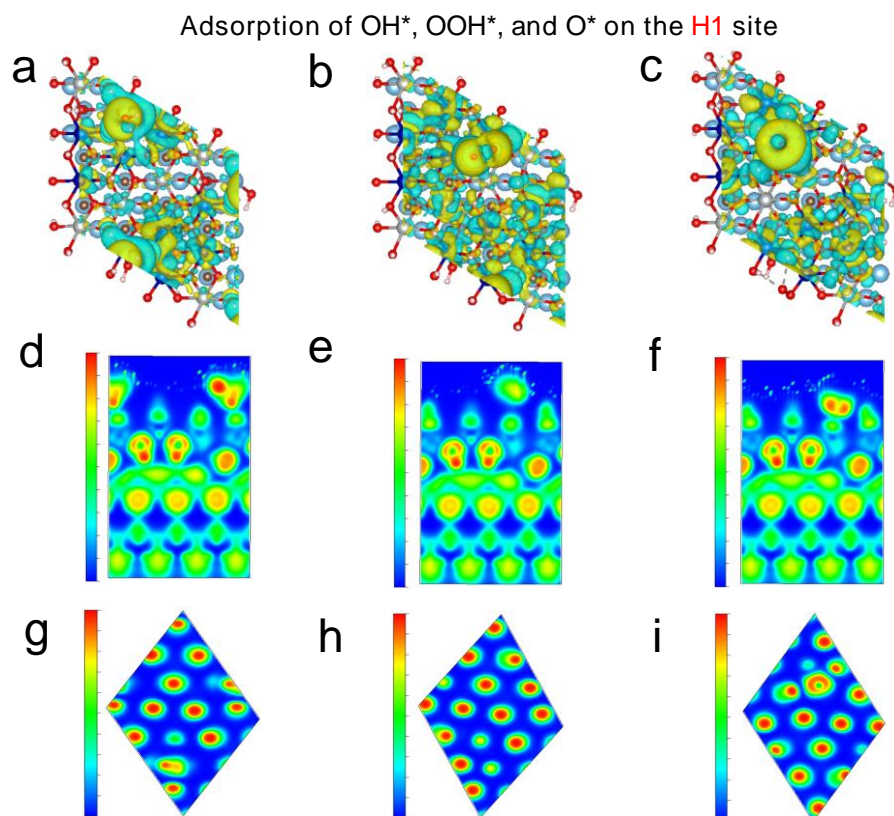


Figure S13. Electron localization function (a-c) and charge density difference (d-i) diagrams for the adsorption of OH* (a, d, f), OOH* (b, e, h), and O* (c, f, i) on the H1 site.

Adsorption of OH*, OOH*, and O* on the H2 site

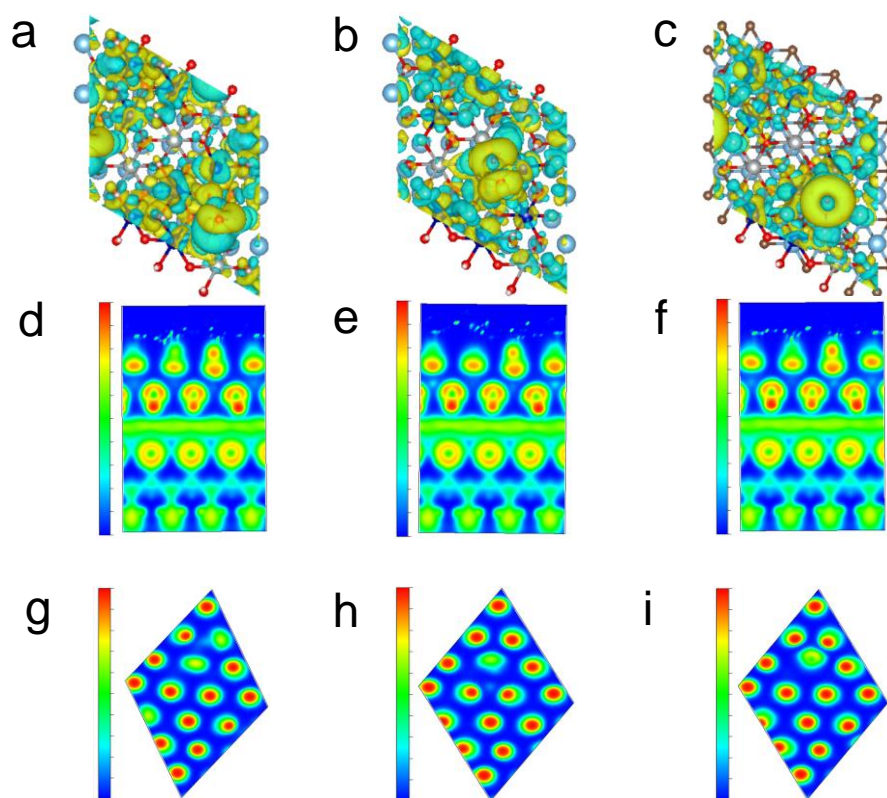


Figure S14. Electron localization function (a-c) and charge density difference (d-i) diagrams for the adsorption of OH* (a, d, f), OOH* (b, e, h), and O* (c, f, i) on the H2 site.

Adsorption of OH*, OOH*, and O* on the O site

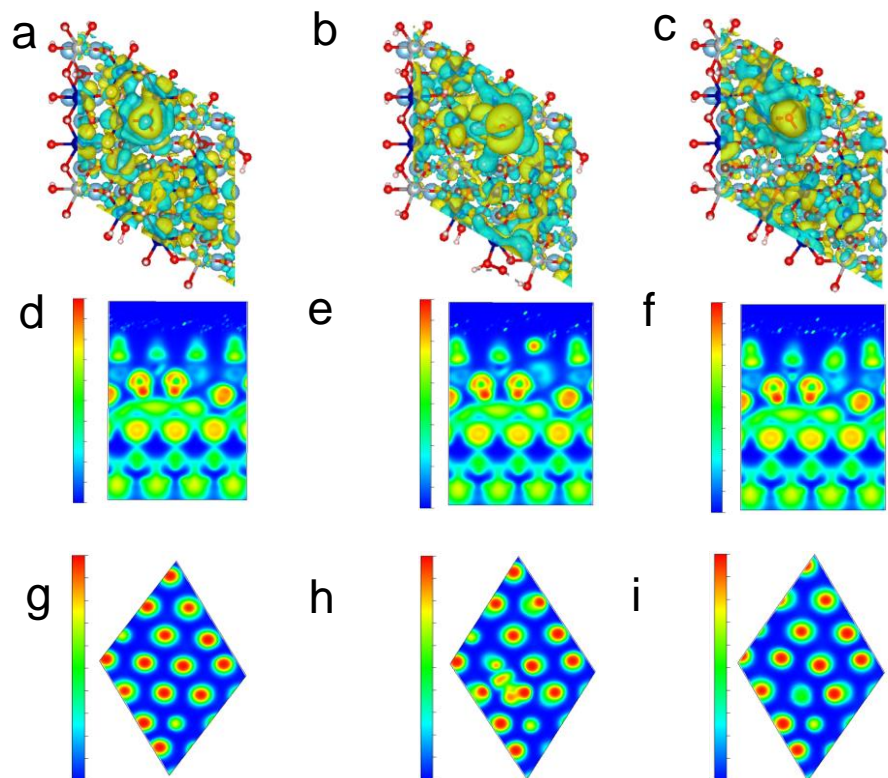


Figure S15. Electron localization function (a-c) and charge density difference (d-i) diagrams for the adsorption of OH* (a, d, f), OOH* (b, e, h), and O* (c, f, i) on the O site.