

## **Vanadium-regulated nickel phosphide nanosheets for electrocatalytic sulfion upgrading and hydrogen production**

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## **Experimental section**

### **Chemicals**

$\text{Ni}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ,  $\text{VCl}_3$ ,  $\text{CO}(\text{NH}_2)_2$ ,  $\text{NH}_4\text{F}$  and  $\text{NaH}_2\text{PO}_4 \cdot \text{H}_2\text{O}$  were bought from Aladdin and  $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$  was supplied Sinopharm group.  $\text{NaOH}$  and nickel foam (NF) were supplied by Adamas-beta<sup>®</sup> and Taiyuan source of power company.

### **Synthesis of V-Ni<sub>2</sub>P**

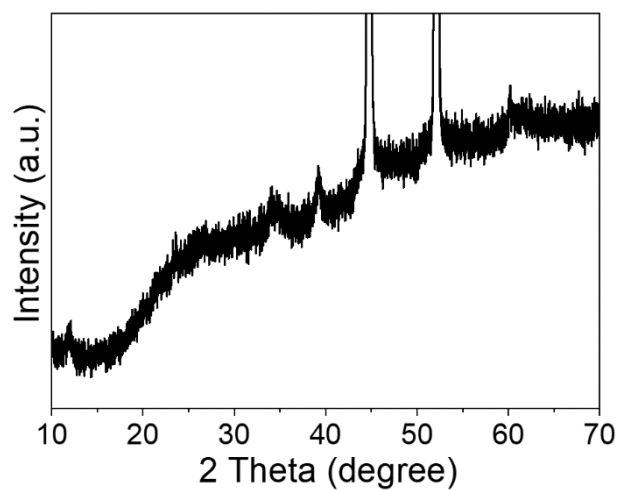
In a typical synthesis, 1.6 mmol  $\text{Ni}(\text{NO}_3)_2$ , 0.4 mmol  $\text{VCl}_3$ , 6 mmol  $\text{NH}_4\text{F}$ , 10 mmol  $\text{CO}(\text{NH}_2)_2$  were dissolved in 35 mL water with stirring, which was transferred to autoclave with the NF put into in the solution. The autoclave was and kept at 120 °C for 6 h in an electric oven to prepare Ni-based layered double hydroxide (NiV LDH) precursor. Then, the NiV LDH precursor and  $\text{NaH}_2\text{PO}_4 \cdot \text{H}_2\text{O}$  were placed in furnace and heated up to 300 °C for 2h to obtain V-Ni<sub>2</sub>P. In contrast, the Ni<sub>2</sub>P, V-Ni<sub>2</sub>P-1 and V-Ni<sub>2</sub>P-2 were synthesized in a similar process without adding  $\text{VCl}_3$  and with adding various molar ratios of Ni and V (9:1, 7:3), respectively.

### **Materials characterization**

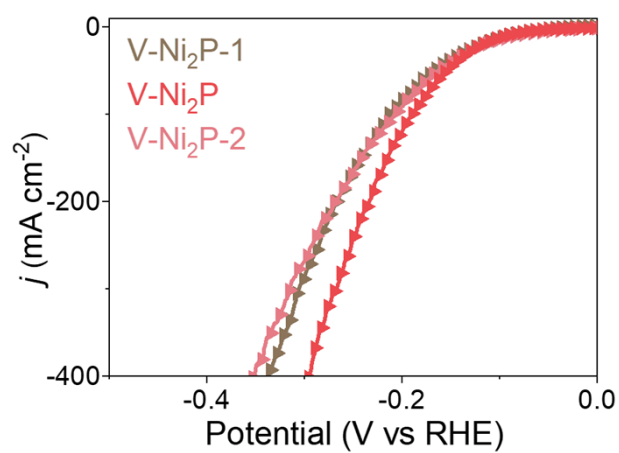
The crystalline structures, morphology, element distribution and chemical states of the samples were examined on the X-ray diffraction (XRD, Bruker D8), scanning electron microscopy (SEM, FIB-SEM GX4) equipped with energy dispersive spectroscopy, transmission electron microscopy (TEM, FEI Tecnai G20) and X-ray photoelectron spectrometer (XPS, Thermo ESCALAB 250Xi).

### **Electrochemical measurements**

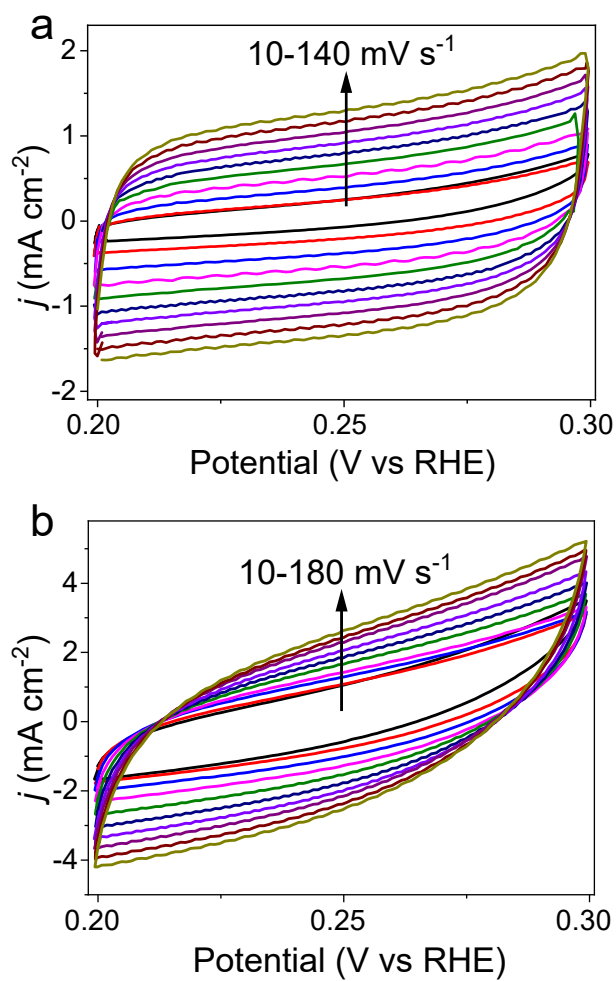
The electrochemical measurements of the samples were carried out on a CHI 760E electrochemical workstation using synthesized samples as working electrodes in a typical three-electrode system to test the electrochemical performances of HER, OER and SOR. The polarization curves were measured at a scan of 2 mV s<sup>-1</sup> and the electrochemical impedance spectroscopy was performed in the frequency range from 0.01 Hz to 100 kHz.



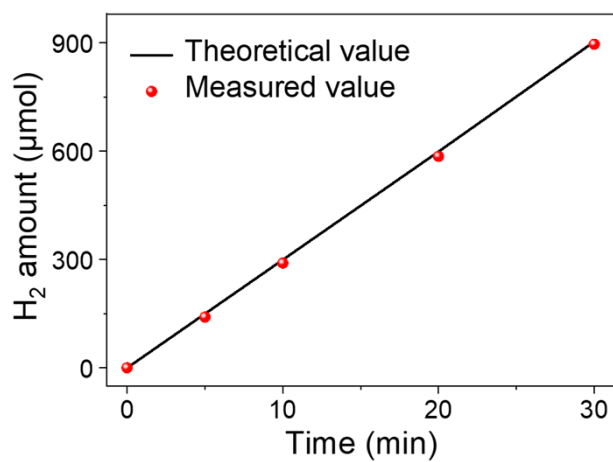
**Fig. S1** XRD pattern of NiV LDH precursor.



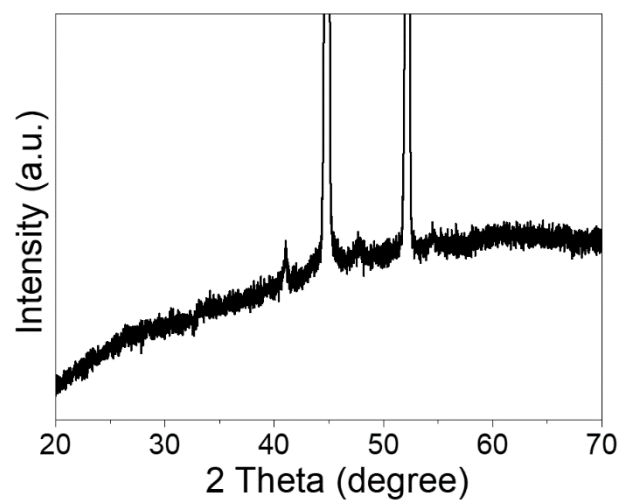
**Fig. S2** LSV curves of V-Ni<sub>2</sub>P-1, V-Ni<sub>2</sub>P and V-Ni<sub>2</sub>P-2 in 1.0 M KOH.



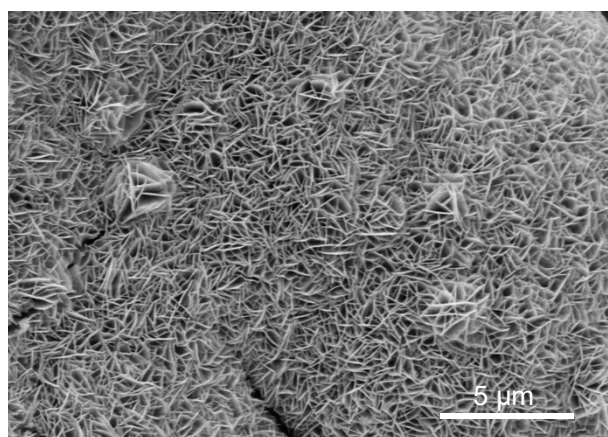
**Fig. S3** Cyclic voltammograms of (a) Ni<sub>2</sub>P and (b) V-Ni<sub>2</sub>P.



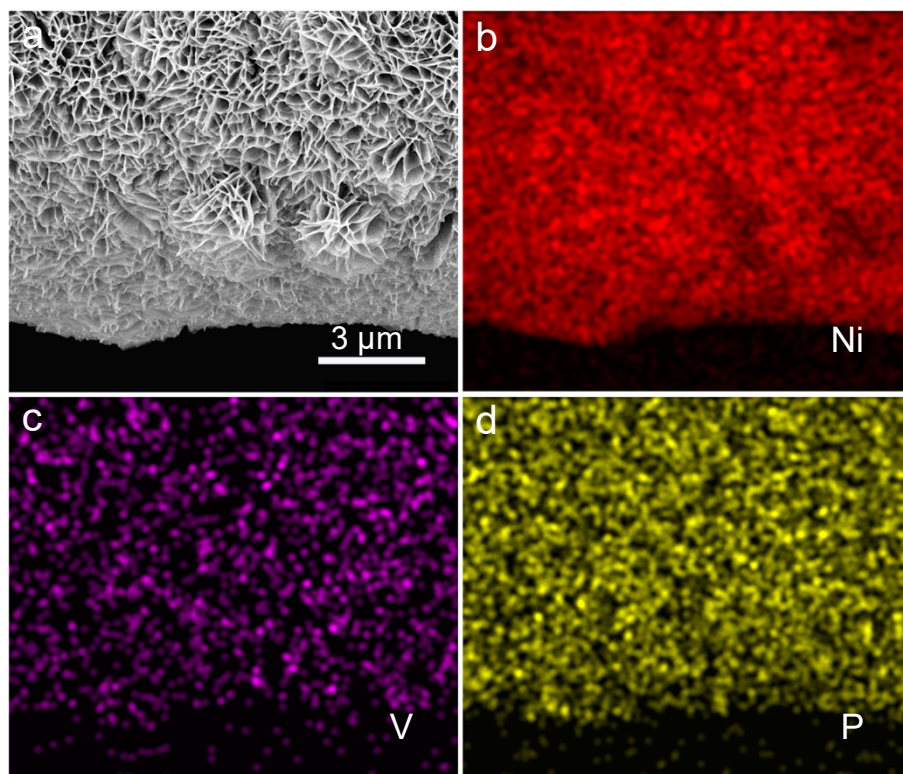
**Fig. S4** Experimental and theoretical calculated amount of H<sub>2</sub> over the V-Ni<sub>2</sub>P.



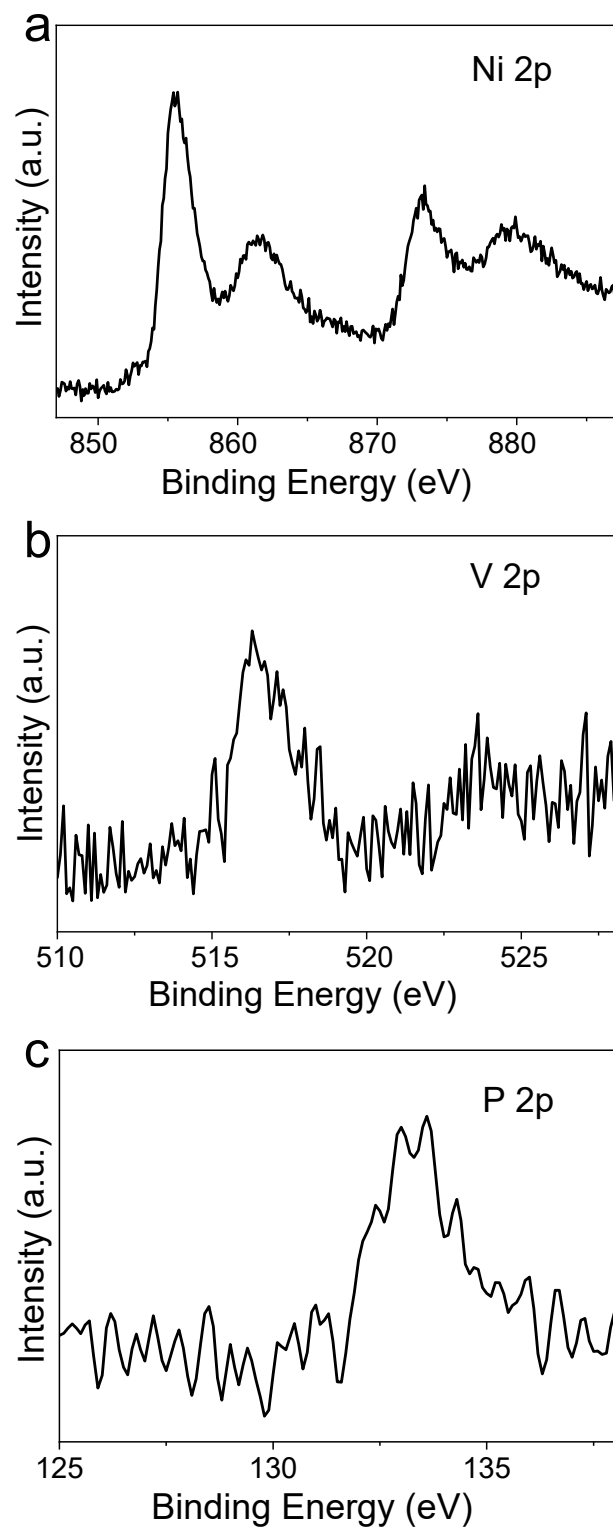
**Fig. S5** XRD pattern of V-Ni<sub>2</sub>P after HER test.



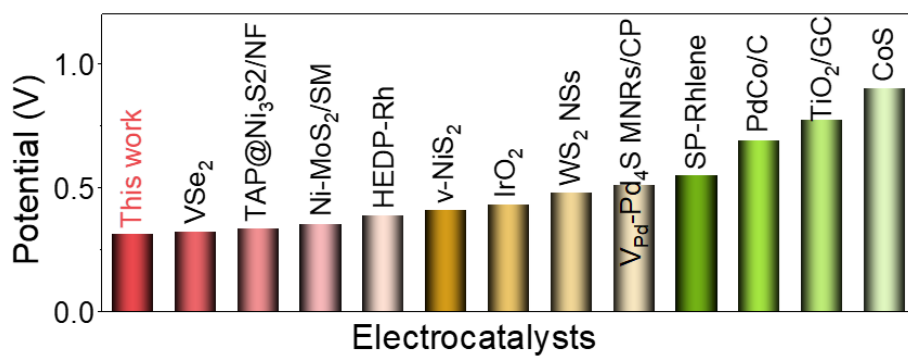
**Fig. S6** SEM image of V-Ni<sub>2</sub>P after HER test.



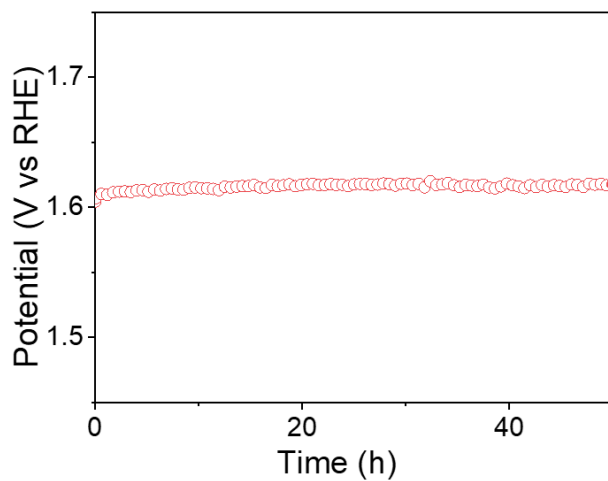
**Fig. S7** SEM and corresponding element mapping images of V-Ni<sub>2</sub>P after HER test.



**Fig. S8** XPS spectra of (a) Ni 2p, (b) V 2p and (c) P 2p of V-Ni<sub>2</sub>P after HER test.

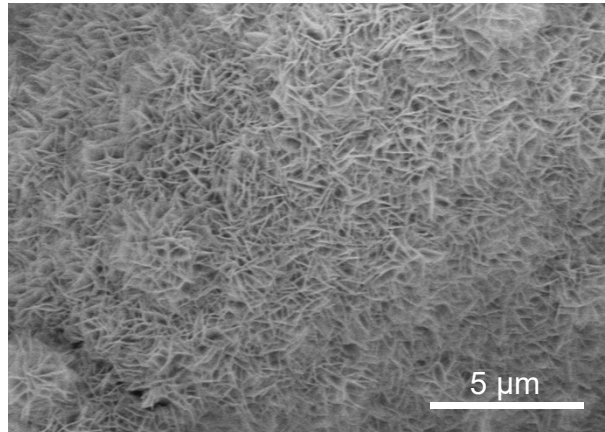


**Fig. S9** Potentials comparison of V-Ni<sub>2</sub>P with reported catalysts.

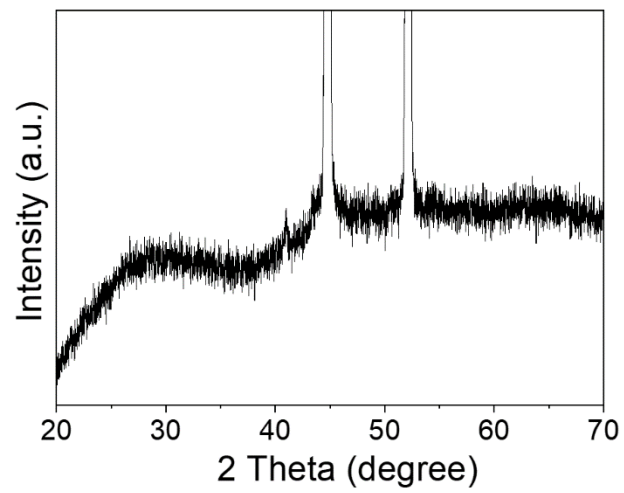


**Fig. S10** Durability test of V-Ni<sub>2</sub>P for OER.

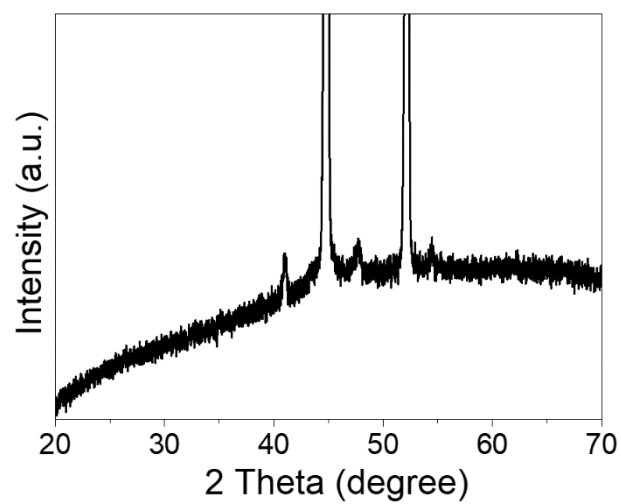




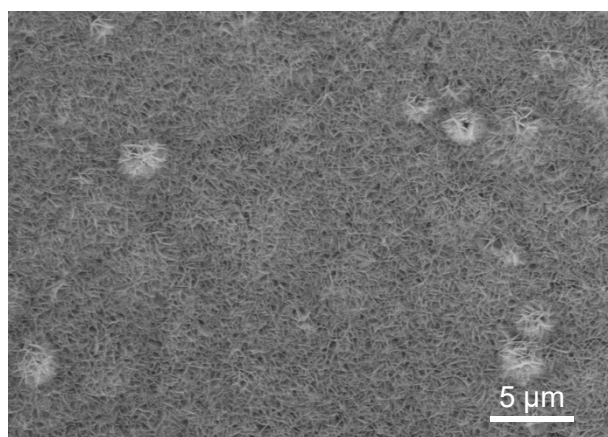
**Fig. S11** SEM image of V-Ni<sub>2</sub>P after OER test.



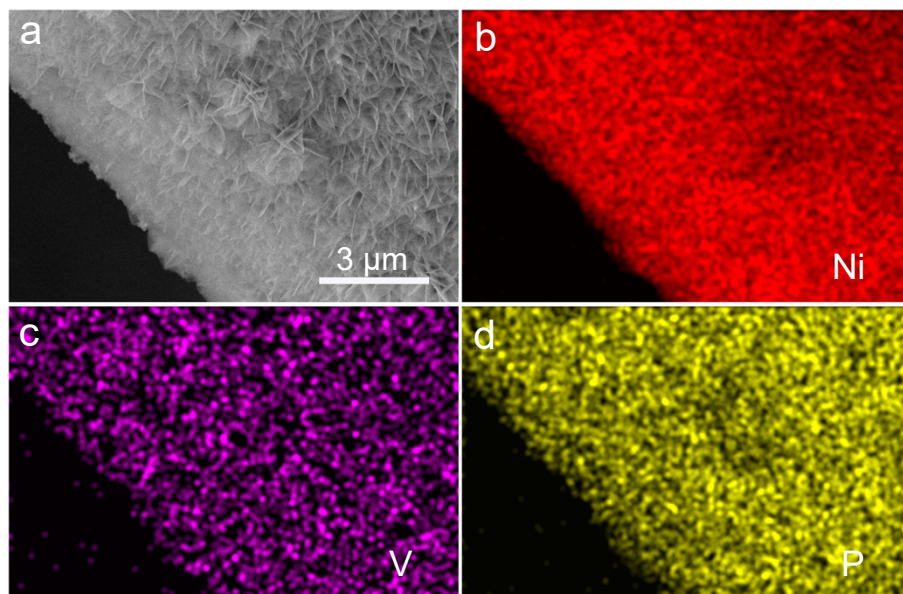
**Fig. S12** XRD pattern of V-Ni<sub>2</sub>P after OER test.



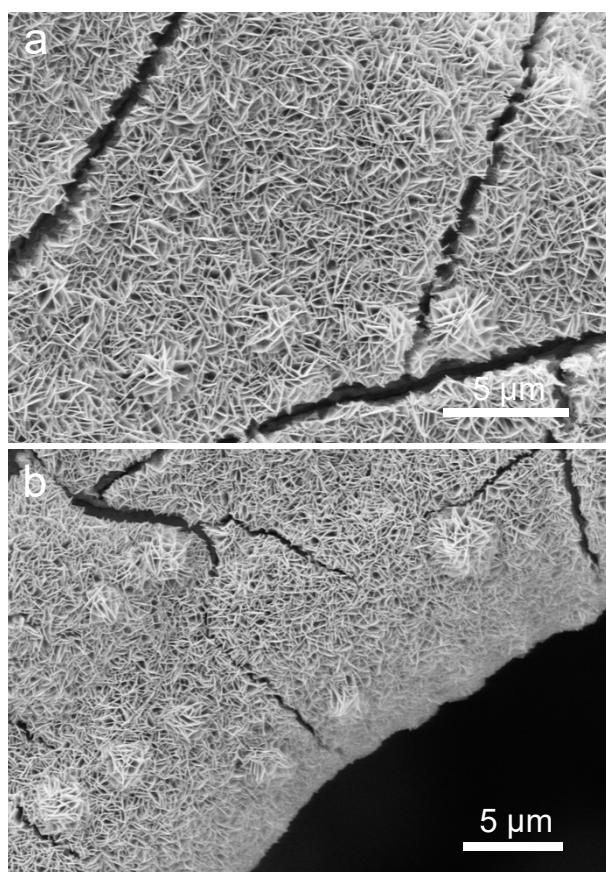
**Fig. S13** XRD pattern of V-Ni<sub>2</sub>P after SOR test.



**Fig. S14** SEM image of V-Ni<sub>2</sub>P after SOR test.



**Fig. S15** SEM and corresponding element mapping images of V-Ni<sub>2</sub>P after SOR test.



**Fig. S16** SEM image of V-Ni<sub>2</sub>P at the cathode (a) and at the anode (b) after the HWE test.

**Table S1** The comparison of HER performance of V-Ni<sub>2</sub>P with reported catalysts.

Catalysts	Overpotential at 10 mA cm <sup>-2</sup> (mV)	Reference
<b>V-Ni<sub>2</sub>P</b>	<b>93</b>	<b>This work</b>
Ni <sub>2</sub> P/CoN-PCP	94	S1
Ni-FeP/C	95	S2
Mo,Fe-NiCoP <sub>x</sub> /NF	99	S3
Ni-N <sub>3</sub>	103	S4
O <sub>3</sub> -V <sub>10</sub> -Ni <sub>2</sub> P	108	S5
Fe(OH) <sub>x</sub> @Cu-MOF	112	S6
NiFeP/SG	115	S7
Ti <sub>3</sub> CN(OH) <sub>x</sub> @MoS <sub>2</sub>	120	S8
Al-Ni <sub>2</sub> P/TM	129	S9
Cu@Co-CoO/Rh	137.7	S10
MoS <sub>2</sub> /CeO <sub>2</sub>	147	S11
1T-MoS <sub>2</sub>	158	S12
CoP	173	S13
Ni-Mo <sub>2</sub> C/NPC	183	S14
MoS <sub>2</sub> @FePS <sub>3</sub>	196	S15
WS <sub>2</sub> -NSs	214	S16
Cu-CoFe/Co/NC	217	S17
H-TaS <sub>2</sub>	230	S18
MoS <sub>2</sub>	248	S19
Ni <sub>1.5</sub> Fe <sub>0.5</sub> P	282	S20

**Table S2** The comparison of SOR performance of V-Ni<sub>2</sub>P with reported catalysts.

<b>Catalysts</b>	<b>Potential (V) at 10 mA cm<sup>-2</sup></b>	<b>Reference</b>
<b>V-Ni<sub>2</sub>P</b>	<b>0.313</b>	<b>This work</b>
VSe <sub>2</sub>	0.32	S21
TAP@Ni <sub>3</sub> S <sub>2</sub> /NF	0.335	S22
Ni-MoS <sub>2</sub> /SM	0.35	S23
HEDP-Rh	0.385	S24
v-NiS <sub>2</sub>	0.41	S25
IrO <sub>2</sub>	0.43	S26
WS <sub>2</sub> NSs	0.48	S27
V <sub>Pd</sub> -Pd <sub>4</sub> S MNRs/CP	0.511	S28
SP-Rhlene	0.550	S29
PdCo/C	0.69	S30
TiO <sub>2</sub> /GC	0.77	S31
CoS	0.9	S32

**Table S3** The comparison of HWE performance of V-Ni<sub>2</sub>P with reported catalysts.

Catalysts	Voltage at 10 mA cm <sup>-2</sup> (V)	Reference
V-Ni <sub>2</sub> P	0.389	This work
Cu <sub>2</sub> S/NF	0.43	S33
a-RhS <sub>2-x</sub>	0.442	S34
v-NiS <sub>2</sub>	0.49	S25
Ni-MoS <sub>2</sub> /SM	0.49	S23
CoS/MoS <sub>2</sub>	0.528	S35
HEDP-Rh	0.585	S24
WS <sub>2</sub> NSs	1.17	S27
CoP/NF	1.27	S36

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