

The search for rigid, tough polyesters with high T_g - Renewable aromatic polyesters with high isosorbide content

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Supporting information

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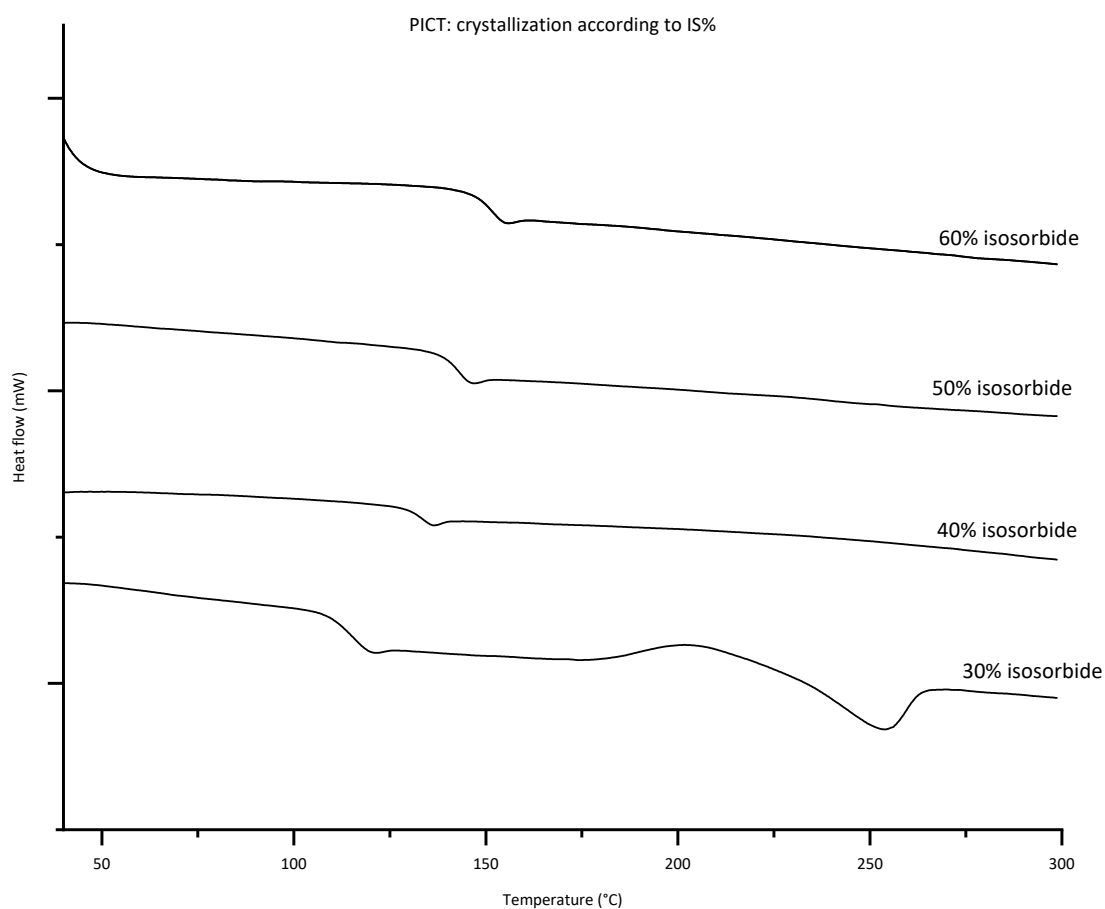
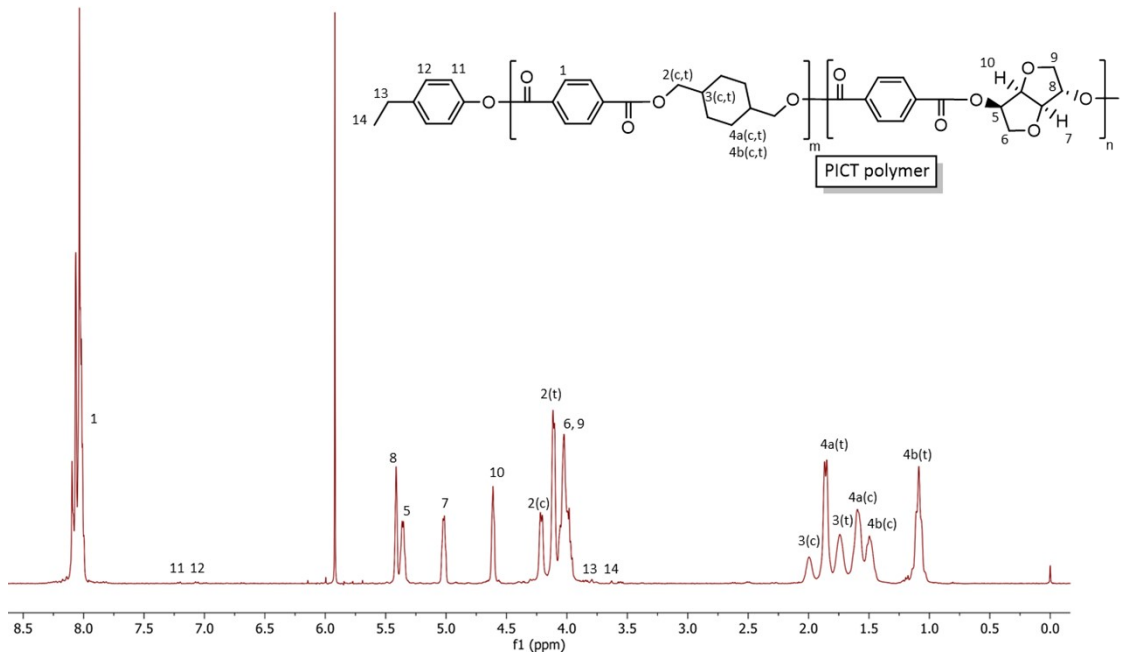
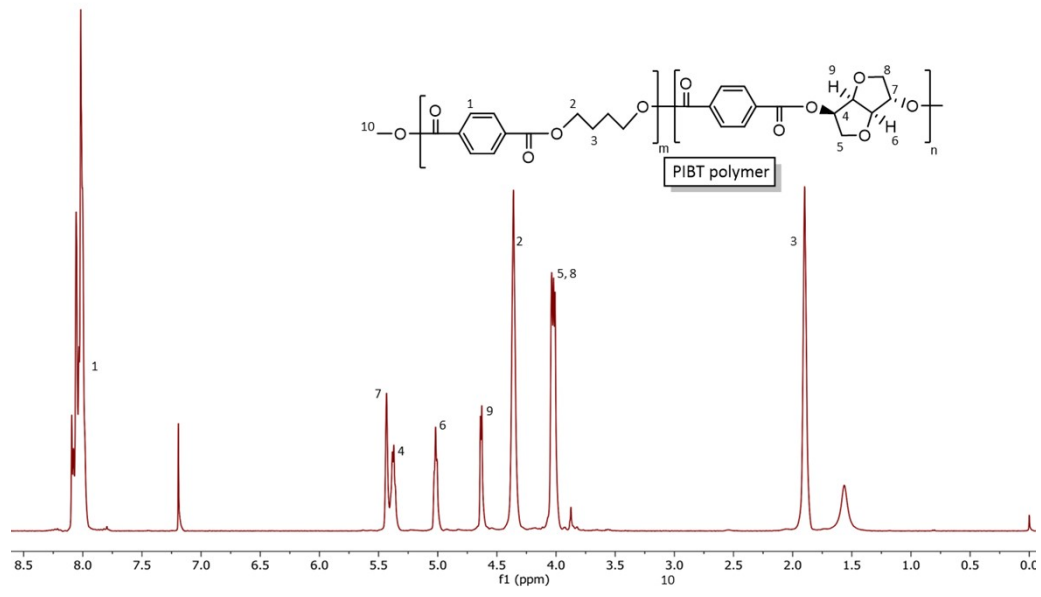
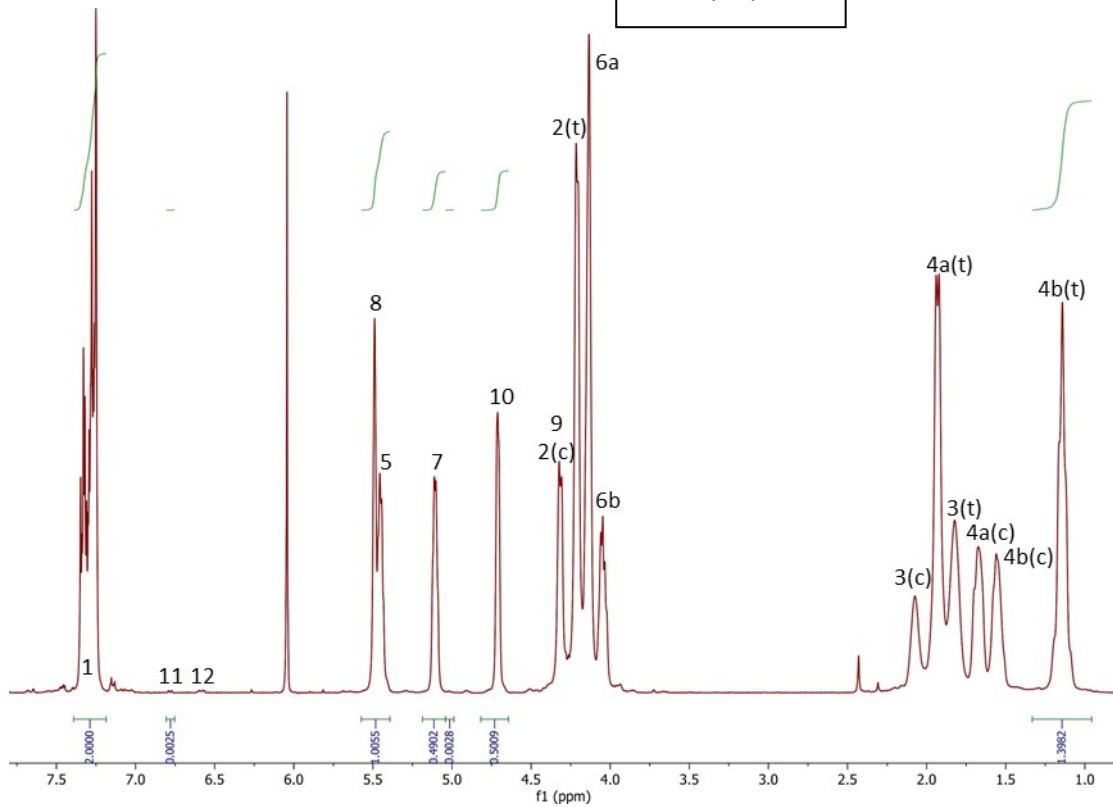
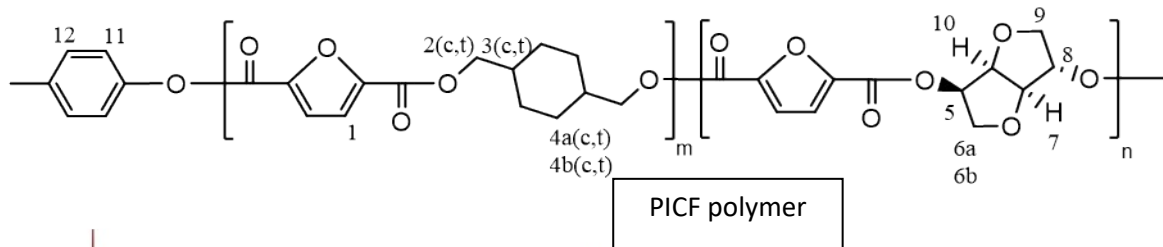
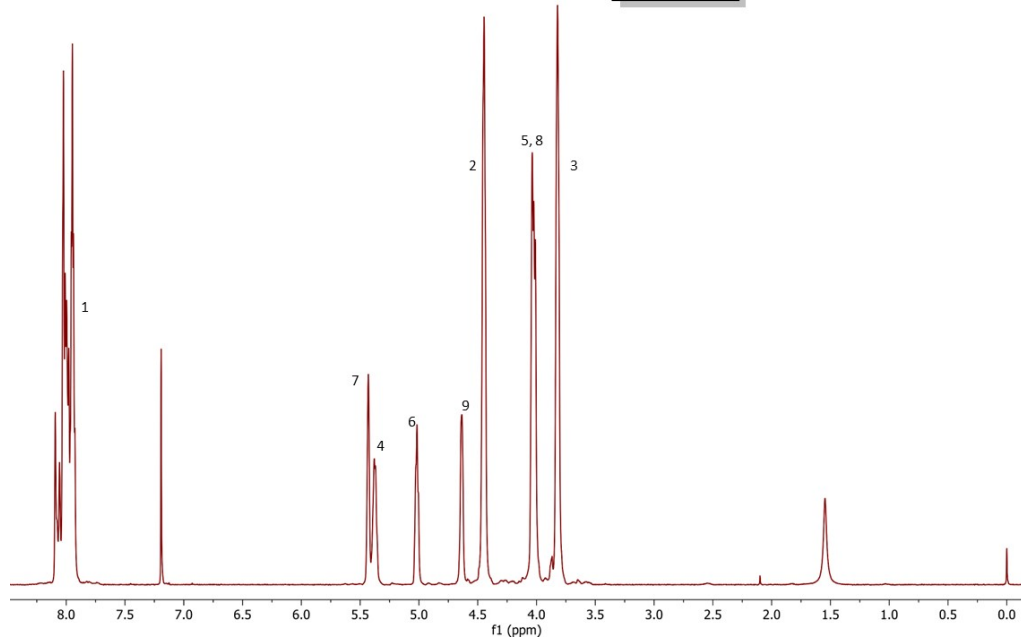
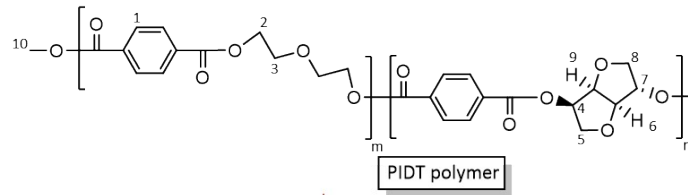


Figure 1 - DSC graphs of PICT with different isosorbide composition.



m n





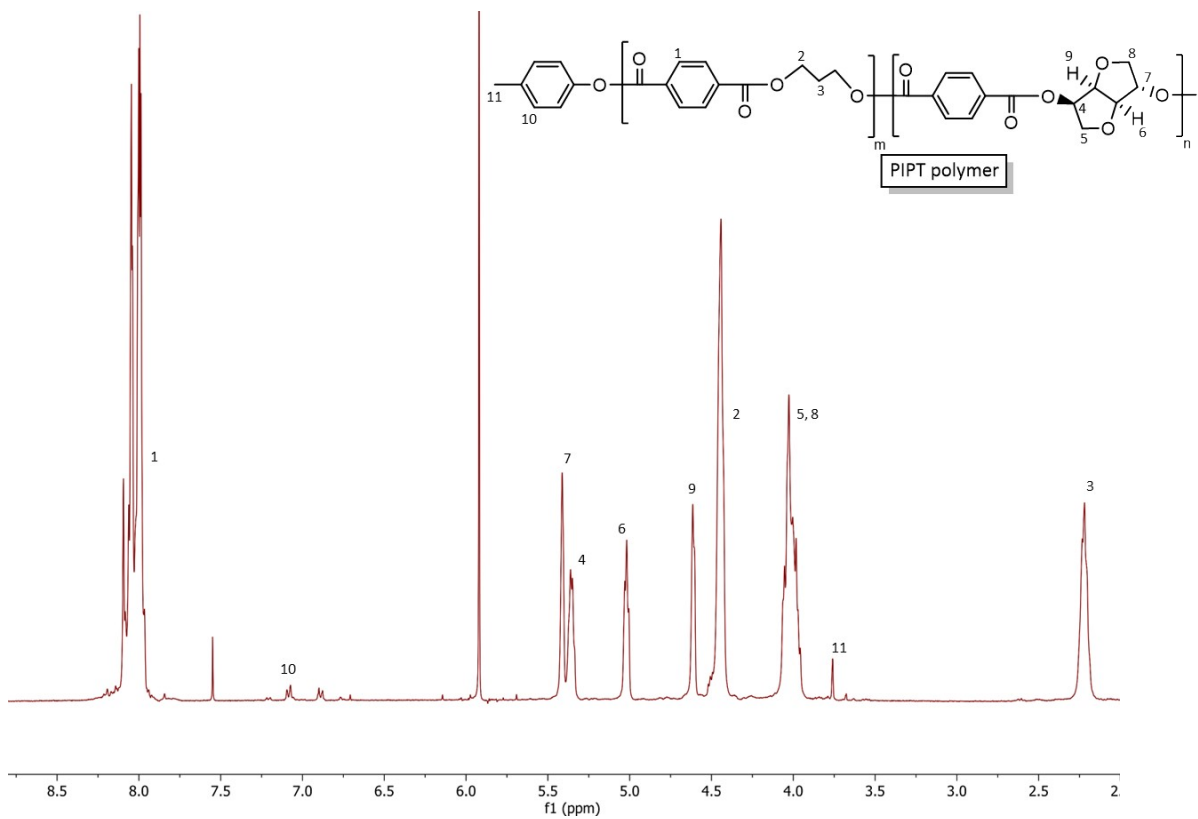
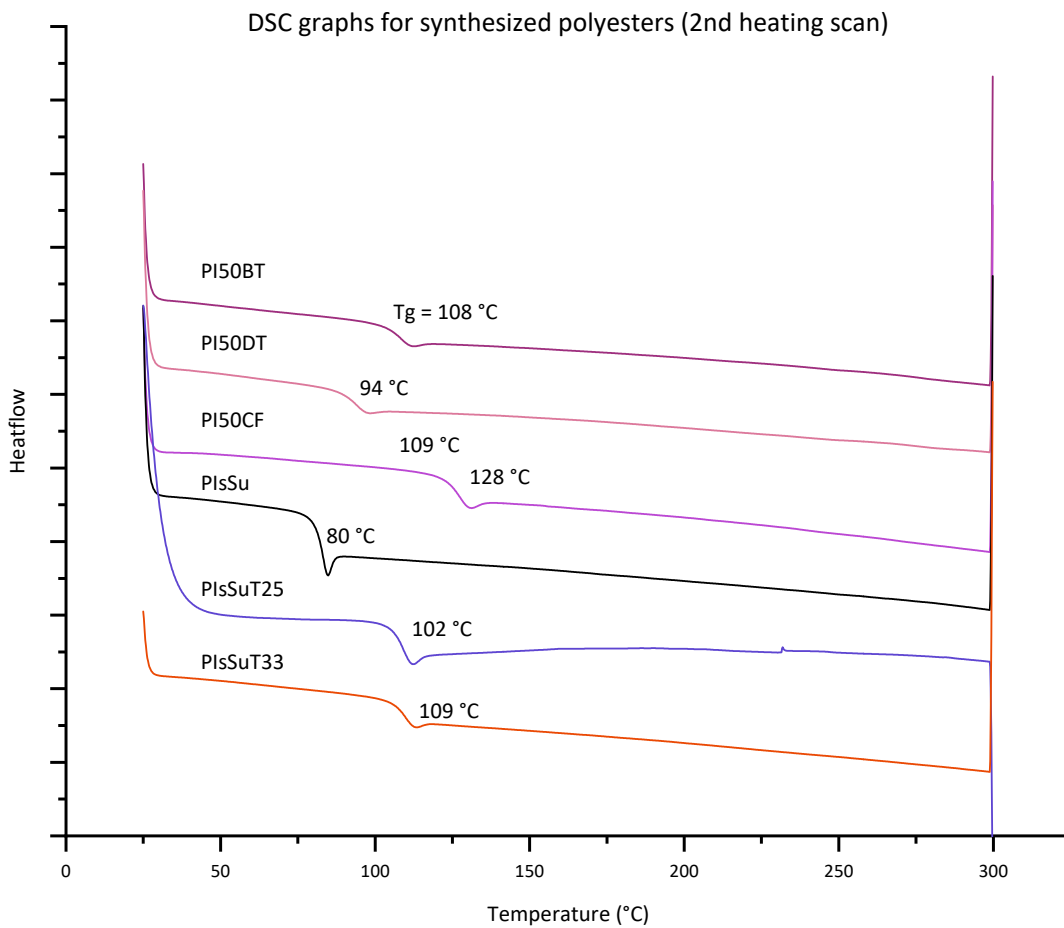


Figure 2 - Proton NMR of synthesized polyesters.



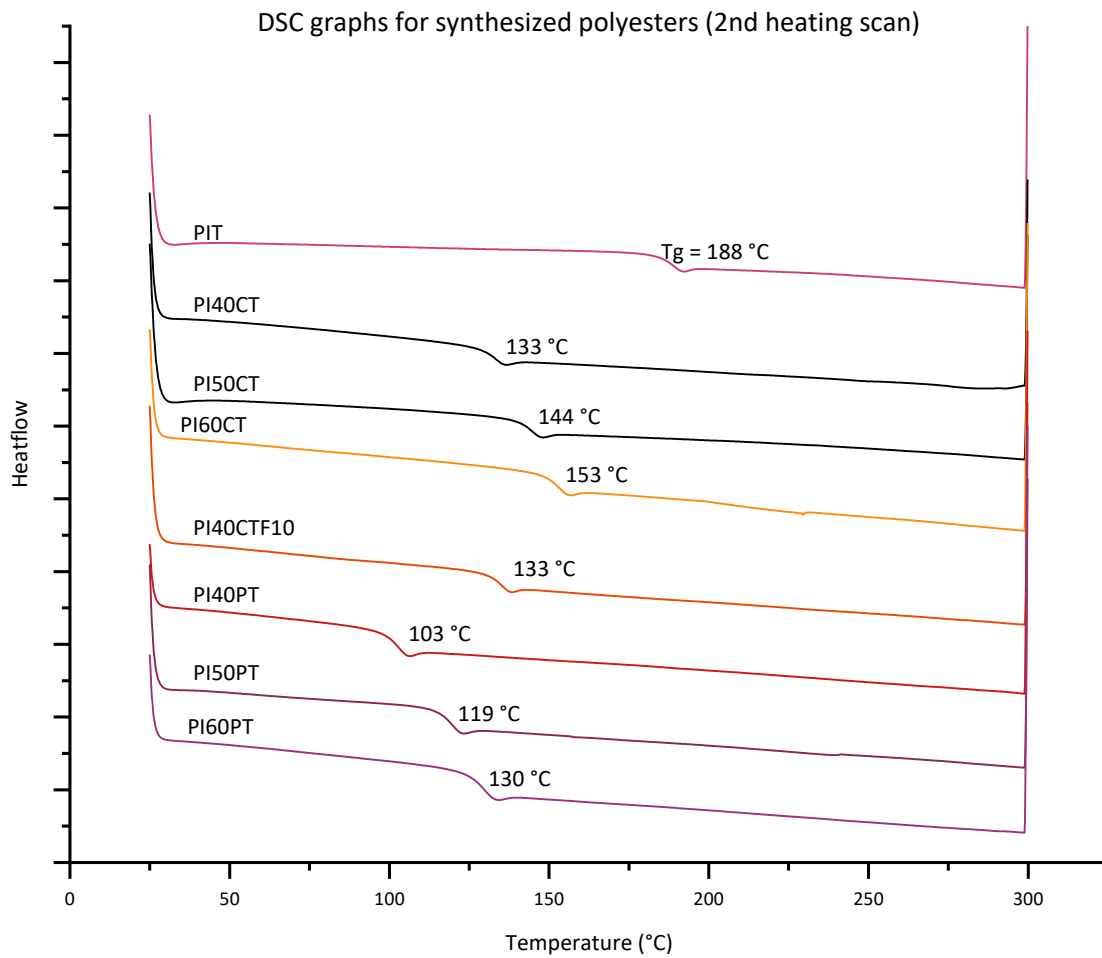
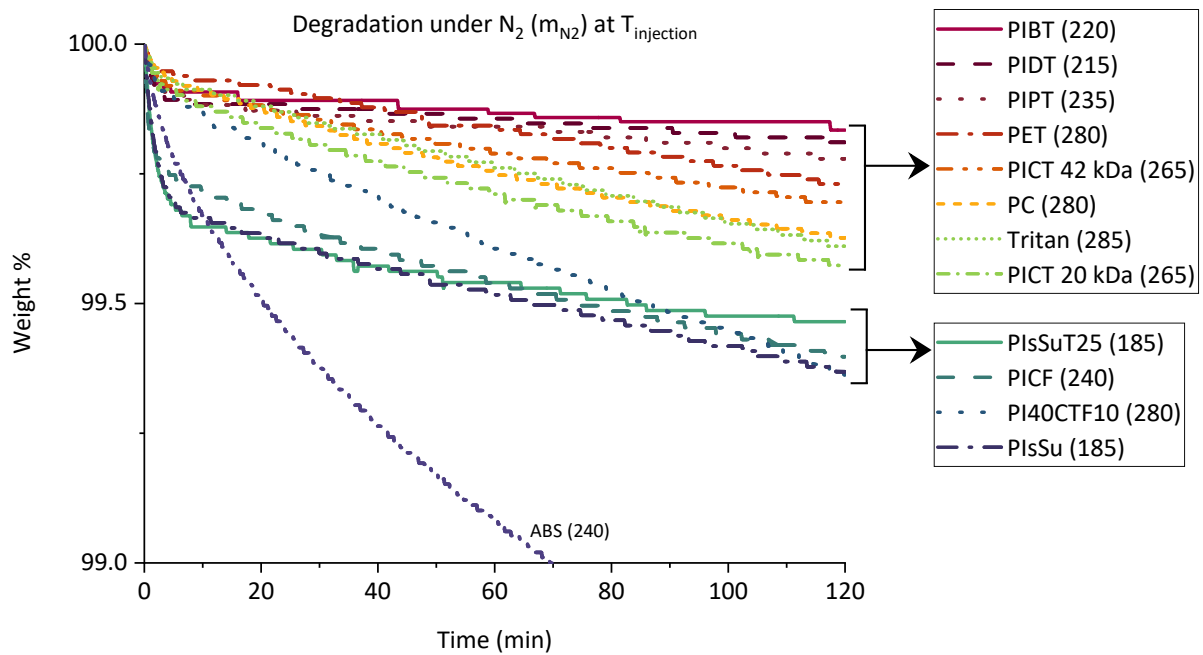


Figure 4 - DSC scans (2nd heating) for synthesized polyesters, with T_g indicated.



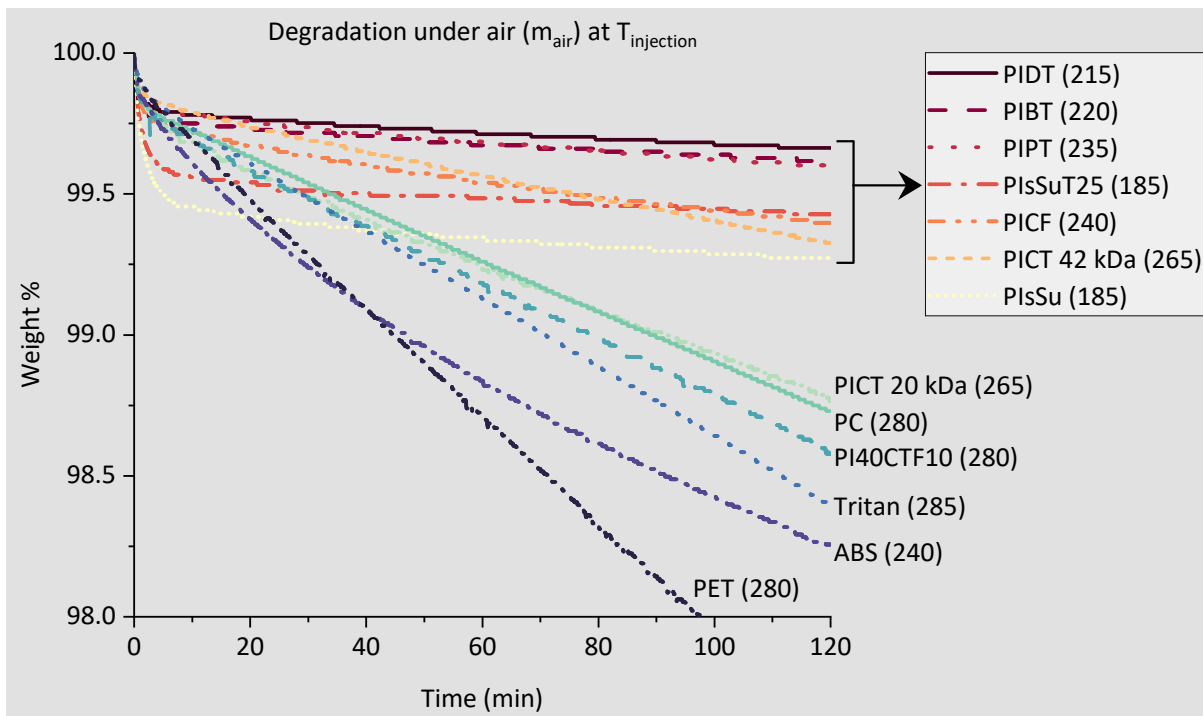
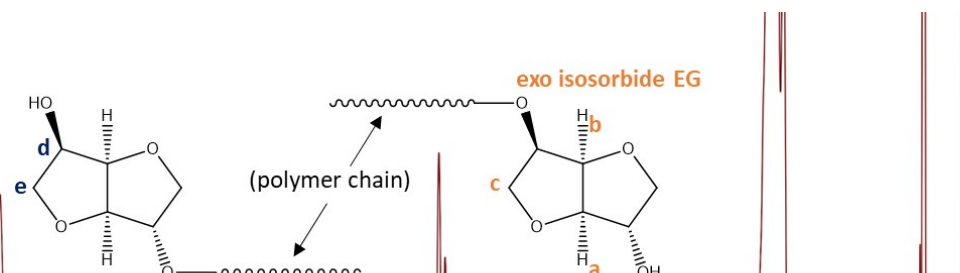


Figure 4 - TGA stability of samples under air (first graph) and nitrogen flow (second). Injection temperatures shown in brackets.



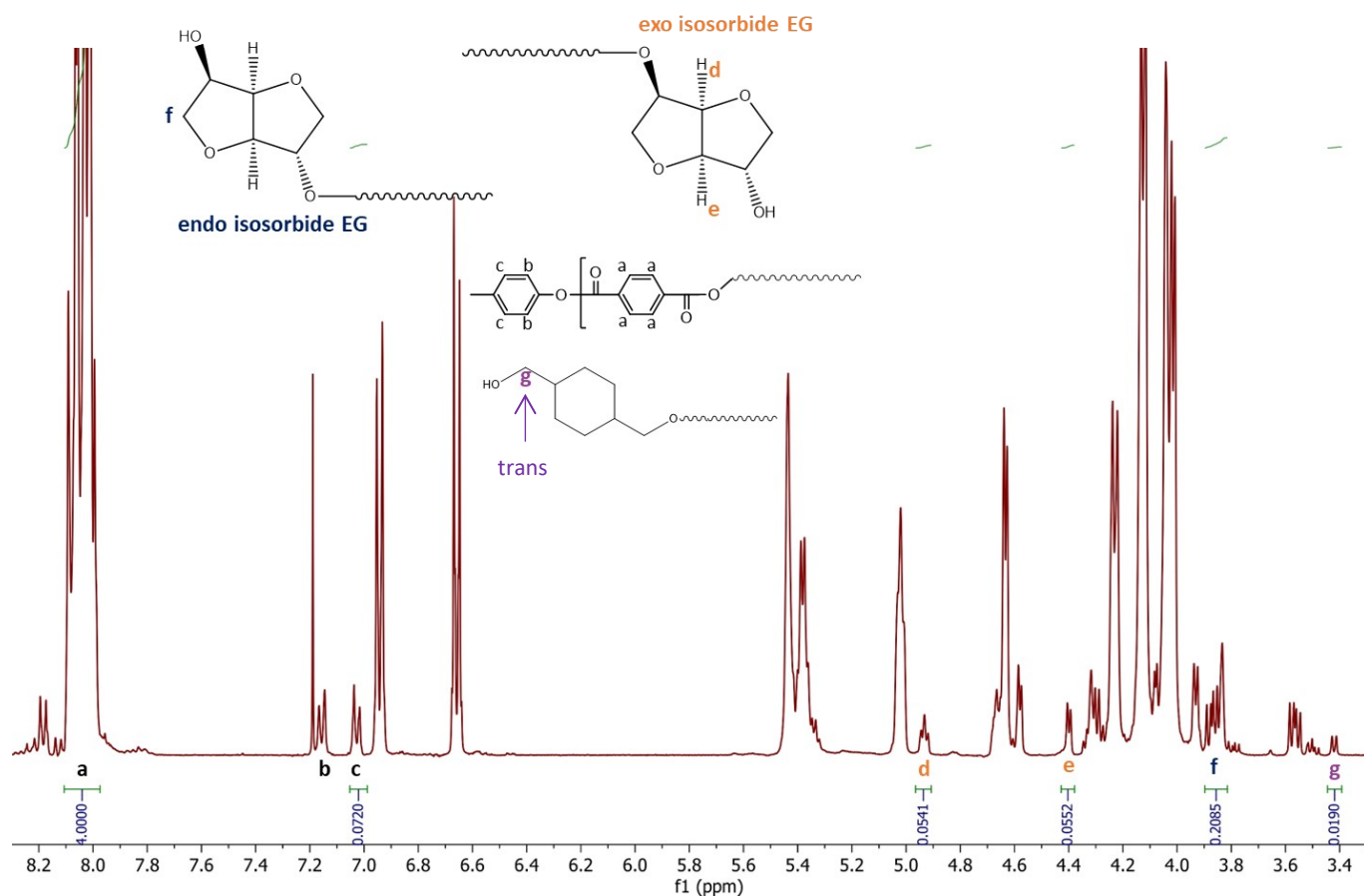


Figure 6 - End groups in PICT highlighted for calculation example. Same reaction as ^{13}C -NMR spectrum, in autoclave with 0.3 equivalent of p-cresol. Current time: 18 h of esterification (pre-vacuum). The integrals are normalized in comparison to the terephthalate group at 8.1 ppm, with 4 protons. This image can be used as an example of the calculations presented in the paper: peak 'c' has 2 protons, so a total of 3.60 mol% of end groups of p-cresol relative to the repeat unit. Analogously, 'd' and 'e' have 1 proton each, totaling ~5.4 mol% of isorbide endo end groups and 'f' has two protons, which totals ~10.4 mol% of exo end groups – summing the two values, isorbide end groups represent ~15.8 mol% compared to the repeat unit. At last, 'g' has 2 protons for the trans CHDM phase (70% of total), so in total $0.95 \text{ mol\%} / 70\% = 1.36 \text{ mol\%}$ total for CHDM.