

Table S1. Protonation constants of **L1** and **L2** in 0.1 M NMe₄NO₃ aqueous solution at 298.1 and 308.1 K

	L1		L2	
	298.1 K	308.1 K	298.1 K	308.1 K
$\mathbf{L} + \mathbf{H}^+ = \mathbf{LH}^+$	10.75(3)	10.90(8)	11.16(4)	11.22(3)
$\mathbf{LH}^+ + \mathbf{H}^+ = \mathbf{LH}_2^{2+}$	10.38(3)	10.45(8)	10.55(4)	10.65(3)
$\mathbf{LH}_2^{2+} + \mathbf{H}^+ = \mathbf{LH}_3^{3+}$	9.83(3)	9.94(1)	10.13(4)	10.20(3)
$\mathbf{LH}_3^{3+} + \mathbf{H}^+ = \mathbf{LH}_4^{4+}$	8.99(5)	8.95(1)	9.22(4)	9.30(3)
$\mathbf{LH}_4^{4+} + \mathbf{H}^+ = \mathbf{LH}_5^{5+}$	8.71(3)	8.87(2)	8.98(4)	9.05(3)
$\mathbf{LH}_5^{5+} + \mathbf{H}^+ = \mathbf{LH}_6^{6+}$	8.33(7)	8.44(2)	8.54(5)	8.64(3)
$\mathbf{LH}_6^{6+} + \mathbf{H}^+ = \mathbf{LH}_7^{7+}$	6.31(1)	6.60(2)	7.62(5)	7.73(3)
$\mathbf{LH}_7^{7+} + \mathbf{H}^+ = \mathbf{LH}_8^{8+}$	4.06(1)	4.14(2)	5.99(5)	6.09(3)
$\mathbf{LH}_8^{8+} + \mathbf{H}^+ = \mathbf{LH}^{9+}$			5.78(5)	6.01(4)
$\mathbf{LH}_9^{9+} + \mathbf{H}^+ = \mathbf{LH}_{10}^{10+}$			5.11(5)	5.18(4)
$\mathbf{LH}_{10}^{10+} + \mathbf{H}^+ = \mathbf{LH}_{11}^{11+}$			4.90(5)	5.04(4)