

Supporting Information

Self-Assembly of Colloidal Sulfur Particles on Glass Surface from Evaporating Sessile Drops: Influence of Different Salts

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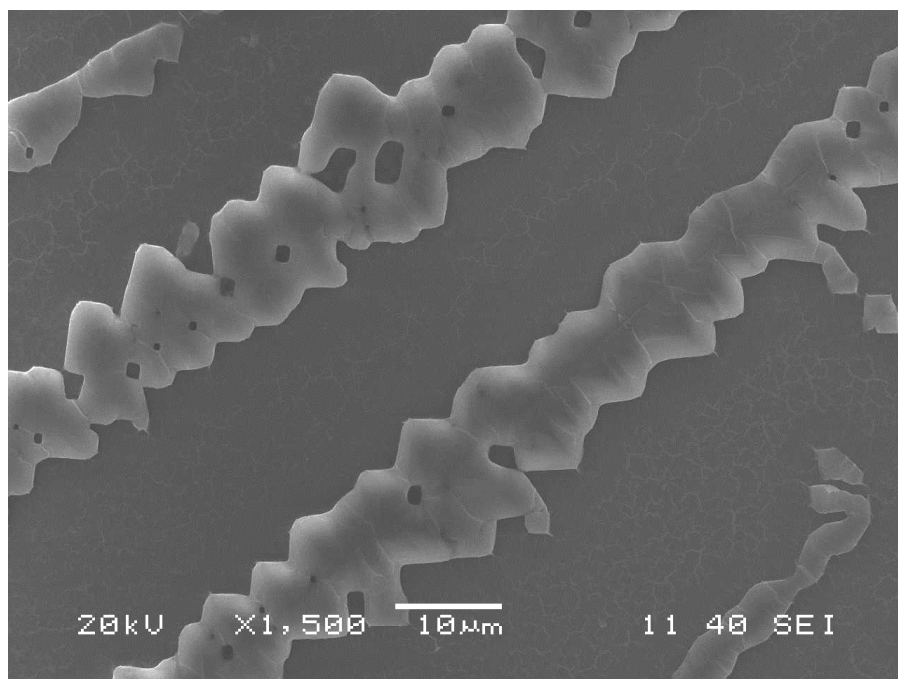


Figure S1. High resolution image of a branch in self-assembled patterns of sodium chloride and sulfur suspension at 10 mM concentration, forming parallel linear structures with uniform jagged edges on both sides. This is usually found near the periphery of the drop.

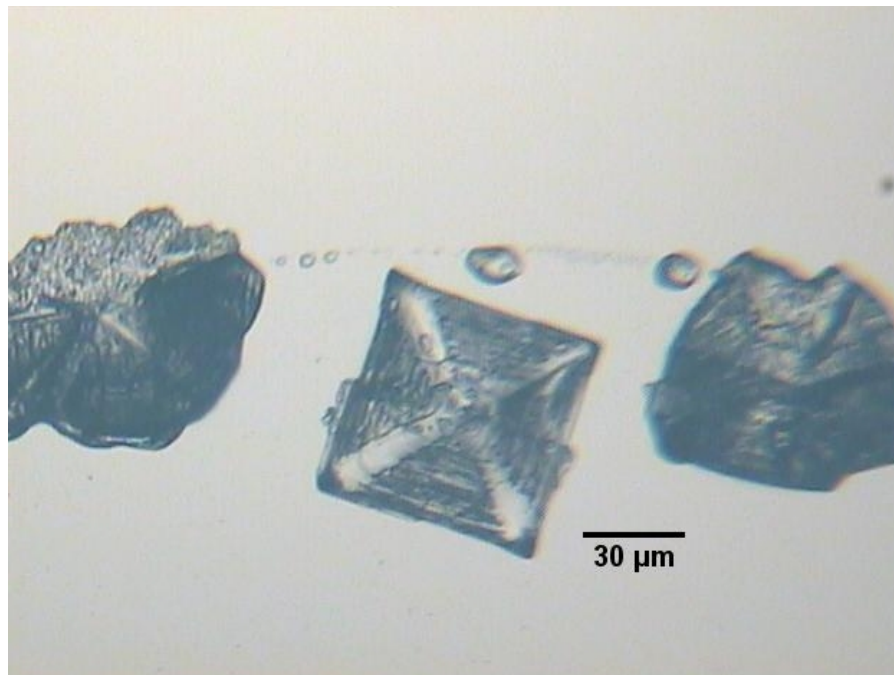


Figure S2. Salt morphology of NaCl at 10 mM concentration. Note the square pyramidal structure formed instead of the octahedral structure because of growth restriction in the XZ plane.

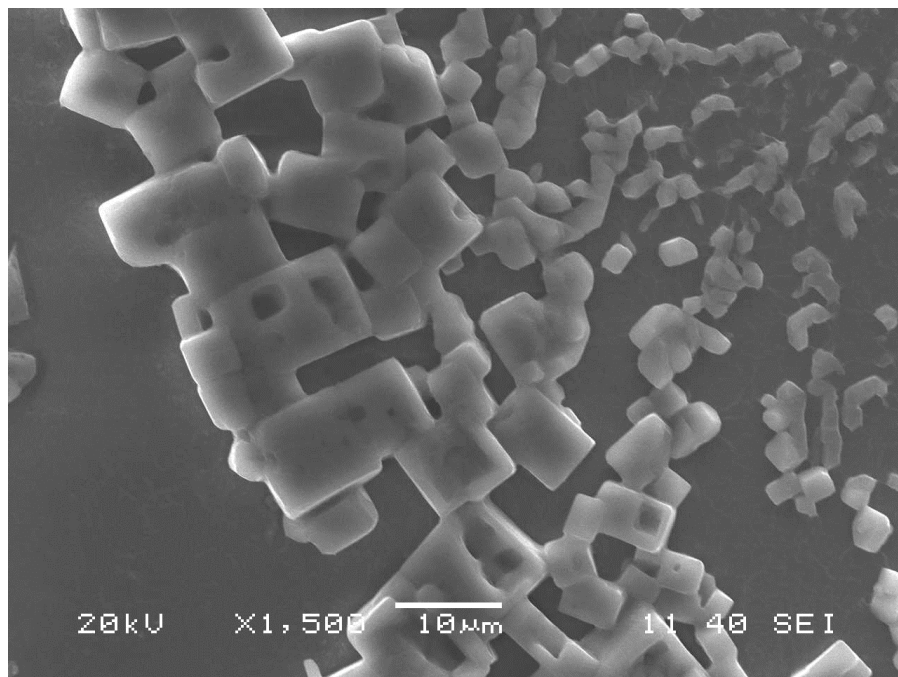


Figure S3. Salt crystal morphology of NaCl and sulfur suspension at 10 mM concentration.

Table: S1 Sulfur particles size in different acid media at different thiosulphate concentration.

Acids	Thioulphate Concentration		
	1 mM	10 mM	30 mM
Hydrochloric acid (HCl)	978.9 ± 21.1 nm	1321.3 ± 30.4 nm	Not Measured
Sulfuric acid (H ₂ SO ₄)	950.6 ± 12.3 nm	1251.8 ± 13.9 nm	
Oxalic acid (C ₂ O ₄ H ₂ ·2H ₂ O)	450 ± 13.2 nm	770 ± 22.7 nm	

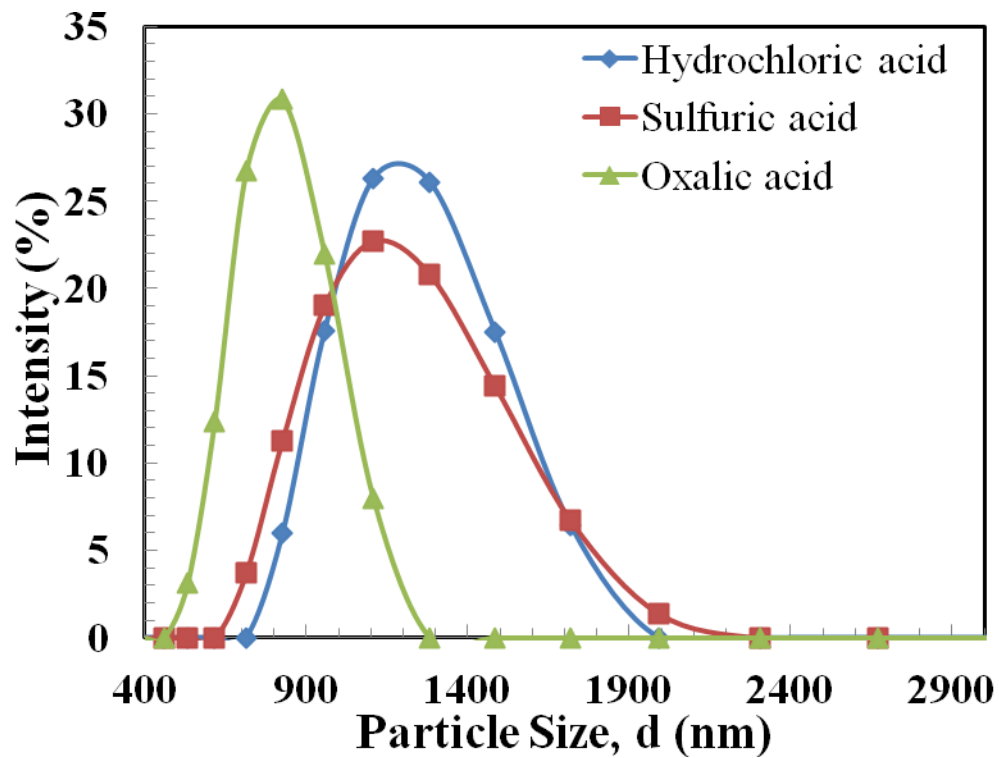


Figure 4: Particle size distribution of sulfur nanoparticles measured by DLS.