Electronic Supplementary Information (ESI) for Organic & Biomolecular Chemistry

Reactions of Dehydrodiferulates with Ammonia

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1. HPLC chromatograms of crude reaction mixtures

After reaction completion (18 h), solvents were removed under vacuum and the product was dissolved in methanol (approximately 15 mL). The crude product, 1 μ L, was analyzed on a Shimadzu LCMS-2020 (using a Kinetex C18 (150 × 4.6 mm, 2.6 μ m, Phenomenex column; detection was at 256 nm via a SPD-M20A photodiode array detector). Aqueous 0.1% formic acid (v/v, solvent system A) and methanol (0.1% formic acid, solvent system B) served as the mobile phase at flow rate of 0.5 mL/min in a gradient mode (Solvent B from 20 to 95% in 35 min, kept at 95% for 7 min, from 45 to 20% in 5 min). Compounds were characterized by their retention times and their ESIMS (in positive and negative modes) and compared with the isolated and structurally characterized compounds (as labeled; see Scheme 1 and Fig. 3 in the main paper).





Fig. S1. HPLC chromatograms of crude reaction mixtures.

2. Proposed mechanisms for some of the reactions

In Scheme S1 (A), acyclic 8–8-diferulate **1** produces a quinone methide intermediate which, after the addition of a hydroxyl group *via* conjugate addition (Michael reaction) and then breaking of the bond through a retro-Aldol type reaction, could form vanillin **1a**.¹⁻³ Intra-molecular attack of the amide nitrogen to the newly formed quinone methide could form compound **1b**. In Scheme S1 (B), similarly, a quinone methide could form from 8–O–4-linked diferulate **3.** Following addition of hydroxide ion to the α -carbon, it could ionize and attack the β -carbon of the β -aryl ether bond (forming an epoxide intermediate through neighboring group participation) to cleave the ether linkage. In Scheme S1 (C), the tetrahydrofuran ring in the 8–5-linked diferulate (compound **4**) could open during the formation of quinone methide followed by abstraction of a proton and re-aromatization. The final retro-Aldol type reaction could yield compounds **4b** and **1a**. In compound **2**, aromatization of ring B to form a naphthalene structure could be the driving force to assist the elimination of H-7' and H-8'. We couldn't explain the formation of some of the products (i.e., **2b-2e**, **1c** and **1d**) by similar nucleophilic type mechanisms; possibly some other types of reactions (e.g., radical reactions via the presence of oxygen as a radical initiator) could promote them.^{4,5}



Scheme S1. Proposed mechanisms for some of the reactions.

3. Photograph of the pressure vessel

Photograph of the used 16 ml stainless steel (grade 316, HEL Inc. USA) pressure vessel with a multi-port lid standing inside a heating block. It was used to perform the AFEX reactions.



Fig. S2. Pressure vessel with a multi-port lid used for AFEX

4. References

- 1. J. Gierer, Holzforschung, 1982, 36, 43-51.
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5. ¹H NMR and ¹³C NMR spectra for all isolated products

¹³C NMR spectrum of **1a** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **1a** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **1b** (125 MHz, DMSO-d₆)



S9

¹H NMR spectrum of **1b** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **1c** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **1c** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **1d** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **1d** (500 MHz, DMSO-d₆)



HMBC spectrum of 1e (DMSO-d₆)



S15

¹H NMR spectrum of **1e** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **2a** (125 MHz, DMSO-d₆)



S17

¹H NMR spectrum of **2a** (500 MHz, DMSO-d₆)



¹H NMR spectrum of **2b** (500 MHz, DMSO-d₆)



HMBC spectrum of 2b (DMSO-d₆)



¹³C NMR spectrum of **2c** (125 MHz, DMSO-d₆)



S21

¹H NMR spectrum of **2c** (500 MHz, DMSO-d₆)



HMBC spectrum of 2d (DMSO-d₆)



S23

¹H NMR spectrum of **2d** (500 MHz, DMSO-d₆)



HMBC spectrum of 2e (DMSO-d₆)



S25

¹H NMR spectrum of **2e** (500 MHz, DMSO-d₆)



S26

HMBC spectrum of 3a (DMSO-d₆)



¹H NMR spectrum of **3a** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **3b** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **3b** (500 MHz, DMSO-d₆)



HMBC spectrum of 3c (DMSO-d₆)



¹H NMR spectrum of **3c** (500 MHz, DMSO-d₆)



HMBC spectrum of 3d (DMSO-d₆)



S33

¹H NMR spectrum of **3d** (500 MHz, DMSO-d₆)



S34

¹³C NMR spectrum of **4b** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **4b** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **4c** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **4c** (500 MHz, DMSO-d₆)



¹³C NMR spectrum of **5a** (125 MHz, DMSO-d₆)



¹H NMR spectrum of **5a** (500 MHz, DMSO-d₆)



HMBC spectrum of **5b** (DMSO-d₆)



¹H NMR spectrum of **5b** (500 MHz, DMSO-d₆)



HMBC spectrum of 5c (DMSO-d₆)



¹H NMR spectrum of **5c** (500 MHz, DMSO-d₆)



6. ESI spectra for 2d and 2e

ESI spectrum of 2d

<Spectrum> Line#:1 R.Time:25:433(Scan#:1527) MassPeaks:220 RawMode:Averaged 25:400-25:467(1525-1529) BasePeak:366(210311) BG Mode:Cale Segment 1 - Event 1 2d mical Formula: C₁₉H₁₅N₃O₅ Molecular Weight: 365 Ch m/z Line#:2 R.Time:25.450(Scan#:1528) MassPeaks:270 RawMode:Averaged 25.416-25.483(1526-1530) BasePeak:364(48436) BG Mode:Cale: Segment 1 - Event 2 30000-m/z

ESI spectrum of 2e



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